

Article

# Photo-Regeneration of Zeolite-Based Volatile Organic Compound Filters Enabled by TiO<sub>2</sub> Photocatalyst

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**Abstract:** Indoor air filtration received significant attention owing to the growing threat to the environment and human health caused by air pollutants such as volatile organic compound (VOC) gases. However, owing to the limited adsorption capacity of VOC adsorbents, such as activated carbon, a rapid breakthrough can occur, reducing the service life of the filter. Therefore, TiO<sub>2</sub>-coated zeolite (TiO<sub>2</sub>/zeolite) was utilized as a photo-regenerative VOC adsorbent to increase the service life of VOC filters. In particular, with photoactive TiO<sub>2</sub> forms on zeolite, efficient and repetitive photo-regeneration is attainable through the dissociation of VOC molecules by the photocatalytic reaction. We optimized the TiO<sub>2</sub> coating amount to obtain TiO<sub>2</sub>/zeolite particles with a high surface area (BET surface area > 500 m<sup>2</sup>/g) and high adsorption capacity. A VOC filter with an adsorption efficiency of 72.1% for formaldehyde was realized using TiO<sub>2</sub>/zeolite as the adsorbent. Furthermore, the TiO<sub>2</sub>/zeolite filter exhibits a photo-regeneration efficiency of >90% for the initial two regeneration cycles using ultraviolet illumination, and >60% up to five cycles. Based on these observations, we consider that TiO<sub>2</sub>/zeolite is a potential adsorbent candidate for photo-regenerative VOC filters.

**Keywords:** volatile organic compounds; photo-regeneration; ultraviolet; zeolite; titanium dioxide

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## 1. Introduction

Recently, indoor air filtration received significant attention owing to the growing threat to the environment and human health caused by small-size air pollutants such as volatile organic compound (VOC) gases [1–5]. However, due to the limited adsorption capacity of VOC adsorbents, including activated carbon, a fast breakthrough can occur, which reduces the service life of the filters [6]. For example, in the case of activated carbon fiber (ACF) filters, full saturation and breakthrough can occur within several hours in a 50–500 ppm toluene environment [7]. Therefore, a new strategy that substantially extends the service life of VOC filters is required; in this regard, the regeneration of VOC filters is a feasible strategy. In particular, through the regeneration process, the adsorbed VOC molecules on the filter media can be desorbed or dissociated, thereby recovering the initial filter performance. Thus, the service life can be extended without increasing adsorption capacity. Various adsorbents and regeneration processes were demonstrated previously. For example, Sidheswaran et al. reported on ACF-based VOC filters regenerated via a thermal desorption process [8]. By heating the filter at approximately 150 °C to provide sufficient energy to remove the VOC molecules from the filter, 70–80% regeneration efficiency was achieved. Lv et al. demonstrated a microwave-based regeneration process using hydrophobically modified NaY zeolite filters [9], wherein microwave irradiation induced heating of water molecules, which desorbed toluene molecules and restored the adsorption capacity. Further, Coss et al. reported microwave-regenerated activated carbon filters [10], in which solvent molecules, such as those of tetrachloroethylene, were desorbed from the inner pores of the active carbon via microwave treatment, recovering

the surface area and adsorption capacity. Next, Liu et al. reported hybrid TiO<sub>2</sub>/zeolite composite synthesized for H<sub>2</sub>S removal and SO<sub>2</sub> capture [11]. Here, regeneration was achieved by washing and calcination. Although these previous approaches for regenerative VOC filters are significant, the thermal regeneration process may cause serious issues, such as overheating of the filter system and large energy consumption, hindering the implementation of these approaches in indoor air filtration. Additionally, the microwave-based regeneration process requires a sophisticated microwave generation unit embedded in the filter system. From this perspective, a simple regeneration process is desired for indoor air filtration to eliminate the possible causes of overheating, and reduce energy consumption and system complexity.

Among the various approaches for regeneration of VOC filters, light-induced photocatalytic regeneration (photo-regeneration) is promising, because the process is simple and does not require complex equipment. The VOC molecules adsorbed on the filter media can be readily desorbed or dissociated by light irradiation, owing to the photocatalytic reaction. Therefore, the use of a heater or a microwave can be excluded. In addition, selecting an appropriate adsorbent that can induce photocatalytic reactions is essential for achieving a high regeneration efficiency. Among the various adsorbents, zeolites have high adsorption efficiency for VOC gases due to their nanoporous structure and charge-compensating cations, which promote the adsorption of polar molecules [11]. However, zeolites are mainly composed of insulating oxides such as Al<sub>2</sub>O<sub>3</sub> and SiO<sub>2</sub> [12]. Therefore, light-induced photocatalytic reactions are significantly limited. To circumvent this limitation and realize sufficient photo-regeneration using zeolites, a combination with photocatalytic materials such as TiO<sub>2</sub> is essential [13]. Furthermore, since TiO<sub>2</sub> can be activated by irradiating ultraviolet (UV) or near-UV visible light, photo-regeneration can be enabled using the TiO<sub>2</sub>/zeolite composite as the adsorbent. For instance, Ichiura et al. prepared a TiO<sub>2</sub>-zeolite sheet by applying a papermaking technique to perform the effective removal of toluene and formaldehyde [14]. Also, Takeuchi et al. suggested TiO<sub>2</sub>/Y-zeolite hybrid photocatalysts by simple impregnation to remove toluene and benzene [15]. Hydrophobic USY zeolite enabled the smooth transfer of aromatic compounds to TiO<sub>2</sub> surfaces, which resulted in high adsorption efficiency. Meanwhile, acid leaching was applied to perform high adsorption and photoactivities of TiO<sub>2</sub>/zeolite composite, according to Zhang et al.'s research [16]. The dealumination and decalcification processes during the acid leaching created more micropores, enhancing the degradation rate of RhB, MO, phenol, and HCHO [16].

