

## Article

# PPO-Inhibiting Herbicides and Structurally Relevant Schiff Bases: Evaluation of Inhibitory Activities against Human Protoporphyrinogen Oxidase

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**Abstract:** The study of human protoporphyrinogen oxidase (hPPO) inhibition can contribute significantly to a better understanding of some pathogeneses (e.g., porphyria, herbicide exposure) and the development of anticancer agents. Therefore, we prepared new potential inhibitors with Schiff base structural motifs (2-hydroxybenzaldehyde-based Schiff bases **9–13** and chromanone derivatives **17–19**) as structurally relevant to PPO herbicides. The inhibitory activities (represented by the half maximal inhibitory concentration (IC<sub>50</sub>) values) and enzymatic interactions (represented by the hPPO melting temperatures) of these synthetic compounds and commercial PPO herbicides used against hPPO were studied by a protoporphyrin IX fluorescence assay. In the case of PPO herbicides, significant hPPO inhibition and changes in melting temperature were observed for oxyfluorten, oxadiazon, lactofen, butafenacil, saflufenacil, oxadiargyl, chlornitrofen, and especially fomesafen. Nevertheless, the prepared compounds did not display significant inhibitory activity or changes in the hPPO melting temperature. However, a designed model of hPPO inhibitors based on the determined IC<sub>50</sub> values and a docking study (by using AutoDock) found important parts of the herbicide structural motif for hPPO inhibition. This model could be used to better predict PPO herbicidal toxicity and improve the design of synthetic inhibitors.

**Keywords:** protoporphyrinogen oxidase; inhibitors; herbicides

## 1. Introduction

Porphyrias are very rare diseases associated with abnormal haem production that lead to the accumulation of porphyrins or their precursors within the liver and other organs [1,2]. Variegate porphyria is caused by mutations to protoporphyrinogen oxidase (PPO, EC 1.3.3.4) and subsequent activity loss. This enzyme catalyzes the oxygen-dependent six-electron oxidation of protoporphyrinogen IX to protoporphyrin IX [3–5]. During acute attacks, patients display abnormal skin reactions to sunlight due to reactive singlet oxygen formation and oxidative stress following photooxidation, mainly in the liver [6,7]. Patients with variegate porphyria can have a significantly higher risk of liver cancer [6,8].

This phenomenon can also be caused by various toxic agents, such as some herbicides [9,10]. Using herbicides is one of the most commonly used methods to control weeds

and preserve crop quality and yield. During the last 30 years, numerous structurally diverse PPO-inhibiting herbicides have been developed and commercialized, such as diphenyl ethers and thiadiazoles [11]. In contrast to herbicides with other modes of action, the application of these classes of herbicides displays some important benefits, such as a low use rate, broad-spectrum herbicidal activity, a quick onset of action, long-lasting effects, and environmentally friendly side effects. Nevertheless, the display of serious toxic effects to animals and humans after exposure to these compounds, especially long-term exposure, cannot be excluded.

At present, some influencing work has implied that these herbicide types also display serious inhibitory effects on human protoporphyrinogen IX oxidase (hPPO) [12] due to porphyrin accumulation in hepatocytes [13]. Butafenacil has long been used as an anti-PPO herbicide. It is expected that some of its toxicity is caused by inhibition of animal PPOs [14,15]. For example, a published study by Leet et al. stated that exposure of zebrafish embryos to butafenacil led to porphyria symptoms, such as porphyrin accumulation [14]. After fomesafen application, higher levels of reactive oxygen species (ROS), weaker immune systems, and reduced numbers of white cells were observed [16]. In a mouse model, a specific dosage caused porphyria syndrome, as evidenced by liver porphyrin accumulation [17,18]. Similarly, oxyfluorfen and oxadiazon treatment have led to higher porphyrin accumulation in rat hepatocytes and mouse livers. These phenomena were associated with decreased PPO activity; however, the inhibition activity (represented by its  $IC_{50}$  values) for these herbicides have not yet been determined.

BaR-2 has displayed significantly higher inhibitory activity against PPO than commercially used herbicides [19]. Nevertheless, its inhibitory activity has not yet been determined for human or mammalian PPOs. The above strongly implies that detailed quantification of their inhibitory effects on PPOs (represented by  $IC_{50}$  values) could precisely predict the toxicity of these compounds in humans and animals.

In addition, obtaining this information could also be useful for the design of new pharmacophores targeting PPOs for the treatment of oncological diseases. Relevant work found that PPOs are overexpressed in some cancer types (e.g., colorectal cancer) [20]. Targeting PPO could be a promising way to enhance photodynamic activity based on the application of 5-aminolevulinic acid [21–23]. However, prospective strategies for the preparation of anticancer agents could combine additional structural motifs, such as Schiff bases, for their molecular design.

Schiff bases possess a wide range of biological activities, including the repression of oncological processes and signaling pathways [24–26]. Schiff bases bearing 2-hydroxyaryl moieties can bind arginine groups [27] and thereby possibly target Arg-98 or Arg-59, which are key amino acids that maintain PPO activity [28,29]. Combining this structural motif with polychlorinated phenyls (for a possible interaction with Met-368 of PPO) [29] could lead to the design of new types of hPPO inhibitors. These compounds are structurally relevant to known PPO herbicides. Thus, we prepared two sets of potential inhibitors: 2-hydroxybenzaldehyde-based Schiff bases and chromanone derivatives. Moreover, Schiff bases bearing 2-hydroxyaryl moieties can bind metal ions and thereby expand their anticancer effects [30–33].

In this work, we decided to study the influence of PPO herbicides and novel PPO inhibitors based on the Schiff base structural motif on hPPO activity. Both sets of compounds could represent new structural motifs for novel anticancer agents. In the case of PPO herbicides, the obtained results could also be used for better environmental risk prediction.

## 2. Materials and Methods

### 2.1. Measurements and Materials

All chemicals and solvents were purchased from commercial suppliers (Sigma-Aldrich, Santiago Chemicals and TCI Europe) and were used without further purification. Nuclear magnetic resonance (NMR) spectra were recorded on a 500 MHz instrument at room temperature ( $\sim 25$  °C) in dimethyl sulfoxide (DMSO)- $d_6$ . The chemical shifts ( $\delta$ ) are presented

in ppm and coupling constants ( $J$ ) are presented in Hz. The program MestReNova ver. 14.1 was used to process the NMR spectra. High-resolution mass spectrometry (HRMS) spectra were obtained using electrospray ionization (ESI) with a triple quadrupole mass spectrometer (TSQ Quantum Access) and an LTQ Orbitrap spectrometer.  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra of prepared compounds are shown in Supplementary Figures S1–S10.

