



Article

Polarizability of Kekulene, Septulene, and Nearest Non-Planar Polycyclic Aromatic Hydrocarbons

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Abstract: The polarizability of polycyclic aromatic hydrocarbons (PAHs) is an important property that relates to their abundance in natural environments. To assess the differences in the mean polarizability of planar and non-planar polycyclic aromatic hydrocarbons (PAHs), we computationally studied the PAH series of circulenes (kekulene C₄₈H₂₄, [13]circulene C₅₂H₂₆, and septulene C₅₆H₂₈), a number of the nearest helicenes, and their “expanded” isomers. We mean under “nearest” a close number of aromatic rings: 11 (C₄₆H₂₆), 12 (C₅₀H₂₈), 13 (C₅₄H₃₀), and 14 (C₅₈H₃₂). For these PAHs, we performed the quantum chemical calculations of thermodynamic and polarizability parameters with the PBE/3ζ density functional theory method, which is widely used in the theoretical chemistry of fullerenes and PAHs. The calculated mean polarizabilities (in Å³) ranged from 80.1 for [11]helicene to 135.5 for septulene, and while the circulenes and expanded helicenes had similar values, the mean polarizability of the normal helicenes was markedly lower. In all four pairs of helical PAHs, the expanded helicene was energetically considerably more favorable than its standard helicene isomer. Herewith, the ratio of their polarizabilities was equal to 1.3.

Keywords: kekulene; septulene; circulenes; polarizability; helicenes; expanded helicenes; polycyclic aromatic hydrocarbons; density functional theory

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1. Introduction

Polycyclic aromatic hydrocarbons (PAHs) make up a large group of organic compounds composed of two or more condensed aromatic rings. The PAHs’ properties, originating mainly from extensive delocalized π -conjugated systems, are one of the reasons for investigating their transformations [1–4]. The correlations between the structure and physicochemical properties were addressed in theoretical works [5–10]. The interest in molecular properties of PAHs is also due to their environmental chemistry: They are usually formed under the incomplete combustion of organic matter and, therefore, relate to ecological problems dealing with their potential cancerogenicity [11]. Various PAHs are expected to be ubiquitous in the interstellar and circumstellar media. This is based on experimental works [12–15] and theoretical models [16].

Generally, most PAHs are planar, and the planar compounds are studied more extensively, compared with the non-planar ones. The latter include helicenes, which have a helical structure [3], and their isomers—the so-called “expanded helicenes”—whose structures also have a helix as a basis, just wider [4]. Currently, helicenes attract increasing attention because of the possibility of their formation in the interstellar medium, which has been predicted with theoretical models [10] and then confirmed with laboratory experiments [14]. We have previously shown that helicenes are less thermodynamically stable than their planar isomers but have lower mean polarizabilities. The latter feature enhances the survivability of helicenes under cosmic conditions that typically involve inter-

actions of matters with irradiations [14,17]. Furthermore, we have formulated a qualitative correlation in the isomeric PAH series: non-planar isomers have higher total energy, lower mean polarizability, and should be more abundant in astronomical objects or their laboratory models (and vice versa in the case of flat PAHs) [17].

Polarizability is an important property of PAHs because it is connected to their structure, physicochemical properties, and chemical reactivity [9]. There are a number of publications dealing with the mean polarizabilities of PAHs computed with different theoretical methods: semiempirical [18], quantitative structure–property relationship (QSPR) [19], coupled cluster [20], and original and modified density functional theory [9,21]. Based on these works, the methods based on the Perdew–Burke–Ernzerhof density functional (PBE) [22] are applicable to study the links between structure, polarizability, and related properties of different PAHs.

In the continuation of our systematic study on the PAH polarizability [17], we investigated the stability and polarizability of kekulene $C_{48}H_{24}$, septulene $C_{56}H_{28}$, and a number of their nearest non-planar PAHs: [13]circulene, [n]helicenes, and expanded [n]helicenes (in the following text, we will mark the latter with the suffix “-exp”), where $n = 11, 12, 13$, and 14. The optimized geometries of these molecules are shown in Figure 1 (the associated Cartesian coordinates of atomic positions can be found in the Supplementary Materials). A helix can be described by its diameter, number of turns, and pitch. [11]Helicene has only one complete turn of the helix, whereas the remaining helicenes have two. Expanded [n]helicenes ([n]helicenes-exp) are in a similar situation, but the first turn of their helix finishes at $n = 12$. The results are presented in this study.

