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Mixed Zinc/Manganese on Highly Reduced Graphene Oxide: A Highly Active Nanocomposite Catalyst for Aerial Oxidation of Benzylic Alcohols

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Abstract: Nanocomposites of highly reduced graphene oxide (HRG) and ZnO_x nanoparticles doped manganese carbonate containing different percentages of HRG were prepared via a facile co-precipitation method. The prepared sample calcined at 300 °C yielded i.e., ZnO_x(1%)–MnCO₃/X%(HRG (where X = 0–7), calcination at 400 °C and 500 °C, yielded different manganese oxides i.e., ZnO_x(1%)–MnO₂/X%(HRG and ZnO_x(1%)–Mn₂O₃/X%(HRG respectively. The prepared catalyst were subjected to catalytic evaluation and a comparative catalytic study between carbonates and oxides for the liquid-phase aerobic oxidation of benzylic alcohols to corresponding aldehydes using molecular oxygen as an eco-friendly oxidant without adding additives or bases. The influence of various parameters such as percentage of HRG, reaction time, catalyst amount, calcination and reaction temperature was systematically examined to optimize reaction conditions using oxidation of benzyl alcohol as a substrate model. It was found that the catalytic performance is remarkably enhanced after using HRG as catalyst co-dopant for the aerobic oxidation of alcohols, possibly owing to the presence of carbon defects and oxygenated functional groups on HRG surface. The as-synthesized catalysts were characterized by SEM, EDX, XRD, Raman, TGA, BET, and FT-IR. Under optimal conditions, the catalyst with composition ZnO_x(1%)–MnCO₃/1%(HRG calcined at 300 °C exhibited remarkable specific activity (57.1 mmol·g^{−1}·h^{−1}) with 100% conversion of benzyl alcohol and more than 99% product selectivity within extremely short time (7 min). The as-prepared catalyst was re-used up to five consecutive times without significant decrease in its activity and selectivity. To the best of our knowledge, the achieved specific activity is the highest so far compared to the earlier reported catalysts used for the benzyl alcohol oxidation. A wide range of substituted benzylic and aliphatic alcohols were selectively oxidized into their corresponding aldehydes with complete convertibility and selectivity in short reaction times without over-oxidation to the acids. Due to their significant low cost, superior reproducibility, excellent catalytic efficiency, the ZnO_x(1%)–MnCO₃/X%(HRG nanocomposites possess several application prospect in other organic chemistry reactions.

Keywords: ZnO_x nanoparticles (NPs); highly reduced graphene oxide; nanocomposite; aerial oxidation

1. Introduction

Selective oxidation of aromatic and aliphatic alcohols to their corresponding carbonyl compounds is a significant and fundamental organic transformation because the carbonyl compound i.e., aldehyde and ketone derivatives formed are important starting material/intermediate in the confectionaries, perfumery, food, dyestuffs, cosmetics, and pharmaceutical industries [1–5]. Nevertheless, oxidation of alcohols into corresponding aldehydes and ketones via conventional methods not only includes harmful, toxic, corrosive oxidizing agents such as (chromate, hypochlorite or permanganate etc.) and harsh condition (such as high temperature and pressure), and also produces a large quantity of contaminant and toxic by-products [6–9]. With increasing environmental concerns, enormous efforts have been made to develop some green processes to overcome these drawbacks [10,11]. Therefore, selective oxidation of alcohols by green approaches using heterogeneous catalysis has gained special attention from the economic and environmental points of view [12–14].

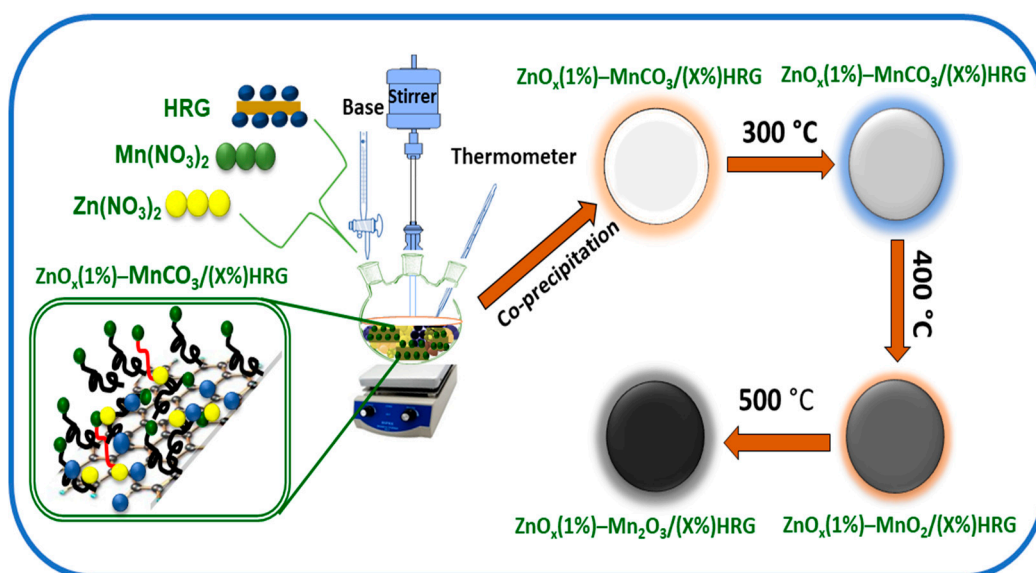
Several studies have been reported using precious metals such as Au [15], Pd [16], Pt [17] and Ru [18] as a catalyst for the aerobic oxidation of alcohols with high catalytic activities and selectivities, often these catalysts have disadvantages owing to being expensive, difficulty in preparation, and rarity of these noble metals that limit their applications [19]. Therefore, use of series of plentiful and cheaper transition metals such as copper [20], cobalt [21], nickel [22], iron [23], vanadium [24], silver [25], chromium [26], molybdenum [27], rhenium [28] and zirconium [29] as catalysts for the oxidation of alcohols to their corresponding aldehydes or ketones is still in demand and also, metal and/or metal oxide nanoparticles catalysts were found to be very effective and efficient for the aerial oxidation of alcohols. Moreover, it has been extensively reported that the catalytic performance of metal oxide nanoparticles catalysts remarkably improved upon doping with other metals probably due to the synergistic effects between metal and metal oxide nanoparticles as well as the significant augmentation of the surface area [30].

Particularly, manganese carbonate (MnCO_3)/manganese oxide is considered one of the inexpensive and most stable metal oxides with high catalytic performance and have been commonly used as a catalyst or catalyst support for the selective oxidation of alcohols into their corresponding carbonyl compounds due to various advantages such as high activity, stable, inexpensive and environmentally friendly catalyst [31,32]. Additionally, various types of manganese oxide, mixed manganese oxide, noble metal doped/supported manganese oxides were employed for the oxidative conversions of several organic moieties such as naphthalene to carbon dioxide (CO_2) [33], carbon monoxide to CO_2 [34], toluene to CO_2 [35], olefins to olefins oxide/epoxide [36], ethylene and propylene to CO_2 [37], cyclohexane to cyclohexanol and cyclohexanone [38], oxidation of benzene to CO_2 and CO [39], oxidation of alkyl aromatics to ketones [40], nitrogen monoxide to nitrogen dioxide [41], 4-tert-butyltoluene to 4-tert-butylbenzaldehyde [42] and formaldehyde to CO_2 [43].

Graphene a single atomically thin layer of sp^2 hybridized carbon atoms in a honeycomb crystal two-dimensional (2D) lattice [44], has been reported for wide range of applications including engineering, electronics, energy storage, drug delivery, lithium ion batteries, gas storage, water purification, sensors, supercapacitor, solar cells, and catalysis [45–51]. In recent years, HRG has received growing interest as promising material for heterogeneous catalyst supports for a plethora of graphene-based metal and metal oxide nanocomposites such as Pd NPs/rGO [52], GO-FePc [53], $\gamma\text{-MnO}_2$ /GO [54] Ag NPs/rGO [55], Au NPs/HRG [56], Pt NPs/HRG [57], rGO/MnCoO [58], Fe_3O_4 -Pt/rGO [59] and TiO_2 /HRG [60], AuPd-GO/ TiO_2 [61] for the liquid phase selective oxidation of alcohols to corresponding carbonyl compounds.

In continuation of our efforts on the use of various mixed metal oxide nanoparticles as an efficient catalyst for the aerobic oxidation of benzylic and aliphatic alcohols [15,25,30,52,62], we report here, a simple and straightforward procedure for the synthesis of ZnO_x nanoparticles doped manganese carbonates and oxides supported with different percentages of highly reduced graphene and employed for aerial oxidation of alcohols in order to understand the effect of presence of HRG in the catalytic system (Scheme 1). Nanocomposites of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(X\%)\text{HRG}$

were synthesized and their catalytic activities examined in selective oxidation of alcohols to corresponding aldehydes and ketones via ideal green procedure. To the best of our knowledge, this is the first report on the nanocomposites of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(X\%)\text{HRG}$ as a catalyst for selective oxidation of alcohols. Moreover, the nanocomposite material was characterized by several techniques such as scanning electron microscopy, energy dispersive X-ray spectroscopy, powder X-ray diffraction, Thermogravimetric analysis, Brunauer–Emmett–Teller analysis, Raman spectroscopy and Fourier-transform infrared spectroscopy.



