

## Article

# Thiourea Derivative Metal Complexes: Spectroscopic, Anti-Microbial Evaluation, ADMET, Toxicity, and Molecular Docking Studies

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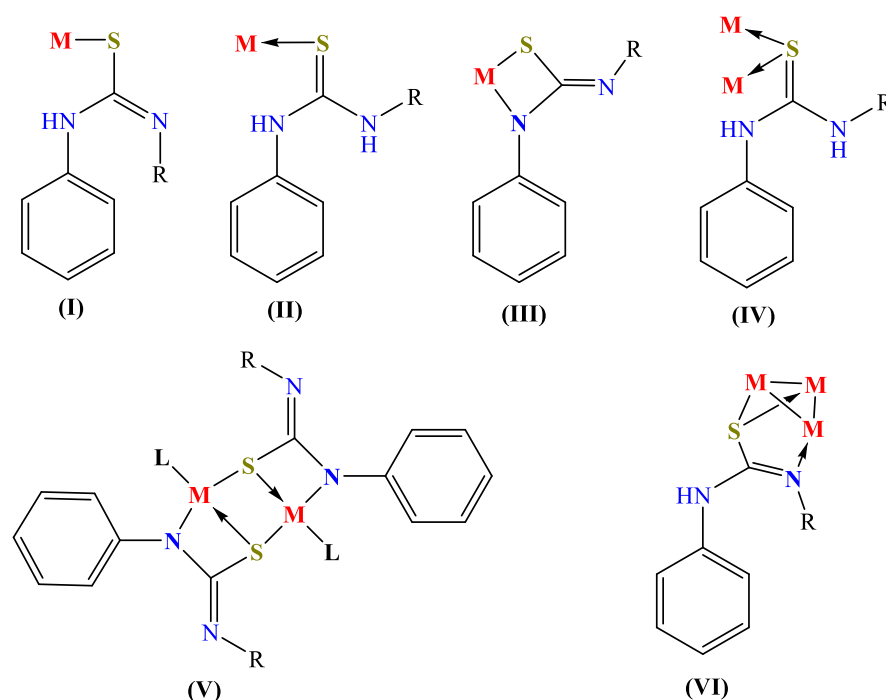
**Abstract:** The treatment of N-Phenylmorpholine-4-carbothioamide (HPMCT) with bivalent metal ions in a 2:1 mol ratio without a base present affords  $[MCl_2(\kappa^2S\text{-HPMCT})_2]$   $\{M = \text{Cu(1)}, \text{Pd(2)}, \text{Pt(3)}, \text{and Hg(4)}\}$  in a good yield. Furthermore, the reaction of two equivalents of HPMCT and one equivalent of bivalent metal ions in the presence of Et<sub>3</sub>N has afforded  $[M(\kappa^2S,N\text{-PMCT})_2]$   $\{M = \text{Ni(5)}, \text{Cu(6)}, \text{Pd(7)}, \text{Pt(8)}, \text{Zn(9)}, \text{Cd(10)}, \text{and Hg(11)}\}$ . Infrared, <sup>1</sup>H, <sup>13</sup>C Nuclear Magnetic Resonance molar conductivity, and elemental analysis were used to characterize the synthesized complexes. The results suggest that HPMCT is bonded as monodentate via an S atom in Complexes (1–4), whereas linkage as a bidentate chelating ligand via S and N atoms gives two chelate rings. Moreover, the synthesized ligand and the complexes were screened for antibacterial activity, which displayed that the very best antibacterial activities for Complexes (1), (6), and (3). In addition, the cytotoxic activity of the HPMCT ligand,  $[PdCl_2(HPMCT)_2]$  (2), and  $[PtCl_2(HPMCT)_2]$  (3) were screened on breast cancer cell lines (MCF-7), and Complex (3) reveals the most promising activity with an IC<sub>50</sub> value 12.72 ± 0.4 μM. Using the B3LYP method and 6-311++G(d,p) basis sets for the ligand and the SDD basis set for the central metal, the synthesized complexes utilizing the prepared ligand were optimized. Various quantum parameters such as hardness, electron affinity, dipole moment, vibrational frequencies, and ionization energy for the ligand and its complexes have been calculated. In general, a favorable agreement was found between the experimental results and the obtained theoretical results.

**Keywords:** thiourea; morpholine; anti-cancer; molecular docking; DFT

## 1. Introduction

In the previous 50 years, thioureas, an important ligand, have received a lot of attention. Their multi-donor sites have contributed to their coordination chemistry. The literature is well-versed in the coordination chemistry of thioureas and their derivatives [1–7]. Different coordination possibilities may arise from attaching various groups to the thiouredio moiety (Figure 1I–VI). Thioureas can be coordinated through their S atom as a

neutral monodentate mode [8] or anionic monodentate mode [9]. They can also be coordinated through both of their N, S donor atoms in a bidentate mode [10]. Moreover, thioureas with a multi-dentate mode have also been reported including an M-M bridging mode (Figure 1) [11]. Due to these interesting coordination behaviors, thioureas have been reported to participate in various fields such as organo-catalysts, anticorrosion, and synthesis of bio-molecule [12–16].



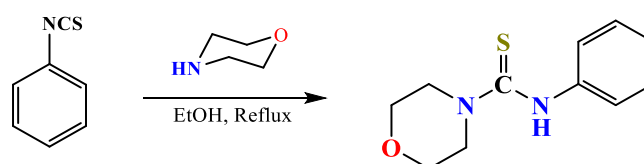
**Figure 1.** Different coordination possibilities of thiourea derivatives.

Even though N-Phenylmorpholine-4-carbothioamide (HPMCT) was reported and its X-ray structure was determined [17], only a few examples of its metal complexes have been reported. For instance, Yuen et al. have reported the synthesis of mixed ligand complexes of Ni derived from HPMCT and dppe [18]. To the best of our knowledge, this ligand has not been well-represented in the literature. We therefore sought to expand on this work by using various metal ions and examining their biological and quantum characteristics. Herein, we report a series of divalent metal complexes ( $M^{II} = \text{Ni, Cu, Pd, Pt, Zn, Cd, and Hg}$ ) derived from HPMCT, studying their quantum and biological applications.

## 2. Results and Discussion

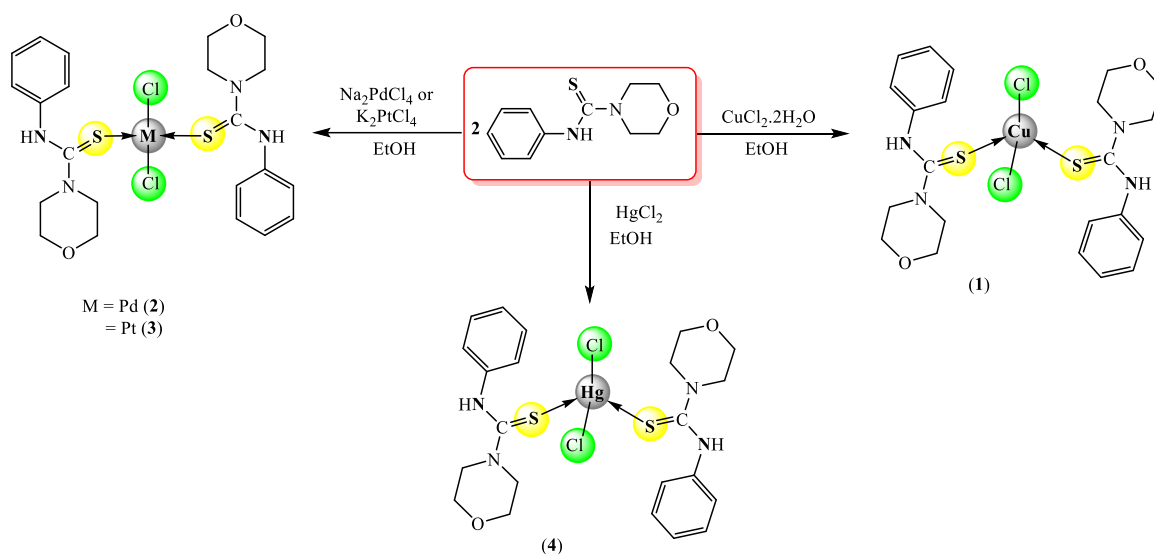
### 2.1. Synthesis

The ligand HPMCT was synthesized by following a modified literature method [19] (Scheme 1). The reaction of morpholine with phenylisothiocyanate in an equal ratio gave the desired ligand as a white solid in a high yield 90% yield.



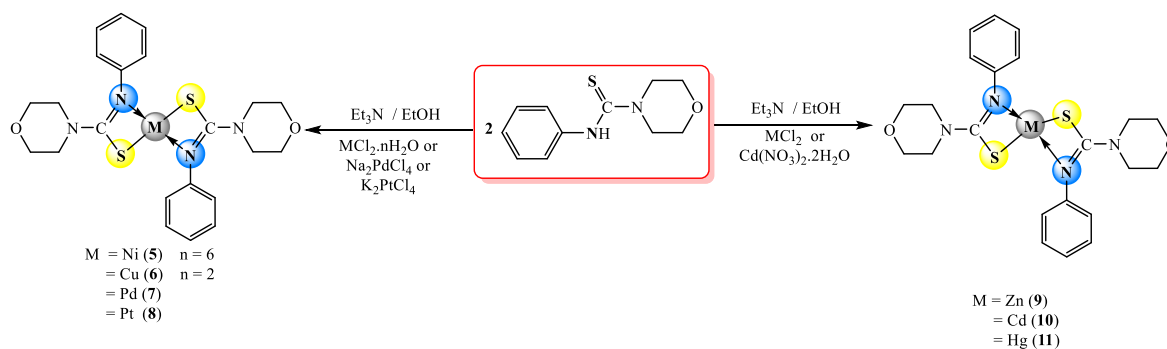
**Scheme 1.** Preparation of N-Phenylmorpholine-4-carbothioamide (HPMCT).

Reaction of the thiourea derivative (HPMCT) ligand with inorganic salts  $\{\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ ,  $\text{Na}_2\text{PdCl}_4$ ,  $\text{K}_2\text{PtCl}_4$  and  $\text{HgCl}_2\}$  in EtOH as a solvent afforded  $[\text{MCl}_2(\kappa^1\text{S-HPMCT})_2]$   $\{\text{M}^{\text{II}} = \text{Cu}$  (1), Pd (2), Pt (3), and Hg (4) in a high yield (73 and 95)% (Scheme 2).



**Scheme 2.** Preparation of Complexes (1–4).

The treatment of two moles of the N-Phenylmorpholine-4-carbothioamide ligand with one mole of  $\{\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ ,  $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ ,  $\text{Na}_2\text{PdCl}_4$ ,  $\text{K}_2\text{PtCl}_4$ ,  $\text{ZnCl}_2$ ,  $\text{Cd}(\text{NO}_3)_2 \cdot 2\text{H}_2\text{O}$ , and  $\text{HgCl}_2\}$  in a basic medium afforded complexes of the type  $[\text{M}(\kappa^2\text{S,N-PMCT})_2]$   $\{\text{M}^{\text{II}} = \text{Ni}$  (5), Cu (6), Pd (7), Pt (8), Zn (9), Cd (10), and Hg (11) in good yields (range 67–91%) (Scheme 3).



**Scheme 3.** Preparation of Complexes (3) and (4).

The synthesized HPMCT ligand and its complexes  $[\text{MCl}_2(\kappa^1\text{S-HPMCT})_2]$   $\{\text{M}^{\text{II}} = \text{Cu}$  (1), Pd (2), Pt (3), and Hg (4) and  $[\text{M}(\kappa^2\text{S,N-PMCT})_2]$   $\{\text{M}^{\text{II}} = \text{Ni}$  (5), Cu (6), Pd (7), Pt (8), Zn (9), Cd (10), and Hg (11) are stable and readily soluble in DMF and DMSO. The analytical data of the HPMCT and its complexes are listed in Table 1. With the help of FT-IR,  $^1\text{H}$ - and  $^{13}\text{C}$  NMR spectroscopic techniques, conductivity measurements, and CHN analysis, the structure of the complexes has been studied further. The nanostructures of the synthesized HPMCT ligand and its complexes were studied by XRD and SEM techniques. The complexes (1–11) prepared in DMSO solvent at  $10^{-3}$  M (3–19)  $\Omega^{-1} \text{ cm}^{-1} \text{ mol}^{-1}$  showed that the complexes are non-electrolytes in data obtained from conductivity measurements [20]. The CHN analysis agreed with the suggested structures of the prepared compounds.

