

Dual/Bi-Stage Curing of Nanocomposites from Renewable Resources upon Volumetric Expansion [†]

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Abstract: Nowadays, polymers used in technical applications are still obtained from petrochemicals, despite the more critical reviews from society. In this work, novel nanodielectrics based on renewable resources were developed. For this purpose, poly(2-oxazoline)s (POx), which can be referred to as pseudo-polyamides, were synthesized from renewable resources and compared with commercially available Nylon 12, which is derived from petrochemicals. The monomers 2-nonyl-2-oxazoline and 2-dec-9'-enyl-2-oxazoline were synthesized from coconut oil and castor oil in solvent-free syntheses according to the Henkel Patent; the corresponding copoly(2-oxazoline)s were synthesized in an energy-efficient fashion in microwave reactors under autoclave conditions. Both types of polyamides (two variations: POx and Nylon 12) were filled with inorganic nanoparticles (four variations: no filler, submicro-scaled BN, nano- and micro-scaled AlN as well as a mixture of nano- and micro-scaled AlN and submicro-scaled BN) and/or expanding monomers, namely spiroorthesters (three variations: 0, 15, and 30 wt.-%), yielding a $2 \times 4 \times 3 = 24$ -membered material library. All polymers were crosslinked according to a newly developed thermally-initiated dual/bi-stage curing system. Intense physicochemical and dielectric characterization revealed that the relative volume expansion was in the range of 0.46 to 2.48 vol.-% for the Nylon 12 samples and in the range of 1.39 to 7.69 vol.-% for the POx samples. Hence, the formation micro-cracks or micro-voids during curing is significantly reduced. The dielectric measurements show competitive dielectric behavior of the “green” POx samples in comparison with the fossil-based Nylon 12 samples at a frequency of 40 Hz.

Keywords: polymer from renewable resources; poly(2-oxazoline); pseudo-polyamide; nanodielectric; expanding monomer

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1. Introduction

The future challenges for polymer-based dielectric materials will be manifold. An important driving-force for material development is the increasing heat development in electronic devices, originating from increasing power density. Recently, ceramic fillers have been evaluated as additives in polymer-based dielectrics because of their intrinsic thermal conductivity, which is of one to two orders of magnitude higher than that of oxide-based fillers, such as alumina Al_2O_3 and silica SiO_2 [1]. It was shown by transmission electron microscopy (TEM) that BN tends to agglomerate in a manifold of types of polymer matrices [2]. It was shown that the usage of hybrid-filler systems composed of Al_2O_3 and AlN with different particle sizes can overcome the problem of forming inflexible voids [3].

This study aimed at the development of novel polymer-based composites for future employment as dielectrics. Based on the state-of-the-art knowledge of inorganic fillers, micro-scaled hexagonal boron nitride hBN, nano- and micro-scaled aluminum nitride n-AlN and μ -AlN particles, as well as mixtures of the types of particles were added to the polymer matrices to ensure a higher thermal conductivity.

An additional challenge is the volumetric shrinkage during crosslinking, which leads to the formation of micro-cracks, micro-voids, and delamination, deteriorating the insulation properties [4]. Therefore, a spiroorthoester (SOE) was added in different contents to the polymer material. The SOE exhibits an unsaturated C=C bond that enabled its covalent attachment to the polymer matrices (see hereinafter). By employment of the SOE, the volumetric shrinkage during the curing reaction should be minimized or in the best case completely counter-balanced, yielding networks that are volume-neutral during curing.

An additional important aspect for the design-of-experiments DoE/material choices originated from the fact that the vast majority of polymers used in microelectronics and high-voltage engineering are obtained from petrochemicals. In society, a more critical view has evolved considering aspects of sustainability and environmental challenges [5]. Therefore, it is highly recommendable to replace the organic part of composites with polymers made from renewable resources in order to contribute to environmental protection and lay the foundation for future generations of non-fossil-based materials. A promising alternative to fossil-based polyamides is offered by poly(2-oxazolines), which can be regarded as “pseudo”-polyamides. Like polyamides, (co-)poly(2-oxazoline)s are known for their high temperature stability [6]. The 2-oxazoline monomers with long aliphatic chains can be derived from fatty acids (Figure 1). Precedent studies have already shown that the values of permittivity, loss factor, and conductivity of this class of material are in the same range as for polyamides, classifying these materials as medium-range insulators [7].