Scheme 1. Graphical representation of the preparation of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(X\%)\text{HRG}$, $\text{ZnO}_x(1\%)\text{--MnO}_2/(X\%)\text{HRG}$, and $\text{ZnO}_x(1\%)\text{--Mn}_2\text{O}_3/(X\%)\text{HRG}$ nanocomposites.

2. Results and Discussion

2.1. Characterization of Catalysts

2.1.1. X-ray Diffraction Analysis (XRD)

The crystalline structures of pristine graphite, graphene oxide (GO), highly reduced graphene oxide (HRG), $\text{ZnO}_x(1\%)\text{--MnCO}_3$, and $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite were examined by XRD. As shown in Figure 1a, graphite powder displays an intense and sharp peak at $2\theta = 26.5^\circ$ which corresponds to the diffraction peak of (002) crystal plane with a interlayer spacing of 3.37 Å [63]. However, GO exhibits broad diffraction peak at $2\theta = 11.8^\circ$ which is a characteristic of GO (Figure 1b). The disappearance of the graphite peak at $2\theta = 26.5^\circ$ and appearance of a new peak at $2\theta = 11.8^\circ$, indicates the formation of GO via graphite oxidation [64,65]. This shift leads to an increase in an inter planar distance from 3.37 Å to 6.41 Å for graphite and GO, respectively, which can be attributed to the inclusion of several oxygenated functional groups and water molecules between the graphite sheets during the oxidation process [64,66]. XRD pattern of HRG exhibited a broad peak at approximately $2\theta = 24.6^\circ$ which is a characteristic structure of HRG and disappearance of diffraction peak at $2\theta = 11.8^\circ$, confirms the complete reduction of GO as described in Figure 1c [67]. The XRD pattern of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite displayed a broad peak around $2\theta = 24.6^\circ$, the characteristic diffraction signal of HRG along with the diffraction pattern of rhodochrosite D- MnCO_3 (JCPDS No.: 1-0981) [68], indicating the formation of a composite (Figure 1e). As expected the characteristic broad peak of HRG at $2\theta = 24.6^\circ$ (Figure 1d) is not found in the XRD pattern of unsupported $\text{ZnO}_x(1\%)\text{--MnCO}_3$ which has the diffraction pattern of rhodochrosite D- MnCO_3 (JCPDS

No.: 1-0981). The used catalyst after 5 cycles of reuse was characterized by XRD and no significant change was observed in the XRD pattern of the used catalyst (Figure S1, Supplementary Materials).

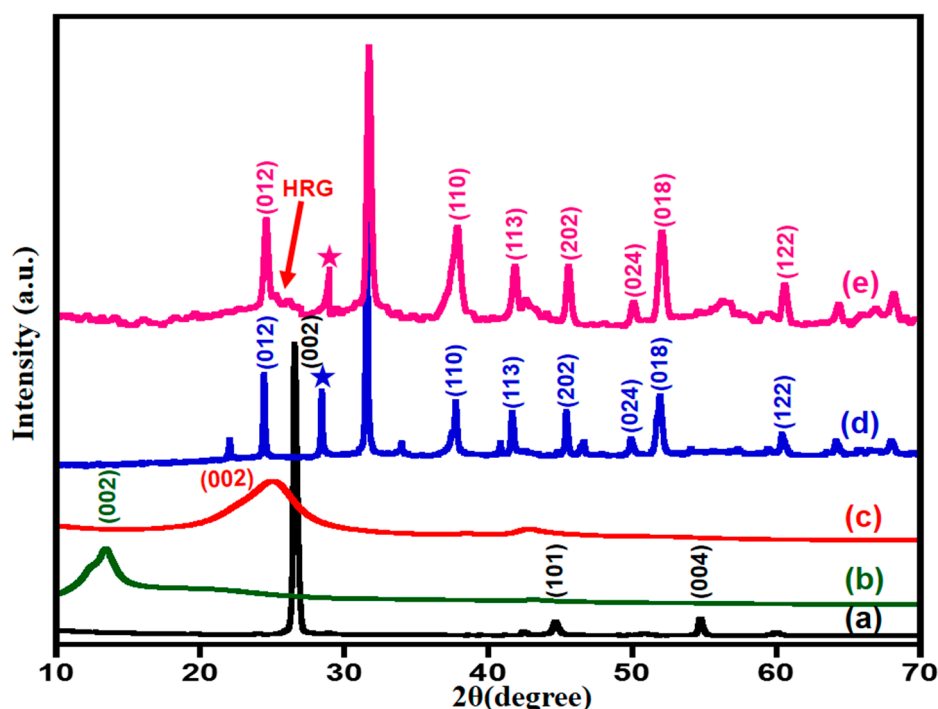


Figure 1. XRD patterns of (a) pristine graphite; (b) GO; (c) HRG; (d) $\text{ZnO}_x(1\%)\text{--MnCO}_3$; and (e) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$. The reflections marked with an asterisk (*) could be attributed to the presence of ZnO_x phase.

2.1.2. Thermogravimetric Analysis (TGA)

Thermogravimetric analysis (TGA) is performed to determine the stability of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite. TGA confirmed that the simultaneous conversion of graphene oxide (GO) to HRG nanosheets via reduction with hydrazine hydrate as reductant. Figure 2 displays the thermal behavior of pristine graphite, graphene oxide, highly reduced graphene oxide, $\text{ZnO}_x(1\%)\text{--MnCO}_3$, and $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite samples. As described in Figure 2a,b the thermal stability of GO nanosheets is extremely lower than the pure graphite. Pure graphite sample exhibits a total weight loss of around 1% in the temperature range (25–800 °C) as illustrated in Figure 2a [69,70]. On the contrary, GO shows ~6% weight loss around 100 °C, obviously due to the evaporation of absorbed water molecules and volatile impurities. It is followed by a main weight loss of about 43% due to pyrolysis of the labile oxygen-containing functional groups (hydroxyl, carbonyl, epoxy, and carboxyl groups) in the temperature range of 200–370 °C [71–73]. Lastly, a weight loss (11%) occurs in the temperature range of 370–800 °C owing to the combustion of the carbon skeleton [74]. These results are in good agreement with earlier publications [64,75]. The TGA curve of HRG exhibited total weight loss (19%) at the identical temperature range because of the reduction of the oxygen containing functional groups (Figure 2c). The $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite exhibited overall weight loss lower than 23% in the same temperature range (Figure 2e). However, the $\text{ZnO}_x(1\%)\text{--MnCO}_3$ is found to have better thermal stability with just 17%, suggesting that the incorporation of HRG in the preparation of the nanocomposites decreased its thermal stability. These results indicate effective reduction of GO to HRG by removing most of the oxygen functionalities such as hydroxyl, carboxyl, epoxy, and carbonyl groups. Further in order to ascertain the conversion of manganese carbonate to manganese(IV) oxide and manganese(III) oxide as obtained from calcination of the salts and which was confirmed by XRD, thermal studies were carried out under

oxygen and the results obtained indicate that the catalyst displays a weight loss of ~22% up to 400 °C which corresponds to the conversion of MnCO_3 to MnO_2 , this further converts of Mn_2O_3 which is confirmed by a weight loss of ~9% around 500 °C, supporting the phases obtained during calcination at these temperatures. The thermogram obtained is given in Supplementary Materials (Figure S2).

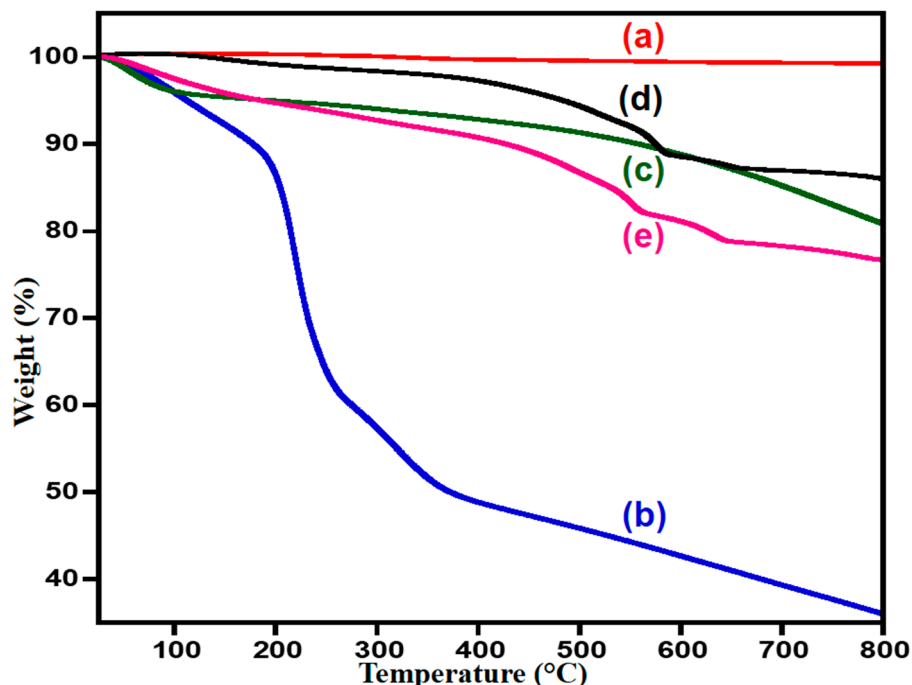


Figure 2. TGA curves of (a) graphite; (b) GO; (c) HRG; (d) $\text{ZnO}_x(1\%)\text{--MnCO}_3$; and (e) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$.