### 2.2. Preparation of Ethyl 2-(2,4-Dichloro-5-nitrophenoxy)acetate (2)

2,4-Dichloro-5-nitrophenol (**1**; 1040 mg; 5 mmol) and potassium carbonate (1382 mg; 10 mmol) were mixed in anhydrous dimethylformamide (15 mL). The mixture was stirred at 25 °C for 15 min, and ethyl bromoacetate (0.9 mL; 8 mmol) was added. This reaction mixture was stirred at 60 °C overnight. Then, water (100 mL) was added, and the mixture was extracted with ethyl acetate ( $2 \times 100$  mL). The organic fraction was then washed with water ( $2 \times 75$  mL) and brine (75 mL), dried over sodium sulfate, and evaporated to dryness. The crude product was purified by column chromatography on silica, eluent: dichloromethane. Ethyl 2-(2,4-dichloro-5-nitrophenoxy)acetate (**2**) was obtained with a yield of 1326 mg (90%) as a white wax.

$^1\text{H}$  NMR: 1.21 (t,  $J = 7.0$  Hz, 3H); 4.18 (q,  $J = 7.0$  Hz, 2H); 5.06 (s, 2H); 7.96 (s, 1H); 8.03 (s, 1H).  $^{13}\text{C}$  NMR: 14.0, 61.0, 66.0, 110.7, 117.1, 126.6, 131.9, 146.6, 152.6, 167.6 ppm. HRMS (ACPI<sup>-</sup>) calcd for  $\text{C}_{10}\text{H}_9\text{Cl}_2\text{NO}_5$  [ $\text{M}$ ]<sup>-</sup>: 292.98578, found: 292.98651.

### 2.3. Preparation of Ethyl 2-(5-Amino-2,4-dichlorophenoxy)acetate (3)

Ethyl 2-(2,4-dichloro-5-nitrophenoxy)acetate (**2**; 1245 mg; 4.19 mmol) was dissolved in ethyl acetate (75 mL), and palladium on carbon (10%; 500 mg) was added. The reaction mixture was stirred under a hydrogen atmosphere (balloon) overnight. Then, the catalyst was filtered off and washed with ethyl acetate (20 mL). The solution was evaporated to dryness to give pure ethyl 2-(5-amino-2,4-dichlorophenoxy)acetate (**3**) with a yield of 1061 mg (96%) as a yellowish wax.

$^1\text{H}$  NMR: 1.22 (t,  $J = 7.1$  Hz, 3H); 4.18 (q,  $J = 7.1$  Hz, 2H); 4.73 (s, 2H); 5.39 (bs, 2H); 6.41 (s, 1H); 7.24 (s, 1H).  $^{13}\text{C}$  NMR: 14.0, 60.8, 65.2, 100.1, 107.8, 109.2, 129.1, 144.6, 152.5, 168.1 ppm. HRMS (ESI<sup>+</sup>) calcd for  $\text{C}_{10}\text{H}_{12}\text{Cl}_2\text{NO}_3$  [ $\text{M}^+\text{H}$ ]<sup>+</sup>: 264.01942, found: 264.01931.

### 2.4. Preparation of Schiff Base 9

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 2-hydroxy-3-methoxybenzaldehyde (**4**; 30 mg; 0.2 mmol) were dissolved in anhydrous ethanol (10 mL) and stirred at 75 °C overnight. Orange precipitate formed almost immediately. After cooling, the solid product was filtered off, washed with ethanol (30 mL) and dried under vacuum. Ethyl-(*E*)-2-(2,4-dichloro-5-((2-hydroxy-3-methoxybenzylidene)amino)phenoxy)acetate (**9**) was obtained with a yield of 63 mg (79%) as an orange solid.

$^1\text{H}$  NMR: 1.22 (t,  $J = 7.1$  Hz, 3H); 3.83 (s, 3H); 4.19 (q,  $J = 7.1$  Hz, 2H); 5.01 (s, 2H); 6.96 (t,  $J = 7.9$  Hz, 1H); 7.19 (dd,  $J = 8.1, 1.4$  Hz, 1H); 7.24 (dd,  $J = 7.9, 1.4$  Hz, 1H); 7.43 (s, 1H); 7.77 (s, 1H); 9.03 (s, 1H); 12.99 (s, 1H).  $^{13}\text{C}$  NMR: 14.0, 55.9, 60.8, 65.6, 105.4, 116.2, 118.9, 120.2, 120.6, 124.0, 130.1, 144.0, 148.0, 150.7, 152.9, 165.3, 167.9 ppm. HRMS (ESI<sup>+</sup>) calcd for  $\text{C}_{18}\text{H}_{18}\text{Cl}_2\text{NO}_5$  [ $\text{M}^+\text{H}$ ]<sup>+</sup>: 398.05620, found: 398.05652.

### 2.5. Preparation of Schiff Base 10

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 2-hydroxy-4-methylbenzaldehyde (**5**; 27 mg; 0.2 mmol) were dissolved in anhydrous ethanol (10 mL) and stirred at 75 °C overnight. A yellowish solid gradually precipitated out of solution. After cooling, the solid product was filtered off, washed with ethanol (30 mL) and dried under vacuum. Ethyl-(*E*)-2-(2,4-dichloro-5-((2-hydroxy-4-methylbenzylidene)amino)phenoxy)acetate (**10**) was obtained with a yield of 57 mg (75%) as a pale-yellow solid.

$^1\text{H}$  NMR: 1.22 (t,  $J = 7.1$  Hz, 3H); 2.39 (s, 3H); 4.19 (q,  $J = 7.1$  Hz, 2H); 5.02 (s, 2H); 6.85 (s, 1H); 6.85 (d,  $J = 8.0$  Hz, 1H); 7.39 (s, 1H); 7.52 (d,  $J = 8.0$  Hz, 1H); 7.74 (s, 1H); 8.95 (s, 1H); 12.93 (s, 1H).  $^{13}\text{C}$  NMR: 14.0, 21.5, 60.8, 65.6, 105.3, 116.7, 117.1, 119.9, 120.5, 120.6, 130.1,

132.7, 144.2, 145.0, 152.9, 160.6, 164.8, 167.9 ppm. HRMS (ESI<sup>+</sup>) calcd for C<sub>18</sub>H<sub>18</sub>C<sub>12</sub>NO<sub>4</sub> [M<sup>+</sup>H]<sup>+</sup>: 382.06129, found: 382.06137.

### 2.6. Preparation of Schiff Base 11

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 5-chloro-2-hydroxybenzaldehyde (**6**; 28 mg; 0.2 mmol) were dissolved in anhydrous ethanol (10 mL) and stirred at 75 °C overnight. A yellowish solid gradually precipitated out of solution. After cooling, the solid product was filtered off, washed with ethanol (30 mL) and dried under vacuum. Ethyl-(*E*)-2-(2,4-dichloro-5-((5-chloro-2-hydroxybenzylidene)amino)phenoxy)acetate (**11**) was obtained with a yield of 65 mg (79%) as a pale-yellow solid.

<sup>1</sup>H NMR: 1.22 (t, *J* = 7.1 Hz, 3H); 4.19 (q, *J* = 7.1 Hz, 2H); 5.01 (s, 2H); 7.04 (d, *J* = 8.8 Hz, 1H); 7.37 (s, 1H); 7.50 (dd, *J* = 8.8, 2.7 Hz, 1H); 7.72 (d, *J* = 2.7 Hz, 1H); 7.76 (s, 1H); 8.95 (s, 1H); 12.78 (s, 1H). <sup>13</sup>C NMR: 14.0, 60.9, 65.6, 105.4, 118.9, 120.4, 120.6, 122.7, 130.2, 131.0, 133.5, 144.2, 152.9, 159.0, 163.6, 167.8 ppm. HRMS (ESI<sup>+</sup>) calcd for C<sub>17</sub>H<sub>15</sub>C<sub>13</sub>NO<sub>4</sub> [M<sup>+</sup>H]<sup>+</sup>: 402.00667, found: 402.00641.

