

# Using the Photo-Piezoelectric Effect of AuPt@BaTiO<sub>3</sub> Oxidase Mimetics for Colorimetric Detection of GSH in Serum

Yiquan Liao, Yichang He, Bin Zhang, Ye Ma \*, Ruiqi Xu, Minggang Zhao \* and Hongzhi Cui

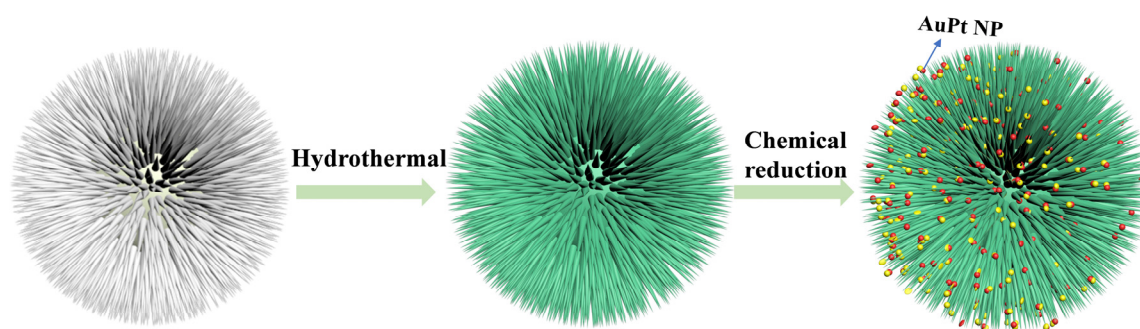
School of Materials Science and Engineering, Ocean University of China, 266100 Qingdao, China

\* Correspondence: maye@ouc.edu.cn (Y.M.); zhaomg@ouc.edu.cn (M.Z.)

## Table of contents

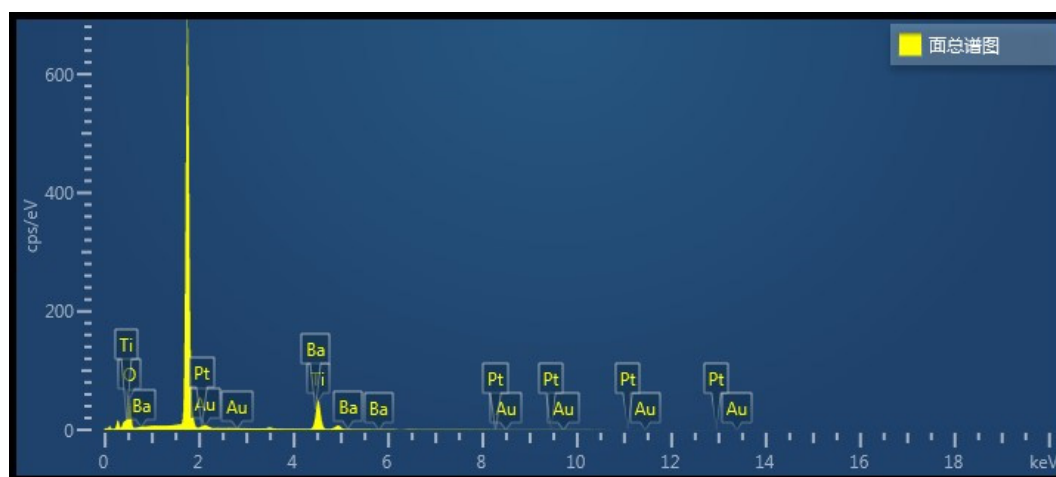
1. Fabrication process of nanozymes .....	3
2. EDX.....	3
3. EDS mapping .....	3
4. XRD.....	4
5. Optimum condition .....	4
6. DMPO-EPR.....	7
7. Steady-state kinetics .....	7
8. Stability.....	8