2.1.3. Fourier Transforms Infrared Spectrometer (FT-IR)

Figure 3 presents FT-IR spectrum of GO, HRG, and $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ samples. The FT-IR spectra of GO (Figure 3a) exhibits an intensive broad peak around 3430 cm^{-1} assigned to hydroxyl groups (O–H) stretching vibrations of (–COOH) groups and water molecules [76]. The sharp peak at 1740 cm^{-1} correspond to (C=O) stretching of (–COOH) on the surface [77], and the strong absorption peak at approximately 1630 cm^{-1} is due to the stretching vibration of carbon backbone (C=C/C–C) [78,79]. In addition, the three absorption bands located at 1397, 1225, and 1060 cm^{-1} are assigned to stretching vibrations of (C–OH), (C–O–C) and (C–O), correspondingly [76,80]. In HRG spectra obtained from the reduction of GO in hydrazine monohydrate (Figure 3b), the broad band around 1214 cm^{-1} corresponds to (C–OH) stretching vibration and weak peak around 1634 cm^{-1} is related to (C=C) group owing to the skeletal aromatic vibration. Other peaks corresponding to oxygenated functional groups disappeared [64]. FT-IR spectrum of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite clearly shows that the complete reduction of most of oxygen containing functional groups in the GO surface (Figure 3c). The strong peaks belong to (C=O), (C–O–C), and (C–O) stretching vibrations at 1740, 1225, and 1060 cm^{-1} were not observed, indicating that the oxygen containing functional groups on GO surfaces were almost removed. In addition to, weak peak at 1634 cm^{-1} corresponding to (C=C) stretching modes attributed to the skeletal aromatic vibration is observed [81,82]. In addition, two sharp absorption peaks located at 730 and 860 cm^{-1} , and a broad vibration band at 1450 cm^{-1} , are the characteristic absorption peaks of MnCO_3 [83,84], these results are in good agreement with XRD data. Meanwhile, absorption band at around 580 cm^{-1} could be due to different oxides of manganese [85].

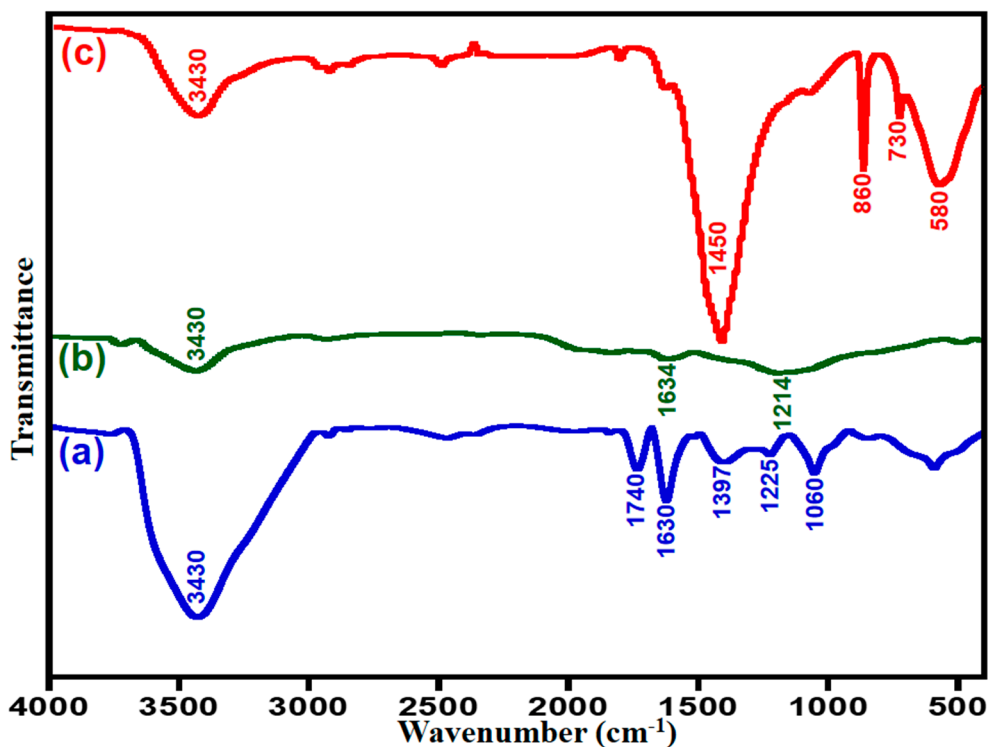


Figure 3. FT-IR spectra of (a) GO; (b) HRG; and (c) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite.

2.1.4. Scanning Electron Microscopy (SEM) and Elemental Analysis (EDX)

The size and morphology of the as-prepared $\text{ZnO}_x(1\%)\text{--MnCO}_3$ and highly reduced graphene oxide supported $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ catalyst obtained by co-precipitation is monitored using scanning electron microscope (SEM). Figure 4 displays the SEM images of the un-supported $\text{ZnO}_x(1\%)\text{--MnCO}_3$ after calcining at 300 °C (Figure 4a) and HRG supported $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ (Figure 4b). The $\text{ZnO}_x(1\%)\text{--MnCO}_3$ catalyst shows micro size but well-defined cuboidal shape particles but HRG supported $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ (Figure 4b), shows aggregation of rather smaller size crystals which appear to have grown on the surface through surface nucleation/growth process. The elemental composition of the as-fabricated catalyst is determined using energy dispersive X-ray spectroscopy (EDX) analysis and stays within experimental error to the theoretical composition (Table 1).

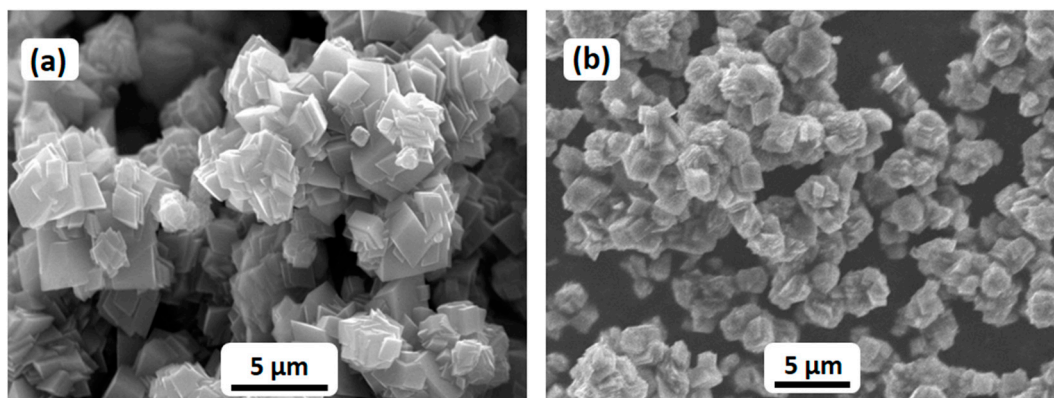


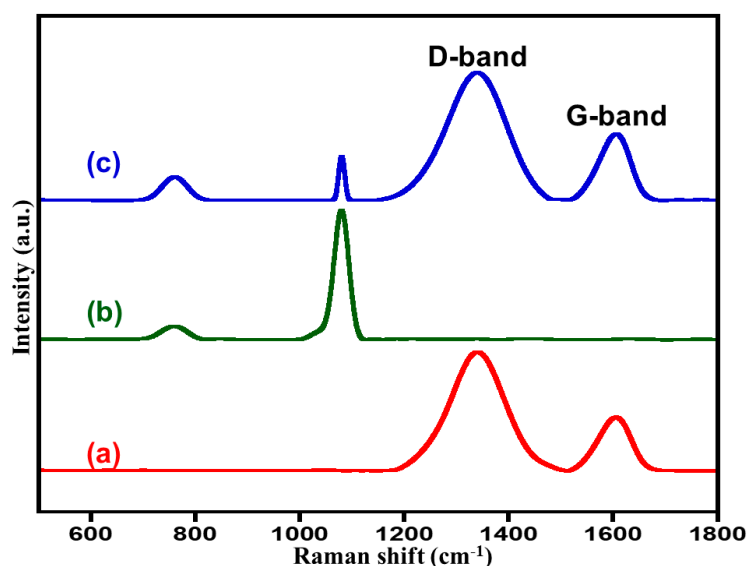
Figure 4. SEM micrographs of (a) $\text{ZnO}_x(1\%)\text{--MnCO}_3$; and (b) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite.

Table 1. Elemental composition of the catalyst $\text{ZnO}_x(1\%)\text{-MnCO}_3$ and $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$.