**Table 1.** Color, yield (%), M.p.(°C), molar conductivity (Cond.), and CHN analysis of the ligand and its complexes.

Comps.	Color	Yield(%)	Mp.°C	Cond.	Elemental Analysis % Calc. (Found)		
					C	H	N
HPMCT	White	90	145–147	--	59.43 (59.29)	6.35 (6.33)	12.60 (12.78)
1	Greenish yellow	73	130–132	10	45.63 (45.66)	4.87 (4.90)	9.68 (9.71)
2	Dark red	95	148–150	17	42.49 (42.52)	4.54 (4.57)	9.01 (9.04)
3	Dark yellow	90	178–180	19	37.18 (37.21)	3.97 (4.07)	7.88 (7.91)
4	White	74	188–190	2	36.90 (36.93)	3.94 (3.97)	7.82 (7.85)
5	Light olive	79	175–178	3	52.71 (52.74)	5.23 (5.26)	11.18 (11.21)
6	Dark yellow	77	158–161	4	52.21 (52.24)	5.18 (5.21)	11.07 (11.10)
7	Dark yellow	76	120–123	8	48.13 (48.16)	4.77 (4.80)	10.20 (10.23)
8	Dark yellow	91	142–144	8	41.44 (41.47)	4.11 (4.14)	8.79 (8.82)
9	White	67	190–193	6	52.02 (52.05)	5.16 (5.19)	11.03 (11.06)
10	White	70	192–194	12	47.61 (47.64)	4.72 (4.75)	10.09 (10.12)
11	Gray	82	118–121	11	41.08 (41.11)	4.07 (4.10)	8.71 (8.74)

## 2.2. Spectroscopic Data

### IR Spectra

IR spectroscopic data of HPMCT and its complexes are measured within the 4000–400  $\text{cm}^{-1}$  range using a KBr disc (Table 2 and Figure S1–S11). A comparison of free HPMCT and its complexes designated the bond of the relevant ligand with ions via the S atom as monodentate or through S, N atoms in a bidentate fashion. Through the coordination, certain changes such as increases or decreases in frequency corresponding to NH, CH aliphatic, CH aromatic, CS, and NCN bonds could be expected [21].

The IR spectrum of the HPMCT ligand (Figure S1) showed an NH peak at 3176  $\text{cm}^{-1}$ ; this band was found in the range (3174–3336)  $\text{cm}^{-1}$  for Complexes (1–4). The  $\nu\text{CS}$  stretching frequency for HPMCT was shown at 707  $\text{cm}^{-1}$ ; this band was shifted towards the lower side (11–42)  $\text{cm}^{-1}$ . This indicates that the HPMCT ligand was bonded to ions through the sulfur atom [1,2,13,19]. The  $\nu\text{C-N}$  band displayed within (1269–1303)  $\text{cm}^{-1}$ ; this band was shifted to higher frequencies indicating the flow of electron density through the C=S bond towards the metal ion further strengthening the bond in the N-C-N fragment. A strong to medium band is shown in the IR spectra of the complexes within the (1022–1031)  $\text{cm}^{-1}$  range assigned to the  $\nu(\text{C-O})$ ; the slight shift of this band compared with that of the free ligand indicates that the C-O group does not participate in bonding. The IR spectra of Complexes (5–11) exhibited the absence of the stretching vibration of  $\nu(\text{NH})$ , which appeared at 3176  $\text{cm}^{-1}$  in the free ligand; this indicates the formation of the proposed complexes. And the frequency of CS stretching appeared within the (665–696)  $\text{cm}^{-1}$  range and shifted to a high frequency in the CN stretching frequency which appeared within (1265–

1274)  $\text{cm}^{-1}$ , indicating that the ligand is coordinated to center ions through sulfur and nitrogen atoms as a bidentate chelating mode [1,2,13,19].

**Table 2.** Selected IR bands of the ligand and its complexes.

Comps.	$\nu\text{NH}$	$\nu(\text{C-H})$		$\delta\text{NH}$	$\nu\text{C=N}$ $\nu\text{C=N}$	$\nu\text{C=C}$	$\nu\text{C-N}$	$\nu\text{C-O}$	$\nu\text{C=S}$	$\nu\text{M-S}$ $\nu\text{M-N}$
		Ar.	Aliph.							
HPMCT	3176s	3030s	2918m 2868m	1595s	--	1531s	1211s	1028s	707s	--
1	3336b	3049w	2968w 2860w	1626s	--	1489s	1276s	1031m	694m	495w
2	3174b	3055w	2960w 2858w	1597s	--	1558s	1269s	1028s	696m	503w
3	3207s	3059w	2966s 2854s	1645s	--	1583s	1303s	1028s	694s	493s
4	3184b	3057w	2966w 2858w	1597s	--	1558s	1269s	1026s	684s	482w
5	--	3024w	2918w 2841w	--	1550s	1481m	1276m	1026s	692m	418w 518w
6	--	3061w	2960w 2848w	--	1529s	1485m	1267m	1022s	694m	470w 515w
7	--	3053w	2968w 2856w	--	1541s	1487w	1267m	1028m	696m	420w 516w
8	--	3051w	2964w 2852w	--	1541s	1485m	1269m	1028s	694m	461w 516w
9	--	3028m	2976m 2854m	--	1531s	1491m	1265m	1028s	665m	472w 507w
10	--	3057w	2962w 2854w	--	1519s	1435m	1269m	1026s	696m	466w 516w
11	--	3051w	2962m 2848w	--	1581s	1485m	1274m	1024s	686m	489w 586w

### 2.3. NMR Spectra

The  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra of the prepared complexes (Figures S12–S22) showed many sets of peaks for the same functional groups. Some of these ligand groupings are disappearing or are more negatively impacted than others. We also took the NMR spectra of the free HPMCT ligand and its complexes to assign resonance signals more precisely. For ease of comparison, Tables 3 and 4 list the chemical shifts of the groups that are more susceptible to changes in the chemical environment (Tables 3 and 4).

**Table 3.** Proton chemical shifts (in ppm) of HPMCT and its complexes.

Comps.	$\delta\text{NH}$	$\delta\text{H}$ Phenyl	$\delta\text{H}$ $\text{OCH}_2$	$\delta\text{H}$ $\text{NCH}_2$
HPMCT	9.29(s, 1H)	7.04–7.28(m, 5H)	3.84(t, J8.00 Hz, 4H)	3.61(t, J8.00 Hz, 4H)
2	10.28(s, 2H)	7.05–7.81(m, 10H)	3.91(t, J7.80 Hz, 8H)	3.75(t, J7.80 Hz, 8H)
3	9.59(s, 2H)	7.33–7.81(m, 10H)	3.92(t, J7.90 Hz, 8H)	3.66(t, J7.90 Hz, 8H)
4	10.37(s, 2H)	7.21–7.36(m, 10H)	3.88(t, J8.00 Hz, 8H)	3.71(t, J8.00 Hz, 8H)
7	-	7.62(d, J8.00 Hz, 4H), 7.29–7.34 (m, 4H), 7.04(t, J7.60 Hz, 2H)	3.92(t, J8.00 Hz, 8H)	3.66(t, J8.00 Hz, 8H)
8	-	7.48–7.76 (m, 10H)	3.56(t, J7.60 Hz, 8H)	3.38(t, J7.60 Hz, 8H)
9	-	7.20–7.67 (m, 10H)	3.82(t, J8.00 Hz, 8H)	3.70(t, J8.00 Hz, 8H)

10	-	7.21–7.36 ( <i>m</i> , 10H)	3.88( <i>t</i> , <i>J</i> 7.80 Hz, 8H)	3.71( <i>t</i> , <i>J</i> 7.80 Hz, 8H)
11	-	6.87–7.24 ( <i>m</i> , 10H)	3.74( <i>t</i> , <i>J</i> 8.00 Hz, 8H)	3.61( <i>t</i> , <i>J</i> 8.00 Hz, 8H)

**Table 4.** Carbons chemical shifts (in ppm) of HPMCT and its complexes.

Comps.	C=S	$\delta$ C Phenyl	$\delta$ C OCH <sub>2</sub>	$\delta$ C NCH <sub>2</sub>
HPMCT	182.35	141.45, 128.56, 125.97, 124.92	66.30	48.96
2	176.78	139.41, 126.23, 124.34, 122.10	64.44	51.45
3	177.31	137.26, 126.11, 124.09, 122.45	63.71	50.61
4	174.80	139.59, 129.59, 127.14, 126.77	65.86	50.52
7	168.51	136.02, 127.81, 124.75, 121.93	66.52	48.22
8	167.37	137.34, 127.63, 124.93, 121.81	67.01	47.39
11	168.89	136.02, 128.94, 125.01, 122.13	66.34	46.19

The  $^1\text{H}$  nmr spectrum of the free ligand showed four separated signals which represent four different proton groups in the ligand structure, and a singlet peak at  $\delta$ 9.29 ppm due to the proton of the NH group, while the methylene protons of the morpholine ring were displayed as two triplet peaks at  $\delta$ 3.84 ppm and  $\delta$ 3.61 ppm, whereas the phenyl protons exhibit as a multiplet within the  $\delta$  7.04–7.28 ppm range.

The  $^{13}\text{C}$ -NMR spectrum of the HPMCT ligand is in good agreement with the suggested structure, since HPMCT has seven different types of carbon atoms (see Figure S20). The significant peaks in the  $^{13}\text{C}$  NMR spectrum are the carbon peaks of the C=S group at  $\delta$ 182.35 ppm and the methylene groups of the morpholine ring at  $\delta$ 66.30 ppm and  $\delta$ 48.96 ppm [19]. Other chemical shifts are listed in Table 4.

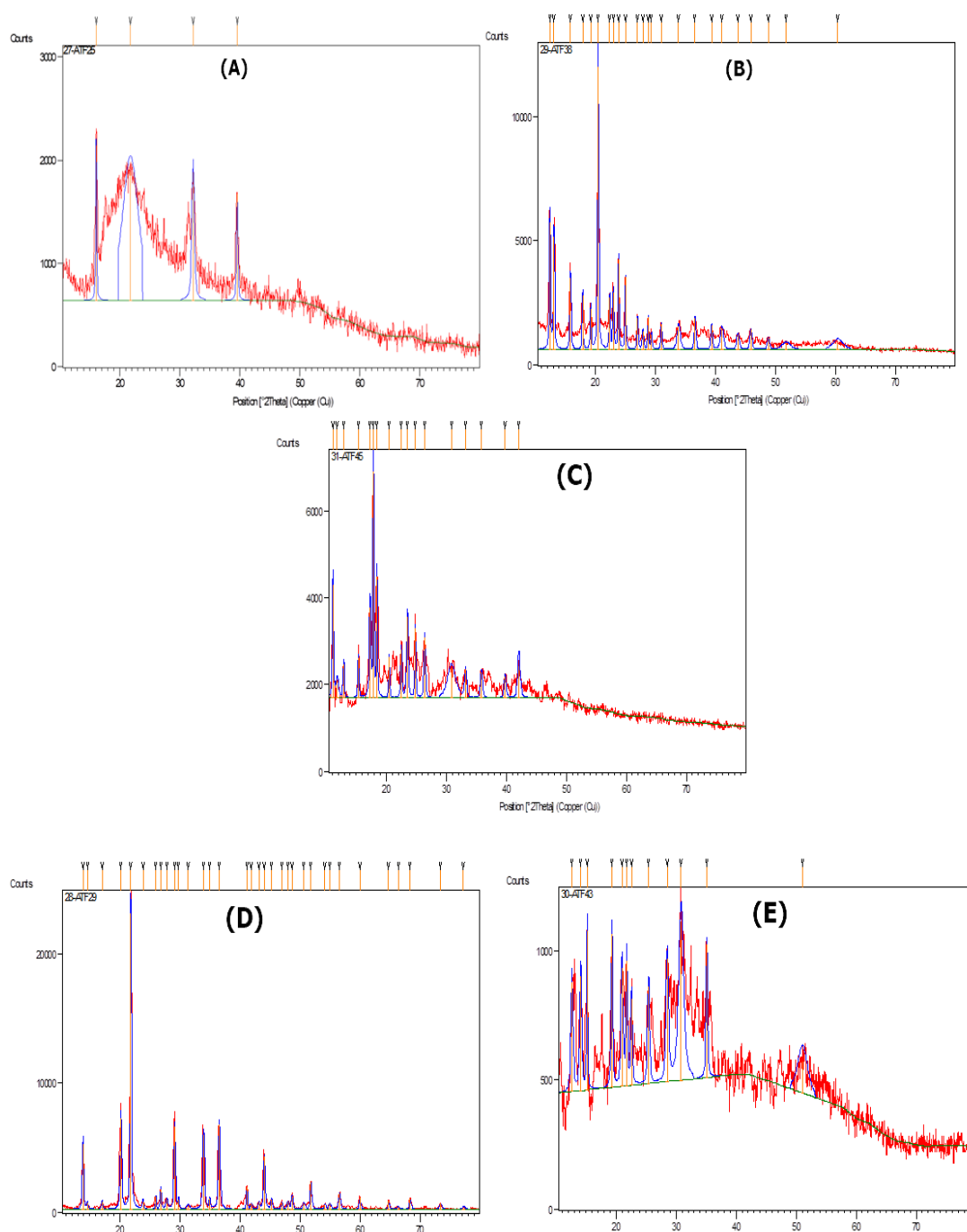
In Complexes (2–4), the  $^1\text{H}$  NMR spectra contain NH signals at  $\delta$ 10.28 ppm (2),  $\delta$ 9.59 ppm (3), and  $\delta$  10.37 ppm (4). It is worth noting that the chemical shifts of the proton of the NH group were shifted to a high chemical shift compared with the free ligand, suggesting increased electron density on this proton and that coordination happened through the sulfur atom of the thiocarbonyl group. This NH peak disappeared in the spectra of Complexes (5, 7–11) which supports the existence of ligands in a deprotonated form on complexation. All spectra of Complexes (2–5, 7–11) displayed two triplet peaks, with the first peak at  $\delta$  3.56–3.92 ppm (approx. *J* = 8.00 Hz) corresponding to four protons assigned to the protons of the methylene groups of morpholine (O-CH<sub>2</sub>). The second triplet peak at  $\delta$ 3.38–3.75 ppm (approx. *J* = 8.00 Hz) corresponding to four protons was assigned to the protons of the second methylene groups of morpholine (N-CH<sub>2</sub>). In addition, the aromatic protons of Complexes (2–5, 7–11) appear in the range of 6.87–7.81 ppm.