2. Computation Details

Quantum chemical calculations were performed in the PRIRODA program (version 11) [23]. We used the Perdew–Burke–Ernzerhof (PBE) exchange–correlation functional with gradient correction [22] in combination with the split-valence, triple-zeta basis set 3 ζ . As shown in previous works, the PBE/3 ζ method reliably reproduces the experimental results on the polarizability, structural, and thermochemical properties of PAHs and fullerenes [10,21,24,25]. There are known limitations when using PBE and other pure DFT methods to describe extensive polyaromatic compounds [26]. These problems were studied in previous works on the examples of oligoacenes and phenacenes [9]. Analyzing our previous experience described in the abovementioned paper, we believe that the chosen method PBE/3 ζ is applicable to the compounds of this work, as they are not in the list of “restricted” compounds.

The structures of the studied molecules were fully optimized without restrictions on symmetry. We calculated vibrational spectra for all optimized structures to ensure that they correspond to the minima on the potential energy surfaces (no imaginary frequencies are observed).

The reaction enthalpies were calculated as the difference between the total energies E_{tot} of the products and the reactants, taking into account the zero-point vibrational energies (ZPVEs) and thermal corrections H_{corr} :

$$\Delta H_{298}^{\circ} = \sum_{products} (E_{tot} + ZPVE + H_{corr}) - \sum_{reactants} (E_{tot} + ZPVE + H_{corr}) \quad (1)$$

The polarizability tensors were calculated according to the finite field approach whereby their elements α_{ij} were defined as the second derivatives of the total energy of the molecule over the homogeneous external electric field:

$$\alpha_{ij} = -\frac{\partial^2 U}{\partial E_i \partial E_j} \quad (2)$$

The mean polarizability α was calculated from the diagonal elements (we express α values in \AA^3 in the present work):

$$\alpha = (\alpha_{xx} + \alpha_{yy} + \alpha_{zz})/3. \quad (3)$$

The reliability of the method has been previously scrutinized. Here, we just note one example of the difference between the PAH polarizabilities. The calculated α values of anthracene (27.0 \AA^3) and phenanthrene molecules (25.7 \AA^3), obtained with the PBE/3 ζ method, are in good agreement with the experimental references of 25.3 and 23.6 \AA^3 , respectively (taken from [27]).