Catalyst	Element Mass (%)			
	C	O	Mn	Zn
$\text{ZnO}_x(1\%)\text{-MnCO}_3$	6.30	23.65	69.04	1.01
$\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$	26.03	40.33	32.52	1.12

2.1.5. Raman Spectra

Raman spectroscopy is a very informative characterization instrument for studying the structure of graphene and its derivatives. Figure 5 shows the Raman spectrum of graphene oxide (GO), $\text{ZnO}_x(1\%)\text{-MnCO}_3$, and $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$. Raman spectra of $\text{ZnO}_x(1\%)\text{-MnCO}_3$, and $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ nanocomposites exhibited two distinct peaks at 724 and 1086 cm^{-1} which are a finger-print for MnCO_3 . This indicates the existence of MnCO_3 in the nanocomposite consistent with XRD analysis as displayed in Figure 5b,c [86]. Likewise, the existence of HRG in the $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ nanocomposite is also confirmed by the Raman spectra. For this purpose, the Raman spectra of $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ is compared with that of GO, as described in Figure 5a. For the synthesized $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ nanocomposite and GO, two characteristic bands are detected at ~ 1350 and ~ 1590 cm^{-1} respectively, generally denoted as D-band and G-band [87]. The D-band is associated to the disorder in carbon structure induced by lattice defects and the G-band is related to well-ordered structure [88,89]. For GO spectra, the G and the D-bands were shifted and appear at 1605 and 1346 cm^{-1} , respectively, due to the destruction of the sp^2 character by the oxidation process and presence of oxygenated functionalities [90]. While, for the spectrum of $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$, the G characteristic band is shifted by ~ 11 cm^{-1} from 1605 to 1594 cm^{-1} , whereas a slight shift is noticed in the D-band from 1346 to 1342 cm^{-1} , which is an important evidence of the effective reduction of graphene oxide to highly reduced graphene oxide [52].

**Figure 5.** Raman spectra for (a) GO; (b) $\text{ZnO}_x(1\%)\text{-MnCO}_3$; and (c) $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$.

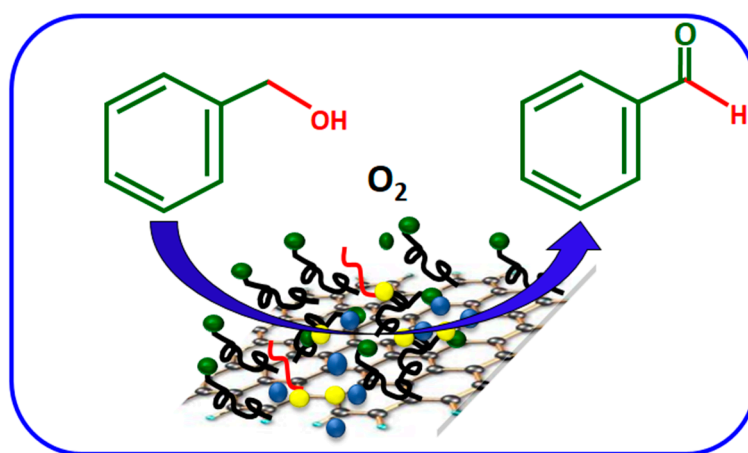
2.1.6. Surface Area Analysis (BET)

In order to determine the surface area of the as-prepared catalyst and study its correlation with the performance of the catalytic system for alcohol oxidation, the BET analysis is studied. The specific surface areas of the catalysts $\text{ZnO}_x(1\%)\text{-MnCO}_3$, $\text{ZnO}_x(1\%)\text{-MnO}_2$, and $\text{ZnO}_x(1\%)\text{-Mn}_2\text{O}_3$ without HRG support is about 120.3, 69.8, and 21.8 $\text{m}^2\cdot\text{g}^{-1}$ correspondingly as mentioned in the Section 2.2.2.

While the surface areas of the nanocomposites after using HRG support have increased to 239.1, 179.3, and 131.1 $\text{m}^2 \cdot \text{g}^{-1}$ respectively. As expected, the surface area of the catalyst prepared using HRG in the catalytic system is higher than the one without HRG and its effect on the catalytic performance on oxidation of benzyl alcohol has been studied and discussed in the later sections.

2.2. Catalytic Evaluation

The primary goal of the present work is the employment of the prepared material for oxidation of organic moieties and benzyl alcohol was selected as substrate for the model reaction, using molecular O_2 as a clean oxidant (Scheme 2). In order to achieve this objective, we have developed nanocomposites of HRG and ZnO_x nanoparticles doped MnCO_3 , which is employed as a catalyst for the aerobic oxidation of benzylic and aliphatic alcohols. It is observed that the oxidation process is significantly influenced by several parameters, such as different percentages of the support (HRG), calcination temperature, catalyst dosage, reaction time, and reaction temperature.



Scheme 2. Aerobic oxidation of benzyl alcohol to benzaldehyde over the as-prepared catalyst.

2.2.1. Role of HRG on the Catalytic Performance

The role of the graphene support in the performance of the heterogeneous catalysts for the oxidation of alcohols has been widely examined in these publications [53,67,85,91–96]. From our earlier publication, it was found that ZnO_x nanoparticles behaved as an excellent promoter in the manganese carbonate catalyst, and the $\text{ZnO}_x(1\%)$ – MnCO_3 catalyst was found to be the best among the various catalysts tested for the aerobic oxidation of aromatic alcohols to their corresponding aldehydes [68]. Hence for the current study, the catalyst $\text{ZnO}_x(1\%)$ – MnCO_3 is selected with an aim to further modify it by adding HRG to the catalytic system to further enhance its catalytic efficiency. Firstly, efforts were exerted in order to find out the suitable amount of HRG (%) in the catalyst to enhance the catalytic efficiency.

An experiment using HRG alone as catalyst, is found to yield negligible conversion product i.e., 2.0%; however when the various% HRG is incorporated in the catalytic system, i.e., $\text{ZnO}_x(1\%)$ – MnCO_3 / $X\%$ (HRG nanocomposites where ($X = 0, 1, 3, 5$, and 7)), it is found to have a significant influence on the catalytic efficiency. The results were summarized in Table 2 and plotted in Figure 6. From the results obtained it is observed that pure MnCO_3 calcined at 300°C gave a 42.6% conversion of benzyl alcohol within 7 min of the reaction, but after doping it with 1% of ZnO_x in MnCO_3 i.e., $\text{ZnO}_x(1\%)$ – MnCO_3 (without HRG), the catalytic performance remarkably improved, and gave 67.5% benzyl alcohol conversion under same circumstances (Table 2, entries 2 and 3). However, after further modification with the incorporation of 1% HRG, the catalyst i.e., $\text{ZnO}_x(1\%)$ – MnCO_3 / (1%)HRG nanocomposite is found to be the best catalyst among all other catalysts including catalysts with increased percentages of HRG. The $\text{ZnO}_x(1\%)$ – MnCO_3 / (1%)HRG catalyst

exhibited a 100% benzyl alcohol conversion after short reaction time only 7 min along with very high specific activity $57.1 \text{ mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$ (Table 2, entry 4). Further increasing the percentage of HRG to 3% led to a slight decrease in the benzyl alcohol conversion to 97.2% (Table 2, entry 4). While the catalyst with 5% and 7% of HRG yielded 85.4% and 62.7% conversion product respectively, under identical circumstances (Table 2, entries 6 and 7). It is also found that all the catalysts tested displayed selectivity towards benzaldehyde, which is found to be >99% (Table 2, entries 1–7). From the above findings, it can be said that the HRG plays an essential role in enhancing the catalytic performance for the oxidation of benzyl alcohol. The higher activity of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite, may be ascribed to the chemical adsorption of the aromatic alcohols onto HRG surface via $\pi\text{--}\pi$ interaction and reaction with the nearest ZnO_x nanoparticles anchored onto the HRG nanosheets as well as the increased surface area that can homogeneously distribute the active sites of mixed metal oxide ($\text{ZnO}_x\text{--MnCO}_3$) nanoparticles. In addition, the presence of carbon defects and oxygenic functionalities on HRG surface, which anchored the ZnO_x nanoparticles, may prevent the agglomeration of graphene layers. Therefore, the $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite is the best catalysts prepared, based on which $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ is utilized for the further studies in order to optimize other parameters.

Table 2. Catalytic results of various catalysts in the selective oxidation of benzyl alcohol.

Entry	Catalyst	Conversion (%)	Specific Activity ($\text{mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$)	Selectivity (%)
1	HRG	2.0	1.2	>99
2	MnCO_3	42.6	24.3	>99
3	$\text{ZnO}_x(1\%)\text{--MnCO}_3$	67.5	38.6	>99
4	$\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$	100.0	57.1	>99
5	$\text{ZnO}_x(1\%)\text{--MnCO}_3/(3\%)\text{HRG}$	97.2	55.6	>99
6	$\text{ZnO}_x(1\%)\text{--MnCO}_3/(5\%)\text{HRG}$	85.4	48.8	>99
7	$\text{ZnO}_x(1\%)\text{--MnCO}_3/(7\%)\text{HRG}$	62.7	35.9	>99

Reaction conditions: catalyst, 300 mg; benzyl alcohol, 2 mmol; calcination temperature, 300 °C; oxygen flow rate, $20 \text{ mL}\cdot\text{min}^{-1}$, toluene, 10 mL; and 7 min reaction time.