The above results were supported by the  $^{13}\text{C}$  NMR spectra which showed low-frequency peaks at  $\delta$  174.80–177.31 ppm. These recorded shifts reflected a down-field shift of about 6 ppm from the free thiourea ligand, suggesting decreased electron density on the thiocarbonyl carbon and that the coordination happened through the sulfur atom. In a similar view, the chemical shift of the methylene group (NCH<sub>2</sub>) was increased upon being coordinated with metal ions, which showed at  $\delta$  50.52–50.61 ppm. Other chemical shifts are listed in Table 4.

#### 2.4. XRD and SEM Data

The XRD pattern of Complexes (6) and (7) (Figure 2C,D) displayed sharp peaks within the 10–45° range which demonstrate that the complexes are well crystalline, whereas in the XRD spectra of Complexes (1), (5), and (10) (Figure 2A,B,E), wide peaks and non-regular spectra indicated that these complexes indicated a low crystallinity and the presence of an amorphous phase. Depending on the Scherrer equation [22], the average particle size of complexes was recorded for the highest five peaks and the obtained results verify that the calculated particle size was within the range of 2.68–49.93 nm, and

the average particle size was equal to 29.26 nm, 23.95 nm, 25.48 nm, 32.06 nm, and 24.70 nm in Complexes (1), (5), (6), (7), and (10), respectively.



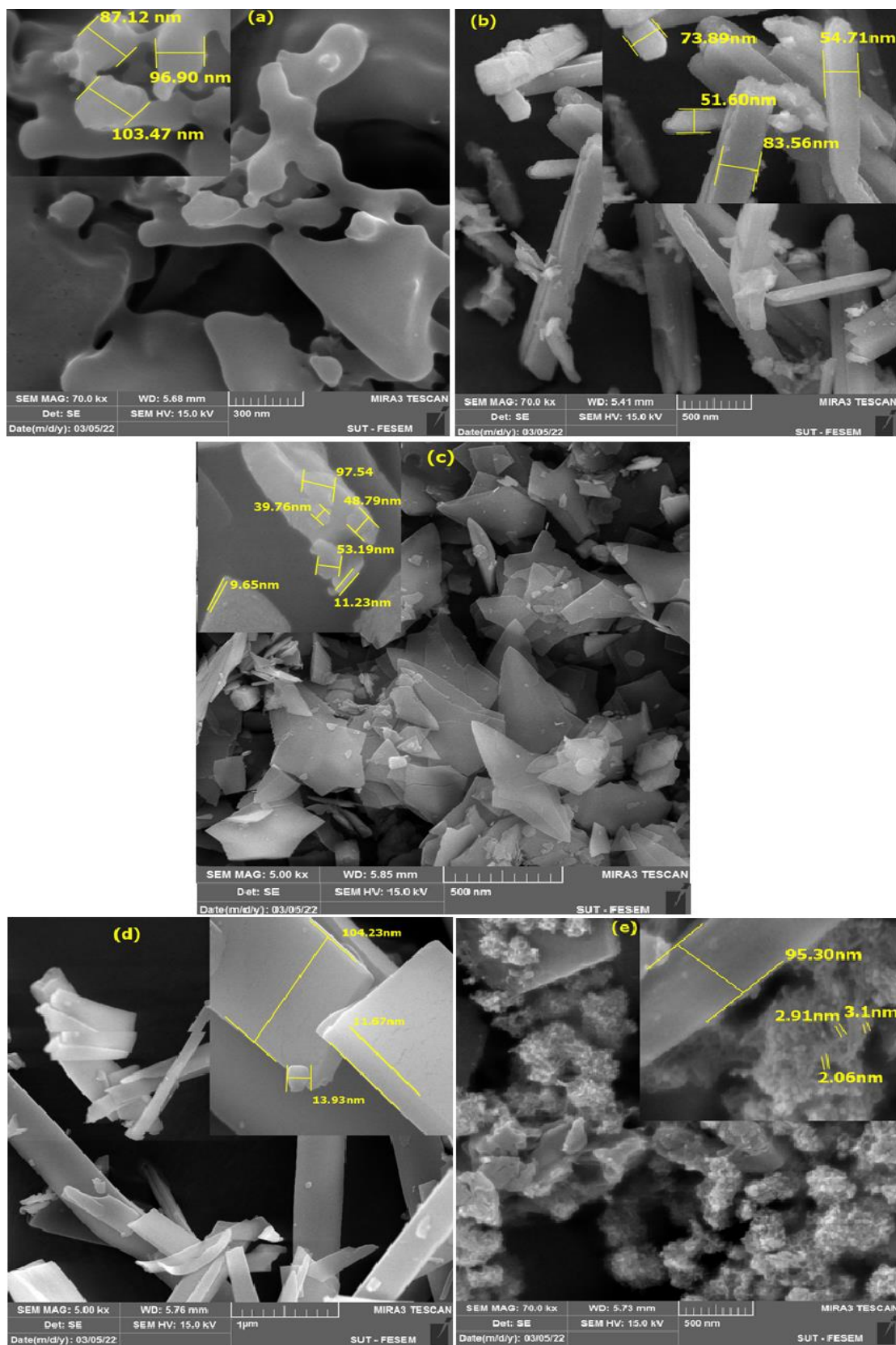
**Figure 2.** XRD patterns of (A) Complex 1, (B) Complex 5, (C) Complex 6, (D) Complex 7, and (E) Complex 10.

The SEM images of Complex (1) (Figure 3a) showed fused irregular shapes with each other at different nano-scale sizes within 87–103 nm, whereas the SEM images of Complexes (5) and (7) (Figure 3b,d) appeared as nano-rods with different nano-scale sizes within 12–104 nm. Also, Complex (5) contains nano-sheets distributed between the nano-rods.

Complex (6) in SEM images (Figure 3c) was displayed as nano-sheets within 40–98 nm, while the thickness of the nano-sheets was within 9–11 nm.



Finally, Complex (10) appeared in SEM images (Figure 3e) as nanostructures like wool with small nano-scale sizes within 2–3 nm. Also, the image contained irregular structures distributed on the complex surface with different nano-scale sizes within 95–105 nm.





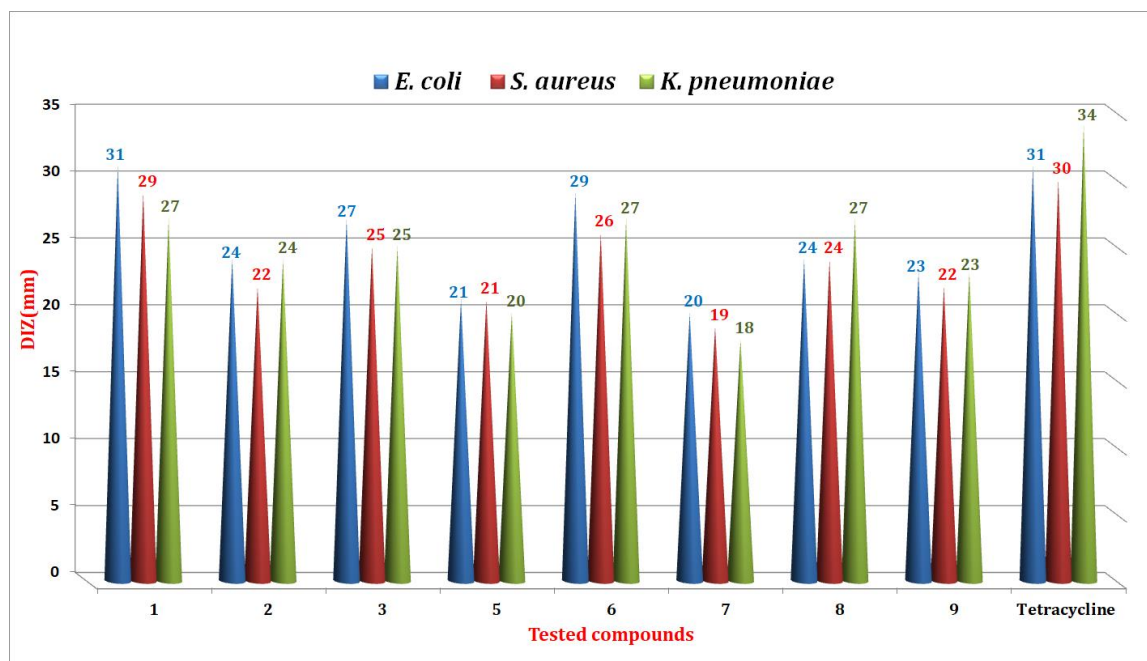
**Figure 3.** SEM images of (a) Complex 1, (b) Complex 5, (c) Complex 6, (d) Complex 7, and (e) Complex 10.

## 2.5. Biological Studies

In view of the importance of thiourea or derivatives in the medical field and their vital efficacy towards many bacteria or fungi and their important role having a good activity against many types of cancer cell lines [23], we have decided to study the effectiveness of the HPMCT ligand and its complexes as antibacterial against three pathogenic bacteria and cytotoxicity activity against the MCF-7 cell lines.

### 2.5.1. Antibacterial Activity

The biological activity of the HPMCT ligand and its complexes was evaluated against three types of bacteria (*Escherichia coli*, *Staphylococcus aureus*, and *Klebsiella pneumoniae*) at a  $10^{-3}$  M concentration and the results were compared with tetracycline as a positive control and DMSO solvent as a negative control. The DIZ (mm) of the obtained results is shown in Figure 4.