To realize photo-regenerative VOC filters, we used TiO<sub>2</sub>-coated zeolite (TiO<sub>2</sub>/zeolite) as the filter medium. TiO<sub>2</sub> was directly formed on the zeolite particles using a sol-gel process, which allowed uniform and conformal coating of the TiO<sub>2</sub> layer on the zeolites. We optimized the TiO<sub>2</sub>-to-zeolite (T:Z) ratio in the TiO<sub>2</sub> coating process to obtain a high surface area and adsorption capacity. UV-regenerative VOC filters were then fabricated on corrugated paper substrates using TiO<sub>2</sub>/zeolite via spray-coating. The TiO<sub>2</sub>/zeolite VOC filters exhibit an adsorption efficiency of 72.1% for formaldehyde; the regeneration efficiency is above 90% for up to two regeneration cycles, and above 60% for five regeneration cycles. These results indicate that TiO<sub>2</sub>/zeolite is a promising candidate for use as a VOC adsorbent in photo-regenerative air filtration systems.

## 2. Materials and Methods

TiO<sub>2</sub> coating on zeolite particles was carried out using a sol-gel method. Initially, 7.5 g of zeolite powder (Zeolite 13X, Jishim Tech, Cheongwon, Chungbuk, South Korea) was added to a mixed solution of ethanol (396 mL) and deionized (DI) water (4 mL). The solution was then stirred at 750 rpm for 30 min to form a zeolite suspension. For the TiO<sub>2</sub> precursor solution, 2.5, 5, 10, or 15 g of titanium (IV) isopropoxide (97% gravimetric, Sigma-Aldrich, St. Louis, MO, United States) was dissolved in 100 mL of ethanol, and mixed using a vortex mixer. The TiO<sub>2</sub> precursor solution was then added dropwise to the zeolite suspension. The mixed solution of the zeolite suspension and the TiO<sub>2</sub> precursor

was stirred at 750 rpm for 30 min, and aged for 3 h to form precipitates. The mass ratio of the TiO<sub>2</sub> (T) precursor solution and zeolite (Z) suspension solution varied in the range 0.5:1–2:1. Subsequently, to remove the excess TiO<sub>2</sub> precursor that was not coated on the zeolite, centrifugation was performed at 6000 rpm for 10 min. Finally, the collected particles were dried in an oven at 80 °C for 12 h to remove the residual solvent and calcined at 350 °C for 3 h on a hot plate. To investigate the surface morphology and specific surface area of the TiO<sub>2</sub>/zeolite particles, field-emission scanning electron microscopy (FESEM; JSM-7600F, JEOL, Akishima, Tokyo, Japan) and the Brunauer–Emmett–Teller (BET) method (ASAP 2460, Micromeritics, Norcross, GA, United States) were used, respectively. The atomic composition of TiO<sub>2</sub>/zeolite was analyzed using energy-dispersive X-ray spectroscopy (EDS).

VOC filters were fabricated using the synthesized TiO<sub>2</sub>/zeolite particles. A 5 cm × 5 cm corrugate-shaped paper was used as the base substrate. To prepare the solution for spray-coating, DI water and styrene butadiene rubber were used as the solvent and binder, respectively. After a brief stirring of DI water and the binder, the prepared TiO<sub>2</sub>/zeolite particles were added to the solution. The TiO<sub>2</sub>/zeolite particles, binder, and solvent mixture had a mass ratio of 10:1:30. A reference solution was fabricated using bare zeolite particles without TiO<sub>2</sub>. Next, using spray-coating, the prepared solutions were coated on corrugated paper substrates. Finally, the coated samples were dried at 150 °C for 15 min in an oven. The coating and drying processes were repeated three times to achieve an adequate thickness of the zeolite or TiO<sub>2</sub>/zeolite particles on the paper substrate.