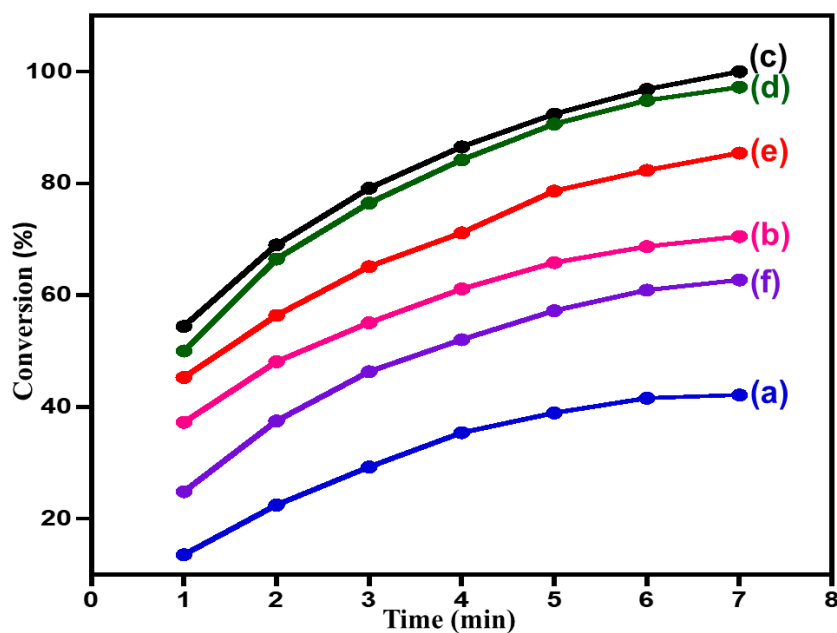


Figure 6. Graphical illustration of benzyl alcohol oxidation using the catalysts (a) MnCO_3 ; (b) $\text{ZnO}_x(1\%)\text{--MnCO}_3$; (c) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$; (d) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(3\%)\text{HRG}$; (e) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(5\%)\text{HRG}$; and (f) $\text{ZnO}_x(1\%)\text{--MnCO}_3/(7\%)\text{HRG}$.

2.2.2. Influence of Calcination Temperature

Calcination temperature has a significant effect on the composition of the synthesized catalysts, such as changes in the oxide composition, surface morphology and surface area. These factors in turn, effect the catalytic efficiency of the as-synthesized catalyst. Therefore, the catalytic activity of the as-prepared catalyst calcined at various temperatures 300 °C, 400 °C, and 500 °C is also investigated in order to get the optimum calcination temperature that gives best catalytic performance, the results are summarized in Table 3 and plotted in Figure 7. It is found that the reactivity of benzyl alcohol is substantially affected by calcination treatment. In our previous study it was found that the prepared $\text{ZnO}_x(1\%)\text{-MnCO}_3$ upon calcination at 400 °C yielded $\text{ZnO}_x(1\%)\text{-MnO}_2$. Calcination at 500 °C resulted in the formation of $\text{ZnO}_x(1\%)\text{-Mn}_2\text{O}_3$. Catalytic evaluation, shows the conversion product obtained within 7 min of reaction time to be 67.5%, 40.8%, and 21.9% respectively (Table 3, entries 1–3). The results indicated that calcination temperature have impeding effect on the oxidation process by decreasing the benzyl alcohol conversion [12]. However in the present study upon the inclusion of HRG in the catalytic system, the catalyst calcined at 300 °C i.e., $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ gave a 100% conversion of benzyl alcohol with 57.1 specific activity within 7 min (Table 3, entry 4). While the ones calcined at 400 °C and 500 °C i.e., $\text{ZnO}_x(1\%)\text{-MnO}_2/(1\%)\text{HRG}$ and $\text{ZnO}_x(1\%)\text{-Mn}_2\text{O}_3/(1\%)\text{HRG}$, respectively, yielded 77.2% and 56.5% conversion product, respectively under similar conditions (Table 3, entries 5 and 6). Additionally, in case of 400 and 500 °C calcination temperatures, there is a decrease in specific activity, which gives only 44.1 and 32.3 $\text{mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$ with no change with the product selectivity during all oxidation experiments. The catalytic performance has remarkably increased after incorporating the HRG in the catalytic system, which indicated that the graphene plays an essential role in enhancing the efficacy of the catalytic system. Interestingly, these observations also suggest that the surface area of the catalyst also plays an essential role in the catalytic performance of the prepared material. Upon BET analysis of the prepared catalyst the surface area of it is found that the catalyst calcined at 300 °C i.e., $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$, which yielded a 100% alcohol conversion product, possesses the highest surface area among the catalyst calcined at 400 °C and 500 °C, which were found to possess lower surface area and also have yielded lesser benzyl alcohol conversion. The specific surface area in case of 400 and 500 °C calcination temperature i.e., $\text{ZnO}_x(1\%)\text{-MnO}_2/(1\%)\text{HRG}$ and $\text{ZnO}_x(1\%)\text{-Mn}_2\text{O}_3/(1\%)\text{HRG}$, respectively have lower surface area, could be due to the sintering. It is noteworthy to mention that sintering not only results inter-particles, but it also minimizes the porosity within a single particle [97]. Thus, it can be concluded that apart from the inclusion of HRG, the catalytic activity is also strongly influenced by calcination treatments of the catalyst. Hence, we used 300 °C calcination temperature in order to optimize other parameters. The obtained results include benzyl alcohol conversion, surface area, specific activity and selectivity over the catalyst were tabulated in Table 3 and plotted in Figure 7.

Table 3. Effect of the calcination temperature on the catalytic properties.

Entry	Catalyst	T (°C)	Surface Area ($\text{m}^2\cdot\text{g}^{-1}$)	Conversion (%)	Specific Activity ($\text{mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$)	Selectivity (%)
1	$\text{ZnO}_x(1\%)\text{-MnCO}_3$	300	120.3	67.5	38.6	>99
2	$\text{ZnO}_x(1\%)\text{-MnO}_2$	400	69.8	40.8	23.3	>99
3	$\text{ZnO}_x(1\%)\text{-Mn}_2\text{O}_3$	500	21.8	21.9	12.6	>99
4	$\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$	300	239.1	100.0	57.1	>99
5	$\text{ZnO}_x(1\%)\text{-MnO}_2/(1\%)\text{HRG}$	400	179.3	77.2	44.1	>99
6	$\text{ZnO}_x(1\%)\text{-Mn}_2\text{O}_3/(1\%)\text{HRG}$	500	131.1	56.5	32.3	>99

Reaction conditions: catalyst, 300 mg; benzyl alcohol, 2 mmol; reaction temperature, 100 °C; oxygen flow rate, 20 $\text{mL}\cdot\text{min}^{-1}$; toluene 10 mL; and 7 min reaction time.

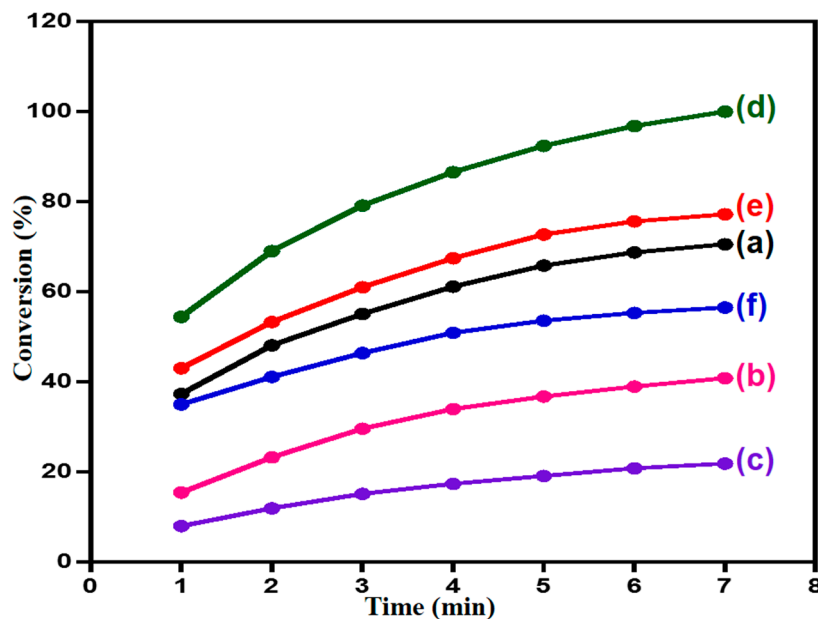


Figure 7. Graphical representation of benzyl alcohol oxidation catalyzed by (a) ZnO_x(1%)–MnCO₃; (b) ZnO_x(1%)–MnO₂; (c) ZnO_x(1%)–Mn₂O₃; (d) ZnO_x(1%)–MnCO₃/(1%)HRG; (e) ZnO_x(1%)–MnO₂/(1%)HRG; and (f) ZnO_x(1%)–Mn₂O₃/(1%)HRG.

2.2.3. Influence of Temperature

Reaction temperature usually plays a fundamental role in enhancing the effectiveness of the catalytic protocol. The influence of reaction temperature on the aerial oxidation of benzyl alcohol is also explored by varying the temperature from 20 °C to 100 °C while using ZnO_x(1%)–MnCO₃/(1%)HRG as catalyst. The obtained results including benzyl alcohol conversion, specific activity, and benzaldehyde selectivity at various temperatures (20, 40, 60, 80, and 100 °C) were compiled in Table 4 and represented in Figure 8. According to Table 4, the benzyl alcohol conversion strongly depends on the reaction temperature. A low conversion of 43.9% was obtained at lower temperature 20 °C (Table 4, entry 1). As expected, the elevated temperature contributed to a higher oxidation rate and led to the remarkable enhance in the catalytic performance of the as-synthesized catalyst. A 100% conversion of benzyl alcohol is achieved at the temperature of 100 °C under similar conditions (Table 4, entry 5), whereas the benzaldehyde selectivity still unchanged during all oxidation experiments (above 99%). The best catalytic activity within the conditions studied is achieved with a reaction temperature of 100 °C, hence is selected for further studies.

Table 4. Effect of reaction temperature on the catalytic evaluation.