**Figure 4.** Histogram of biological activities of synthesized compounds.

Apart from the ligand (HPMCT), all of the compounds demonstrated strong antibacterial action against harmful bacteria, suggesting their potential for development. These compounds have not, however, surpassed the ordinary tetracycline's antibacterial effectiveness. The Cu(II) (1 and 6) and Pt(II) (3 and 8) complexes have also been found to be particularly effective at inhibiting the chosen microorganisms. On the other hand, the sequence of complexes in terms of inhibition was as follows:

$$1 > 6 > 3 > 8 > 2 > 9 > 5 > 7 > \text{HPMCT}$$

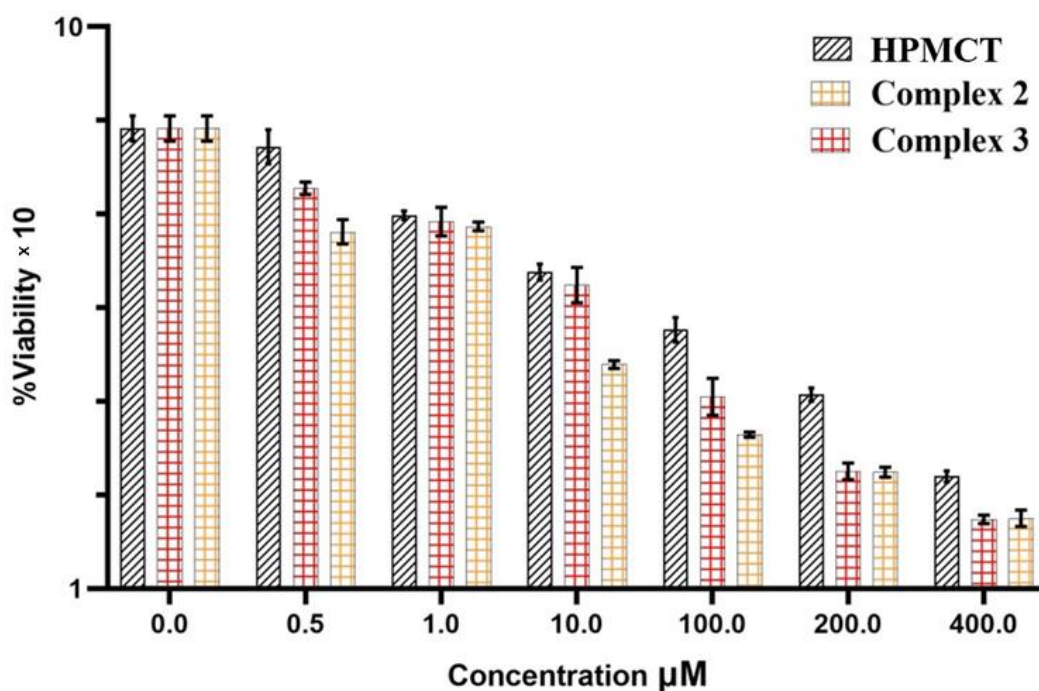
All complexes displayed a good activity against *E. coli* in comparison with other two species when compared against the tetracycline standard (Figure 4). Thus, from the above study, it was concluded that these synthesized compounds may be targeted as antibacterial agents for further investigations.

### 2.5.2. Anti-Cancer Activity

The cytotoxic activity of the free HPMCT ligand and its complexes (2) and (3) was evaluated using the MTT colorimetric assay [24] against the MCF-7 cell line in 0.01  $\mu\text{M}$ –100  $\mu\text{M}$ . The free HPMCT ligand showed a negligible effect, whereas the Pt(II) complex (3), exhibited a good inhibitory effect and the highest activity compared with the Pd(II) complex (2) against the MCF-7 cell line. The  $\text{IC}_{50}$  value of the HPMCT LH ligand is  $80.30 \pm 1.905 \mu\text{M}$ . The complexes  $[\text{PdCl}_2(\text{HPMCT})_2]$  (2) and  $[\text{PtCl}_2(\text{HPMCT})_2]$  (3) appeared to have a good inhibitory effect against the MCF-7 cell line with  $\text{IC}_{50} = 28.93 \pm 0.795 \mu\text{M}$  and  $12.72 \pm 0.376 \mu\text{M}$ , respectively (see Figure 5 and Table 5). These results indicate that Complexes (2) and (3) have a good cytotoxicity but are less effective than cis-platin.

**Table 5.**  $\text{IC}_{50}$  values of HPMCT,  $[\text{PdCl}_2(\text{HPMCT})_2]$ , and  $[\text{PtCl}_2(\text{HPMCT})_2]$  against the MCF-7 cell line.

Compounds	$\text{IC}_{50}$ ( $\mu\text{M}$ )
HPMCT	$80.30 \pm 1.905$
$[\text{PdCl}_2(\text{HPMCT})_2]$	$28.93 \pm 1.105$
$[\text{PtCl}_2(\text{HPMCT})_2]$	$12.72 \pm 0.576$
cis-platin	$3.19 \pm 0.097$



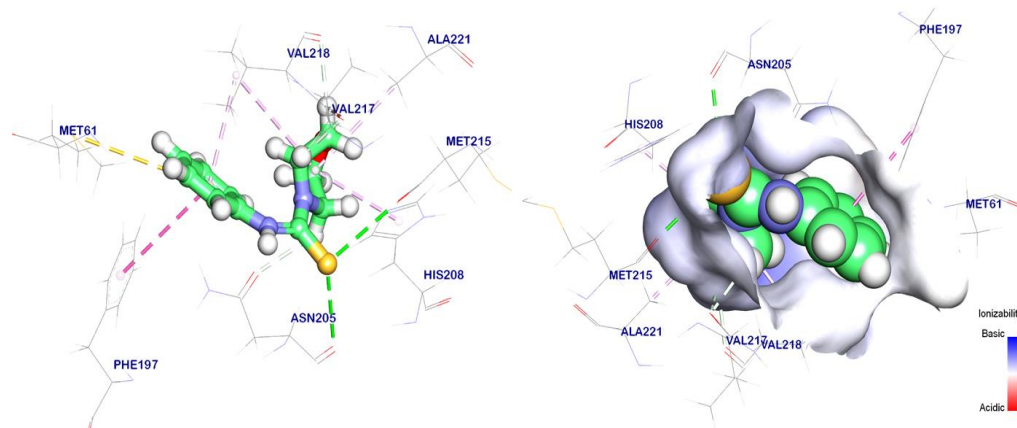
**Figure 5.** Cytotoxicity diagram of HPMCT,  $[\text{PdCl}_2(\text{HPMCT})_2]$  (2), and  $[\text{PtCl}_2(\text{HPMCT})_2]$  (3) against the MCF-7 cell line at different concentrations ( $\mu\text{M}$ ).  $p$  value = 0.005 in each case.

### 2.6. Molecular Docking

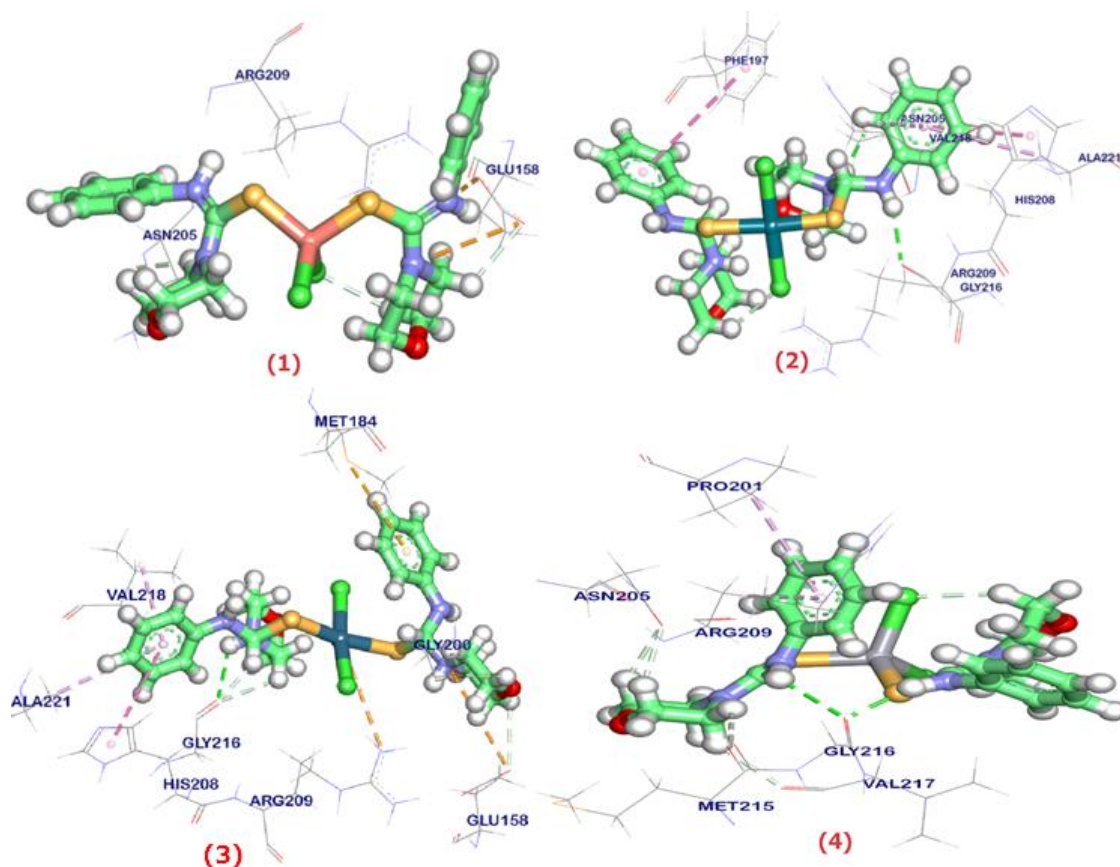
Bacterial tyrosinase was bound by the free HPMCT ligand with an energy of  $-7.01$  kcal/mol, and five  $\pi$ - $\pi$  and  $\pi$ -alkyl interactions with Pro201, His60, Ala221, His208, and Val2018 were formed. Two hydrogen bonds with Arg209 and Asn205 are formed at distances of 1.67 and 1.79 Å, as well as halogen interaction binding with His231 (Figure 6).

The binding mode of Complexes (1, 2, 3, and 4) exhibited an energy binding of  $-6.11$ ,  $-6.15$ ,  $6.32$ , and  $-5.86$  Kcal/mol against bacterial tyrosinase (Figure 7 and Table 6). Complex (1) creating pi-cation and attractive interactions with Glu158 and Arg202. Complex (2) interacted with Asn205 and Gly216 by two hydrogen bonds with  $\pi$ - $\pi$  and  $\pi$ -alkyl interactions with His208, Phe197, Val221, and Val218, while Complex (3) formed  $\pi$ - $\pi$ ,  $\pi$ -sulfur, and  $\pi$ -alkyl interactions with His208, Met184, and Ala221. Moreover, it interacted with

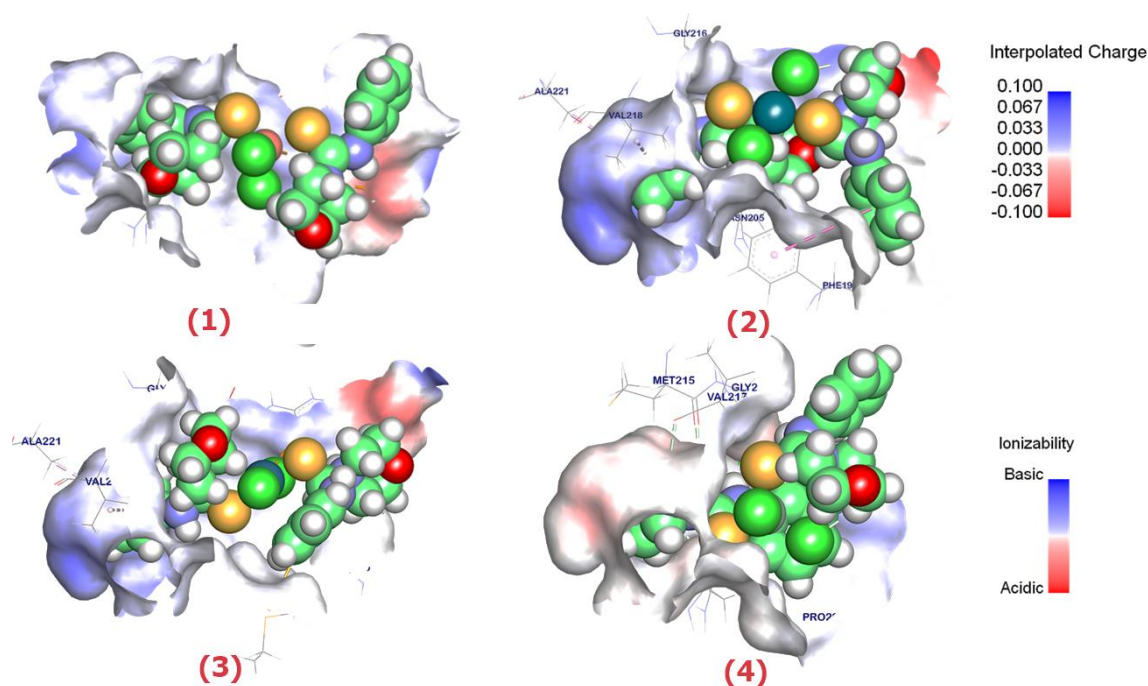
Gly216, Arg209, and Glu158 by hydrogen bond and attractive interactions. Complex (4) interacted with Arg209, Pro201, and Phe197 by  $\pi$ -alkyl,  $\pi$ - $\pi$ , and attractive interactions and additionally formed hydrogen bonds with Gly216 and Arg209, respectively. Furthermore, we studied the interaction of the synthesized complexes (1–4) with the active pocket of the bacterial tyrosinase enzyme, and Figure 8 shows this interference through the surface mapping in addition to the color gradient of the interpolated charge and ionizability during the interference of the prepared complex with the active pocket of the enzyme.



