

# Engineering of Halide Cation in All-Inorganic Perovskite with Full-Color Luminescence

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**Abstract:** All-inorganic halide perovskites are emerging as a class of superstar semiconductors with excellent optoelectronic properties and show great potential for a broad range of applications in solar cells, lighting diodes, X-ray imaging, and photodetectors. Tremendous research about their device performance has been performed since 2015. In this study, we synthesized the all-inorganic perovskite by the hot-injection method and particularly investigated their crystal structural and photoluminescence properties. By halide anion engineering, the all-inorganic perovskites showed a high-symmetry cubic phase. They also showed a tunable optical bandgap, and almost the full color luminescence was achieved (434 to 624 nm). These basic optoelectronic properties could give a guide for further development of this area.

**Keywords:** perovskite solar cell; thin film; crystal structure; photoluminescence



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## 1. Introduction

Recently, halide perovskites have been extensively studied due to their fascinating optoelectronic properties, such as tunable bandgaps, high absorption range and coefficients (up to  $10^5 \text{ cm}^{-1}$ ), long minority carrier lifetimes, low-cost fabrication process, flexible device, etc. [1,2]. A wide range of optoelectronic applications have been developed with them as the active layer, e.g., perovskite solar cells (PSCs), perovskite light-emitting diodes (PeLEDs), perovskite photodetectors, and perovskite sensors. For instance, Miyasaka et al. first discovered PSCs in 2009 [3], and the power conversion efficiency (PCE) has nowadays rapidly boosted to 25.5%, already surpassing that of commercial silicon solar cells [4]. Dan et al. reported perovskite organic light-emitting diodes (PeOLEDs) in 2014 for the first time [5], and the maximum external quantum efficiency (EQE) has today increased to over 20%, which is comparable to that of conventional OLEDs or quantum dot light-emitting diodes (QLEDs) [6]. In general, halide perovskites have a chemical formula of  $\text{ABX}_3$ , where A represent a large monovalent cation, such as organic methylammonium (MA:  $\text{CH}_3\text{NH}_3$ ) and inorganic cesium (Cs). B is a bivalent metal cation, such as Pb and Sn. The X anion, as the name implies, is a halogen anion: Cl, Br, I, and mixed Cl/Br/I systems. With the inorganic Cs cation at A position,  $\text{CsPbX}_3$  perovskite shows a significantly improve stability, which has already proven in the solid-state dye sensitized solar cells and previous PSC study [7,8].

One of the most fascinating properties of  $\text{CsPbX}_3$  perovskite materials is their tunable bandgaps, which have allowed them to achieve astonishing breakthrough in a variety of optoelectronic applications. For instance, the ideal bandgap for PSCs is 1.34 eV [9–11]; the ideal bandgap for tandem photovoltaic devices (top sub-cell) is 1.70 eV [12]; and the ideal bandgaps for PeLED are 1.80, 2.30, and 2.70 eV for red, green, and blue emission,

respectively. It has been proposed that in the CsPbX<sub>3</sub> system, the Cs cation acts to fulfill charge neutrality within the lattice and only slightly affects the electronic states [13]. The valance band of CsPbX<sub>3</sub> is predominately formed by mixing the halide  $np^6$  orbitals ( $n$  is the principal quantum number, Cl:  $n = 3$ , Br:  $n = 4$ , and I:  $n = 5$ ) and  $ns^2$  orbitals from the lead ( $n = 5$ ), while the conduction band mainly arises from antibonding mixing of the lead and the halide  $np^6$  orbital; thus, the X cation plays the main role in the determination of the bandgap [14,15]. Therefore, with the mixed halogen cation, a wide broad bandgap CsPbX<sub>3</sub> could be achieved. Although the CsPbX<sub>3</sub> perovskites were reported as early as 1958 [16], investigation of their application in optoelectronic devices had not been conducted until 2015 [17,18]. After that, there has been tremendous research on optoelectronic device performance. Noh et al. demonstrated excellent perovskite solar cells based on organic–inorganic perovskite MAPbI<sub>3-x</sub>Br<sub>x</sub> [17]. Guhrenz et al. demonstrated an excellent light-emitting device based on a mix of CsPbCl<sub>3</sub>, CsPbBr<sub>3</sub>, and CsPbI<sub>3</sub> [18]. Motivated by this, in this study, we systematically varied the halogen anion and synthesized CsPbCl<sub>3</sub>, CsPbCl<sub>2.5</sub>Br<sub>0.5</sub>, CsPbCl<sub>2</sub>Br, CsPbCl<sub>1.5</sub>Br<sub>1.5</sub>, CsPbClBr<sub>2</sub>, CsPbCl<sub>0.5</sub>Br<sub>2.5</sub>, CsPbBr<sub>3</sub>, CsPbBr<sub>2.5</sub>I<sub>0.5</sub>, CsPbBr<sub>2</sub>I, CsPbBr<sub>1.5</sub>I<sub>1.5</sub>, CsPbBrI<sub>2</sub>, CsPbBr<sub>0.5</sub>I<sub>2.5</sub>, and CsPbI<sub>3</sub> by the hot-injection method, and their crystal structural and photoluminescence properties were systematically studied.