### 2.7. Preparation of Schiff Base 12

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 2,5-hydroxybenzaldehyde (**7**; 28 mg; 0.2 mmol) were dissolved in anhydrous ethanol (10 mL) and stirred at 75 °C overnight. After cooling, the dark yellow solution was concentrated under reduced pressure to approximately half the volume, and water (10 mL) was added. The yellow precipitate was filtered off, washed with an ethanol-water mixture (1:2 v/v, 30 mL) and dried under vacuum. Ethyl-(*E*)-2-(2,4-dichloro-5-((2,5-dihydroxybenzylidene)amino)phenoxy)acetate (**12**) was obtained with a yield of 63 mg (88%) as a yellow solid.

<sup>1</sup>H NMR: 1.23 (t, *J* = 7.1 Hz, 3H); 4.19 (q, *J* = 7.1 Hz, 2H); 5.02 (s, 2H); 6.83 (d, *J* = 8.8 Hz, 1H); 6.92 (dd, *J* = 8.8, 3.0 Hz, 1H); 7.03 (d, *J* = 3.0 Hz, 1H); 7.42 (s, 1H); 7.77 (s, 1H); 8.87 (s, 1H); 9.40 (s, 1H); 12.23 (s, 1H). <sup>13</sup>C NMR: 14.0, 60.8, 65.6, 105.4, 116.8, 117.5, 119.0, 119.9, 120.5, 122.2, 130.1, 144.7, 149.7, 152.9, 153.4, 164.8, 167.9 ppm. HRMS (ESI<sup>+</sup>) calcd for C<sub>17</sub>H<sub>16</sub>C<sub>12</sub>NO<sub>5</sub> [M<sup>+</sup>H]<sup>+</sup>: 384.04055, found: 384.04102.

### 2.8. Preparation of Schiff Base 13

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 2-hydroxy-1-naphthaldehyde (**8**; 35 mg; 0.2 mmol) were dissolved in anhydrous ethanol (10 mL) and stirred at 75 °C overnight. Yellow precipitate formed almost immediately. After cooling, the solid product was filtered off, washed with ethanol (30 mL) and dried under vacuum. Ethyl-(*E*)-2-(2,4-dichloro-5-(((2-hydroxynaphthalen-1-yl)methylene)amino)phenoxy)acetate (**13**) was obtained with a yield of 77 mg (92%) as a yellow solid.

<sup>1</sup>H NMR: 1.24 (t, *J* = 7.1 Hz, 3H); 4.22 (q, *J* = 7.1 Hz, 2H); 5.15 (s, 2H); 7.13 (d, *J* = 9.1 Hz, 1H); 7.43 (t, *J* = 7.6 Hz, 1H); 7.63 (t, *J* = 7.8 Hz, 1H); 7.72 (s, 1H); 7.79 (s, 1H); 7.87 (d, *J* = 8.0 Hz, 1H); 8.03 (d, *J* = 9.1 Hz, 1H); 8.54 (d, *J* = 8.4 Hz, 1H); 9.72 (s, 1H); 15.50 (s, 1H). <sup>13</sup>C NMR: 14.0, 60.8, 65.8, 105.2, 109.2, 119.5, 119.8, 120.7, 120.8, 123.9, 127.1, 128.2, 129.1, 130.1, 132.7, 137.0, 141.9, 153.0, 158.4, 167.3, 167.9 ppm. HRMS (ESI<sup>+</sup>) calcd for C<sub>21</sub>H<sub>18</sub>C<sub>12</sub>NO<sub>4</sub> [M<sup>+</sup>H]<sup>+</sup>: 418.06129, found: 418.06126.

### 2.9. Preparation of Chromanone 17

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 3-formylchromone (**14**; 35 mg; 0.2 mmol) were dissolved in anhydrous isopropanol (8 mL) and stirred at 75 °C overnight. Then, the reaction mixture was evaporated to dryness, and diethyl ether (10 mL) was added followed by hexane (10 mL). The solid product was filtered off, washed with a diethyl ether-hexane mixture (1:1 v/v, 20 mL) and dried under vacuum. Ethyl-(*Z*)-2-(2,4-dichloro-5-(((2-isopropoxy-4-oxochroman-3-ylidene)methyl)amino)phenoxy)acetate (**17**) was obtained with a yield of 63 mg (66%) as a pale-yellow solid.

<sup>1</sup>H NMR: 1.04 (d, *J* = 6.2 Hz, 3H); 1.20 (d, *J* = 6.1 Hz, 3H); 1.23 (t, *J* = 7.1 Hz, 3H); 4.12 (hept, *J* = 6.1 Hz, 1H); 4.20 (q, *J* = 7.1 Hz, 2H); 5.02 (s, 2H); 6.01 (s, 1H); 7.08 (d, *J* = 8.1 Hz,

1H), 7.14 (t,  $J = 7.5$  Hz, 1H); 7.38 (s, 1H); 7.56 (m, 1H); 7.74 (s, 1H); 7.87 (m, 1H); 8.16 (d,  $J = 11.9$  Hz, 1H); 12.21 (d,  $J = 11.9$  Hz, 1H).  $^{13}\text{C}$  NMR: 14.0, 22.1, 23.2, 60.8, 65.8, 70.0, 98.6, 101.2, 105.9, 113.3, 116.3, 118.1, 122.0, 122.2, 125.9, 130.2, 135.0, 136.1, 143.2, 153.1, 155.8, 167.7, 181.1 ppm. HRMS (ACPI<sup>-</sup>) calcd for C<sub>23</sub>H<sub>22</sub>Cl<sub>2</sub>NO<sub>6</sub> [M-H]<sup>-</sup>: 478.08242, found: 478.08345.

#### 2.10. Preparation of Chromanone 18

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 6-fluoro-3-formylchromone (**15**; 38 mg; 0.2 mmol) were dissolved in anhydrous isopropanol (8 mL) and stirred at 75 °C overnight. A yellow solid gradually precipitated out of the solution. After cooling, the solid product was filtered off, washed with isopropanol (20 mL) and dried under vacuum. Ethyl-(Z)-2-(2,4-dichloro-5-(((6-fluoro-2-isopropoxy-4-oxochroman-3-ylidene)methyl)amino)phenoxy)acetate (**18**) was obtained with a yield of 69 mg (69%) as a yellow solid.