## 1. Fabrication process of nanozymes



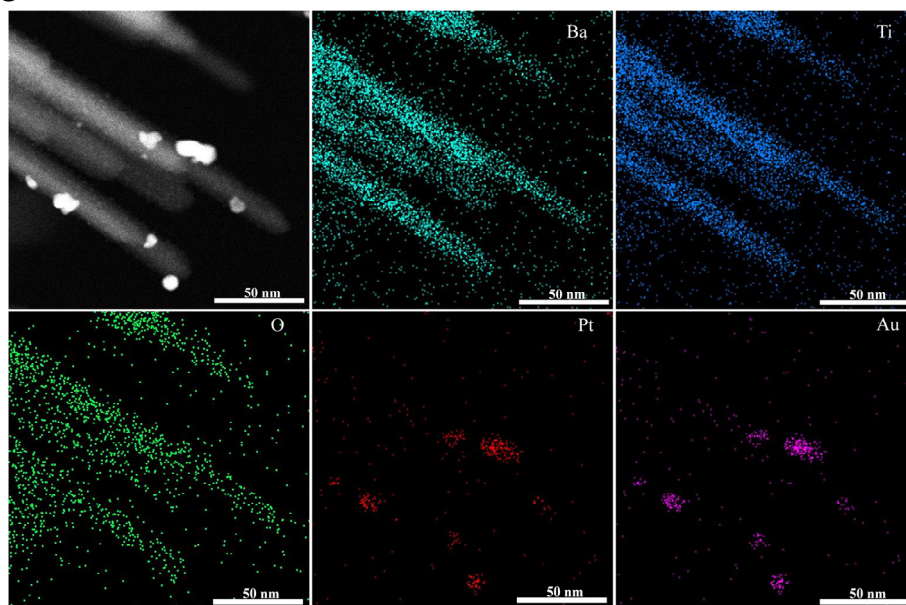
**Figure S1.** Fabrication process of the AuPt@BaTiO<sub>3</sub> SUMs.

## 2. EDX



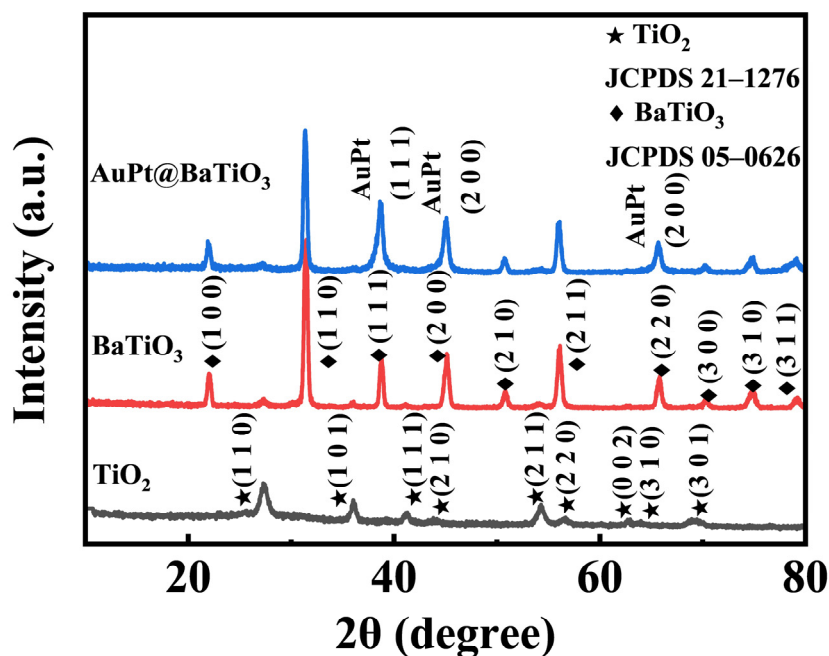
**Figure S2.** EDX photo of the prepared AuPt@BaTiO<sub>3</sub> SUMs.

## 3. EDS mapping



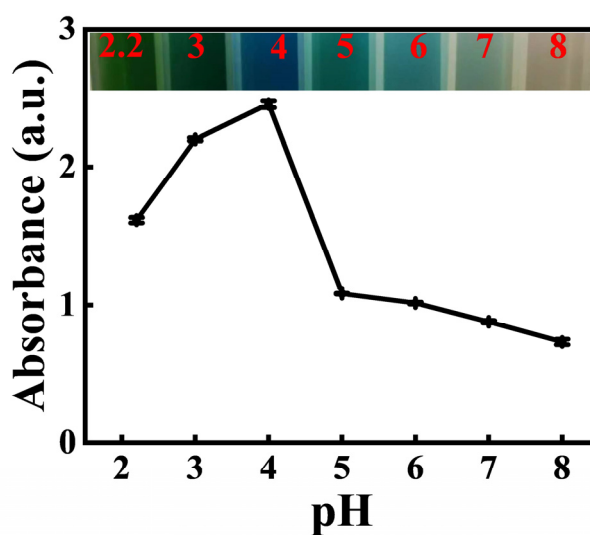
**Figure S3.** EDS mapping of the AuPt@BaTiO<sub>3</sub> SUMs.