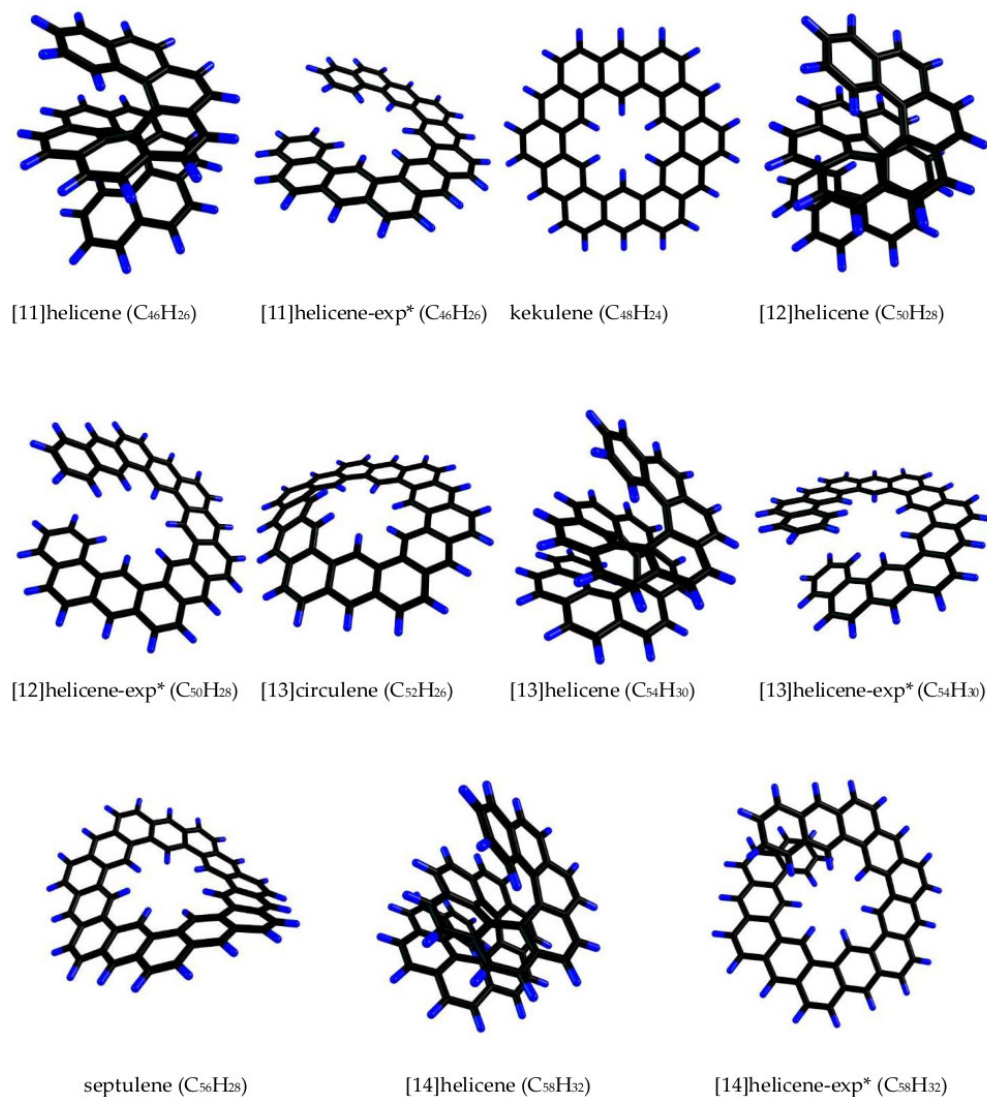


Figure 1. All polycyclic aromatic hydrocarbons under study. The energetically unfavorable conformations are not shown. Note: * Suffix “-exp” indicates expanded helicene structures.

3. Results and Discussion

3.1. Objects under Study

Kekulene (first synthesized in 1978 [28]) is a flat molecule consisting of 12 aromatic rings (Figure 1). It could be considered as a loop of six phenanthrene fragments [29,30]. The highest possible point symmetry group for kekulene is D_{6h} , which could be achieved in the case of full delocalization of the π -electrons. The real kekulene structure has a lower symmetry, viz. D_{3h} , because its π -electrons are partially localized within rings, so kekulene consists of alternating aromatic and non-aromatic rings (this is in accordance with Clar’s rule). The alternation reduces idealized symmetry [29,30]. The chosen PBE/3 ζ method reproduces this feature, and the optimization of the geometry results in a D_{3h} symmetry structure.

As the nearest kekulene benzenoids, we considered the isomeric pairs of helicenes and expanded helicenes consisting of 11 ($C_{46}H_{26}$) and 12 ($C_{50}H_{28}$) aromatic rings. Both regular and expanded helicenes are optically active, though devoid of chiral centers. In our calculations, we used the (*P*)-enantiomers, which means that their helices are right-handed (plus), i.e., they are dextrorotatory. Note that both enantiomers have the same mean polarizabilities, and we plan to focus on the types of enantiomers in future studies.

Expanded helicenes with the odd number of aromatic rings (11 and 13 in our work) theoretically have two forms: (i) the three terminal rings at each end of the helix may form a kinked pattern as in phenanthrene or (ii) a straight one as in anthracene. The forms with the kinked ends of the helix are noticeably more favorable in terms of their total energy: by 11.5 and 11.9 kJ/mol for [11]helicene-exp and [13]helicene-exp, respectively.