$^1\text{H}$  NMR: 1.04 (d,  $J = 6.2$  Hz, 3H); 1.20 (d,  $J = 6.1$  Hz, 3H); 1.23 (t,  $J = 7.1$  Hz, 3H); 4.11 (hept,  $J = 6.1$  Hz, 1H); 4.20 (q,  $J = 7.1$  Hz, 2H); 5.02 (s, 2H); 6.02 (s, 1H); 7.15 (dd,  $J = 8.9, 4.2$  Hz, 1H); 7.39 (s, 1H); 7.44 (td,  $J = 8.7, 3.1$  Hz, 1H); 7.57 (dd,  $J = 8.5, 3.1$  Hz, 1H); 7.75 (s, 1H); 8.21 (d,  $J = 11.9$  Hz, 1H); 12.20 (d,  $J = 12.0$  Hz, 1H).  $^{13}\text{C}$  NMR: 14.0, 22.1, 23.1, 60.8, 65.8, 70.1, 98.6, 101.4, 105.5, 111.1 (d,  $J = 24.0$  Hz), 113.4, 116.6, 120.1 (d,  $J = 7.8$  Hz), 122.1 (d,  $J = 24$  Hz), 123.0 (d,  $J = 6.6$  Hz), 130.3, 135.9, 143.9, 151.9 (d,  $J = 1.5$  Hz), 153.1, 157.0 (d,  $J = 239.2$  Hz), 167.7, 180.1 ppm. HRMS (ACPI<sup>-</sup>) calcd for C<sub>23</sub>H<sub>21</sub>Cl<sub>2</sub>FNO<sub>6</sub> [M-H]<sup>-</sup>: 496.07300, found: 496.07404.

#### 2.11. Preparation of Chromanone 19

Ethyl-2-(5-amino-2,4-dichlorophenoxy)acetate (**3**; 53 mg; 0.2 mmol) and 6-nitro-3-formylchromone (**16**; 44 mg; 0.2 mmol) were dissolved in anhydrous isopropanol (8 mL) and stirred at 75 °C overnight. A yellow solid gradually precipitated out of solution. After cooling, the solid product was filtered off, washed with isopropanol (20 mL) and dried under vacuum. Ethyl-(Z)-2-(2,4-dichloro-5-(((2-isopropoxy-6-nitro-4-oxochroman-3-ylidene)methyl)amino)phenoxy)acetate (**19**) was obtained with a yield of 98 mg (93%) as a yellow solid.

$^1\text{H}$  NMR: 1.07 (d,  $J = 6.2$  Hz, 3H); 1.20 (d,  $J = 6.4$  Hz, 3H); 1.24 (t,  $J = 7.1$  Hz, 3H); 4.19 (m, 3H); 5.03 (s, 2H); 6.19 (s, 1H); 7.36 (d,  $J = 9.0$  Hz, 1H); 7.42 (s, 1H); 7.82 (s, 1H); 8.31 (d,  $J = 12.0$  Hz, 1H); 8.40 (dd,  $J = 9.0, 2.8$  Hz, 1H); 8.63 (d,  $J = 2.8$  Hz, 1H); 12.26 (d,  $J = 12.2$  Hz, 1H).  $^{13}\text{C}$  NMR: 14.0, 22.1, 23.1, 60.9, 65.8, 70.9, 99.7, 101.6, 104.8, 113.7, 117.1, 119.7, 121.6, 122.0, 129.4, 130.3, 135.6, 142.1, 144.9, 153.1, 160.3, 167.6, 178.8 ppm. HRMS (ACPI<sup>-</sup>) calcd for C<sub>23</sub>H<sub>21</sub>Cl<sub>2</sub>N<sub>2</sub>O<sub>8</sub> [M-H]<sup>-</sup>: 523.06750, found: 523.06861.

#### 2.12. Preparation of BAR-2

Bar-2 was made according to Wang et al. [19].

#### 2.13. PPO Expression and Purification

A plasmid for bacterial expression of human PPO (hPPO) was kindly provided by Dr. Harry A. Dailey (University of Georgia, Athens, GA, USA) [34]. For heterologous expression, *E. coli* JM109 cells were grown at 30 °C as reported previously [12,34]. Circle growth media was supplemented with riboflavin to a final concentration of 0.75 µg/mL for 2 h prior to harvesting. Cells were collected by centrifugation (10,000× *g*, 10 min), resuspended in breaking buffer (50 mM Tris pH 8, 100 mM NaCl, 10% glycerol, 0.5% Tween 20 (*v/v*)) containing EDTA-free protease inhibitor cocktail (Roche Diagnostics GmbH, Mannheim, Germany), sonicated on ice and centrifuged at 40,000× *g* for 30 min. The supernatant was applied onto a Ni-NTA column (Ni-NTA Superflow, IBA, Germany), washed with 20 column volumes of equilibration buffer (50 mM Tris-HCl, 100 mM NaCl, 30 mM imidazole, 10% glycerol, 0.02% Tween 20, pH 8) and eluted with equilibration buffer supplemented with 300 mM imidazole. The eluted fractions were pooled and concentrated. The final purification step encompassed size-exclusion chromatography on a Superdex

200 16/60 column (GE Healthcare Life Sciences, Uppsala, Sweden) using 50 mM Tris-HCl, 100 mM NaCl, 5% glycerol, and 0.02% Tween 80, pH 8.0, as the mobile phase.

#### 2.14. Inhibition Assay

IC<sub>50</sub> values were determined using a fluorescence-based assay quantifying an increase in fluorescence signal upon conversion of the non-fluorescent protoporphyrinogen IX to fluorescent protoporphyrin IX [12]. Briefly, PPO (final concentration 4 nM) was preincubated with a dilution series of tested inhibitors (concentration range 0.2–100 μM) in a 384-well plate in a total volume of 14 μL for 10 min at 37 °C in reaction buffer comprising 100 mM KH<sub>2</sub>HPO<sub>4</sub>, 0.3% (*w/v*) Tween 80, 5 mM DTT, and 1 mM EDTA, pH 7.2. Reactions were initiated by the addition of 7 μL of 10 μM protoporphyrinogen IX into the PPO/inhibitor mixture. The fluorescence signal of protoporphyrin IX was monitored with a CLARIOstar fluorimeter (BMG Labtech GmbH, Ortenberg, Germany) at λ<sub>EX</sub>/λ<sub>EM</sub> = 410/632 nm using continuous readout mode at 37 °C for 1 h. The reaction velocity was calculated from the linear portion of the measured signal against a standard calibration curve of defined protoporphyrin IX concentrations. Non-linear regression analysis was employed to calculate the IC<sub>50</sub> values using GraphPad Prism software [35]. Reactions without enzyme or inhibitor were used to define 0% and 100% PPO activity, respectively.

#### 2.15. Calculation of the Binding Energy of the PPO Complex with Schiff Bases

Three-dimensional (3D) structural models for human protoporphyrinogen IX oxidase (hPPO) and for the orthologue from *Nicotiana tabacum* (ntPPO) were downloaded from the Protein Data Bank. 3D structural models of compounds **9** to **19** were drawn using MolView (<https://molview.org>, accessed on 19 May 2020), and the corresponding mol files containing 3D coordinates were generated. These files were converted to PDB format using NCI's online SMILES translator and structure file generator (<https://cactus.nci.nih.gov/translate/>, accessed on 19 May 2020). For docking calculations, software from the AutoDock Vina suite was used [36]. The protein molecules were first placed in an orthorhombic box of suitable size for their total enclosure for the first round of docking calculations. Since the docking poses were found in the vicinity of the flavin adenine dinucleotide (FAD) isoalloxazine ring system, a second round of docking was performed using an orthorhombic box of size 20 × 18 × 22 Å<sup>3</sup> centered at a location on the accessible side of the isoalloxazine ring.