#### 4. XRD

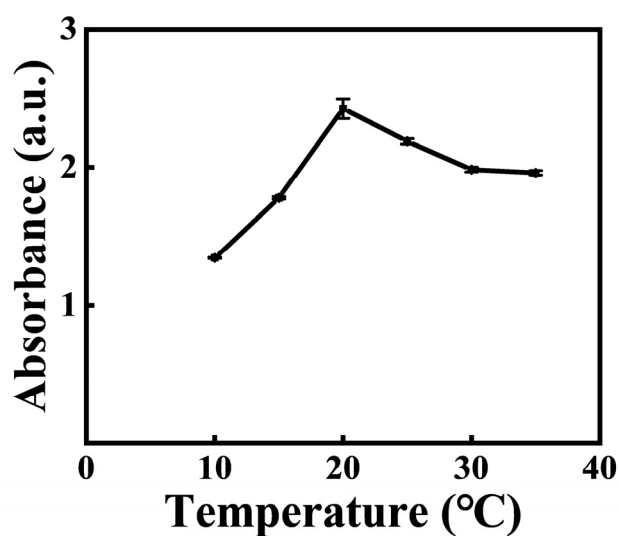


**Figure S4.** XRD patterns of the TiO<sub>2</sub>, BaTiO<sub>3</sub> SUMs, and the AuPt@BaTiO<sub>3</sub> SUMs.

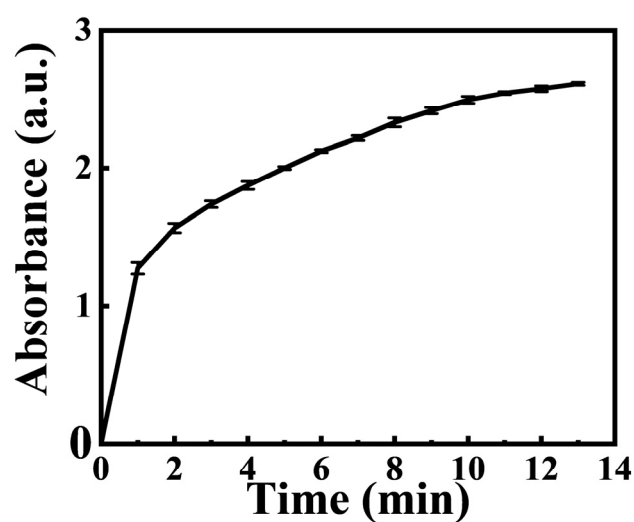
#### 5. Optimum condition



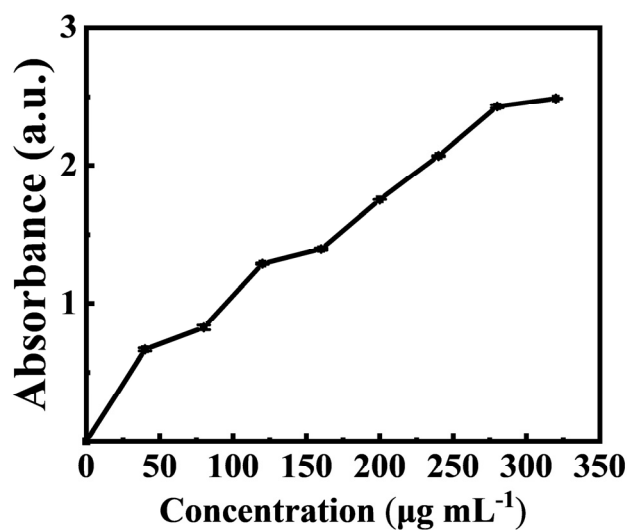
**Figure S5.** The oxidase-like activity of AuPt@BaTiO<sub>3</sub> SUMs depends on pH. Reaction condition: 0.28 mg mL<sup>-1</sup> AuPt@BaTiO<sub>3</sub> SUMs, 0.2 M Na<sub>2</sub>HPO<sub>4</sub>-CA buffer, 0.5 mM TMB, 10 min, and 20 °C.



**Figure S6.** The oxidase-like activity of AuPt@BaTiO<sub>3</sub> SUMs depends on temperature. Reaction condition: 0.28 mg mL<sup>-1</sup> AuPt@BaTiO<sub>3</sub> SUMs, 0.2 M Na<sub>2</sub>HPO<sub>4</sub>-CA buffer (pH=4.0), 0.5 mM TMB, and 10 min.