## 2. Materials and Methods

**Preparation of Cs-oleate:** Oleic acid (OA, 90%), oleylamine (OAm, 90%), octadecene (ODE, 99%), Cs<sub>2</sub>CO<sub>3</sub> (98%), lead iodide (99%), cesium bromide (>98%), cesium chloride (98%), and cesium iodide (99%) were purchased from Sigma-Aldrich (Shanghai, China). Cs<sub>2</sub>CO<sub>3</sub> (0.4 g, 1.23 mmol), 15 mL ODE, and 1.25 mL OA were loaded into a 50 mL three-necked flask and dried for 0.5 h at 100 °C with stirring and then heated under N<sub>2</sub> to 150 °C until all Cs<sub>2</sub>CO<sub>3</sub> reacted with OA (transparent solution). Because Cs-oleate precipitated out of ODE at below 100 °C, it had to be annealed to 150 °C before CsPbX<sub>3</sub> perovskite synthesis.

**Synthesis of CsPbX<sub>3</sub>:** All-inorganic perovskites were synthesized by a hot-injection method according to Protesescu's work [17]. In this study, 0.36 mmol PbX<sub>2</sub> (X = Cl, Br, and I), OAm (1.0 mL), oleylamine (1.0 mL), and octadecene (10 mL) were added to a three-necked round-bottom flask (25 mL). The resulting mixture was heated to 100 °C with stirring and maintained for 0.5 h. At this time, the water residue was removed by nitrogen purging and vacuum aspiration, and the mixture was heated to 160 °C until the PbX<sub>2</sub> precursor was completely dissolved. Then, the hot cesium oleate precursor solution (1 mL) was quickly injected into the above reaction mixture. After 5 s reaction, the flask was quickly transferred to the ice bath, and the obtained CsPbX<sub>3</sub> was kept by centrifugation at 10,000 rpm for 10 min and stored in cyclohexane (4 mL) before further use.

**Synthesis of CsPbX<sub>3</sub> (X = Cl, Br, or I) nanocrystals containing mixed halogens:** a mmol of PbX<sub>2</sub> (X = Cl, Br, or I) and (0.36 – a) mmol of another PbX<sub>2</sub> (X = Cl, Br, or I), OA (1.0 mL), OAm (1.0 mL), and ODE (10 mL) were added to a three-necked round-bottom flask (50 mL). Then, it was annealed to 100 °C with stirring and maintained for 30 min. Then, it was heated to 160 °C until the PbX<sub>2</sub> precursor was completely dissolved. The hot cesium OA solution (1 mL) was quickly injected into the above solution, and after 5 s reaction, the flask was quickly transferred to the ice bath. The obtained CsPbX<sub>3</sub> perovskite was kept by centrifugation at 10,000 rpm for 5 min and stored in cyclohexane prior to further use. The added amount were as follows: CsPbCl<sub>3</sub> (PbCl<sub>2</sub>: 100.12 mg), CsPbCl<sub>2.5</sub>Br<sub>0.5</sub> (PbCl<sub>2</sub>: 83.43 mg, PbBr<sub>2</sub>: 22.02 mg), CsPbCl<sub>2</sub>Br (PbCl<sub>2</sub>: 66.72 mg, PbBr<sub>2</sub>: 44.04 mg), CsPbCl<sub>1.5</sub>Br<sub>1.5</sub> (PbCl<sub>2</sub>: 50.06 mg, PbBr<sub>2</sub>: 66.06 mg), CsPbClBr<sub>2</sub> (PbCl<sub>2</sub>: 33.37 mg, PbBr<sub>2</sub>: 88.08 mg), CsPbCl<sub>0.5</sub>Br<sub>2.5</sub> (PbCl<sub>2</sub>: 16.69 mg, PbBr<sub>2</sub>: 132.03 mg), CsPbBr<sub>3</sub> (PbBr<sub>2</sub>: 132.12 mg), CsPbBr<sub>2.5</sub>I<sub>0.5</sub> (PbBr<sub>2</sub>: 110.01 mg, PbI<sub>2</sub>: 27.66 mg), CsPbBr<sub>2</sub>I (PbBr<sub>2</sub>: 88.08 mg, PbI<sub>2</sub>: 55.32 mg), CsPbBr<sub>1.5</sub>I<sub>1.5</sub> (PbBr<sub>2</sub>: 66.06 mg, PbI<sub>2</sub>: 82.99 mg), CsPbBrI<sub>2</sub> (PbBr<sub>2</sub>: 43.67 mg, PbI<sub>2</sub>: 110.64 mg), CsPbBr<sub>0.5</sub>I<sub>2.5</sub> (PbBr<sub>2</sub>: 22.02 mg, PbI<sub>2</sub>: 138.30 mg), and CsPbI<sub>3</sub> (PbI<sub>2</sub>: 165.96 mg).