#### 2.16. Acifluorfen and Fomesafen Docking

This docking simulation was performed with AutoDock Vina software [36] using the X-ray crystal structure of hPPO (3NKS, [37]). Ligand PDB files were prepared with the AceDRG stereochemical description generator [38]. The docking site was defined as a 30 × 30 × 30 Å<sup>3</sup> cube centered on the original co-crystallized ligand acifluorfen. Prior to the docking experiment, the docking workflow was validated by re-docking acifluorfen and comparing the calculated pose to the conformation of acifluorfen from the crystal structure. The re-docking results are shown in Supplementary Figure S11. The results of the docking experiment of acifluorfen and fomesafen were visualized and analyzed by PyMOL 2.4.1. [39].

#### 2.17. Viability Assay

RPMI-8226 lymphoblasts were used to assess the cytotoxicity of the tested herbicides. Cells were diluted in 90 μL of RPMI-1640 medium supplemented with 10% fetal bovine serum (FBS) to 2.5 × 10<sup>5</sup> cells/mL in a 96-well plate. Herbicides (50 mM stock solutions in dimethylsulfoxide (DMSO)) were diluted in PBS to 10× the desired concentration, and 10 μL was added to the cells followed by incubation for 2 days at 37 °C. Following incubation, 10 μL of (3-(4,5-dimethylthiazol-2-yl)-2,5-diphenyl tetrazolium bromide (MTT reagent) (5 mg/mL in PBS, Sigma-Aldrich) was added followed by incubation for 120 min at 37 °C. The metabolized formazan crystals were dissolved by the addition of 100 μL of

solubilization solution (40% dimethyl formamide (DMF), 2% acetic acid, 16% SDS, pH 4.7). Then, the absorbance at 570 nm was measured using a CLARIOstar plate reader (BMG Labtech), and data were plotted using GraphPad Prism software.

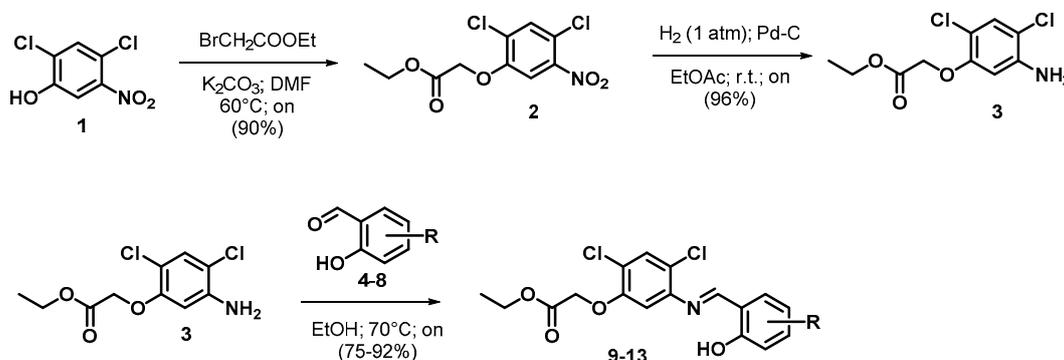
### 2.18. hPPO Stability Determined by the ThermoFAD Assay

The ThermoFAD stability assay according to Forneris et al. [40] was used to determine the melting temperature ( $T_M$ ) of the PPO complexes with the herbicides and Schiff bases. hPPO solution (0.5 mg/mL final concentration) in 50 mM Tris, 100 mM NaCl, 5% glycerol, and 0.2% Tween 20, pH 8, was incubated with/without 50  $\mu$ M herbicide in a total volume of 100  $\mu$ L for 15 min at room temperature. Samples were denatured using a RT-PCR cyclor (CFX96-Touch, Bio-Rad, Hercules, CA, USA) with temperature ramping of 0.5  $^{\circ}$ C per step for 30 s from 20 to 95  $^{\circ}$ C. The fluorescence of the FAD cofactor was monitored using a 450–490 excitation filter and a 515–530 nm emission filter typically used for SYBR Green-based RT-PCR. The maximum of the first derivative of the observed flavin fluorescence was taken as the apparent  $T_M$ .

## 3. Results

### 3.1. Preparation and Characterization of Schiff Bases

Two sets of target compounds were prepared by a three-step synthesis. In the first step, 2,4-dichloro-5-nitrophenol (**1**) was reacted with ethyl bromoacetate in the presence of  $K_2CO_3$  in DMF at 60  $^{\circ}$ C. Ethyl 2-(2,4-dichloro-5-nitrophenoxy)acetate (**2**) was obtained in a yield of 90%. In the next step, the nitro group of **2** was reduced with hydrogen with a palladium catalyst in ethyl acetate and the corresponding amine **3** was obtained in 96% yield. In the last step, amine **3** was reacted with different substituted 2-hydroxybenzaldehydes (**4–8**) in ethanol at 75  $^{\circ}$ C. These reactions gave corresponding Schiff bases **9–13** in yields of 75–92% (Scheme 1). The *E*-isomer was exclusively formed because it is stabilized by intramolecular hydrogen bonds (Figure 1). Thus, the  $^1H$  NMR spectra showed only one singlet from the  $-CH=N-$  hydrogen.



Scheme 1. Preparation of Schiff bases 9–13.

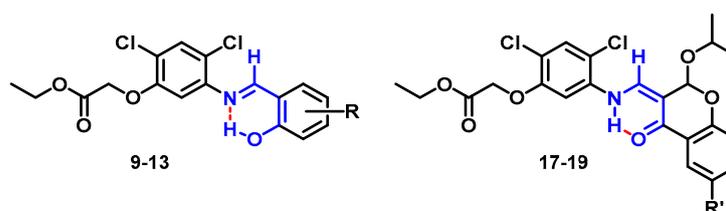
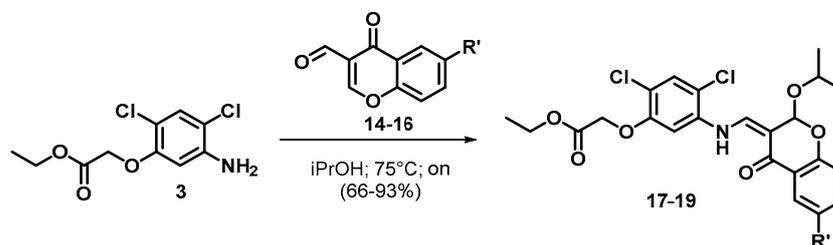


Figure 1. Intramolecular hydrogen bonds in Schiff bases 9–13 and chromanones 17–19.

Amine **3** was also reacted with substituted 3-formylchromones **14–16**. The reaction of 3-formylchromones with primary aromatic amines gives both Schiff base derivatives and 4-chromanone derivatives, depending on the solvent used [41–43]. In our case, we used isopropanol as the solvent; thus, we exclusively obtained the corresponding 4-chromanones

17–19 in yields of 66–93% (Scheme 2). The Z-isomer was exclusively formed because it is stabilized by intramolecular hydrogen bonds (Figure 1). Thus, the  $^1\text{H}$  NMR spectra showed only one doublet from the =CH-N- hydrogen.