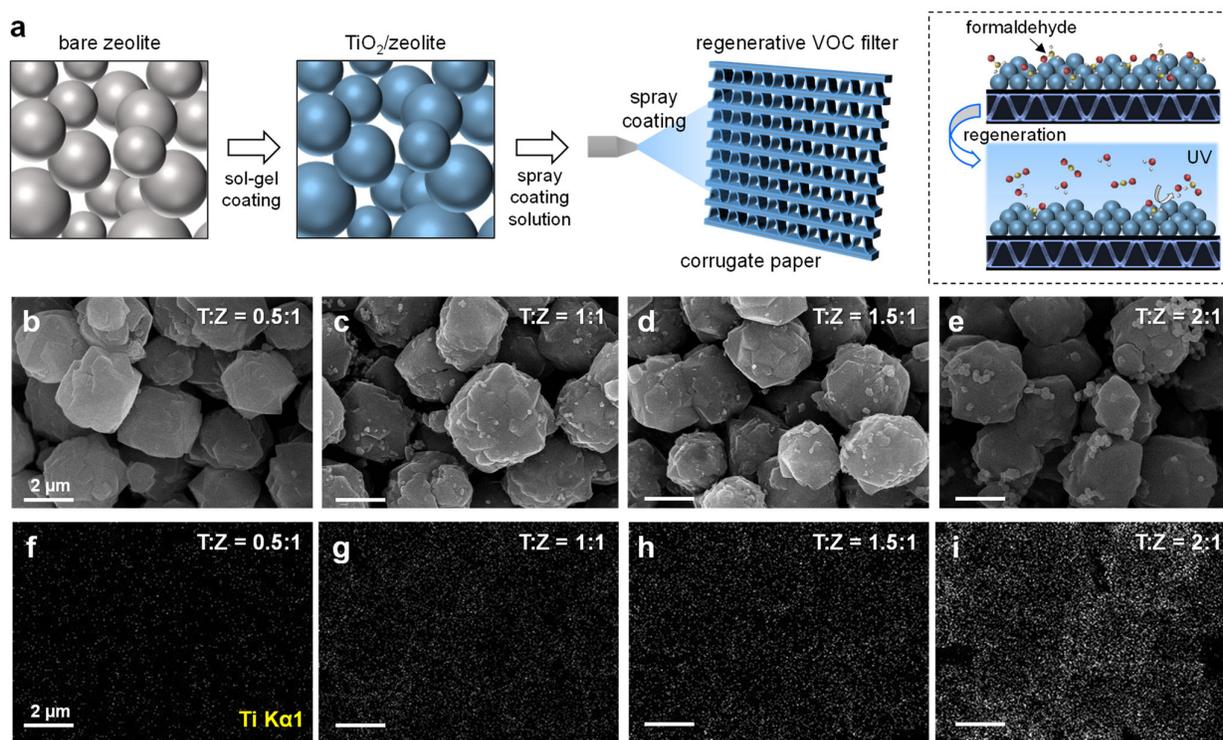
The adsorption efficiencies of formaldehyde and toluene were analyzed using a closed test chamber with a volume of ~0.19 m<sup>3</sup>. First, the fabricated VOC filter was placed inside the chamber, where a fan was installed to circulate air. The formaldehyde or toluene solution was evaporated by placing the solution on a hot plate (80 °C). The initial concentrations of formaldehyde and toluene were measured using a gas detection tube (Gastec detector tube, Gastech, Wangara, WA, Australia). Then, the fan was operated to initiate the air circulation and adsorption of VOC gases. The gas concentration was measured periodically to evaluate the concentration change, and the adsorption efficiency was calculated. For the photo-regeneration of VOC filters, the tested VOC filters were exposed to additional formaldehyde or toluene until the samples became saturated. Then, UV light with a peak emission wavelength of ~253 nm was irradiated onto the filter using a UV lamp (UVT-8CC, Boteck, Gunpo, Gyeonggi, South Korea) for 3 h. After regeneration, adsorption and regeneration efficiencies were evaluated.

### 3. Results

#### 3.1. Fabrication and Characterization of TiO<sub>2</sub>/Zeolite Particles

Figure 1a shows the fabrication procedure and schematic photo-regeneration process for TiO<sub>2</sub>/zeolite-based VOC filters. The TiO<sub>2</sub> layer was coated on bare zeolite particles using a sol–gel process. Then, the fabricated TiO<sub>2</sub>/zeolite particles were dispersed in DI water and spray-coated on a corrugated paper substrate. As described, the TiO<sub>2</sub> layer plays a crucial role in the photo-regeneration process. However, the surface structure of zeolite can be altered during the TiO<sub>2</sub> coating process, which may degrade its adsorption properties [17]. Therefore, we first investigated the influence of TiO<sub>2</sub> coating on the surface morphology, surface area, and pore structure of the zeolites. Figure 1b–e show FESEM images of TiO<sub>2</sub>/zeolite particles synthesized with different TiO<sub>2</sub>-to-zeolite (T:Z) ratios. Here, the T:Z ratio represents the mass ratio of the TiO<sub>2</sub> (T) precursor solution to the zeolite (Z) suspension solution used for the synthesis, which varies in the range of 0.5:1–2:1. As shown in the FESEM images, the zeolite surface becomes rougher as the T:Z ratio increases. In particular, when the T:Z ratio is 0.5:1, the surface is relatively smooth, without large precipitates (Figure 1b). EDS elemental mapping for Ti Ka1 (Figure 1f–i) suggests that a relatively small amount of TiO<sub>2</sub> is coated on the surface. By increasing the