**Figure 6.** The mapping surface displays the crystal ligand filling the active pocket of bacterial tyrosinase, with hydrogen bonds (green) and the pi interactions shown by purple lines. The HPMCT ligand is docked in the enzyme.



**Figure 7.** Complexes (1–4) docked in the bacterial enzyme tyrosinase are shown with hydrogen bonds (green) and pi interactions as purple lines.



**Figure 8.** Mapping the surface of the complexes (1–4) that the bacterial tyrosinase uses as its active pocket.

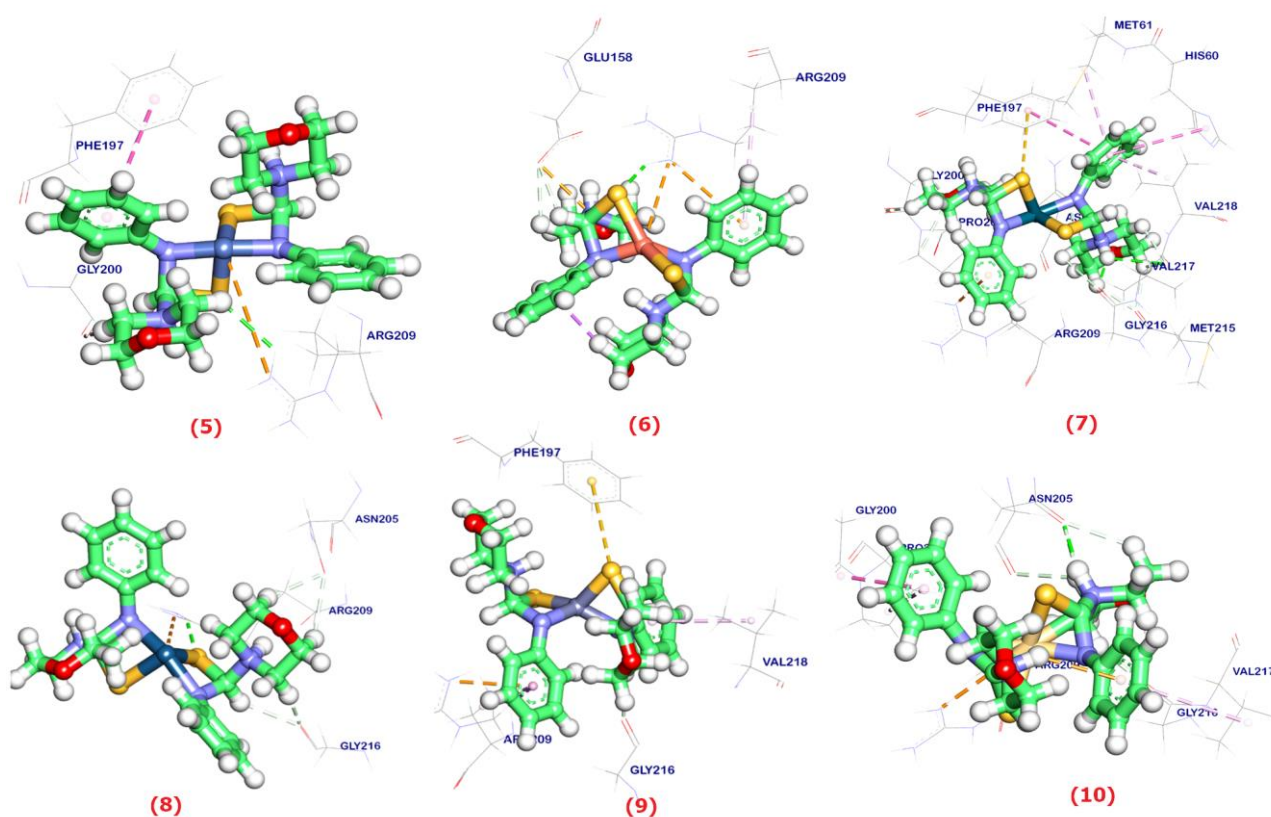
Similarly, Complexes (5–11) were docked against the bacterial tyrosinase enzyme and the energy binding values are −5.20, −5.10, −5.13, −5.56, −5.36, −4.02, and −5.20 Kcal/mol, for Complexes (5–11), respectively (Figure 9 and Table 6). Complex (5) interacted with Arg209, Pro201, and Phe197 by  $\pi$ -alkyl,  $\pi$ - $\pi$ , and attractive interactions, and one hydrogen bond with Gly216 and Arg209, respectively, whereas Complex (6) showed  $\pi$ -cation interactions and hydrogen bonding with Glu158 and Arg209. Also, Complex (7) exhibited two hydrogen bonds with Gly216 and Val217 and  $\pi$ - $\pi$ ,  $\pi$ -cation with His60, Val218, Met61, Phe197, and Arg209, respectively. Complex (8) enacted attractive and hydrogen bonding with Arg209, while Complex (9) bound with Phe197, Arg, and Val218; moreover, Complexes (10) and (11) formed a hydrogen bond with Asn205 or Met215 and Asn205 and  $\pi$ -alkyl,  $\pi$ - $\pi$ , and attractive interactions with Val217, Pro201, Gly200, Arg209, Phe197, Val218, Ala221, His208, and Met61. Moreover, the mapping surface studied the interaction of the synthesized complexes (5–11) with the active pocket of the bacterial tyrosinase enzyme, and Figure 10 shows this interference through the color gradient of the interpolated charge and ionizability during the interference of the prepared complex with the active pocket.

**Table 6.**  $\Delta G$  (kcal/mol) of synthesized compounds against the *bacterial tyrosinase* target site PDB ID: 6EI4.

Ligand	RMSD Value (Å)	Docking Score (kcal/mol)	Interaction	
			H.B	$\pi$
HPMCT	1.35	−7.01	2	5
1	1.97	−6.11	-	-
2	1.69	−6.15	2	5
3	2.10	−6.32	1	5
4	2.40	−5.86	1	2
5	1.23	−5.20	1	1
6	1.26	−5.10	1	1
7	1.13	−5.13	2	5

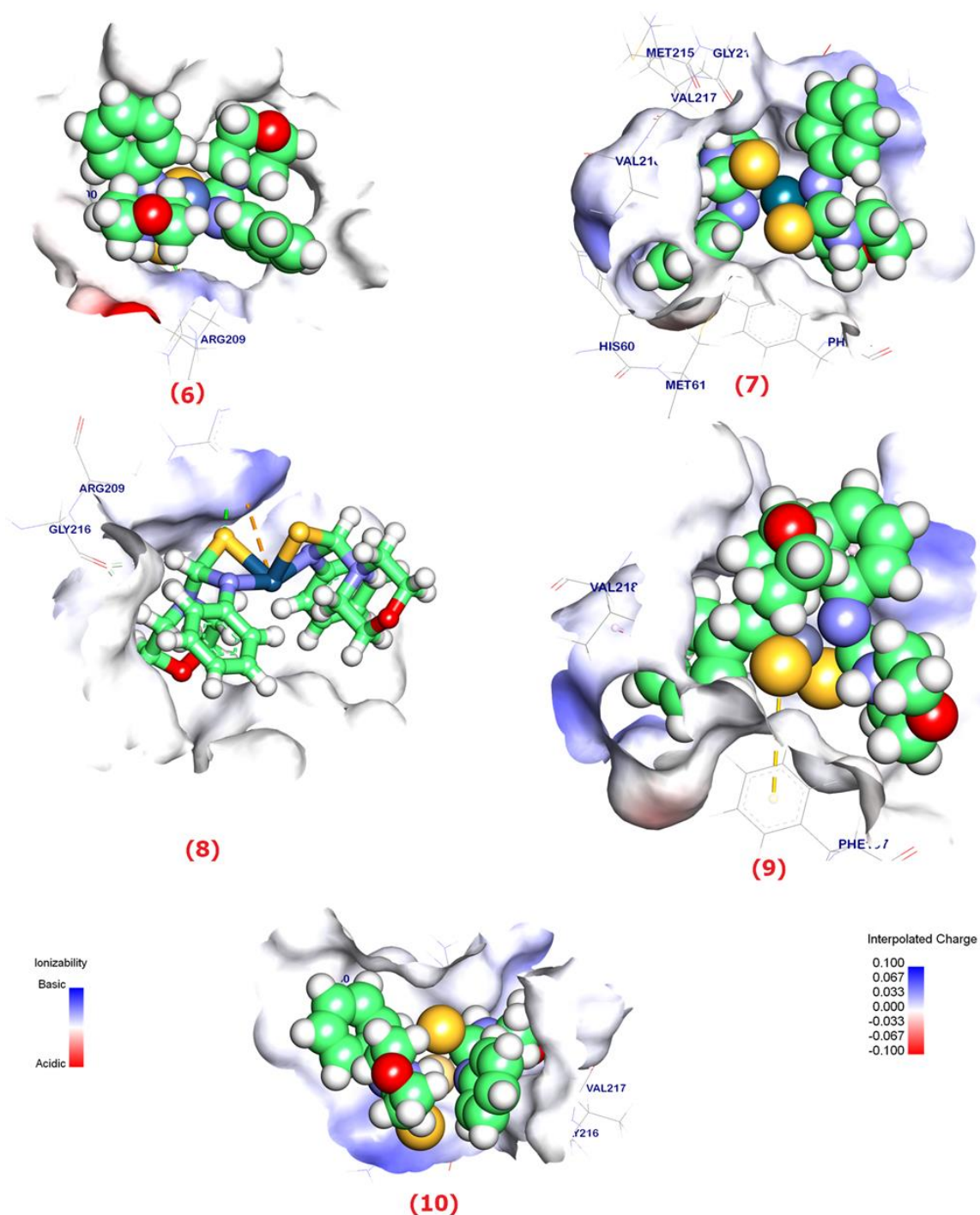


8	1.75	−5.56	1	−
9	1.72	−5.36	−	3
10	2.03	−4.02	1	4
11	1.32	−5.20	2	6



**Figure 9.** Complexes (5–10) docked in the bacterial tyrosinase are shown, together with hydrogen bonds (green) and  $\pi$ -interactions (purple lines).

According to the molecular docking of HPMCT, the reference ligand demonstrated a fitted binding to bacterial tyrosinase with a docking score of  $-7.01$  kcal/mol, Compounds (1–4) exhibited binding energies ranging from  $-5.86$  to  $-6.32$  kcal/mol which bound with the essential amino acids in the target pocket with a similar mode of the reference ligand; additionally, Compounds (5–11) had binding energies ranging from  $-4.02$  to  $-5.56$  kcal/mol when docked with bacterial tyrosinase. They exhibited diverse binding modes, involving interactions like  $\pi$ -alkyl,  $\pi$ - $\pi$ ,  $\pi$ -cation, and hydrogen bonding with various key residues in the enzyme's active site.