Scheme 2. Preparation of chromanones 17–19.

The structures of target compounds 9–13 and 17–19 and preparative yields are shown in Figure 2. The structures of the tested PPO herbicides are shown in Figure 3.

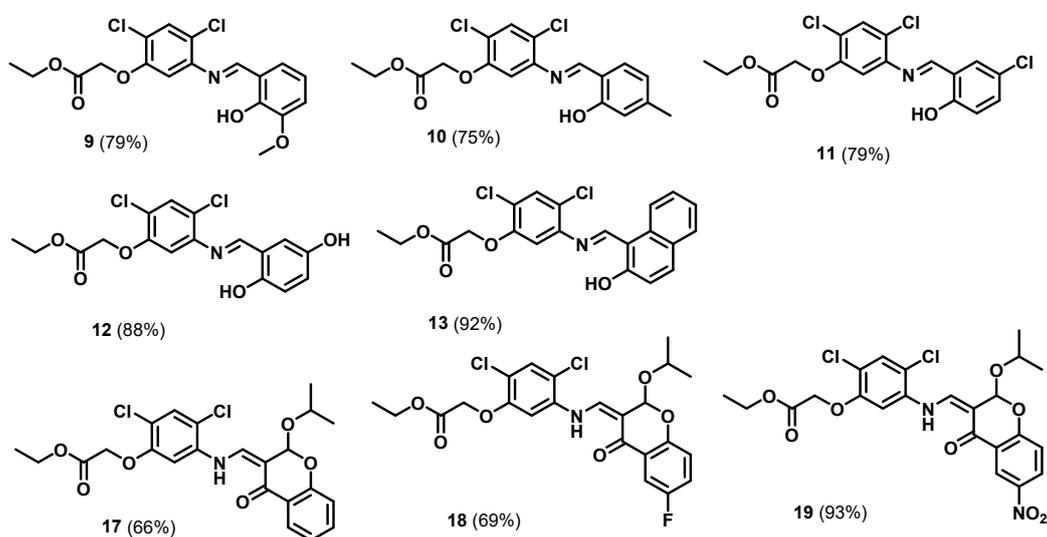


Figure 2. Structures and preparative yields of 9–13 and 17–19.

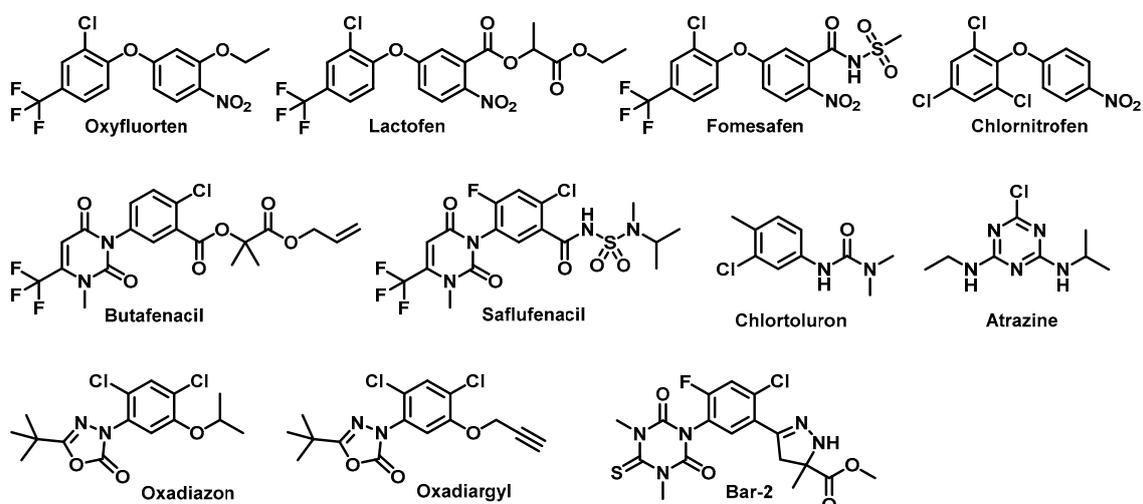


Figure 3. Structures of the tested commercial herbicides and Bar-2.

### 3.2. Determination of the Inhibitory Activities

The IC<sub>50</sub> values of the commercial herbicides (Figure 3) and prepared Schiff bases were determined by a fluorescence protoporphyrin IX assay. The results are shown in Table 1. Some of these compounds, such as oxyfluorten, lactofen, saflufenacil, fomesafen and chlornitrofen, displayed IC<sub>50</sub> values in the submicromolar range. Nevertheless, for oxyfluorten, this value is comparable to published activities for plant PPOs [44]. In addition, the determined inhibitory concentrations of saflufenacil, lactofen and oxadiazon were significantly higher than the values published for plant PPOs [45–47], and their values are shown in Supplementary Table S1. For example, saflufenacil displayed nanomolar inhibitory activity against various plant PPOs [46]. Published values for lactofen (IC<sub>50</sub> = 25 nM) [45] and oxadiazon (IC<sub>50</sub> = 60 nM) [47] are in the tens of nanomolar range; however, their inhibition activity against hPPO was significantly lower. Only fomesafen displayed an IC<sub>50</sub> value against hPPO that was comparable to those of plant PPOs [44]. However, the value of fomesafen against hPPO (110 nM) strongly implies that fomesafen is a very potent inhibitor of mammalian PPOs. For example, the concentration of acifluorfen that elicits 50% of the total change in activity is 1.12 μM. Similarly, a value of 1.48 μM was published by Shepherd et al. [12].

**Table 1.** IC<sub>50</sub> values against hPPO for the tested inhibitors and melting temperatures for the hPPO inhibitor complexes.