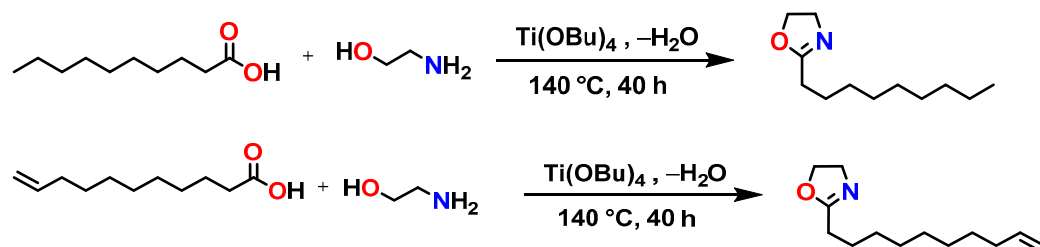


Figure 1. Reaction scheme of the synthesis of 2-oxazoline monomers from fatty acids (of renewable resources).

In summary, in this study, different samples of polymer blends and composites from renewable polymers and fossil-based polymers were prepared. Two 12-membered material libraries (one for the Nylon 12- and one for the copoly(2-oxazoline)-based matrices) were fundamentally characterized by physicochemical and dielectric properties in order to establish the correlations between the structure and the resulting properties.

2. Experiments

2.1. Materials and Methods

Allyl glycidyl ether, boron trifluoride diethyl etherate, γ -butyrolactone, dicumyl peroxide, ethanolamine, methyl tosylate, titanium(IV) butoxide, trimethylamine, decanoic acid, and undec-10-enoic acid were purchased from Sigma Aldrich (Vienna, Austria). Aluminum(III)nitride (particle size 1.0–15.0 μ m), 1-butylpyridin-1-ium hexafluorophosphate(V) were received from ABCR (Karlsruhe, Germany). Aluminum(III)nitride (particle size 50 nm) was purchased from IoLiTec (Heilbronn, Germany). Boron nitride (particle size 70 nm) was purchased from MKnano (Mississauga, Canada). Nylon 12 (particle size 25–30 μ m) was purchased from Goodfellow Limited (Hamburg, Germany). All chemicals were used as received, except for methyl tosylate that was distilled prior to use.

The microwave-assisted polymerization was performed with the Synthos 3000 Microwave Synthesis Platform from Anton Paar GmbH in screwable Wheaton® glass vials (15 × 46 mm) with a Rotor 64. All NMR measurements were performed with a Bruker Advance III 300 MHz spectrometer. As solvent, CDCl₃ with an internal standard of 0.03% trimethylsilane TMS were used. Infrared spectroscopy was performed with an Alpha Fourier-Transform Infrared Spectrometer with attenuated total reflection ATR support. The scan area was from 400 to 4000 cm⁻¹, and 48 scans were performed for each measurement. For the determination of the average molar mass of the polymers, gel permeation chromatography (GPC) analyses were performed with a Merck Hitachi L-6000A pump, a column from Polymer Standards Service column (8 × 300 mm STV linear XL 5 µm grade size), and a differential refractometer Waters 410 detector. As solvent, THF was used. For the preparation of the PA and POx test specimens, a Collin P 200 Laboratory Platen Press was used. The thermogravimetry and differential scanning calorimetry DSC measurements were performed with the NETZSCH STA 449C Jupiter device coupled with quadrupole mass spectrometry. The temperature range was from 20 to 550 °C with a heating rate of 10 K·min⁻¹. The measurements were performed under helium atmosphere. Dielectric characterization was performed with a SPECTANO 100 from OMICRON Lab. All samples were characterized in a frequency range of 0.1 Hz to 5 kHz.