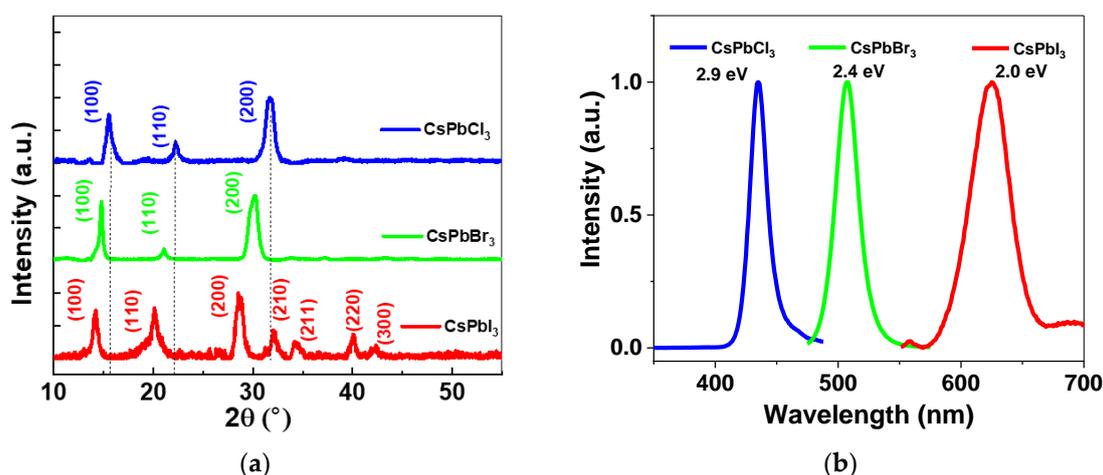
**X-ray diffraction (XRD) pattern characterization:** The experiment was performed by XRD diffraction device (Bruker D8 Advanced diffractometer with Cu K $\alpha$  radiation, Bruker,

Billerica, MA, USA). First, 100  $\mu\text{L}$  of the sample was taken onto the slide and dried under nitrogen atmosphere to make cyclohexane volatilize, which was then tested and analyzed.

Photoluminescence characterization: The experiment was carried out using a fluorescence spectrometer model Fluorolog-3. First, 100  $\mu\text{L}$  of the sample was taken and added to a quartz cuvette. Then, 4 mL of cyclohexane was added, and the luminescence wavelengths under 365 nm excitation were tested and analyzed.

