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First-Row Transition Metal Complexes

Guest Editor:

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Message from the Guest Editor

Dear Colleagues,

First row transition metals and their complexes are ubiquitous in chemistry and biochemistry and are technologically useful as well. Many living organisms contain enzymes that are comprised of metal complexes; for example, nitrogenases, responsible for the reduction of nitrogen (N_2) to ammonia (NH_3), utilize iron (Fe) as part of their molecular machinery. Coordination complexes are also widely used in industrial settings. Molecules such as metal phthalocyanines find use as dyes and pigments. Metal complex formation itself is used as a technique for extracting metals from ores. The production of commercially important polymers relies on the use of coordination or organometallic complexes of metals such as titanium (Ti) or chromium (Cr), while metals can be separated from each other by differences in the solubilities of their resultant metal complexes with various ligands. This Special Issue of *Inorganics* highlights various chemistries of first row transition metal complexes and the relevance that these molecules have with respect to our daily lives.

Dr. John C. Gordon

Guest Editor

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Editor-in-Chief

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Message from the Editor-in-Chief

Inorganic chemistry remains a lynchpin of modern chemistry, not only embracing the function and reactivity of combinations of most elements of the periodic table, but also providing a footing for studies of materials, catalysts, drugs, fuels and industrial chemicals. Arguably, the role and reach of inorganics in society have never been as great as today. Adventurous research at the heart and at the extremes of inorganic chemistry is vital to further advances and Inorganics offers authors the opportunity to publish exciting new research in an open access format.

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