2.2. Synthesis of the Monomers Dec=Ox, NonOx, and the SOE

NonOx: NonOx was synthesized according to the Henkel Patent [8]. In a 500 mL three-neck round-bottom flask equipped with a magnetic stirring bar and a reflux condenser, 135 g of decanoic acid (0.78 mol, 1 equivalent), 71 mL of ethanol-2-amine (1.18 mol, 1.5 equivalents), and 1.4 mL of the catalyst titanium(IV) *n*-butoxide (3.92 mmol, 0.005 equivalent) were stirred under reflux at 160 °C for 24 h. Further amounts of 1.4 mL of the catalyst titanium(IV) *n*-butoxide (3.92 mmol, 0.005 equivalent) were added after 20 and 24 h. After that, the reflux condenser was removed, and the mixture was stirred for additional 12 h in order to remove the reaction water. The crude NonOx product was obtained via fractional distillation under vacuum (125 °C, 6 mbar), yielding a yellow liquid. Purification by column chromatography with chloroform yielded a colorless viscous liquid of pure NonOx. The yield was 53.3% (82.36 g).

¹H-NMR (300 MHz, 298 K, CDCl₃): δ (ppm) = 0.88 (3H, t), 1.27 (12H, s), 1.63 (2H, t), 2.26 (2H, t), 3.82 (2, t), 4.21 (2H, t). ¹³C-NMR (75 MHz, 298 K, CDCl₃): δ (ppm) = 14.1, 22.6, 25.9, 28.0, 29.2, 29.4, 31.8, 54.3, 67.1, 168.7. IR (ATR): ν (cm⁻¹) = 2923, 2854, 1688, 1464, 1362, 1231, 1165, 986, 950, 908, 754, 722.

Dec=Ox: Dec=Ox was synthesized according to the Henkel Patent [8]. In a 500 mL three-neck round-bottom flask equipped with a magnetic stirring bar and a reflux condenser, 135 g of undec-10-enoic acid (0.73 mol, 1 equivalent), 70 mL of ethanol-2-amine (1.17 mol, 1.6 equivalents), and 1.3 mL of the catalyst titanium(IV) *n*-butoxide (3.66 mmol, 0.005 equivalent) were stirred under reflux at 140 °C for 24 h. Further amounts of 1.3 mL of the catalyst titanium(IV) *n*-butoxide (3.66 mmol, 0.005 equivalent) were added after 20 and 24 h. After that, the reflux condenser was removed, and the mixture was stirred for additional 12 h to remove the reaction water. The crude Dec=Ox product was obtained via fractional distillation under vacuum (116 °C, 6 mbar), yielding a yellow liquid. Purification by column chromatography with chloroform yielded a colorless viscous liquid of pure Dec=Ox. The yield was 68.4% (104.9 g).

¹H-NMR (300 MHz, 298 K, CDCl₃): δ (ppm) = 1.23 (10H, s), 1.55 (2H, t), 1.95 (2H, d), 2.19 (3H, t), 3.74 (2H, t), 4.14 (2H, t), 4.88 (2H, m), 5.72 (1H, m). ¹³C-NMR (75 MHz, 298 K, CDCl₃): δ (ppm) = 25.9, 27.9, 28.9, 29.0, 29.2, 29.3, 31.8, 54.3, 67.1, 114.1, 139.1, 168.6. IR (ATR): ν (cm⁻¹) = 2924, 2853, 1668, 1640, 1461, 1432, 1385, 1361, 1227, 1167, 987, 952, 907, 723.

SOE: The spiroorthoester was synthesized according to the method described by Marx et al. [9]. 11.41 g (0.1 mmol, 1.0 eq.) of allyl glycidyl ether were added with a dropping funnel over a period of 40 min to a mixture of 0.46 g (3.24 mmol) of BF₃·OEt₂ and 52

g (0.6 mol, 6.0 eq.) of γ -butyrolactone under a nitrogen atmosphere and cooling with a water bath. After stirring at room temperature for 1 h, the reaction was quenched by the addition of 0.36 g (3.56 mmol) of NEt_3 . The mixture was extracted with CH_2Cl_2 (40 mL) and H_2O (20 mL). The organic layer was dried over Na_2SO_4 , and the solvent was evaporated under reduced pressure. For purification, the crude product was distilled under reduced pressure yielding 19.2 g (48 mmol, 48%) of a colorless liquid (boiling point 115 °C/0.1 mbar).