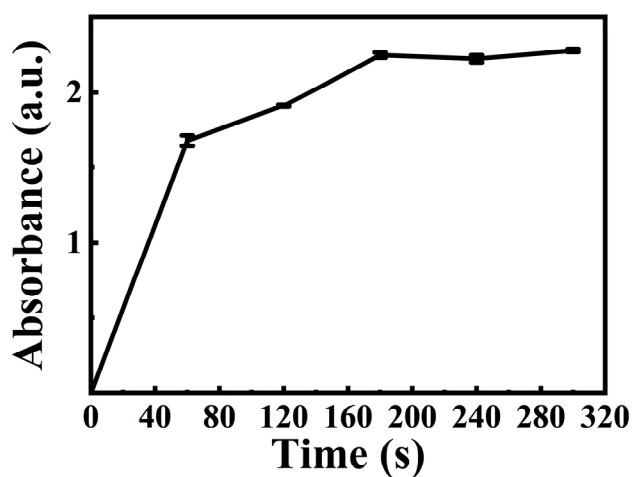


**Figure S7.** Time-dependent absorbance changes of TMB oxidation at 652 nm. Reaction condition: 0.28 mg mL<sup>-1</sup> AuPt@BaTiO<sub>3</sub> SUMs, 0.2 M Na<sub>2</sub>HPO<sub>4</sub>-CA buffer (pH=4.0), 0.5 mM TMB, and 20 °C.



**Figure S8.** The concentration effects on the oxidase-like activity of the AuPt@BaTiO<sub>3</sub> SUMs.

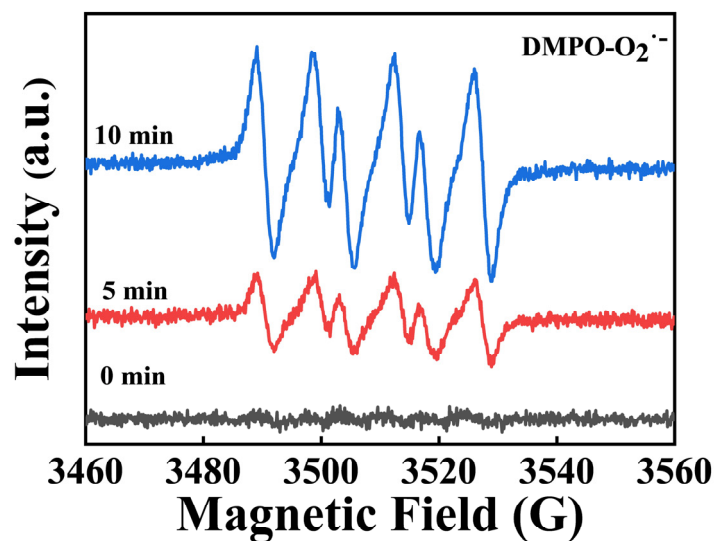
Reaction condition: 0.2 M Na<sub>2</sub>HPO<sub>4</sub>-CA buffer (pH=4.0), 0.5 mM TMB, 10 min, and 20 °C.



**Figure S9.** Time-dependent absorbance changes of TMB oxidation at 652 nm. Reaction condition:

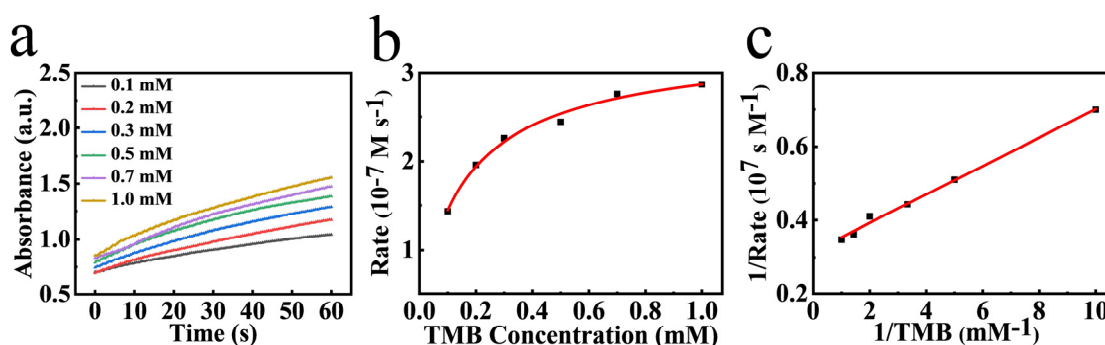
0.28 mg mL<sup>-1</sup> AuPt@BaTiO<sub>3</sub> SUMs, 0.2 M Na<sub>2</sub>HPO<sub>4</sub>-CA buffer (pH=4.0), Hg lamp and ultrasound, 0.5 mM TMB, and 20 °C.