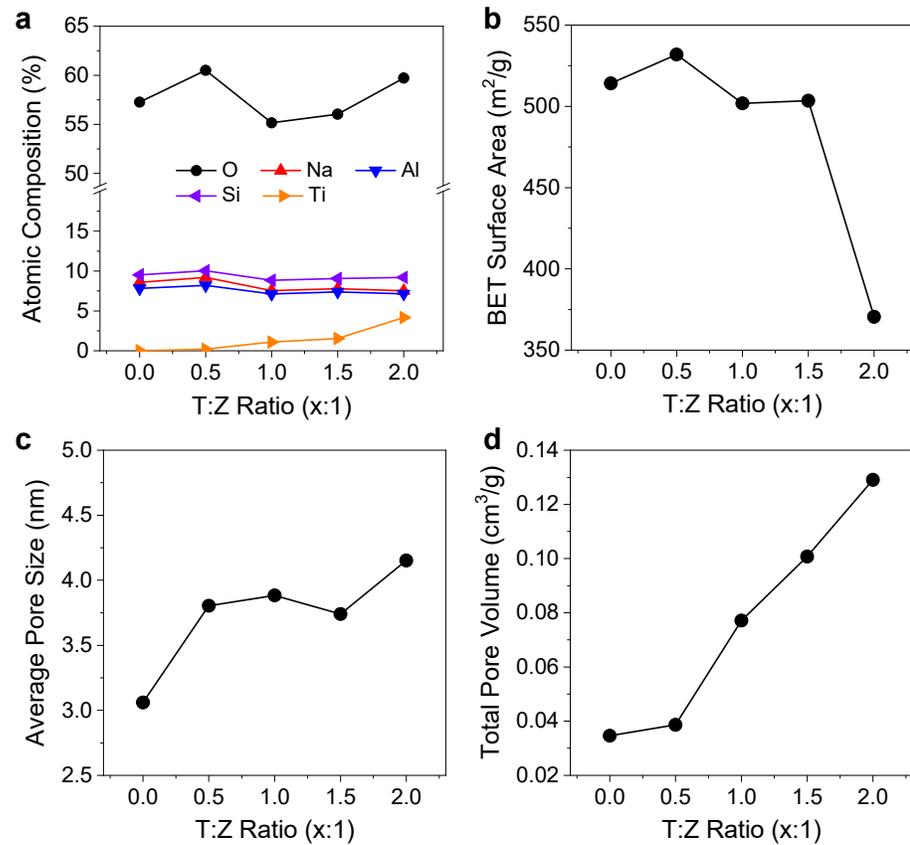
T:Z ratio, the amount of  $\text{TiO}_2$  coated on the surface increases, which is confirmed by the EDS analysis, and the surface of the zeolite becomes rougher, owing to the precipitates formed on the surface. The size of the precipitates is a few hundred nanometers. Furthermore, the number of precipitates tends to increase with the T:Z ratio, and at T:Z = 2:1, the partial agglomeration of  $\text{TiO}_2$  particles is observed, as shown in Figure 1e.



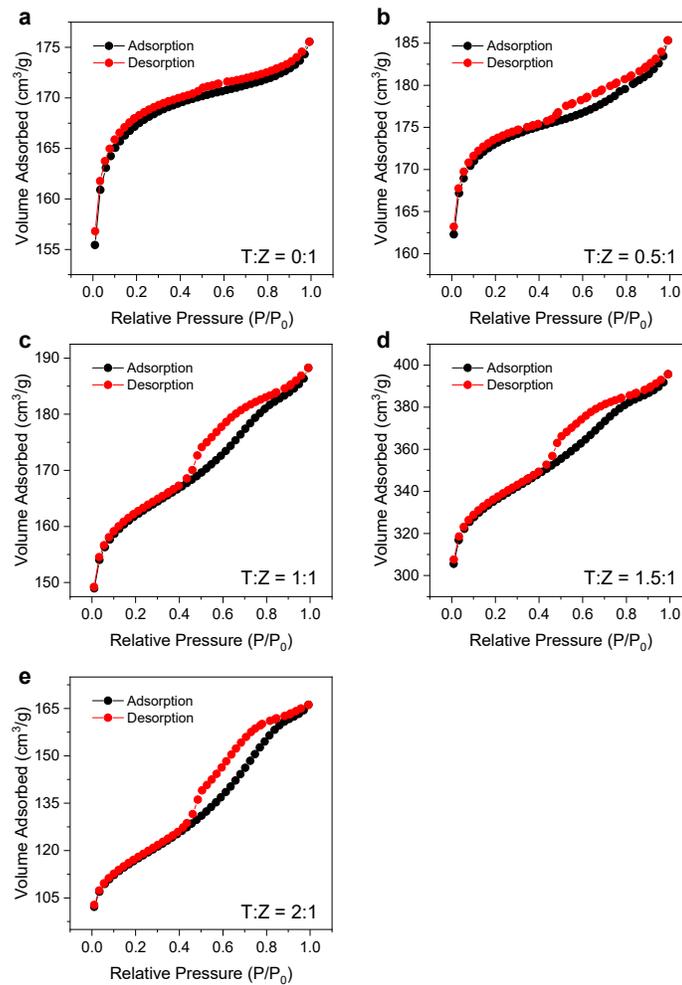
**Figure 1.** (a) Fabrication process and schematic photo-regeneration process of  $\text{TiO}_2/\text{zeolite}$ -based VOC filters. FESEM images of  $\text{TiO}_2/\text{zeolite}$  particles fabricated with different T:Z ratios, (b) T:Z = 0.5:1, (c) 1:1, (d) 1.5:1, and (e) 2:1. Corresponding EDS elemental mapping images (Ti Ka1) of  $\text{TiO}_2/\text{zeolite}$  particles fabricated with different T:Z ratios, (f) T:Z = 0.5:1, (g) 1:1, (h) 1.5:1, and (i) 2:1. The scale bars in the images indicate 2  $\mu\text{m}$ .