Septulene (first synthesized in 2012) has two additional aromatic rings compared with kekulene (14 in total). Similar to kekulene, it can be considered as a loop of seven phenanthrene fragments where π -electrons remain localized over the rings. The molecule is non-planar and, as shown in synthetic research [31], has different conformations in the gas phase and crystalline form. The septulene crystal consists of molecules in the chair conformation, which is evidenced by X-ray analysis and force-field calculations. However, in the gas phase, the chair form is 22–25 kJ/mol less favorable, and one is left with a choice between different saddle conformations (because the flat D_{7h} structure is similarly unfavorable, and its vibrational spectrum had imaginary frequencies in our calculations). We focused on the chair conformation of C_s symmetry because both C_2 and C_1 symmetry structures have small imaginary frequencies, and the energy differences between the three structures are less than 2.6 kJ/mol.

Similar to kekulene, we considered the isomeric pairs of regular and expanded helicenes consisting of 13 ($C_{54}H_{30}$) and 14 ($C_{58}H_{32}$) aromatic rings as the benzenoids nearest to septulene. All chosen helicenes are (*P*)-enantiomers.

Between kekulene and septulene, there should be [13]circulene with similar properties. This compound has not been synthesized thus far. Its molecule is non-planar and has a symmetry plane that corresponds to the C_s symmetry point group.

3.2. Thermodynamic Parameters of the Studied PAHs

When comparing the regular and expanded helicene isomers, we found the expanded helicene energetically more favorable in all studied pairs. The calculated differences between the isomers were 156.6, 169.9, 197.7, and 214.0 kJ/mol as the number of benzene rings increased from 11 to 14.

To assess the relative stabilities of our circulenes, we used the following hypothetical chemical reactions with ethylene that result in the corresponding helicenes:



and



where $n = 12, 13$, and 14 . As found, all reactions (5) were exothermic, whereas reactions (4) were endothermic (Table 1). The enthalpies of both reactions of [13]circulene were noticeably lower than of the other circulenes, i.e., [13]circulene was less favorable than its closest (with the same number of rings) helicenes. This means that either [13]circulene is a less energetically favorable circulene, or both [13]helicene and [13]helicene-exp are especially favorable helicenes.

Table 1. The calculated thermochemical parameters and mean polarizabilities of the studied PAHs.

Molecule	α (\AA^3)	$\Delta\alpha_R$ (\AA^3) ^a	ΔH_{298}° (kJ/mol) ^b
[11]helicene	80.1	n/a	n/a
[11]helicene-exp	102.9	n/a	n/a
kekulene	111.1	n/a	n/a
[12]helicene	86.3	−28.6	126.9
[12]helicene-exp	112.6	−2.3	−43.0
[13]circulene	123.0	n/a	n/a
[13]helicene	92.3	−34.5	90.5
[13]helicene-exp	119.1	−7.7	−107.2
septulene	135.5	n/a	n/a
[14]helicene	98.5	−40.8	141.4
[14]helicene-exp	127.8	−11.4	−72.6

^a The change in the mean polarizability for a reaction is calculated in the same way as the change in total energy; see Equation (6). ^b Heat effects of reactions (4) and (5).

3.3. Mean Polarizabilities of the Studied PAHs

The calculated mean polarizabilities α for all studied PAHs were in the range from 80.1 \AA^3 for [11]helicene to 135.5 \AA^3 for septulene (Table 1). Since the mean polarizability of a molecule depends on its size, we calculated the changes in mean polarizability for the hypothetical reactions (4) and (5) using the following formula [32]:

$$\Delta\alpha_R = \sum_{products} \alpha_i - \sum_{reactants} \alpha_j. \quad (6)$$

If we group the studied PAH molecules by the number of aromatic rings and consider the size effects (Equation (6)), then circulene had the highest mean polarizability, and helicene had the lowest one inside each group.

As previously found, the mean polarizability relates to the yields of the products of chemical reactions (see, e.g., works [10,21,24,33,34]). In addition, there is the minimum polarizability principle that states that less polarizable isomers should be more thermodynamically stable [35,36]. In the classical theory of chemical structure, polarizability is associated with the volume of the electron cloud of a molecule and its deformability under external electric fields [37]. Hence, we can explain the low polarizability of helicenes by the smaller volume of their electron cloud, which is caused by the compactness of their structure relative to circulenes and expanded helicenes.

Considering the change in mean polarizability upon the chemical reactions of the molecules of the same type (Figure 2), we found that it was linearly correlated with the number of aromatic rings for all three types. Another finding was that the ratio of the polarizabilities of helicenes and helicenes-exp was constant and equaled 1.3 for all four pairs of helicene isomers.