## 6. DMPO-EPR



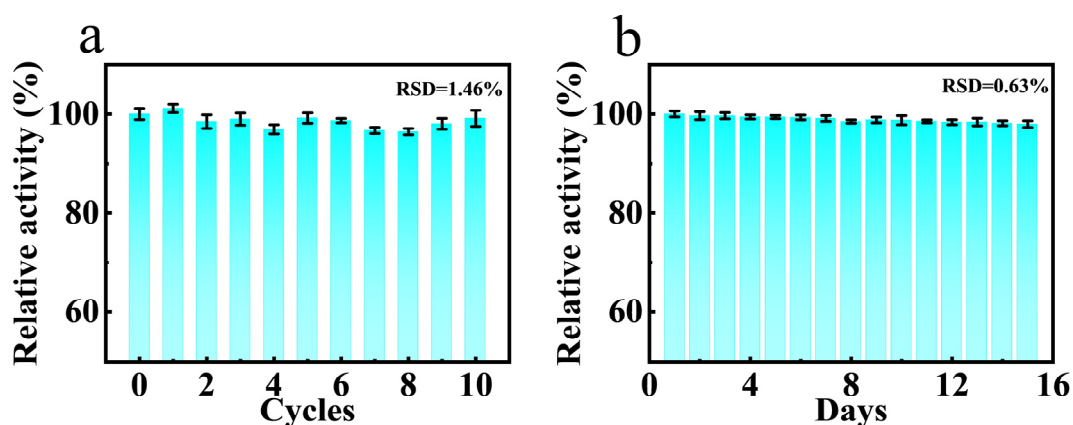
**Figure S10.** DMPO-EPR spin-trapping spectra for  $O_2^{\bullet-}$ .

## 7. Steady-state kinetics

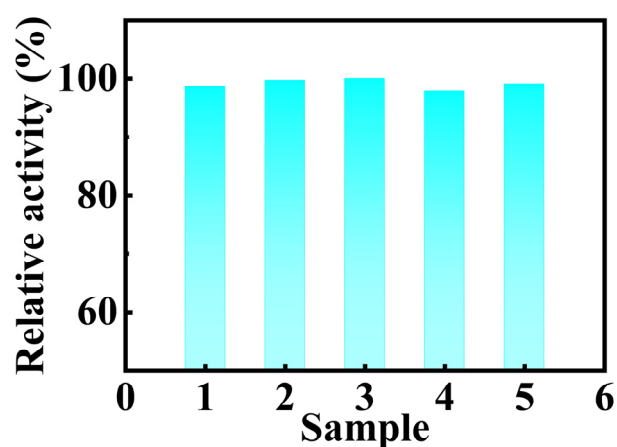


**Figure S11.** Steady-state kinetics of TMB oxidation using the AuPt@BaTiO<sub>3</sub> SUMs: (a) Typical absorbance spectra of different reaction systems for 1 min. (b) Michaelis–Menten curves for different TMB concentrations. (c) Lineweaver–Burk plot for different TMB concentrations.

## 8. Stability



**Figure S12.** (a) The short-term stability of the AuPt@BaTiO<sub>3</sub> SUMs. (b) The long-term storage stability of the AuPt@BaTiO<sub>3</sub> SUMs.



**Figure S13.** The reproducibility of the AuPt@BaTiO<sub>3</sub> SUMs.

**Table S1.** Data of catalytic kinetic parameters.

TMB concentration (mM)	k	initial reaction rate (V)	Michaelis-Menten curve		Lineweaver-Burk	
			x	y	x	y
0.1	0.00558	1.43E-07	0.1	1.43	10	0.6993
0.2	0.00765	1.96E-07	0.2	1.96	5	0.5102
0.3	0.00883	2.26E-07	0.3	2.26	3.333	0.4425
0.5	0.00953	2.44E-07	0.5	2.44	2	0.4098





0.7	0.01078	2.76E-07	0.7	2.76	1.428	0.3623
1.0	0.01119	2.87E-07	1.0	2.87	1	0.3484

---