The amount of  $\text{TiO}_2$  coating based on the T:Z ratio was investigated using atomic composition analysis. Figure 2a and Table 1 show the atomic composition ratios of zeolite and  $\text{TiO}_2/\text{zeolite}$  particle surfaces as a function of T:Z ratio. Ti shows an increasing tendency with the T:Z ratio. For instance, when the T:Z ratio changes from 0.5:1 to 2:1, the Ti composition ratio increases from 0.21% to 4.18%. This indicates that the amount of  $\text{TiO}_2$  formed on the zeolite can be controlled by the T:Z ratio, which affects the degree of photocatalytic reaction under UV illumination. Despite this, we examined the variation in the specific surface area, total pore volume, and pore size of  $\text{TiO}_2/\text{zeolite}$  particles depending on the T:Z ratio, since the formation of  $\text{TiO}_2$  can alter the surface structures of zeolite. Figure 2b shows the BET surface areas of the bare zeolite (T:Z = 0:1) and  $\text{TiO}_2/\text{zeolite}$  fabricated with different T:Z ratios. Corresponding nitrogen adsorption and desorption isotherm data are shown in Figure 3a–e. The BET surface area of the  $\text{TiO}_2/\text{zeolite}$  particles shows a relatively small variation when the T:Z ratio is in the range of 0.5:1–1.5:1, maintaining a value over 500  $\text{m}^2/\text{g}$ , which is similar to that of bare zeolite (514  $\text{m}^2/\text{g}$ ). However, when the T:Z ratio is further increased to 2:1, the BET surface area decreases to 370  $\text{m}^2/\text{g}$ , implying that the surface structure of the zeolite is significantly altered. Furthermore, the average pore size shows a significant increase from  $\sim 3.06$  nm to  $\sim 4.15$  nm when the T:Z ratio is increased to 2:1 (Figure 2c). This indicates that the surface structure is modified at this ratio, which can be attributed to the overloading of the  $\text{TiO}_2$

precursor. Further, the total pore volume shows a monotonous increase with the T:Z ratio, as shown in Figure 2d. Therefore, considering the efficiency of VOC adsorption and the regeneration capability, we consider 1.5:1 as the optimal T:Z ratio for regeneratable VOC filters.



**Figure 2.** (a) Variation in atomic composition of zeolite and TiO<sub>2</sub>/zeolite particle surfaces synthesized with different T:Z ratios (elements: O, Na, Al, Si, and Ti). T:Z ratio of 0:1 indicates bare zeolite particles. Variations in (b) BET surface area, (c) average pore size, and (d) total pore volume of TiO<sub>2</sub>/zeolite particles as a function of T:Z ratio.



**Figure 3.** BET nitrogen adsorption and desorption isotherm data of zeolite and  $\text{TiO}_2/\text{zeolite}$  particles fabricated with different T:Z ratios, (a) T:Z = 0:1, (b) 0.5:1, (c) 1:1, (d) 1.5:1, and (e) 2:1.

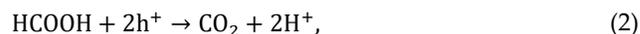
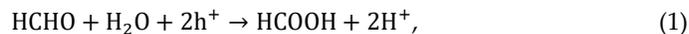
**Table 1.** Atomic compositions of zeolite and  $\text{TiO}_2/\text{zeolite}$  particle surfaces synthesized with different T:Z ratios (in %).

T:Z Ratio	O	Na	Al	Si	Ti
0:1	57.27	8.59	7.82	9.52	0
0.5:1	60.50	9.21	8.20	10.03	0.21
1:1	55.16	7.55	7.13	8.85	1.10
1.5:1	56.04	7.78	7.39	9.06	1.54
2:1	59.71	7.52	7.15	9.20	4.18