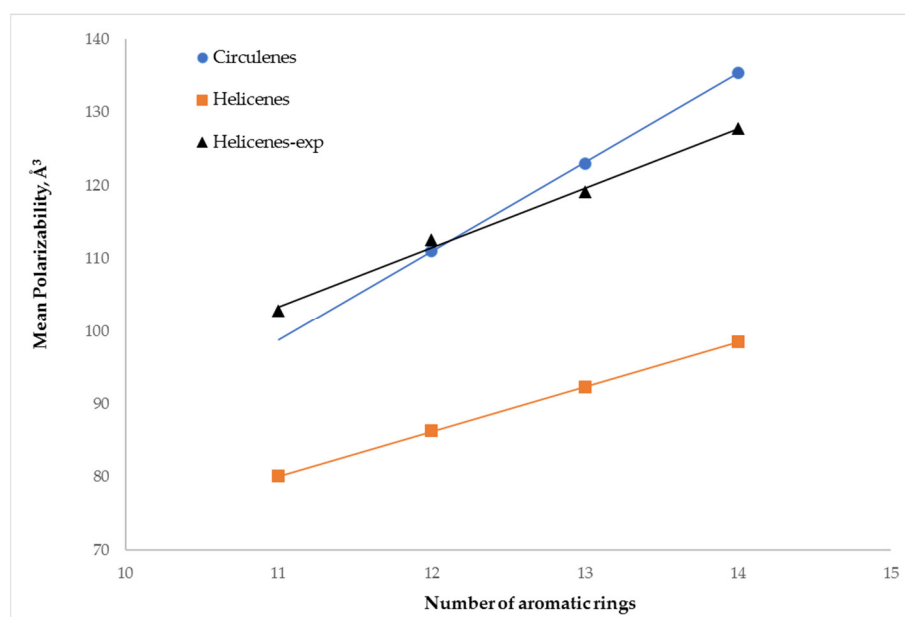


Figure 2. The relationship between the mean polarizability and the number of aromatic rings for selected PAHs.

3.4. General Remarks and Prospective

Very recently, we formulated a hypothesis that connects the structure, stability, and polarizability of PAHs depending on their planar or non-planar shape [17]. In general, non-planar PAHs have higher total energy and lower mean polarizability, compared with their planar isomers. Then, we started a systematic computational study to substantiate this statement with relevant examples. The case considered in the present study fits into the regularity. Indeed, if we consider isomers, the planarity increased in the following series: helicene < expanded helicene < planar PAHs. Accordingly, the stability and mean polarizability increased.

We assume that, in addition to astrochemical applications [10], our results could be useful for studying heavy oil fractions enriched with high-molecular aromatic hydrocarbons [38] and applications of PAHs in material science [39]. In a fundamental aspect, we will continue our research to quantify our qualitative correlations. We plan to find the regularities between structural descriptors and molecular properties [40]. There should be a more general regularity because our above-formulated statement is not exclusive to isomeric compounds (e.g., the ratio of “high planarity–high polarizability” is fulfilled in the case of the related PAHs connected with the transformations according to the Scholl reaction [41]).

4. Conclusions

Based on the performed density functional theory computations, we found that in a series of PAHs including kekulene, septulene, and the nearest benzenoids, the mean polarizability decreased with the molecule’s planarity, even though circulenes themselves were not always flat. This decrease was observed between both isomers and non-isomeric structurally related molecules. These findings serve as another example in favor of a previously found correlation that revealed higher mean polarizability for more planar PAHs than for non-planar isomers or structurally close molecules. The mean polarizability of helicenes was much lower than that of all other PAH types, which is explained by the significantly reduced volume of their electron clouds due to the compactness of the helical structure.

Another intriguing finding was that the expanded helicenes were found to be considerably more energetically favorable than the well-known and well-studied regular helicenes. This observation violates the minimum polarizability principle. We hope this theoretical study will enhance future interest in the expanded helicenes as a class of PAHs with non-trivial molecular properties.

Supplementary Materials: Cartesian coordinates of the optimized molecules, key energy, and polarizability parameters can be downloaded at: www.mdpi.com/article/10.3390/c8040061/s1.

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