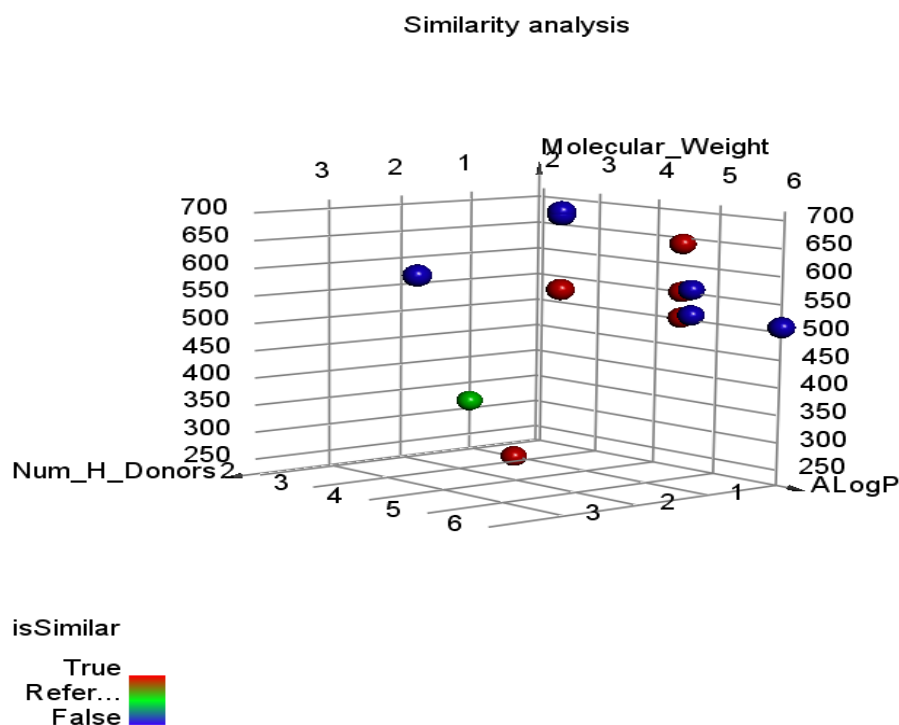


**Figure 10.** Mapping of the complexes (6–10) that reside on the surface of the bacterial tyrosinase's active pocket.

### Molecular Similarity

Drug design has always been a focus area of interest that has motivated a continued effort for developing new strategies toward finding similar molecules. The relationship between a certain chemical structure and its biological activity has always been related. These studies have been improved by including the idea of similar chemical structures when creating new drug-like molecules. Studies on molecular similarity have considered a wide range of descriptors. The molecular descriptors that are being studied include partition coefficient, hydrogen bond acceptors, hydrogen bond donors, and molecular weight (M.W.) (ALog p). In addition, numerous parameters have also been investigated such as

the number of rotatable bonds [25], number of rings, and aromatic rings [26], as well as the molecular fractional polar surface area (MFP SA) [27] (Table 7). A binary array (the total number of binary bits) is used to represent the compound under examination. The descriptions were created using the software Discovery Studio for the candidates that had undergone testing, and they were then contrasted with the reference ligand. Figure 11 shows reference ligands (green balls), compounds with reduced similarities (blue balls), and compounds with good similarities as red balls. The level of molecular similarity or likeness between two substances is assessed using a similarity coefficient that generates a quantitative score. That resulting score, which is equal to the degree of similarity, was developed based on the computed values of a variety of structural descriptors. How similar two molecules are to one another is inversely associated with the computed distance between them in the descriptor space [28]. To compare the descriptor similarity between the tested compounds and the reference ligands, the distances between the various descriptors were computed in this study. The calculated distances describe how far apart two places are from one another. The overall degree to which the tested candidates' actions reflect those of the reference ligands is shown by typed graph distances (Figure 10). Five of the examined options were picked as the favorites by the study.



**Figure 11.** Comparative comparison of the compounds that were tested and colchicine. Reference ligands are represented by green balls, comparable complexes by red balls, and dissimilar ligands by blue balls.

The comparative analysis of molecular properties and structural similarity revealed distinct profiles among the compounds. Several compounds exhibited structural similarity to the reference compound such as Compounds (7, 1, 5, 8 and 11) which showed a high similarity with the reference ligand based on the calculated minimum distances and molecular fingerprint-based similarity (MFP SA) values.



**Table 7.** ALog p, Mwt, HBA, HBD, rotatable bonds, MFPSA, minimal separation, and resemblance of the produced complexes to reference compound.

Comp.	A-Log p	Mwt.	HBA	HBD	Rotatable Bonds	Rings	Aromatic Rings	MFPSA	Minimum Distance	Is Similar
7	4.407	549.017	6	0	4	6	2	0.185	1.81736	TRUE
1	4.691	579.065	6	2	8	4	2	0.147	1.82746	TRUE
5	4.407	501.291	6	0	4	6	2	0.197	1.82961	TRUE
8	4.407	637.675	6	0	4	6	2	0.184	1.89913	TRUE
11	2.735	222.307	2	1	3	2	1	0.243	1.90748	TRUE
Reference	1.915	313.389	1	1	3	3	2	0.074	-	Reference
6	6.004	508.159	8	0	4	6	2	0.186	2.27628	FALSE
10	4.567	555.008	8	0	4	6	2	0.228	2.18775	FALSE
9	4.567	508.006	8	0	4	6	2	0.232	2.17541	FALSE
2	4.807	625.971	4	4	8	4	2	0.168	2.01066	FALSE
4	4.691	716.109	6	2	8	4	2	0.144	1.9746	FALSE
3	4.691	710.597	6	2	8	4	2	0.146	1.97191	FALSE

### 2.7. ADMET Studies

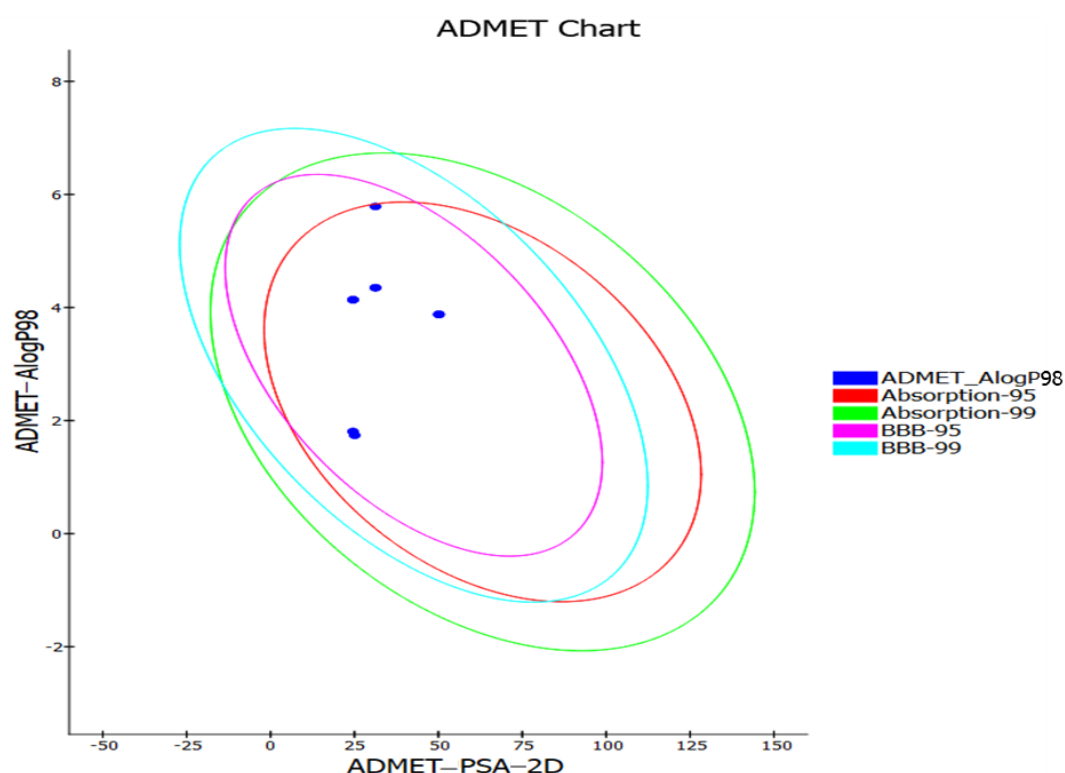
Pharmacokinetics is one of the important indications about how drugs interact with the human body; from the previous study, analysis provides valuable insights into the pharmacological properties and predictive factors of the tested compounds. Differences in BBB permeability, solubility, absorption, hepatotoxicity, and predictive factors like CYP2D6 and PPB predictions are indicative of variations in how these compounds may interact with biological systems and exhibit pharmacological effects.

HPMCT was used as a reference drug in ADMET studies for the synthesized molecules using Discovery Studio 2019 software and the summarized results are shown in Figure 12 and listed in Table 8.

**Table 8.** Predicted ADMET for the synthesized compounds.

Comp.	BBB Level <sup>a</sup>	Solubility Level <sup>b</sup>	Absorption Level <sup>c</sup>	Hepatotoxicity	CYP2D6 Prediction <sup>d</sup>	PPB Prediction <sup>e</sup>
10	1	1	0	False	false	true
6	0	1	0	False	false	true
5	0	2	0	False	false	true
7	0	2	0	False	false	true
8	0	2	0	False	false	true
9	1	1	0	False	false	true
1	1	2	0	False	false	true
4	1	2	0	False	false	true
2	1	3	0	False	true	false
3	1	2	0	False	false	true
11	2	3	0	False	false	false

<sup>a</sup> BBB level, blood brain barrier level, 0 = very high, 1 = high, 2 = medium, 3 = low, 4 = very low. <sup>b</sup> Solubility level, 1 = very low, 2 = low, 3 = good, 4 = optimal. <sup>c</sup> Absorption level, 0 = good, 1 = moderate, 2 = poor, 3 = very poor. <sup>d</sup> CYP2D6, cytochrome P2D6, TRUE = inhibitor, FALSE = non-inhibitor. The classification of whether a compound is a CYP2D6 inhibitor using the cutoff Bayesian score of 0.161. <sup>e</sup> PPB, plasma protein binding, FALSE means less than 90%, TRUE means more than 90%. The classification of whether a compound is highly bounded ( $\geq 90\%$  bound) to plasma proteins using the cutoff Bayesian score of  $-2.209$ .



**Figure 12.** ADMET image for the synthesized compounds.

The toxicity data for a set of compounds (Comp. 1–11) reveal significant variations in the toxicological properties and carcinogenic potential of the tested compounds. Compounds (1, 4, 3, and 11) were classified as single carcinogens, indicating their potential to induce cancer. The remaining compounds (Comp. 2, 5, 6, 7, 8, 9, and 10) were categorized as non-carcinogens, suggesting a lower risk of carcinogenicity. There were variations in LD50 values among the compounds, with some compounds having a lower toxicity than others. Additionally, all compounds were classified as non-toxic in terms of developmental toxicity potential, suggesting they were not able to cause harm during development.

### 2.8. Toxic Effective

Toxic effective calculations of the synthesized compounds were examined by Discovery Studio 2019 software and listed in Table 9, which were founded on validated and assembled models as follows: FDA rat carcinogenicity [29,30], carcinogenic potency TD50 [31], rat maximum tolerated dose (MTD) [32,33], rat oral LD50 [34], rat chronic LOAEL [35,36], ocular irritancy [37], and skin [38].

**Table 9.** In silico toxicity properties of the synthesized compound.

Comp.	FDA Rodent Carcinogenicity (Male, Mouse)	Carcinogenic Potency TD50 (Rat) <sup>a</sup>	Rat Maximum Tolerated Dose (Feed) <sup>b</sup>	Developmental Toxicity Potential	Rat Oral LD50 <sup>b</sup>	Rat Chronic LOAEL <sup>b</sup>	Ocular Irritancy (Rat)	Skin Irritancy (Rat)
10	Non-Carcinogen	0.599	0.0075	Non-Toxic	0.0735	0.0156	None	Mild
6	Non-Carcinogen	0.655	0.0185	Non-Toxic	0.0698	0.0099	Moderate	Mild
5	Non-Carcinogen	4.134	0.0127	Non-Toxic	0.0935	0.0298	None	Mild
7	Non-Carcinogen	3.013	0.0068	Non-Toxic	0.1071	0.0192	Moderate	Mild
8	Non-Carcinogen	3.136	0.0027	Non-Toxic	0.0998	0.0201	Moderate	Mild
9	Non-Carcinogen	0.581	0.0120	Non-Toxic	0.0756	0.0151	None	Mild
1	Single-Carcinogen	5.481	0.0225	Non-Toxic	0.5264	0.0075	Mild	None

4	Single-Carcinogen	5.721	0.0054	Non-Toxic	0.4634	0.0079	Mild	None
2	Non-Carcinogen	11.41	0.0306	Non-Toxic	1.1776	0.0050	Severe	Mild
3	Single-Carcinogen	5.716	0.0057	Non-Toxic	0.4662	0.0079	Mild	None
11	Single-Carcinogen	231.5	0.0666	Non-Toxic	0.3213	0.0249	Mild	None

<sup>a</sup> Unit: mg/kg body weight/day. <sup>b</sup> Unit: g/kg body weight.

## 2.9. DFT Data Analysis

### 2.9.1. Geometry Optimization of HPMCT

The mother ligand N-Phenylmorpholine-4-carbothioamide (HPMCT) was optimized and its structure was determined (Figure 13). It was found that the total energy of the HPMCT is (−634,215.3996 Kcal/mol). Moreover, the energies of the HOMO and LUMO were found to be (−5.793 eV, −1.232 eV), respectively, while the energy gap difference was (4.561 eV) (Figure 14). Additionally, the two structures' predicted electron affinity, ionization energy, dipole moment, and hardness values were 1.232, 5.793, 3.817 Debye, and 2.2805, respectively. Additionally, the bond lengths and angles for the HPMCT were assessed and contrasted with the experimental values. In the case of the (C=S) group, the bond distance is (1.67318 Å), which is extremely near to the experimental result (1.690 Å) [39]. The following bond distances were shown: (N7-C8 = 1.37945 Å, C10-N8 = 1.38048 Å, N7-C3 = 1.41745 Å, N9-C11 = 1.46629 Å). On the other hand, the angles of the following bonds were as shown: (S23-C8-N7 = 123.613, S23-C8-N9 = 123.463, N7-C8-N9 = 112.888 and C8-N9-C10 = 120.221).