^1H -NMR (300 MHz, 298 K, CDCl_3) δ (ppm) = 1.90–2.06 (2H, m), 2.11–2.20 (2H, m), 3.42–3.43 (0.5H, m), 3.44–3.50 (1H, m), 3.52–3.64 (0.5H, m), 3.67–3.75 (1H, m), 3.85–4.09 (4H, m), 4.11–4.30 (1H, m), 4.32–4.43 (0.5H, m), 4.43–4.48 (0.5H, m), 5.17–5.30 (2H, m), 5.85–5.91 (1H, m). ^{13}C -NMR (300 MHz, 298 K, CDCl_3) δ (ppm) = 24.1, 24.2, 32.7, 32.9, 66.5, 66.9, 67.2, 70.3, 71.5, 72.4, 72.5, 74.3, 75.3, 117.3, 117.4, 129.5, 129.7, 134.4. IR (ATR) ν (cm^{-1}) = 3081, 2955, 2879, 1776, 1737, 1647, 1458, 1332, 1244, 1195, 1130, 1040, 954, 922, 867, 807, 736.

2.3. Synthesis of the Copolymer Poly(2-nonyl-2-oxazoline)-stat-poly(2-dec-9'-enyl-2-oxazoline)

The copolymerization of Dec^oOx and NonOx was performed via microwave-assisted polymerization in a scale of 25 g. 20 equiv. of Dec^oOx and 80 equiv. of NonOx were mixed. As initiator, methyl tosylate MeOTs (1 equiv.) was added; as solvent, acetonitrile ACN was used. Under autoclave condition at 140 °C and stirring for 2 h in sealed reaction vials, the copolymer pNonOx₈₀-stat-pDec^oOx₂₀ was obtained via a cationic ring-opening polymerization (CROP). A white powder was obtained after rotary evaporation of the solvent, grinding in the mortar and further drying.

^1H -NMR (300 MHz, 298 K, CDCl_3): δ (ppm) = 0.81 (274H, m), 1.19 (1275H, m), 1.51 (211H, m), 1.97 (45H, m), 2.17 (208H, m), 3.37 (400H, m), 4.88 (40H, m), 5.74 (20H, m). ^{13}C NMR (75 MHz, 298 K, CDCl_3): δ (ppm) = 14.1, 22.7, 25.2, 25.4, 29.0, 29.2, 29.4, 29.6, 32.0, 33.0, 33.8, 43.3, 45.3, 114.2, 139.1, 173.2, 173.7. IR (ATR): ν (cm^{-1}) = 2921, 2852, 1640, 1462, 1430, 1183, 1177, 907, 773, 722. GPC: M_w = 17.6 kDa; M_n = 8.87 kDa; \bar{D} = 1.81.

2.4. Preparation of the Test Specimens

For the POx system, poly(2-nonyl-2-oxazoline)-stat-poly(2-dec-9'-enyl-2-oxazoline), and for the PA-system, Nylon 12 was used. For the crosslinking of the POx sample, 5 mol-% of dicumyl peroxide (referred to the double bonds of the copolymer) were added. For the crosslinking of the PA samples, 2 wt.-% of dicumyl peroxide (referred to the total amount of Nylon 12) were added. For the preparation of the POx/PA samples with SOE 15 or 30 wt.-% of the additive (referred to the total amount of POx/PA) were added. The amount of the additional initiator 1-butylpyridin-1-ium hexafluorophosphate(V) was 5 wt.-% (referred to the amount of SOE). All mixtures were homogenized by grinding.

Test specimens of the PA and POx were prepared in a platen press. A temperature of 180 °C and a pressure of 40 bar were applied for 20 min. A steel plate was covered with a 0.2 mm thick PTFE-foil from Haberkorn and a stainless-steel form was placed on it. The respective amount of the powder mixture was added in the stainless-steel form (squared, circled or rectangular shape). The top was again covered with a PTFE-foil and a steel plate. The composite samples were dried in a vacuum oven at 60 °C and 1 mbar for 72 h.