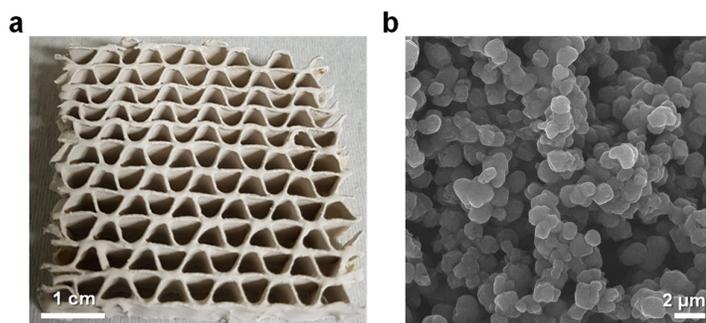
### 3.2. Fabrication of $\text{TiO}_2/\text{Zeolite}$ -Based VOC Filters

Using the optimized  $\text{TiO}_2/\text{zeolite}$  particles (T:Z ratio of 1.5:1), VOC filters were fabricated using the spray-coating method. Here, corrugated paper was used as the filter substrate. Owing to the web-like flute structure of the corrugated paper, a relatively large area for VOC adsorption can be provided, allowing a higher adsorption capacity compared to flat substrates. Figure 4a shows the optical images of the fabricated  $\text{TiO}_2/\text{zeolite}$ -based VOC filter.  $\text{TiO}_2/\text{zeolite}$  particles were uniformly coated on the corrugated paper, including the inner parts of the core flute, owing to the advantage of the spray-coating method. Figure 4b shows the FESEM image of the filter surface. The  $\text{TiO}_2/\text{zeolite}$  formed on the corrugated paper has an open stacking structure. This enabled

easy penetration of VOC gases into the bottom parts of the TiO<sub>2</sub>/zeolite layer, and enlarged the adsorption area. Furthermore, concerning the mechanism of photocatalytic regeneration process, the following photocatalytic reaction can take place. Initially, when the TiO<sub>2</sub>/zeolite filter is exposed to VOC gases, including formaldehyde (HCHO), VOC molecules are adsorbed physically or chemically on the surface depending on the molecular structure of the VOC gas [18,19]. For instance, in the case of formaldehyde, the molecules adsorbed on the TiO<sub>2</sub> surface can be explained with monodentate and bidentate adsorption configurations [18]. Generally, a monodentate configuration depicts the O atom of HCHO (O<sub>F</sub>) bound to the surface under-coordinated Ti atom. For the five-fold Ti<sub>5</sub> atom, binding with the O<sub>F</sub> atom shows the adsorption energy of 0.472 eV and the bond length of 2.271 Å. For the four-fold Ti<sub>4</sub> atom, adsorption energy of 0.907 eV and bond length of 2.168 Å are shown [18]. Based on the Bader charge analysis, part of the electron diffuses during the interaction between HCHO molecules and Ti atom surface. This process results in the weakening of the π bond and the elongation of C=O<sub>F</sub>. Further, chemical changes in the O and C atoms of formaldehyde differ with coverages [19]. At low coverages below the 0.25 monolayer, as in the case of VOC gases, the HCHO molecule binds to surface 5-coordinated Ti atoms via the O atom. However, after continuous exposure to formaldehyde, adsorption saturates, and a breakthrough occurs; thus, the VOC filtration ability is lost. UV illumination was applied to the filter to regenerate it and recover its adsorption capability. Under UV irradiation, TiO<sub>2</sub> induces the oxidation of formaldehyde molecules via the following reactions [20]:



dissociating the adsorbed formaldehyde molecules. In particular, under UV irradiation, electron–hole (e–h) pairs are generated in the TiO<sub>2</sub> layer. The photogenerated holes then contribute to the conversion of formaldehyde into formic acid (HCOOH), and finally to CO<sub>2</sub> [20]. In addition, partial oxidation byproducts such as HCOOH are oxidized to CO<sub>2</sub> with sufficient UV illumination.



**Figure 4.** (a) Optical images of TiO<sub>2</sub>/zeolite VOC filter fabricated using the spray-coating method. (b) FESEM images of TiO<sub>2</sub>/zeolite particles coated on the VOC filter.

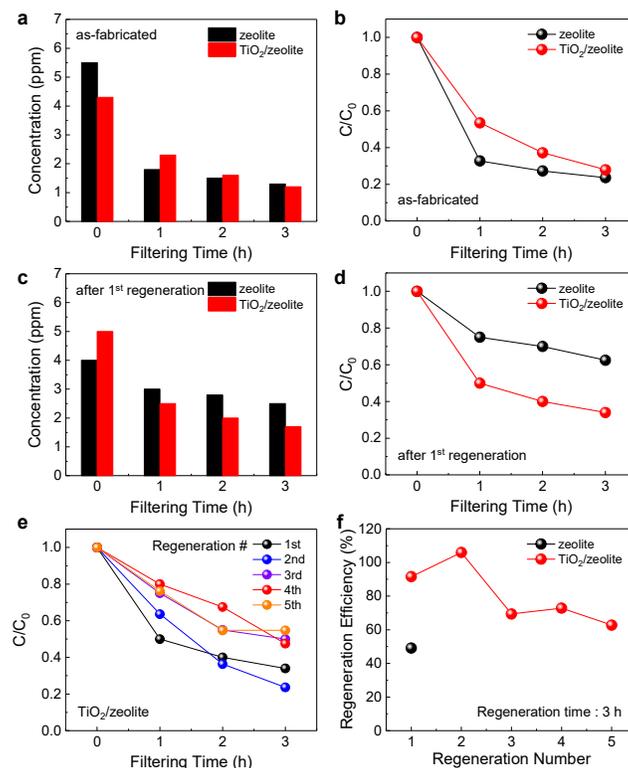
### 3.3. Adsorption Efficiency and Regeneration Efficiency for Formaldehyde