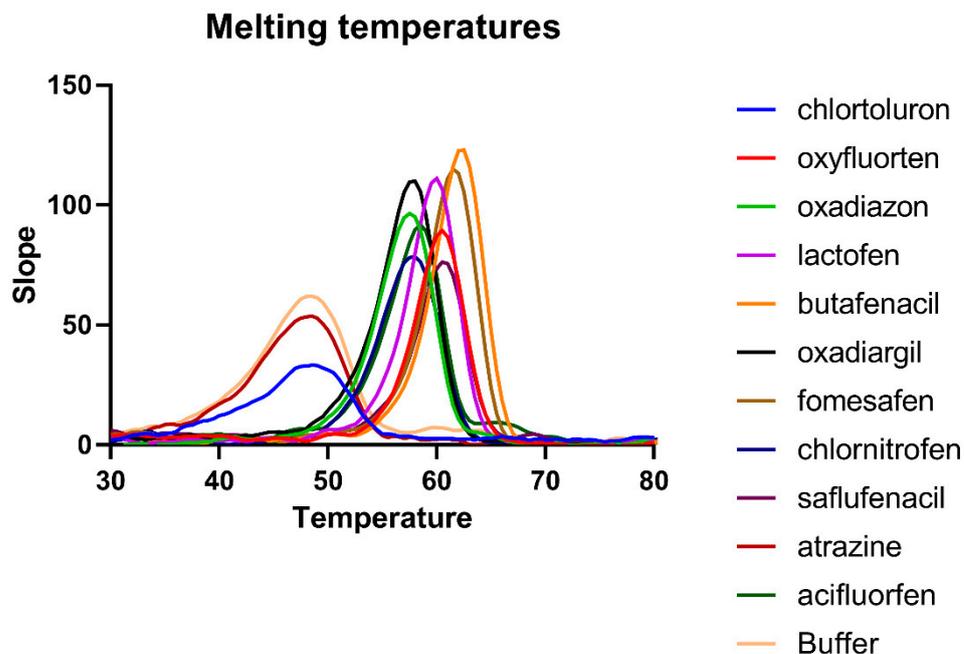
Compound	IC <sub>50</sub> (μM)	T <sub>M</sub> (°C) <sup>1</sup>
No inhibitor		48.5
Acifluorfen	1.12 ± 0.14	58.5
Chlortoluron	>100	48.5
Oxyfluorten	0.13 ± 0.06	60.5
Oxadiazon	1.92 ± 0.39	57.5
Lactofen	0.33 ± 0.18	60.0
Butafenacil	0.15 ± 0.05	62.5
Saflufenacil	0.25 ± 0.02	60.5
Oxadiargyl	0.85 ± 0.14	58.0
Fomesafen	0.11 ± 0.01	61.5
Chlornitrofen	0.420 ± 0.001	58.0
Atrazine	>100	48.5
Schiff base 9	>100	48.5
Schiff base 10	>100	48.5
Schiff base 11	>100	48.5
Schiff base 12	>100	48.5
Schiff base 13	>100	48.5
Chromone 17	>100	48.5
Chromone 18	>100	48.5
Chromone 19	>100	48.5
Bar-2	3.86 ± 0.01	57.5

<sup>1</sup> Melting temperature of hPPO with and without tested inhibitors.

Against expectations, no significant activity was observed from the prepared Schiff bases against hPPO. Published works have implied that these compounds could display high affinity for and inhibitory activity against PPOs. Compounds with dihalogenated phenyl cores (e.g., a dichloro-substitution) display inhibitory activity against PPOs and are thus usually used as herbicides. Although the binding docking studies showed that these compounds had a high affinity (absolute value of interaction energy was 7 kcal/mol or higher (Supplementary Table S2)) for hPPO, incorporation of the Schiff base into the herbicide design led to a loss in inhibitory activity. In this case of BaR-2, the determined value was several times weaker than the published activity against plant mitochondrial PPO2 from *N. tabacum*. [19].

In vitro inhibition data were further corroborated by a biophysical assay that determined the T<sub>M</sub> of hPPO and hPPO/herbicide complexes (Figure 4, Table 1). Here, upon

hPPO binding, the herbicide stabilizes the three-dimensional structure of the protein, resulting in a substantial increase in the  $T_M$  of the complex. However, compounds that do not inhibit hPPO (e.g., the prepared Schiff bases and chromones, chlortoluron and atrazine) and thus likely do not interact with the enzyme do not show any effect on the  $T_M$  of hPPO. As the ThermoFAD assay is better suited for high-throughput sample processing, it can be used to prescreen putative hPPO inhibitors before engaging in demanding inhibitory testing that requires more resources.



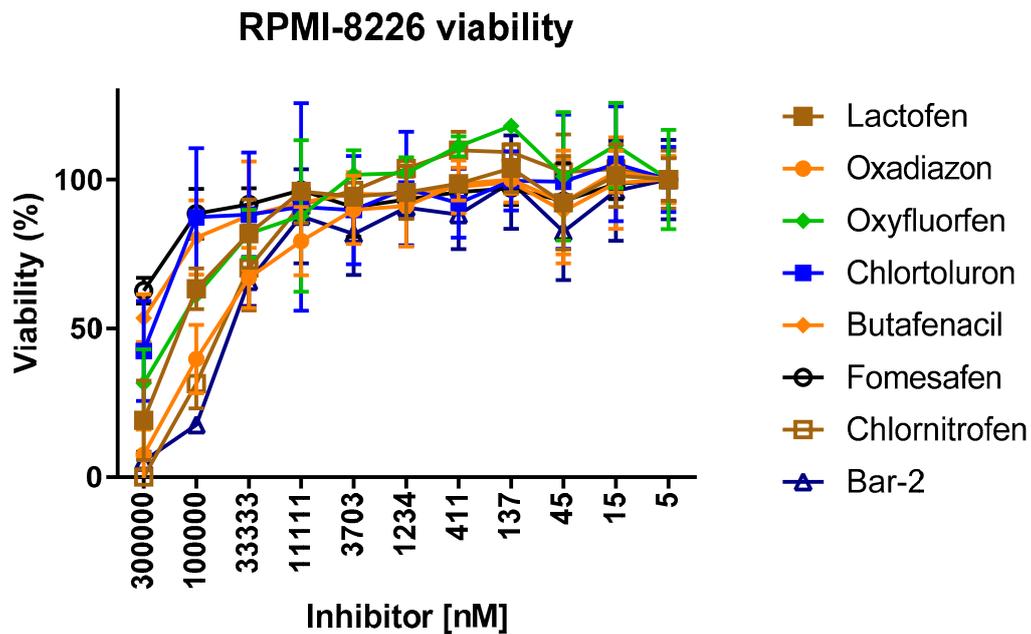
**Figure 4.** Thermal denaturation profiles of the hPPO complexes. The  $T_M$  values of the hPPO/herbicide complexes were determined using the ThermoFAD assay, and the peak of the first derivative of the FAD fluorescence represents the  $T_M$  of a given complex.

### 3.3. Herbicide Toxicity in Human Cells

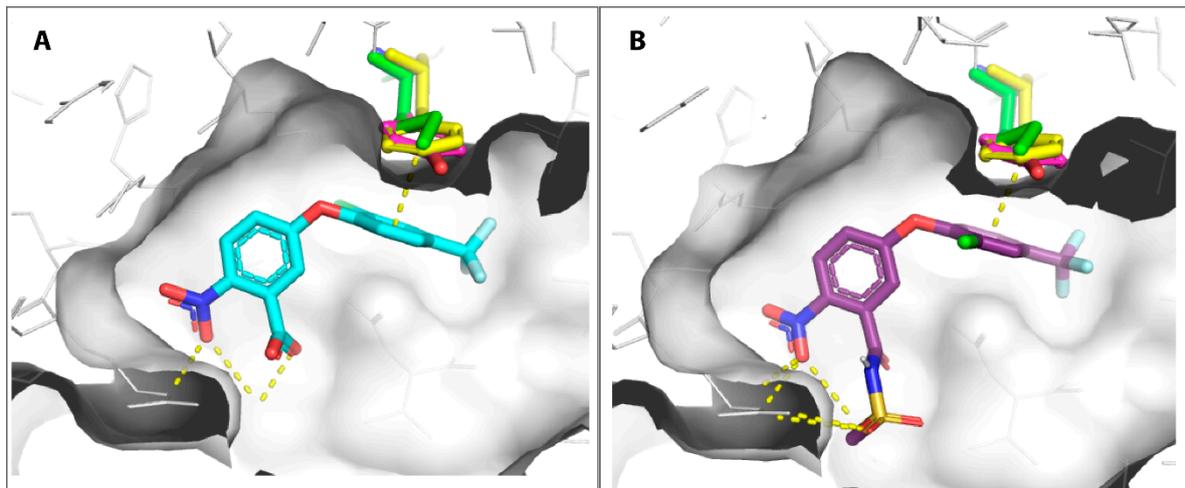
As the tested herbicides inhibit hPPO in the high nanomolar to low micromolar range, we wondered about their toxicity to human cells. To this end, we assayed the viability of human RPMI-8226 lymphoblasts after 2 days of treatment with the herbicides using an MTT assay. Overall, their toxicity to human cells was limited, with half maximal effective concentration ( $EC_{50}$ ) values in the mid-micromolar range (Figure 5). It is interesting to note that cellular toxicity does not strictly correlate with inhibitor affinity for hPPO as determined *in vitro*. This observation can result from differences in cell permeability and cellular metabolism of the individual compounds and/or possible off-target effects.