3. Results and Discussion

3.1. Copolymer Synthesis

For the synthesis of polymers from renewable resources, the synthesis described in the Henkel patent was chosen to obtain the two monomers NonOx and Dec^oOx from fatty acids. Via a microwave-assisted copolymerization, colorless poly(2-nonyl-2-oxazoline)-stat-poly(2-dec-9'-enyl-2-oxazoline) was obtained (Figure 2). The polymer was characterized with ^1H -NMR, ^{13}C -NMR, FT-IR, GPC, TGA and DSC measurements. The NMR reveals high product purity and confirms the targeted stoichiometric amounts of the repetition units of NonOx and Dec^oOx. The TGA shows that the copoly(2-oxazoline) exhibits

a high thermal stability; by DSC measurements, the glass-transition temperature $T_g = 40.5$ °C and the melting point $T_m = 138.0$ °C were determined.

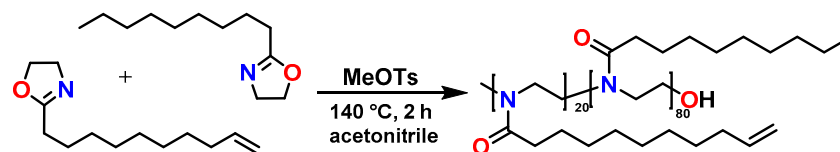


Figure 2. Schematic representation of the copolymerization of NonOx and DecOx.

For the (combined) radical-induced crosslinking of POx/PA, respectively, and the radical-induced covalent attachment of the SOE to the polymer matrix with simultaneous cationic ring-opening (and expansion) of the SOE, a thermally-triggered dual/bi-stage curing system needed to be developed. In addition to the radical initiator dicumyl peroxide, the cationic initiator 1-butylpyridin-1-ium hexafluorophosphate(V) to was used. As SOE with pendant C=C double bond, 2-((allyloxy)methyl)-1,4,6-trioxaspiro[4.4] nonane was added to the polymer matrix.

3.2. Library Design

Composite test specimens of Nylon 12 and POx were successfully produced (two variations), with no SOE or SOE in contents of 15 or 30 wt.-% (three variations), containing no fillers or inorganic fillers in a content of 40 wt.-% hBN, 20/20 wt.-% of n-AIN/ μ -AIN, and 20/10/10 wt.-% of hBN/n-AIN/ μ -AIN (four variations). This DoE yielded two $3 \times 4 = 12$ -membered material libraries (Figure 3). Test specimens were produced in squared shape with an area of 8×8 cm and a height of 300 μ m, and circular shape with diameters of 4 cm and a height of 300 μ m.

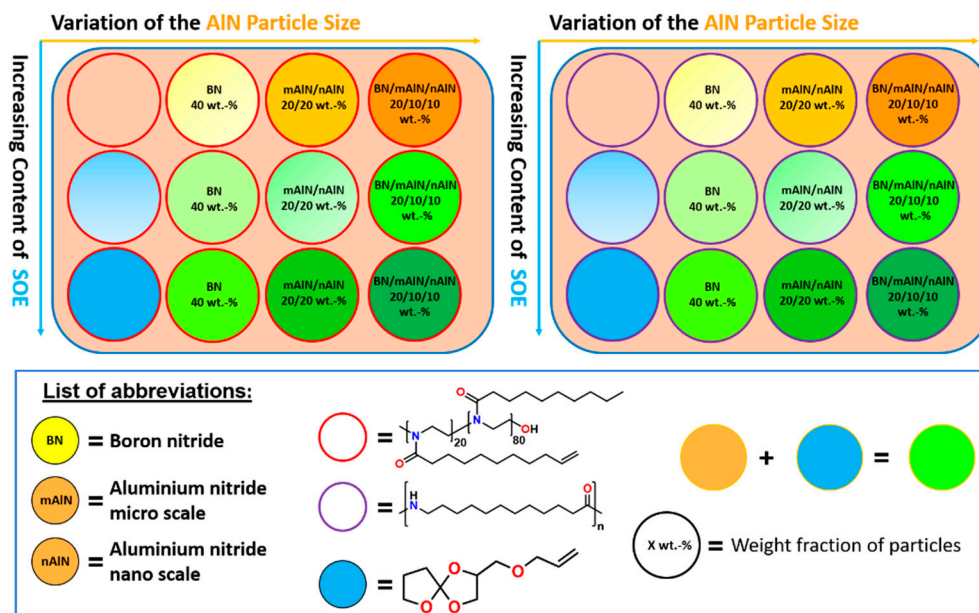


Figure 3. Schematic representation of the material libraries of the composite materials.