For the TiO<sub>2</sub>/zeolite VOC filter, the adsorption efficiency for formaldehyde was evaluated. To analyze the adsorption efficiency, the filter was placed inside a chamber, and the change in formaldehyde concentration (C) was measured every 1 h for up to 3 h. Figure 5a shows the variation in formaldehyde concentration as a function of filtering time for bare zeolite- and TiO<sub>2</sub>/zeolite-based VOC filters. Figure 5b shows the relative concentration (C/C<sub>0</sub>) change as a function of the filtering time, where C<sub>0</sub> denotes the initial concentration of formaldehyde before filtering. For both filters, the concentration of formaldehyde decreases as the filtering time increases, indicating the effective capture of formaldehyde by the zeolite-based filters. After 3 h of filtering, the C/C<sub>0</sub> values decrease

to 0.236 and 0.279 for the zeolite and TiO<sub>2</sub>/zeolite VOC filters, respectively. The slightly higher C/C<sub>0</sub> value for the TiO<sub>2</sub>/zeolite VOC filter can be attributed to the decrease in the BET surface area after TiO<sub>2</sub> coating. To evaluate the regeneration efficiency, both the zeolite and TiO<sub>2</sub>/zeolite VOC filters were first saturated with formaldehyde gas by placing the filters in a formaldehyde-rich environment. Next, UV exposure was performed for 3 h, and the filtering tests were repeated. Figure 5c,d show the formaldehyde concentration and C/C<sub>0</sub> as functions of the filtering time after the first regeneration. The decrease in C/C<sub>0</sub> for the bare zeolite filter is significantly reduced, indicating degradation of the adsorption efficiency. In contrast, in the case of the TiO<sub>2</sub>/zeolite VOC filter, the decrease in C/C<sub>0</sub> is similar to that before regeneration. To evaluate the repetitive regeneration of TiO<sub>2</sub>/zeolite VOC filters, the regeneration process was performed for five cycles. Figure 5e,f show the C/C<sub>0</sub> and regeneration efficiencies, respectively, for five regeneration cycles. Regeneration efficiency is defined as the following equation:

$$\text{Regeneration efficiency (\%)} = \frac{\text{adsorption efficiency after regeneration}}{\text{Initial adsorption efficiency}} \quad (3)$$

In the absence of TiO<sub>2</sub> (bare zeolite filter), the regeneration efficiency after the first regeneration cycle is below 50% (Figure 5f). This indicates that UV irradiation may have limited effects on dissociating or desorbing the formaldehyde molecules adsorbed on the zeolite surface. However, with TiO<sub>2</sub>, the adsorption efficiency is restored close to its initial value, particularly in the first two regeneration processes, indicating the efficient removal of formaldehyde by the photocatalytic reaction. The regeneration efficiency is over 90% during the first two regeneration cycles, and decreases to approximately 60% afterwards. The decrease in regeneration efficiency after repeated cycles is attributed to non-uniform UV irradiation on the VOC filter owing to the corrugated structure of the paper substrate. Nevertheless, the results show that TiO<sub>2</sub> on zeolite has a significant effect on enhancing the regeneration efficiency of zeolite-based VOC filters.