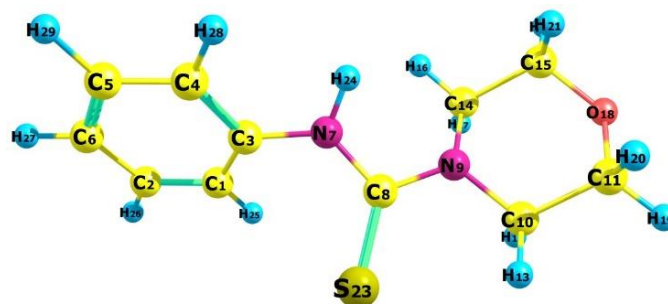


Figure 13. The optimized structure of HPMCT.

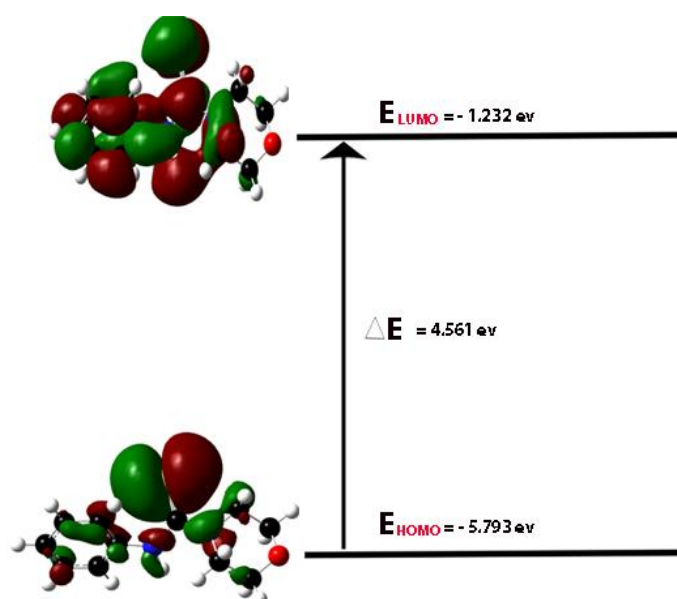
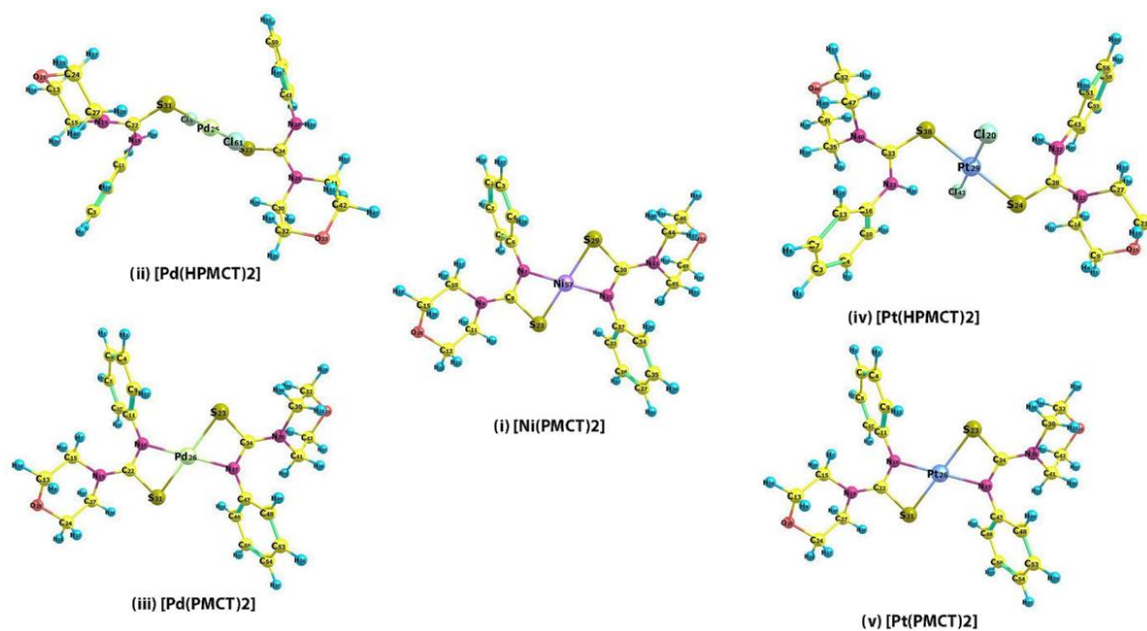


Figure 14. The energies of HOMO – LUMO of HPMCT.

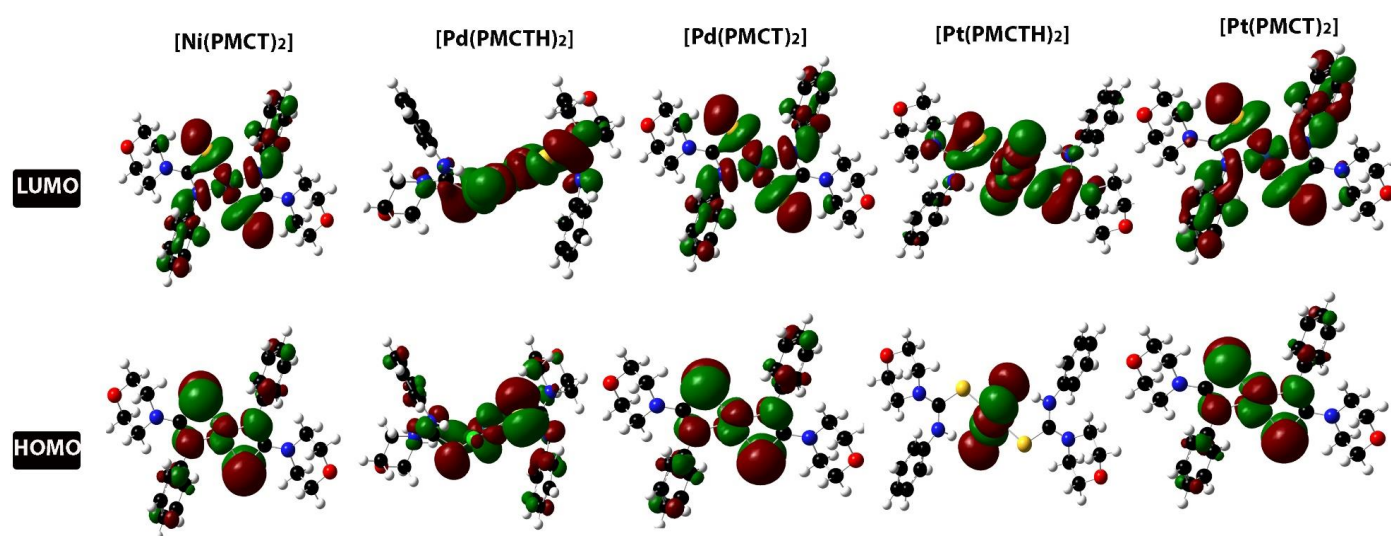
In order to study the geometry changes, the structures of the synthesized complexes (5–9) were optimized (Figures 15 and 16). Following complexation, each compound's bond length, bond angles, and overall energy were examined (Table 10), with particular attention paid to elements like sulfur and nitrogen that are coordinated to the core metals. Two distinct traits can be seen in the resulting complexes. In the absence of a base, for instance, the geometric characteristics of the reaction between the ligand HPMCT and Pd(II) or Pt(II) demonstrate a mono coordination behavior of the ligand. In a basic media, however, the ligand forms a bidentate manner of attachment to the core metal. The center metal binds to the ligand via the sulfur atom in a neutral medium and to the sulfur and nitrogen atoms in a basic medium.

**Table 10.** Global reactivity descriptors of M-HPMCT or M-PMCT complexes.

Parameters	[Ni(PMCT) <sub>2</sub> ]	[Pd(HPMCT) <sub>2</sub> ]	[Pd(PMCT) <sub>2</sub> ]	[Pt(HPMCT) <sub>2</sub> ]	[Pt(PMCT) <sub>2</sub> ]
E <sub>total</sub> (Kcal/mol)	−1,374,759.30	−1,926,136.774	−1,347,751.571	−1,920,786.055	−1,342,397.088
E <sub>HOMO</sub> , ev	−5.365	−5.829	−5.413	−5.805	−5.128
E <sub>LUMO</sub> , ev	−1.497	−2.419	−1.815	−1.807	−1.211
E <sub>g</sub> , ev (E <sub>LUMO</sub> −E <sub>HOMO</sub> )	3.868	3.41	3.598	3.998	3.917
Hardness ( $\frac{1}{2}(E_{LUMO}-E_{HOMO})$ )	1.934	1.705	1.799	1.999	1.9585
Dipole Moment (debye)	0.000735	2.291495	0.003571	0.004176	0.001631
LE (−E <sub>HOMO</sub> ev)	5.365	5.829	5.413	5.805	5.128
EA (−E <sub>LUMO</sub> ev)	1.497	2.419	1.815	1.807	1.211



**Figure 15.** Optimized geometry of; (i) [Ni(PMCT)<sub>2</sub>], (ii) [PdCl<sub>2</sub>(HPMCT)<sub>2</sub>], (iii) [Pd(PMCT)<sub>2</sub>], (iv) [Pt(HPMCT)<sub>2</sub>], and (v) [Pt(PMCT)<sub>2</sub>].



**Figure 16.** An illustration of the molecules' highest occupied and lowest vacant molecular orbitals of  $[\text{Ni}(\text{PMCT})_2]$ ,  $[\text{PdCl}_2(\text{HPMCT})_2]$ ,  $[\text{Pd}(\text{PMCT})_2]$ ,  $[\text{Pt}(\text{HPMCT})_2]$ , and  $[\text{Pt}(\text{PMCT})_2]$ .

For Pd and Pt complexes in monodentate mode, the calculated bond length of C34-S23 is 1.738 Å and 1.727 Å, respectively, while for Pd and Pt complexes in bidentate mode, the calculated bond distance for C8-S23 is 1.769 Å. These numbers show that the CS group has a single bond character for the latter complex and a double bond character for the first complex. Further evidence for a resonance structure between the NCS moiety is provided by the bond distances of C8-N7, which are 1.333 Å for Pd and Pt, respectively, and 1.335 Å for Pd (Table 11).