3.3. Determination of the Volumetric Expansion

In order to determine the influence of the SOE on the volumetric change during the crosslinking of the polymer, the mass (m) and the dimensions of each sample were measured. The thickness and length of the test specimens were determined with a caliper at eight different measuring points. The volume (V) of the test specimens was calculated

from the mean values of the thickness and the length. The density (ρ) was calculated by the following formula (Equation (1); Tables 1 and 2):

$$\rho = \frac{m}{V} \quad (1)$$

The volumetric expansions were calculated relative to the density of the specimen without SOE (Tables 3 and 4). The addition of the SOE induces a volume expansion compared to the sample without the SOE. The relative volumetric expansion of the POx samples due to the addition of SOE in a content of 15 (30) wt.-% is in the range of 1.23–1.54% (4.29–7.69%). The relative volume expansion of the Nylon 12 samples due to the addition of SOE in a content of 15 (30) wt.-% is smaller than 1% (in the range of 1.41–2.48%).

Hence, the POx samples show higher volumetric expansion than the Nylon 12 samples. A possible explanation can be the steric hindrance in the condensed phase due to the hydrophobic (nonyl and decenyl) side-chains in the POx system, which, in final consequence, induces different separation among the polyamide and polyester/polyether phases than in the Nylon 12. This phenomenon needs to be subjected to further investigation. Nevertheless, volumetric shrinkage could be reduced in both polymer systems.

Table 1. Density in g·cm^{−3} of the POx samples.

SOE wt.-%	No Particles: ρ (g·cm ^{−3})	hBN, 40 wt.-%: ρ (g·cm ^{−3})	n-AlN/ μ -AlN, 20/20 wt.-%: ρ (g·cm ^{−3})	hBN/n-AlN/ μ -AlN, 20/10/10 wt.-%: ρ (g·cm ^{−3})
0	1.05	1.23	1.29	1.29
15	1.04	1.21	1.24	1.23
30	1.01	1.17	1.19	1.19

Table 2. Density in g·cm^{−3} of the Nylon 12 samples.

SOE wt.-%	No Particles: ρ (g·cm ^{−3})	hBN, 40 wt.-%: ρ (g·cm ^{−3})	n-AlN/ μ -AlN, 20/20 wt.-%: ρ (g·cm ^{−3})	hBN/n-AlN/ μ -AlN, 20/10/10 wt.-%: ρ (g·cm ^{−3})
0	1.02	1.21	1.26	1.28
15	1.01	1.20	1.25	1.27
30	0.99	1.19	1.24	1.26

Table 3. Relative volumetric expansion of the POx samples.

SOE wt.-%	No Particles: ΔV_{rel} (%)	hBN, 40 wt.-%: ΔV_{rel} (%)	n-AlN/ μ -AlN, 20/20 wt.-%: ΔV_{rel} (%)	hBN/n-AlN/ μ -AlN, 20/10/10 wt.-%: ΔV_{rel} (%)
0	–	–	–	–
15	1.54	1.39	3.70	4.97
30	4.39	4.84	7.69	7.50

Table 4. Relative volumetric expansion of the Nylon 12 samples.

SOE wt.-%	No Particles: ΔV_{rel} (%)	hBN, 40 wt.-%: ΔV_{rel} (%)	n-AlN/ μ -AlN, 20/20 wt.-%: ΔV_{rel} (%)	hBN/n-AlN/ μ -AlN, 20/10/10 wt.-%: ΔV_{rel} (%)
0	–	–	–	–
15	0.98	0.84	0.46	0.84
30	2.48	1.41	1.92	1.60

3.4. Dielectric Properties of the Polymer Networks and the Corresponding Composites

By dielectric spectroscopy with a SPECTANO 100 from OMICRON Lab, all specimens were characterized in a frequency range of 0.1 Hz to 5 kHz. The permittivity (real part, as well as the imaginary part) increases with decreasing frequencies for all samples. In general, the permittivity increases due to the presence of the SOE, which has been referred to the increasing free volume in the course of the volumetric expansion. The dielectric properties of the Nylon 12 and the POx samples were compared at a low frequency of 40 Hz on the example of the loss factor $\tan\delta$ (Tables 5 and 6). The unfilled POx sample show nearly the same loss factors as the Nylon 12 samples. Additionally, the copoly(2-oxazoline)s filled with hBN and n-AlN/ μ -AlN are competitive with the PA samples as well. However, the Nylon 12 samples with the hBN/n-AlN/ μ -AlN filler combination show the lowest loss factors. All specimens exhibit loss factors $\tan\delta \ll 1$ and can be considered as good dielectrics [7].