Entry	Reaction Temperature (°C)	Conversion (%)	Specific Activity (mmol·g ^{−1} ·h ^{−1})	Selectivity (%)
1	20	43.9	25.1	>99
2	40	58.7	33.6	>99
3	60	73.1	41.8	>99
4	80	86.6	49.5	>99
5	100	100.0	57.1	>99

Reaction conditions: benzyl alcohol, 2 mmol; calcination temperature, 300 °C; amount the catalyst, 300 mg; oxygen flow rate, 20 mL·min^{−1}; toluene, 10 mL; and 7 min reaction time.

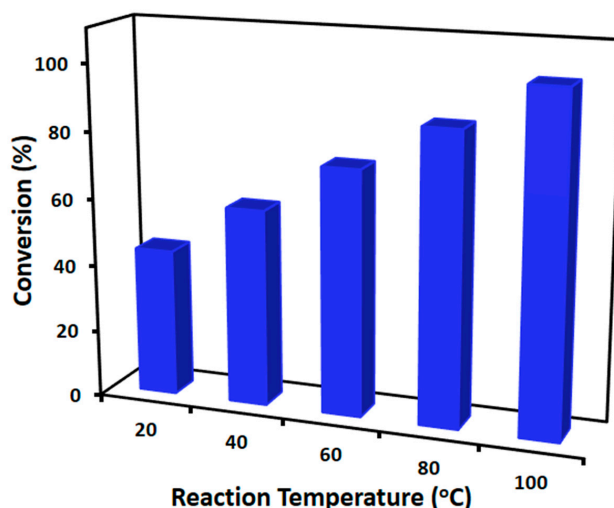


Figure 8. Catalytic activity of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ catalyst as a function of reaction temperature.

2.2.4. Influence of Catalyst Amount

The influence of the catalyst dosage upon the oxidation of benzyl alcohol is also assessed and the obtained results are illustrated in Figure 9 and listed in Table 5. The amount of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite is optimized by employing the reaction with 50, 100, 150, 200, 250 and 300 mg of catalyst under the optimized reaction conditions mentioned earlier. The selectivity towards benzaldehyde is almost unchanged over all oxidation reactions (more than 99%). It is noteworthy that effectiveness of catalytic system is directly proportional to the amount of the catalyst. A poor conversion of 19.4% is obtained in presence of low catalyst amount (50 mg) (Table 5, entry 1), probably due to the fewer catalytic active sites. As expected, by increasing the catalyst amount to 100 mg, the conversion of benzyl alcohol also increase to 36.9% (Table 5, entry 2). When the catalyst amount reaches to 300 mg, the catalyst exhibited superior catalytic activity by yielding a complete conversion with the specific activity of $57.1 \text{ mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$ after short reaction time (7 min) (Table 5, entry 6). While the rest gave lower than 100% conversion under same conditions. This study demonstrates that only 300 mg of the catalyst is required for this oxidation reaction to reach a complete conversion of benzyl alcohol to benzaldehyde within short reaction period.

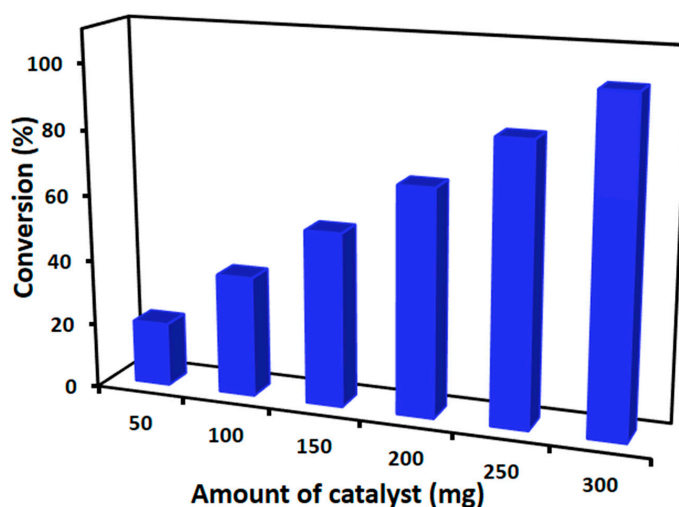


Figure 9. Catalytic performance of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ catalyst as a function of catalyst amount.

Table 5. Effect of the catalyst $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ amount on the catalytic performance.

Entry	Catalyst Amount (mg)	Conversion (%)	Specific Activity ($\text{mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$)	Substrate/Catalyst Molar Ratio	Selectivity (%)
1	50	19.4	11.1	4.56	>99
2	100	36.9	21.1	2.30	>99
3	150	53.2	30.4	1.53	>99
4	200	69.1	39.5	1.15	>99
5	250	85.0	48.6	0.92	>99
6	300	100.0	57.1	0.76	>99

Reaction conditions: benzyl alcohol, 2 mmol; calcination temperature, 300 °C; reaction temperature, 100 °C; oxygen flow rate, 20 $\text{mL}\cdot\text{min}^{-1}$; toluene, 10 mL; and 7 min reaction time.

Under similar conditions, a blank experiment performed over the catalyst with only solvent (toluene) in absence of substrate benzyl alcohol in order to prove that there is no role of toluene in the catalytic oxidation of benzyl alcohol and the benzaldehyde product in all oxidation reactions is obtained only from the aerobic oxidation of benzyl alcohol. No formation of benzaldehyde indicates that the aldehyde is from the catalytic oxidation of benzyl alcohol and not from the oxidation of toluene. Similarly, a blank reaction is carried out without the catalyst at the optimum conditions and which also yielded, no conversion product, indicating that the catalyst is necessary for the benzyl alcohol oxidation. In addition, in order to show the significance of molecular O_2 , the reaction was carried out using $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ as catalyst in presence of air without bubbling. Under the optimized reaction conditions, the results indicate the formation of 27.3% conversion of benzyl alcohol, which is far less than the 100% alcohol conversion obtained when the reaction is performed with molecular O_2 bubbling through the reaction mixture.

A comparison of the catalytic performance of present catalyst with the Zn-based catalysts that already reported in the literature is collected in Table 6 to show that the present catalytic system is indeed superior to several of the other systems. By using $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ nanocomposite, the selective oxidation of benzyl alcohol to benzaldehyde is completed within extremely short reaction time (7 min) at 100 °C with (>99%) benzaldehyde selectivity and higher specific activity ($57.1 \text{ mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$) with respect other Zn-based catalysts.

Table 6. Comparison of the $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ catalyst for the oxidation of benzyl alcohol with previously reported Zn-based catalysts.

Catalyst	Conversion (%)	Selectivity (%)	T (°C)	Time	Specific Activity ($\text{mmol}\cdot\text{g}^{-1}\cdot\text{h}^{-1}$)	Reference
$\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$	100	>99	100	7 min	57.1	This work
$\text{ZnO}_x(1\%)\text{-MnCO}_3$	100	>99	100	0.33 h	16.0	[68]
TCP/PP/Zn- Fe_2O_4 @ZnO	96	100	reflux	1.5 h	32.0	[98]
Ag/ZnO	98	>99	100	8 h	24.5	[99]
Zn-Al LDH	100	100	reflux	2.5 h	8.0	[100]
ZnIn_2S_4	100	>99	room temperature	2 h	6.25	[101]
Au/CuO-ZnO	94	>99	reflux	1 h	18.8	[102]
ZnWO_4	70	100	70	10 h	0.7	[103]
[bmim] ₅ [PW ₁₁ ZnO ₃₉]	100	100	reflux	1.25 h	16.0	[104]
PW ₁₁ Zn@AC	100	>99	reflux	0.75 h	6.7	[105]
ZnSA4	100	>99	80	8 h	2.5	[106]
ZnBr_2	100	91	room temperature	16 h	5.3	[107]
ZnPW_{12}	80.1	89.1	90	3.5 h	21.7	[108]
ZnFe_2O_4	6.9	95.9	90	10 h	0.6	[109]
Zn-Cr-LDH/HD	41.9	91.5	94	5 h	8.7	[110]
1-POM(Zn)	100	100	reflux	4.5 h	4.4	[111]
Zn/ Fe_2O_4	100	60	70	6.5 h	7.7	[112]

According to data presented in Table 6, other Zn-based catalysts listed take a longer reaction time to complete oxidation of benzyl alcohol, higher reaction temperature or have a lower specific activity. For instance, Golafzani et al. [98] reported liquid phase selective oxidation of benzyl alcohol using TCP/PP/Zn- Fe_2O_4 @ZnO catalyst via aqueous hydrogen peroxide as a green oxidant

in CH_3CN as solvent with benzyl alcohol conversion of 96%, excellent selectivity about 100%, and the specific activity of this conversion was about $32 \text{ mmol} \cdot \text{g}^{-1} \cdot \text{h}^{-1}$ within 1.5 h under reflux conditions. Another example, Sarvari et al. [99] have prepared Nano Ag/ZnO catalyst and used for oxidant-free oxidation of benzyl alcohol to benzaldehyde. The nano Ag/ZnO catalyst exhibited 98% benzyl alcohol conversion and more than 99% benzaldehyde selectivity along with $24.5 \text{ mmol} \cdot \text{g}^{-1} \cdot \text{h}^{-1}$ specific activity after long reaction time 8 h at 100°C . On the basis of these findings, it can be said that $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ nanocomposite has found to be the best choice for the alcohol oxidation.