**Table 11.** Selected bond lengths and bond angles of M-HPMCT or M-PMCT complexes.

Bond Length	$[\text{PdCl}_2(\text{HPMCT})_2]$	$[\text{PtCl}_2(\text{HPMCT})_2]$	Bond Length	$[\text{Ni}(\text{PMCT})_2]$	$[\text{Pd}(\text{PMCT})_2]$	$[\text{Pt}(\text{PMCT})_2]$
S23-M	2.402	2.377	S23-M	2.277	2.399	2.405
-----	-----	-----	N7-M	1.913	2.071	2.079
-----	-----	-----	N31-M	1.913	2.399	2.079
S31-M	2.379	2.377	S29-M	2.277	2.071	2.405
C34-N37	1.365	1.360	C8-N7	1.334	1.333	1.335
N38-C34	1.354	1.366	N9-C8	1.365	1.366	1.365
C34-S23	1.738	1.727	C8-S23	1.766	1.769	1.769
C22-S31	1.723	1.727	C30-S29	1.766	1.769	1.769
Bond Angles	$[\text{PdCl}_2(\text{HPMCT})_2]$	$[\text{PtCl}_2(\text{HPMCT})_2]$	Bond Length	$[\text{Ni}(\text{PMCT})_2]$	$[\text{Pd}(\text{PMCT})_2]$	$[\text{Pt}(\text{PMCT})_2]$
S23-M-S31	178.966	179.994	S23-M-S29	179.994	179.986	179.995
Cl60-M-Cl61	177.999	179.996	N7-M-N31	179.995	179.993	179.991
S23-M-Cl60	85.741	87.216	S23-M-N7	73.928	69.254	68.497
S23-M-Cl61	93.792	92.781	N31-M-S29	73.930	69.248	68.499
S31-M-Cl60	93.405	92.790	S23-M-N31	106.076	110.739	111.50
S31-M-Cl61	87.043	87.213	N7-M-S29	106.066	110.759	111.505

### 2.9.2. Vibration Analysis

Calculated and experimental data were compared for the ligand and its complexes' vibrational frequencies. In contrast to the experimental results, which showed the N-H stretching band to be present at 3176  $\text{cm}^{-1}$ , the ligand exhibits the unique N-H stretching band at 3140  $\text{cm}^{-1}$ . The four unique thioamide absorption bands also appeared at the following wavelengths: 1570  $\text{cm}^{-1}$ , 1350  $\text{cm}^{-1}$ , 1210  $\text{cm}^{-1}$ , and 550  $\text{cm}^{-1}$  (supplementary data) [1,13,40]. The four thioamide absorption bands were discovered at 1530  $\text{cm}^{-1}$ , 1462  $\text{cm}^{-1}$ ,

1211  $\text{cm}^{-1}$ , and 601  $\text{cm}^{-1}$  in the experimental results. The neutral ligand is anticipated to maintain its absorption after complexation with minimal changes. For instance, the complex  $[\text{MCl}_2(\text{HPMCT})_2]$  estimated FT-IR absorptions showed a strong peak at 3260  $\text{cm}^{-1}$ , which is in the N-H absorption band. At 1591  $\text{cm}^{-1}$ , 1450  $\text{cm}^{-1}$ , and 1210  $\text{cm}^{-1}$ , respectively, the thioamide absorptions were visible. On the other hand, the ligand coordinates to the central metal in a bidentate mode when it is deprotonated through its anionic sulfur and nitrogen atoms. According to theoretical vibrational frequencies, the  $[\text{M}(\text{PMCT})_2]$  complex does not absorb N-H in the range of 3000–3500  $\text{cm}^{-1}$ . Additionally, the (C=S) absorption band, which was seen between 1200 and 1300  $\text{cm}^{-1}$ , disappeared. The fact that the quantum calculations were carried out in the gas phase, excluding any input from the surrounding molecules and the solvent molecules, is the cause of the discrepancies between the experimental and theoretical values.

### 3. Experimental Part

#### 3.1. Instrumental and Materials

All the reagents bought were of correct commercial fulfillment and have been utilized without purification. The ligand N-Phenylmorpholine-4-carbothioamide (HPMCT) was prepared following a modified protocol from the literature [19]. All chemical materials were obtained from Merck (USA), Alfa Aesar (China), and BDH (India) companies. Nuclear Magnetic resonance (NMR), Fourier-transform infrared spectroscopy (FT-IR), Mass, Ultra violet–Visible (UV–vib.), XRD, and Scanning electronic microscope (SEM) techniques were used to characterize the synthesized complexes. The FT-IR spectra (KBr pellets; 400–4000  $\text{cm}^{-1}$ ) were recorded on a Shimadzu FT-IR 8400 spectrophotometer, NMR spectra on a Bruker Advance II 400 MHz spectrometer, Tehran - Iran, and mass spectra on a Xevo G2-XS QT of Mass Spectrometer, Tehran - Iran. UV–Visible spectra were recorded on the Shimadzu UV–Vis spectrophotometer, Japan within the range of 200–1000 nm in dimethyl sulfoxide (DMSO).

#### 3.2. Preparation of N-Phenylmorpholine-4-Carbothioamide (HPMCT)

To a stirred solution of morpholine (1.742 g, 20 mmol) in 40 mL ethanol, phenylisothiocyanate was added (2.700 g, 20 mmol). The reaction mixture was refluxed for 4 h, which resulted in a white solid. Cooled ethanol was used to wash the obtained solid after filtering, and then it was dried in a vacuum oven. White crystals (3.981 g, 90% yield) were obtained after recrystallization from ethanol.

#### 3.3. Synthesis of Complexes

##### 3.3.1. General Procedure for the Synthesis of $[\text{MCl}_2(\kappa^1\text{S-HPMCT})_2]$ (1–4)

An ethanolic solution (20 mL) of (HPMCT) (0.500 g, 2.200 mmol) was added to a solution of metal salts ( $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ ,  $\text{Na}_2\text{PdCl}_4$ ,  $\text{K}_2\text{PdCl}_4$ , and  $\text{HgCl}_2$ ) (1.100 mmol) in ethanol (15 mL) with stirring. The precipitate obtained from the filtrate of the stirred mixture was dried for 3 h, in a vacuum after washing with distilled  $\text{H}_2\text{O}$ /EtOH.

##### 3.3.2. Synthesis of $[\text{Ni}(\text{PMCT})_2]$ (5–11)

An ethanolic solution (10 mL) of the HPMCT ligand (0.050 g, 0.210 mmol), containing some drops of triethyl amine, was added to an aqueous solution (10 mL) of  $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$  (0.050 g, 0.210 mmol) with stirring. A light olive precipitate was formed. The mixture was stirred for 3 h, then filtered off, washed with ethanol and distilled water, and dried in a vacuum (yield: 0.05 g, 80%, m.p  $^\circ\text{C}$  = 175–178). The following complexes were identified:  $[\text{Cu}(\kappa^2\text{S,N-mpt})_2]$  (6),  $[\text{Pd}(\kappa^2\text{S,N-PMCT})_2]$  (7),  $[\text{Pt}(\kappa^2\text{S,N-PMCT})_2]$  (8),  $[\text{Zn}(\kappa^2\text{S,N-PMCT})_2]$  (9),  $[\text{Cd}(\kappa^2\text{S,N-PMCT})_2]$  (10), and  $[\text{Hg}(\kappa^2\text{S,N-PMCT})_2]$  (11).



### 3.4. Antibacterial Studies

The antibacterial activities of the synthesized compounds were tested against three bacteria species, *Escherichia coli*, *Staphylococcus aureus*, and *Klebsiella pneumoniae*. At a  $10^{-3}$  M concentration of the prepared complexes, these three bacteria species were used to study the antibacterial activity of the synthesized complexes. The result was in contrast with tetracycline as an effective antibiotic drug and DMSO as a negative control. The activity effect of the complexes used was examined using a technique called agar disc diffusion, which was first described by Bauer [41].

### 3.5. Anti-Cancer Studies

The researchers tested the anti-tumor activity of free ligands (HPMCT) and their complexes (**2,3**) toward breast cell lines (MCF-7) using MTT (3-[4,5-dimethylthiazol-2-yl]-2,5-diphenyl-2H-tetrazolium bromide) as described by Tan et al. [24]. The results of the anti-cancer study of the free ligand and its complexes were compared with *cis-platin*.

Briefly,  $0.8 \times 10^4$  cells for the (MCF-7) cell line were seeded in different 96-well culture plates (Eppendorf Products Pvt. Ltd.) at the density of  $1 \times 10^4$  viable cells/well with Dulbecco's Modified Eagle Medium (DMEM) and incubated overnight to allow cell adherence in the CO<sub>2</sub> incubator. After 24 h of incubation, the spent medium was discarded, and the cells were incubated with different concentrations of each sample (0.01 to 100  $\mu$ M) for 24 h. After 24 h, the drug-containing medium was discarded followed by the addition of 90  $\mu$ L medium and 10  $\mu$ L of 5 mg/mL MTT in Phosphate-buffer solution (PBS), pH 7.4, into each well of the plate followed by another 2 h of incubation at 37 °C. Then, the MTT dissolved in the medium was removed and 100  $\mu$ L of DMSO was added to dissolve the MTT formosan crystals. Finally, the absorbance of samples was taken at 570 nm by a microplate reader and cytotoxicity was measured by using the percentage inhibition.

### 3.6. Molecular Docking

#### 3.6.1. Preparation of Targeted Proteins

The site where bacteria tyrosinase attaches was produced using synthetic complexes that are in the shape of the protein [PDB codes: 6EI4 (<https://www.rcsb.org>)]. This was achieved with the help of software called MOE 19.0901. Initially, water molecules were deleted from the structure. Then, crystallographic disorders and empty valence atoms were corrected using protein reports and utility and smooth protein options. Protein energy was once minimized by making use of MMFF94 force fields. The inflexible shape of the binding site is created by making use of a fixed atom constraint. The protein essential amino acids are described and organized for the docking performance [42].

#### 3.6.2. Preparation of Tested Ligands and Receptor

The 2D structures of the compounds have been drawn by using Chem-Bio Draw Ultra17.0 as an MDL-SD format; this file was opened with the aid of MOE 19.0901 software [43]. The 3D structure used was protonated, and the partial charge was fixed, and the energy used was minimized by making use of a 0.1 Root mean-square-deviation RMSD kcal/mol MMFF94 force field. Next, the minimized structures were prepared for docking using a ligand preparation protocol.

### 3.7. DFT Studies

All computations were carried out using the Gaussian 09 W software package [44]. The ligand and its complexes were successfully optimized by utilizing the B3LYP method [45–47], utilizing the 6-311++G(d,p) [48,49] basis set for the ligand while the SDD basis set was implemented for Pd and Pt atoms [50]. No symmetry constraints were applied during the geometry optimization [51]. The vibrational frequencies of each molecule have been determined at the same level of theory. The obtained structures were assessed in the Chemcraft v1.6 package [52] and GaussView 5.0.9 version [53].



#### 4. Conclusions

We describe the synthesis, characterization, and in vitro antibacterial and cytotoxic activity of new 1-morpholyl-3-phenyl thiourea ligands and associated metal complexes in this paper. The thiourea ligand coordinates as monodentate through a sulfur atom in Complexes (1–4), or as bidentate chelating in Complexes (5–11) where it coordinates to the central ions via N and S atoms. In comparison to the free ligand, all of the investigated compounds exhibit a good antibacterial effectiveness. Compounds (2) and (3) are observed as having a good cytotoxic activity against MCF-7 cell lines with appropriate  $IC_{50}$  values, respectively. The bioactivities of the produced compounds range from modest to good; however, by adding active groups to the ligand structure, new and more potent thiourea derivatives can be created. The prepared compounds exhibited diverse binding modes, involving interactions like  $\pi$ -alkyl,  $\pi$ - $\pi$ ,  $\pi$ -cation, and hydrogen bonding with key amino acid residues within the enzyme's active site. The compounds also showed a high similarity with the reference ligand based on the calculated minimum distances and MFPSPA values. Moreover, DFT calculations were also performed for Complexes (5–9) showing bond distance and bond angle changes upon complexation between the metal and the ligand. As expected, the five complexes adopt a square planar geometry with the metal being centered. The calculated vibrational frequencies of the five complexes showed similar absorption bands to those of the reported results.

**Supplementary Materials:** The following supporting information can be downloaded at: <https://www.mdpi.com/article/10.3390/inorganics11100390/s1>, Figure S1: IR spectrum of complex HPMCT ligand, Figure S1A: Theoretical IR spectrum of complex HPMCT ligand, Figure S2 : IR spectrum of complex 1, Figure S3 : IR spectrum of complex 2, Figure S3A: Theoretical: IR spectrum of complex 2, Figure S4 : IR spectrum of complex 3, Figure S4A: Theoretical: IR spectrum of complex 3, Figure S5 : IR spectrum of complex 4, Figure S6 : IR spectrum of complex 5, Figure S6A: Theoretical: IR spectrum of complex 5, Figure S7 : IR spectrum of complex 6, Figure S8 : IR spectrum of complex 7, Figure S7A: Theoretical: IR spectrum of complex 7, Figure S9 : IR spectrum of complex 8, Figure S9A: Theoretical: IR spectrum of complex 8, Figure S9 : IR spectrum of complex 9, Figure S10 : IR spectrum of complex 10, Figure S11 : IR spectrum of complex 11, Figure S12:  $^1\text{H}$  NMR spectrum of HPMCT, Figure S13:  $^1\text{H}$  NMR spectrum of complex 2, Figure S14:  $^1\text{H}$  NMR spectrum of complex 3, Figure S15:  $^1\text{H}$  NMR spectrum of complex 4, Figure S16 :  $^1\text{H}$  NMR spectrum of complex 7, Figure S17 :  $^1\text{H}$  NMR spectrum of complex 8, Figure S18 :  $^1\text{H}$  NMR spectrum of complex 10, Figure S19 :  $^1\text{H}$  NMR spectrum of complex 11, Figure S20:  $^{13}\text{C}$  NMR spectrum of HPMCT, Figure S21 :  $^{13}\text{C}$  NMR spectrum of complex 4, Figure S22 :  $^{13}\text{C}$  NMR spectrum of complex 11.

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