Table 5. Dielectric loss factor $\tan\delta$ of the POx samples at frequency of 40 Hz.

SOE wt.-%	No Particles: $\tan\delta$	hBN, 40 wt.-%: $\tan\delta$	n-AlN/ μ -AlN, 20/20 wt.-%: $\tan\delta$	hBN/n-AlN/ μ -AlN, 20/10/10 wt.-%: $\tan\delta$
0	0.10	0.09	0.11	0.16
15	0.16	0.18	0.25	0.46
30	0.17	0.44	0.18	0.26

Table 6. Dielectric loss factor $\tan\delta$ of the Nylon 12 samples at frequency of 40 Hz.

SOE wt.-%	No Particles: $\tan\delta$	hBN, 40 wt.-%: $\tan\delta$	n-AlN/ μ -AlN, 20/20 wt.-%: $\tan\delta$	hBN/n-AlN/ μ -AlN, 20/10/10 wt.-%: $\tan\delta$
0	0.10	0.07	0.07	0.07
15	0.24	0.24	0.15	0.12
30	0.14	0.13	0.13	0.06

4. Conclusions

The aim of this study was to develop nanodielectrics from renewable resources. The 2-oxazoline monomers were synthesized from renewable resources (fatty acids) according to the Henkel patent, and copolymerized via a microwave-assisted cationic ring-opening polymerization. The material properties of the copoly(2-oxazoline) poly(2-nonyl-2-oxazoline)_{80-stat}-poly(2-dec-9'-enyl-2-oxazoline)₂₀ were compared with commercially available, fossil-based Nylon 12.

These polymers were blended with the expanding monomer SOE, and compounded with inorganic fillers with high thermal conductivity. The samples were crosslinked and expanded by a thermally triggered dual/bi-stage curing system, comprising the crosslinking of the polyamides and SOEs, according to a radical mechanism (showing volumetric shrinkage) and the (double) ring-opening of the SOE, according to a cationic mechanism (showing volumetric expansion).

Due to the addition of the SOE, the relative volume expansion of the Nylon 12 samples was in the range of 0.46–2.48%, and in the range of 1.39–7.69% for the POx samples. Hence, the formation micro-cracks or micro-voids during curing is significantly reduced.

The unfilled POx samples show nearly the same values of $\tan\delta$ as the unfilled Nylon 12 samples; as the POx was synthesized from renewable resources, it can be argued that this “green” alternative shows competitive dielectric behavior in comparison with fossil-based Nylon 12.

Author Contributions: F.B. and F.W. designed the experiments; P.M. synthesized the SOEs; F.B. and P.M. performed the experiments; F.B. and F.W. analyzed the data; F.B., P.M., and F.W. wrote the paper. All authors have read and agreed to the published version of the manuscript.

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Data Availability Statement: The data presented in this study are available on request from the corresponding author. The data are not yet publicly available due to the performance of on-going studies of the scientific findings that have been reported.

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Conflicts of Interest: The authors declare no conflict of interest.

Abbreviations

The following abbreviations are used in this manuscript:

AlN	aluminum nitride
Al ₂ O ₃	aluminum oxide
BN	boron nitride
CROP	cationic ring-opening polymerization
DCP	dicumyl peroxide
DoE	design of experiments
Dec ^o Ox	2-dec-9'-enyl-oxazoline
hBN	hexagonal boron nitride
MeOTs	methyl tosylate
NonOx	2-nonyl-oxazoline
PA	polyamide
pNonOx ₈₀ -stat-pDec ^o Ox ₂₀	poly(2-nonyl-2-oxazoline)-stat-poly(2-dec-9'-enyl-2-oxazoline)
POx	poly(2-oxazoline)
SOE	spiroorthoester
wt.-%	weight percent

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