### 3.4. Homology Modeling

Our data revealed that the tested herbicides are more potent against Plantae PPOs than the human enzyme. To rationalize these findings, we performed structural alignment of the crystal structures of hPPO (PDB: 3NKS) and ntPPO (PDB: 1SEZ) together with an *Arabidopsis thaliana* PPO (atPPO) homology model prepared using the SWISS-MODEL server (Figure 6). The most intriguing difference between the inhibitor binding pockets of these orthologues is the substitution of Met-368 (human numbering) in hPPO by Tyr and Phe in atPPO and ntPPO, respectively (Figure 6). The trifluoromethoxy group of acifluorfen, the herbicide co-crystallized with hPPO, is in van der Waals distance from Met-368. Substitution at this position with an aromatic amino acid (present in both atPPO and ntPPO) by Met-368 in the human enzyme leads to the disruption of the original  $\pi$ - $\pi$  stacking interactions, which likely results in lower inhibitor potency for the human orthologue.



**Figure 5.** Cytotoxicity of the tested herbicides. Cells were treated with varying concentrations of herbicides for 48 h, and cell viability was determined using the MTT assay.



**Figure 6.** Methionine substitution is linked to the lower potency of herbicides against hPPO. The hPPO binding pocket shown in grey representation (A) co-crystallized with acifluorfen (stick representation, cyan carbon atoms) and (B) docked with fomesafen (stick representation, violet carbon atoms). hPPO Met-368 is highlighted by green sticks. The atPPO and ntPPO structures were superimposed onto hPPO, but only the corresponding Tyr (stick representation, magenta carbon atoms) and Phe (stick representation, yellow carbon atoms) are highlighted.

As there were no crystal structures of hPPO with herbicides used in this study, we were interested in whether the tested compounds share the same binding mode as acifluorfen and would therefore act on human and plant PPOs in a similar manner. To this end, we selected fomesafen a representative compound, which displayed the most potent hPPO inhibition, and docked it into hPPO. The docking experiment result, as depicted in Figure 6B, points towards a similar binding orientation of both herbicides. We thus hypothesize that all herbicides of this chemotype act similarly and that the Met-368 substitution is responsible for their lower potency against hPPO.



with inhibition activity for hPPO, aliphatic parts of synthetic compounds could be enabled to bind to hPPO. Most likely, in our opinion, the corresponding part of the enzyme active site is narrow and rigid for intercalation of the hydroxyaryl moiety of synthetic inhibitors.

Based on the determined  $IC_{50}$  values, fomesafen should be the most toxic of the tested inhibitors. However, the toxicity of the tested herbicides to lymphoblasts did not correlate with their inhibitory activity against hPPO. It is possible that a significant part of their toxic effects in vitro cannot be associated with hPPO inhibition, such as coproporphyrin accumulation [13]. Krijit et al. reported that after herbicide dosage, liver, fecal and bile porphyrin IX accumulation increased in the following order: fomesafen, oxyfluorfen and oxadiazon [50]. However, based on our results, we would expect the opposite trend.

This result implies that other phenomena, such as cellular uptake and herbicide transport in the blood, can play a role in the physiological effects of these compounds. For example, Dong et al. observed that fomesafen can form a complex with serum albumin [51]. Serum albumin is a long-circulating and highly abundant protein (40 mg/mL) [52]. Transported agents have a longer half-life than their free form in the blood and thereby lower cellular uptake. On the other hand, serum albumin has been intensively studied as a carrier for the targeted transport of anticancer agents, and PPO inhibition can significantly increase the effects of photodynamic therapy.

Although we did not design new synthetic inhibitors of hPPO, based on the structural motifs of the tested herbicides (mainly fomesafen), we found prospective structural motifs as starting points for redesign. Based on docking studies, we found that commercial herbicides interact with the carbonyl group of valine and their close surroundings (e.g., carboxy, ester and ether groups) and Met-365 (e.g., halogenated, mostly fluorinated aromatic groups of herbicides). The obtained information could also be used to design new synthetic inhibitors to better predict the toxicity of commercial PPO herbicides.

## 5. Conclusions

Compounds with Schiff base structural motifs (9–13 and 17–19) were prepared and characterized by  $^1H$  NMR,  $^{13}C$  NMR and mass spectroscopy (MS). The inhibition activity of these compounds and of PPO herbicides and the influence of these compounds on the hPPO  $T_M$  were determined by a protoporphyrin IX fluorescence assay and the ThermoFAD assay, respectively. Some commercial herbicides, such as fomesafen, displayed strong inhibition of hPPO activity. In the case of Schiff bases 9–13 and chromones 17–19, no significant inhibitory activity was observed. However, we discovered the hPPO binding model with inhibitors based on a docking study (by using AutoDock Vina software) and determined the  $IC_{50}$  values for their redesign and participation in the accurate prediction of the toxicity of PPO herbicides.

**Supplementary Materials:** The following are available online at <https://www.mdpi.com/2227-9717/9/2/383/s1>, Figures S1–S10:  $^1H$  and  $^{13}C$  NMR spectra of compounds 2 and 3, Schiff bases 9–13 and chromones 17–19. Figure S11: Re-docking of acifluorfen into the binding pocket of hPPO (PDB: 3NKS). Table S1: Calculated interaction energy for hPPO and tobacco PPO. Table S2: Influence of PPO herbicides on the activity of plant PPOs.

**Author Contributions:** Conceptualization, project administration and funding were achieved by P.M. Original draft preparation and supervision of the manuscript were performed by M.J. Writing, discussion and introduction, and design of the synthetic inhibitors were performed by M.M. MS and  $^1H$  and  $^{13}C$  NMR analyses of prepared synthetic inhibitors were performed by K.V. Synthesis of BAR-2 and compounds 2 and 3 was performed by T.B. Synthesis of Schiff bases 9–13 was performed by R.K. Preparation of chromones 17–19 was performed by N.A. All authors have read and agreed to the published version of the manuscript.

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