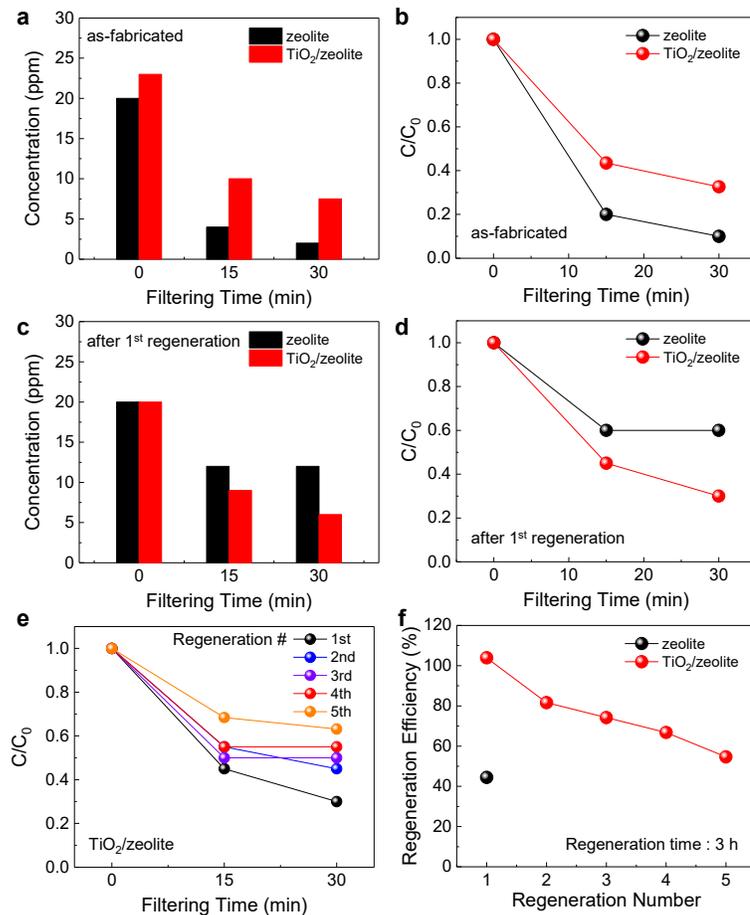


**Figure 5.** (a) Variation in formaldehyde concentration for bare zeolite and TiO<sub>2</sub>/zeolite VOC filters as a function of filtering time. (b) Relative formaldehyde concentration (C/C<sub>0</sub>) changes as a function

of filtering time. Here,  $C_0$  and  $C$  indicate formaldehyde concentration before and during the filtering process, respectively. (c) The variation in formaldehyde concentration and (d)  $C/C_0$  as a function of filtering time after the first regeneration process. (e) Relative formaldehyde concentration ( $C/C_0$ ) change for five regeneration processes (TiO<sub>2</sub>/zeolite filter). (f) Variation in regeneration efficiency as a function of regeneration number.

### 3.4. Adsorption Efficiency and Regeneration Efficiency for Toluene

In addition to formaldehyde, the toluene adsorption and regeneration efficiencies were investigated. Similarly, adsorption efficiency was measured using a closed test chamber. However, in this case, the change in toluene concentration was measured every 15 min for up to 30 min, because of the relatively rapid adsorption of toluene molecules. Figure 6a,b show the toluene concentration change and relative concentration ( $C/C_0$ ) variation for the bare zeolite and TiO<sub>2</sub>/zeolite VOC filters, respectively, as a function of filtering time. Similar to formaldehyde, the bare zeolite filter exhibits a relatively higher adsorption efficiency before regeneration. However, after regeneration by UV exposure for 3 h, the TiO<sub>2</sub>/zeolite VOC filter shows higher adsorption efficiency, as shown in Figure 6c,d. In particular, the bare zeolite filter has a limited adsorption capability, with the  $C/C_0$  value saturated at approximately 0.6. In contrast, the adsorption efficiency of the TiO<sub>2</sub>/zeolite filter is close to that of the as-fabricated filter. The repetitive regeneration performance for toluene was also evaluated. Figure 6e,f show the  $C/C_0$  and corresponding regeneration efficiencies of the bare zeolite and TiO<sub>2</sub>/zeolite filters for five regeneration cycles, respectively. Similar to that of formaldehyde, the regeneration efficiency is significantly lower for the bare zeolite filter, owing to the absence of the TiO<sub>2</sub> photocatalytic layer. In particular, after the first regeneration, the regeneration efficiency of the bare zeolite filter is 44.4%, whereas in the case of TiO<sub>2</sub>/zeolite filters, the regeneration efficiency is above 80% during the initial two cycles, and above 50% up to five cycles. Although further improvements in achieving a high regeneration efficiency are still required, for example, through the optimization of the regeneration process, the implementation of the TiO<sub>2</sub> photocatalytic layer has a substantial influence on improving the regeneration efficiency for various VOC gases. Also, it is of note that the regeneration efficiency obtained by photo-regeneration is comparable to that using thermal- or microwave-assisted regeneration processes, indicating that the synergetic combination of TiO<sub>2</sub>/zeolite adsorbent and photo-regeneration is promising for extending the service life of VOC filters.



**Figure 6.** (a) Variation in toluene concentration for bare zeolite and TiO<sub>2</sub>/zeolite VOC filters as a function of filtering time. (b) Relative toluene concentration ( $C/C_0$ ) changes as a function of filtering time. Here,  $C_0$  and  $C$  indicate toluene concentration before and during the filtering process, respectively. (c) Variation in toluene concentration and (d)  $C/C_0$  as a function of filtering time after the first regeneration process. (e) Relative toluene concentration ( $C/C_0$ ) change for five regeneration processes (TiO<sub>2</sub>/zeolite filter). (f) Variation in regeneration efficiency as a function of regeneration number.

#### 4. Conclusions

In this study, we demonstrate UV-regenerative VOC filters using TiO<sub>2</sub>/zeolite adsorbents. TiO<sub>2</sub>/zeolite particles with surface properties similar to those of bare zeolite are achieved by employing a solution-based coating method and an optimized synthesis protocol. Furthermore, using the spray-coating method, VOC filters were fabricated on a corrugated paper membrane. From the VOC adsorption and UV-mediated regeneration tests for formaldehyde and toluene, the TiO<sub>2</sub>/zeolite filters show ameliorated regeneration efficiency compared to the bare zeolite-based filters. Based on these results, we propose that the TiO<sub>2</sub>/zeolite VOC filter is a potential adsorbent candidate for multi-use UV-regenerative air filtration systems.

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