2.3. Catalyst Recyclability

The catalyst recovery has considerable importance from the both industrial and academic point of view. The recyclability of $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ for the selective oxidation of benzyl alcohol via molecular oxygen is studied under optimized circumstances. In order to examine the reusability and stability of the as-synthesized catalyst after completion of the reaction, the toluene is evaporated and fresh toluene (solvent) is added and the mixture is filtered by simple filtration to recover the catalyst. The filtered catalyst is washed sequentially with toluene and was dried at 100°C for 4 h. The results obtained, (Figure 10), it is elucidated that $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ catalyst can be reused for at least 5 cycles without any appreciable loss of catalytic activity and benzaldehyde selectivity for all the recycle tests is found to be $>99\%$. During the five runs investigation, the benzyl alcohol conversion decreased in a range from 100% to 93%, which is probably due to the mass loss during the filtration method [67,113]. Therefore, results obtained indicate that our catalyst has a good reproducibility and stability.

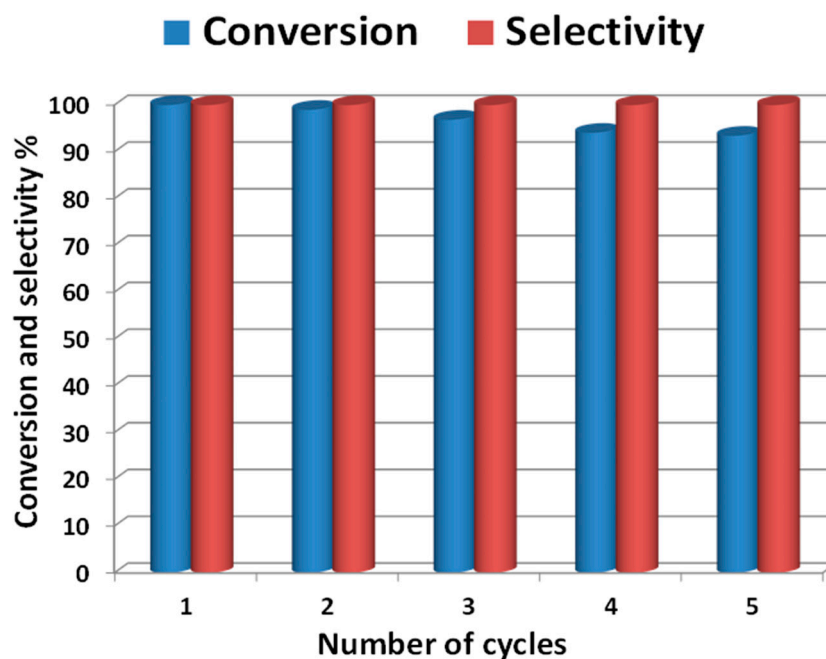


Figure 10. The recyclability $\text{ZnO}_x(1\%)\text{--MnCO}_3/(1\%)\text{HRG}$ catalyst in the selective oxidation of benzyl alcohol. (Reaction conditions: catalyst, 0.3 g; benzyl alcohol, 2 mmol; calcination temperature, 300°C ; reaction temperature, 100°C ; oxygen flow rate, $20 \text{ mL} \cdot \text{min}^{-1}$; toluene, 10 mL; and 7 min reaction time).

2.4. Oxidation of Wide Range of Benzylic and Aliphatic Alcohols

In order to examine the general applicability and efficiency of the present catalytic protocol, oxidation of various benzylic and aliphatic alcohols were performed under the optimized conditions (Table 7). A wide range of substituted benzyl alcohols, containing 4- CH_3 , 3- CH_3 , 2- CH_3 , 4- OCH_3 , 3- OCH_3 , 2,3,4-TriOMe, 4-Cl, 2-Cl, 4-Br, 4-Ph, 2,4-DiCl, 4- NO_2 , 3- NO_2 , 2- NO_2 , 4-OH, 4- $\text{C}(\text{CH}_3)_3$,

4-CF₃, 4-F and 2,3,4,5,6-Pentafluoro groups on the phenyl ring have been easily oxidized to their corresponding aldehydes in relatively short reaction times (Table 7, entries 1–20). As mentioned in Table 7, all benzylic alcohols have selectively oxidized to corresponding aldehydes in various reactions times. A >99% selectivity toward corresponding aldehydes have accomplished in all oxidation experiments and no by-products such as carboxylic acids were detected in the reaction mixture. It has reported that in the earlier literatures the electron-rich aromatic alcohols with electron-donating groups were found to be most reactive and exhibit shorter oxidation times whereas times for electron-withdrawing groups are longer [112,114]. Herein, the oxidation of benzylic alcohols bearing electron releasing substituents such as (–CH₃, –OCH₃, and –OH) resulted in formation of the corresponding aldehydes with complete conversions in relatively short reaction times (Table 7, entries 2, 3 and 6). The oxidation rate of the reaction is slower for derivatives of benzyl alcohol containing electron-withdrawing substituents such as (–Cl, –F, –NO₂, and –CF₃) (Table 7, entries 9, 10, 14 and 17). Steric hindrance is also a significant factor that affects the rate of the oxidation processes, the bulky groups such as (4-CF₃, 2,3,4-TriOMe, 2,4-DiCl, and 2,3,4,5,6-Pentafluoro) attached to the benzyl alcohol decreases the rate of oxidation reaction and required longer reaction time owing to the steric resistance that hinder the oxidation of the bulky substituents benzylic alcohols (Table 7, entries 17–20) [115]. Moreover, the results clearly revealed that the position (*para*-, *ortho*- and *meta*-) of electron donating groups in benzylic alcohols played also an important role in the activity of the catalytic oxidation of alcohols [85,98]. Benzylic alcohols bearing electron donating groups at *para*-position displayed relatively higher reactivity than the substitution in *meta*- and *ortho*-positions, because *para*-substituent have minimum steric hindrance in comparison to *ortho* and *meta*-substituents. For instance, the complete oxidation of *para*-methylbenzyl alcohol occurred within 7 min (Table 7, entry 3), while for *meta*- and *ortho*-methylbenzyl alcohol were completely transformed to their corresponding aldehydes after relatively longer reaction times 8 and 10 min, respectively by compared to *para*-position (Table 7, entries 4 and 7). In addition, *para*-substituted alcohols with electron-withdrawing groups needed longer time than those with *meta*- and *ortho*-positions, such as the complete oxidation of *para*-nitrobenzyl alcohol required shorter reaction time than those of *meta*- and *ortho*-nitrobenzyl alcohol (Table 7, entries 14–16). Hence, it can be said that the aerobic oxidation of substituted benzylic alcohols in presence of ZnO_x(1%)–MnCO₃/(1%)HRG nanocomposite is influenced by two factors, electronic density and steric hindrance. Conversely, the oxidation of aliphatic alcohols is much more difficult than that of benzylic alcohols [100,105]. In this regard, the oxidation of citronellol into citronellal occurs in longer time by compared to aromatic ones, may be owing to the absence of conjugation in the β-position of the OH group (Table 7, entry 21). Therefore, it can be concluded that the prepared catalyst exhibited high selectivity for aromatic alcohols over aliphatic alcohols.

Table 7. Catalytic performance in the aerobic oxidation of benzylic and aliphatic alcohols.

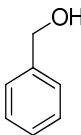
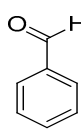
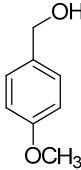
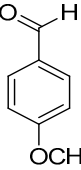
Entry	Substrates	Products	Time (min)	Conversion (%)	Selectivity (%)
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2			7	100	>99

Table 7. Cont.

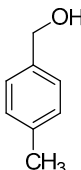
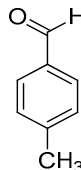
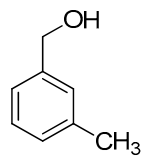
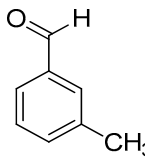
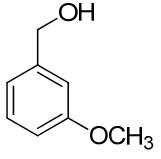
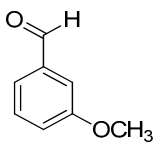
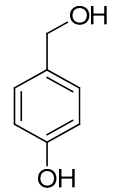
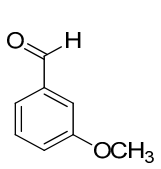
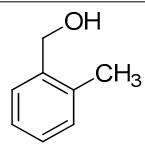
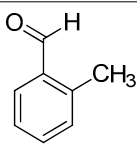
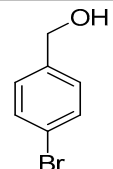
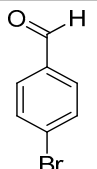
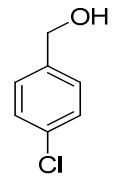
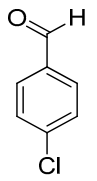
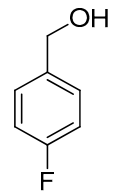
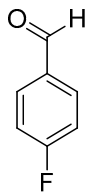
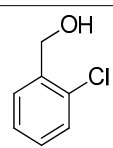
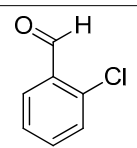
Entry	Substrates	Products	Time (min)	Conversion (%)	Selectivity (%)
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4			8	100	>99
5			8	100	>99
6			8	100	>99
7			10	100	>99
8			9	100	>99
9			9	100	>99
10			10	100	>99
11			10	100	>99

Table 7. Cont.

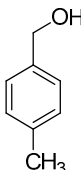
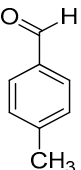
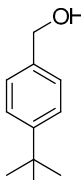
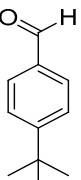
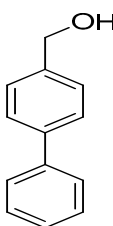
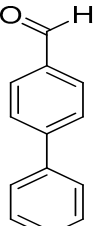
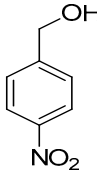
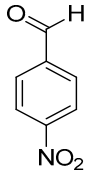
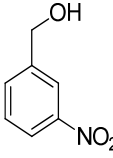
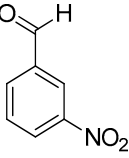
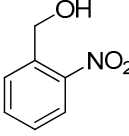
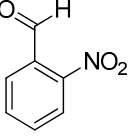
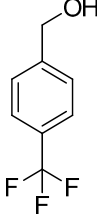
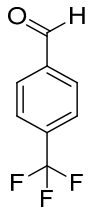
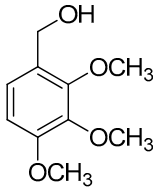
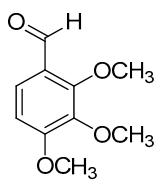
Entry	Substrates	Products	Time (min)	Conversion (%)	Selectivity (%)
3			7	100	>99
12			10	100	>99
13			11	100	>99
14			10	100	>99
15			12	100	>99
16			16	100	>99
17			11	100	>99
18			13	100	>99

Table 7. Cont.

Entry	Substrates	Products	Time (min)	Conversion (%)	Selectivity (%)
19			15	100	>99
20			19	100	>99
21			100	100	>99

Reaction conditions: alcohol, 2 mmol; toluene, 10 mL; calcination temperature, 300 °C; catalyst amount, 0.3 g; oxygen flow rate, 20 mL·min⁻¹; and reaction temperature, 100 °C.

3. Experimental Section

3.1. Materials

Manganese(II) nitrate-tetrahydrate (97%), zinc nitrate hexahydrate (99%), Sodium bicarbonate (99%), Toluene (98%), Benzyl alcohol (99.5%), 4-Methylbenzyl alcohol (99%), 3-Methylbenzyl alcohol (99%), 2-Methylbenzyl alcohol (99%), 4-Methoxybenzyl alcohol (99%), 3-Methoxybenzyl alcohol (99%), 2,3,4-Trimethoxybenzyl alcohol (99%), 4-Nitrobenzyl alcohol (99%), 3-Nitrobenzyl alcohol (99%), 2-Nitrobenzyl alcohol (99%), 4-Chlorobenzyl alcohol (99%), 2-Chlorobenzyl alcohol (99%), 2,4-Dichlorobenzyl alcohol (99%), 4-Bromobenzyl alcohol (99%), 4-Fluorobenzyl alcohol (99%), 4-Hydroxybenzyl alcohol (99%), 2,3,4,5,6-Pentafluorobenzyl alcohol (99%), 4-(Trifluoromethyl)benzyl alcohol (99%), 4-tert-Butylbenzyl alcohol (99%) and β -Citronellol were purchased from Sigma Aldrich (St. Louis, MO, USA).

3.2. Synthesis of Highly Reduced Graphene Oxide (HRG)

Firstly, graphene oxide (GO) is synthesized from graphite by the Hummers method [116]. Then, graphene oxide (200 mg) is first dispersed in 40 mL distilled water and sonicated for 30 min. Then the suspension will be heated to 100 °C subsequently 6 mL of hydrazine hydrate is added into the suspension. The suspension is then kept at 98 °C for 24 h. After that, the reduced graphene is collected by filtration in the form of black powders. The obtained material will then be washed using distilled water for several times to remove the excessive hydrazine and will be redistribute into water by sonication. Then the suspension will be centrifuged at 4000 rpm for 3 min to remove bulk graphite. The final product is collected by vacuum filtration and dried in vacuum.

3.3. Synthesis of ZnO_x(1%)–MnCO₃/)X%(HRG Nanocomposites)

ZnO_x(1%)–MnCO₃/)X%(HRG nanocomposite is synthesized by co-precipitation method (where X = 0, 1, 3, 5 and 7). In a typical synthesis, stoichiometric amount of manganese (II) nitrate-tetra-hydrate (Mn(NO₃)₂·4H₂O) and zinc nitrate hexahydrate (Zn(NO₃)₂·6H₂O) and highly reduced graphene oxide (HRG) were mixed in a round bottomed flask and subjected to ultrasonication for 30 min. Then,

the resulting solution is heated to 80 °C, while vigorously stirring using a mechanical stirrer and 0.5 M solution of sodium bicarbonate (NaHCO_3), is added drop wise by using burette until the solution attained at pH 9. The solution is continued to stir at the same temperature for about 3 h and then left under stirring over night at room temperature. The solution is filtered using a Buchner funnel under vacuum and then dried at 70 °C overnight in the oven. The resulting powder is then calcined at various temperatures (i.e., 300 °C, 400 °C, and 500 °C) in the furnace for 24 h.

3.4. Catalyst Characterization

Scanning electron microscopy (SEM) and elemental analysis (energy dispersive X-ray analysis: EDX) were carried out using Jeol SEM model JSM 6360A (Jeol, Akishima-shi, Japan). This is used to determine the morphology of nanoparticles and its elemental composition. Powder X-ray diffraction studies were carried out using Altima IV [Make: Regaku] X-ray diffractometer (Altima IV [Make: Regaku], Shibuya-ku, Tokyo, Japan). Fourier Transform Infrared Spectroscopy (FT-IR) the infrared spectra were recorded as KBr pellets using a Perkin-Elmer 1000 FT-IR spectrophotometer (Perkin-Elmer 1000, Waltham, MA, USA). BET surface area is measured on a NOVA 4200e surface area & pore size analyzer (Quantachrome Instruments, FL, USA). Thermogravimetric Analysis is carried out using Perkin-Elmer Thermogravimetric Analyzer 7 (PerkinElmer, Waltham, MA, USA).

3.5. Typical Procedure for Aerobic Oxidation of Alcohols

Liquid-phase oxidation of benzyl alcohol as example of primary benzylic alcohols has carried out in glass flask equipped with a magnetic stirrer, reflux condenser, and thermometer. In a typical experiment, a mixture of the benzyl alcohol (2 mmol), toluene (10 mL), and the catalyst (300 mg) is transferred in a glass three-necked round-bottomed flask (100 mL), the resulting mixture is then heated to desired temperature with vigorous stirring. The oxidation experiment is started by bubbling oxygen gas at a flow rate of $20 \text{ mL} \cdot \text{min}^{-1}$ into the reaction mixture. After the reaction, the solid catalyst is filtered off by centrifugation and the liquid products were analyzed by gas chromatography to determine the conversion of the alcohol and product selectivity by GC, 7890A (Agilent Technologies Inc., Santa Clara, CA, USA) Agilent Technologies Inc., equipped with a flame ionization detector (FID) and a 19019S-001 HP-PONA column (Agilent Technologies Inc., Santa Clara, CA, USA).

4. Conclusions

In the present work, we have developed an efficient ZnO_x NPs doped MnCO_3 supported on highly reduced graphene composite as an efficient catalyst for the selective oxidation of benzylic and aliphatic alcohols to corresponding carbonyls via ideal green process, i.e., under mild conditions and with molecular O_2 as a green oxidant under base-free conditions. The results of oxidation prove that presence of HRG significantly enhance the catalytic efficiency of $\text{ZnO}_x(1\%)\text{-MnCO}_3$. Thus, the superior activity of the $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ in the aerobic oxidation of alcohols should be due to the promotion role of graphene support (HRG) in the adsorption of reactant alcohol and oxygen. Typically, extremely high specific activity of $57.1 \text{ mmol} \cdot \text{g}^{-1} \cdot \text{h}^{-1}$ and complete conversion with more than 99% benzaldehyde selectivity has achieved in very short reaction time using $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ as catalyst, which exhibits superior catalytic performance than that obtained in previous literatures. Moreover, wide range of benzylic and aliphatic alcohols has examined for selective oxidation into their corresponding carbonyl compounds and showed 100% alcohols conversion within relatively short reaction times under mild reaction conditions. The present catalytic system is very selective giving only aldehydes without over-oxidation to the carboxylic acids using only molecular O_2 as oxidant without adding any additives and bases. It is noteworthy that, the oxidation rate of benzylic alcohols bearing electron donating groups higher than that containing electron withdrawing groups and bulky groups attached to benzyl alcohol decreases the oxidation rate and required longer reaction time for complete oxidation. In contrast, the oxidation of aliphatic alcohols, like citronellol, provides a relatively low conversion to citronellal, probably because of the lack of the conjugation. Additionally,

this catalyst protocol have many advantages including, complete alcohol conversions, extremely high specific activities and selectivity towards aldehydes in very short reaction time by using an a green oxidizing agent, eco-friendly catalyst, mild conditions, low cost oxidant, low percentage of graphene, facile synthesis, heterogeneous and recyclable catalyst, will give rise to this catalyst to be a useful and applicable for the selective oxidation of alcohols.

Supplementary Materials: The following are available online at www.mdpi.com/2073-4344/7/12/391/s1, Figure S1: XRD patterns of $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ catalyst (a) fresh and (b) recycled, Figure S2: TGA isothermal pattern of $\text{ZnO}_x(1\%)\text{-MnCO}_3/(1\%)\text{HRG}$ catalyst.

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Conflicts of Interest: The authors declare no conflict of interest.

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