### 3. Results

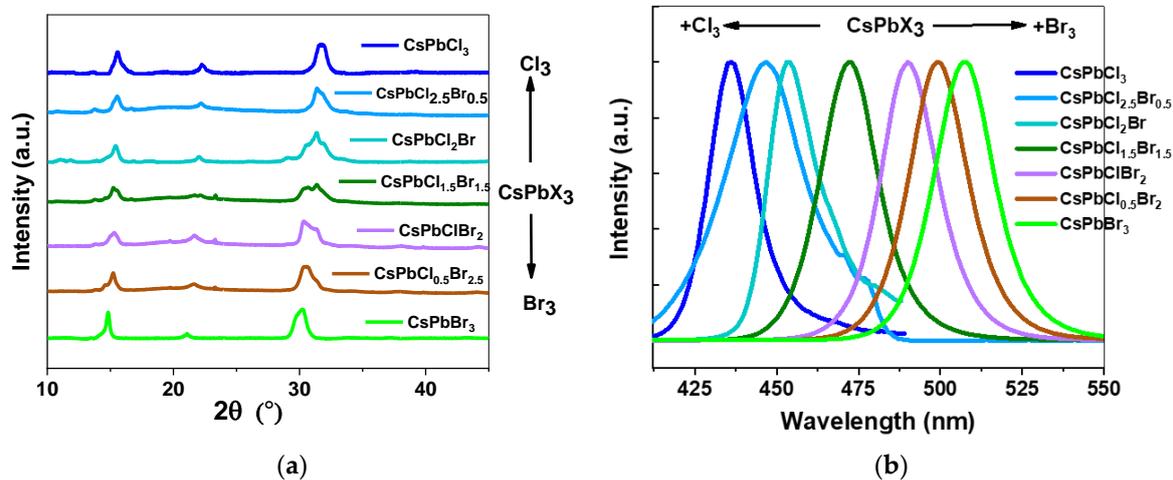
We began by studying how the X cation influenced the optoelectronic properties based on  $\text{CsPbX}_3$  with monohalogen (Cl, Br, and I). Figure 1a exhibits their XRD patterns monitored in the  $2\theta$  range of  $10^\circ$  to  $45^\circ$ , which matched well with the literature. With hot-injection preparation methods, the halogen cation engineering did not change the crystal structure. All  $\text{CsPbX}_3$  showed a high-symmetry cubic phase with the space group  $pm\bar{3}m$  due to the high reaction temperature in the preparation process and also the contribution from the surface energy of the nanocrystal [17,19,20].  $\text{CsPbCl}_3$  showed three dominant diffraction peaks: (100) at  $2\theta = 15.5^\circ$ , (110) at  $2\theta = 22.1^\circ$ , and (200) at  $2\theta = 31.3^\circ$  (ID: mp-23037).  $\text{CsPbBr}_3$  also showed three dominant diffraction peaks: (100) at  $2\theta = 14.8^\circ$ , (110) at  $2\theta = 20.9^\circ$ , and (200) at  $2\theta = 29.7^\circ$  (ID: mp-600089).  $\text{CsPbI}_3$  showed more diffraction peaks than those of the above two: (100) at  $2\theta = 13.9^\circ$ , (110) at  $2\theta = 19.6^\circ$ , (200) at  $2\theta = 27.9^\circ$ , (210) at  $2\theta = 31.2^\circ$ , (211) at  $33.9^\circ$ , (220) at  $40.2^\circ$ , and (300) at  $42.3^\circ$  (ID: mp-1069538). The optical properties of these  $\text{CsPbX}_3$  were studied by stable photoluminescence, as shown in Figure 1b. We found that an increase in the cation size resulted in a reduction in the optical bandgap, which was 2.90 eV for  $\text{CsPbCl}_3$ , 2.40 eV for  $\text{CsPbBr}_3$ , and 2.00 eV for  $\text{CsPbI}_3$ . The difference in optical properties originated from the difference in the ionic radius of halide ions with six-fold coordination, which was 1.81, 1.96, and 2.20  $\text{\AA}$  for  $\text{Cl}^-$ ,  $\text{Br}^-$ , and  $\text{I}^-$ , respectively. The  $\text{CsPbCl}_3$  perovskite exhibited the PL peak at 434 nm with full width at half maximum (FWHM) = 16 nm (blue color emission); the  $\text{CsPbBr}_3$  perovskite showed PL peak at 507 nm and FWHM = 23 nm (green color emission); and the  $\text{CsPbI}_3$  perovskite showed PL peak at 625 nm with FWHM = 37 nm (red color emission).



**Figure 1.** (a) X-ray diffraction patterns from  $\text{CsPbX}_3$  perovskite synthesized by hot-injection methods. (b) Photoluminescence emission spectra of  $\text{CsPbX}_3$  perovskite. Note that their main diffraction peaks and optical bandgaps are shown in inset of (a) and (b), respectively.

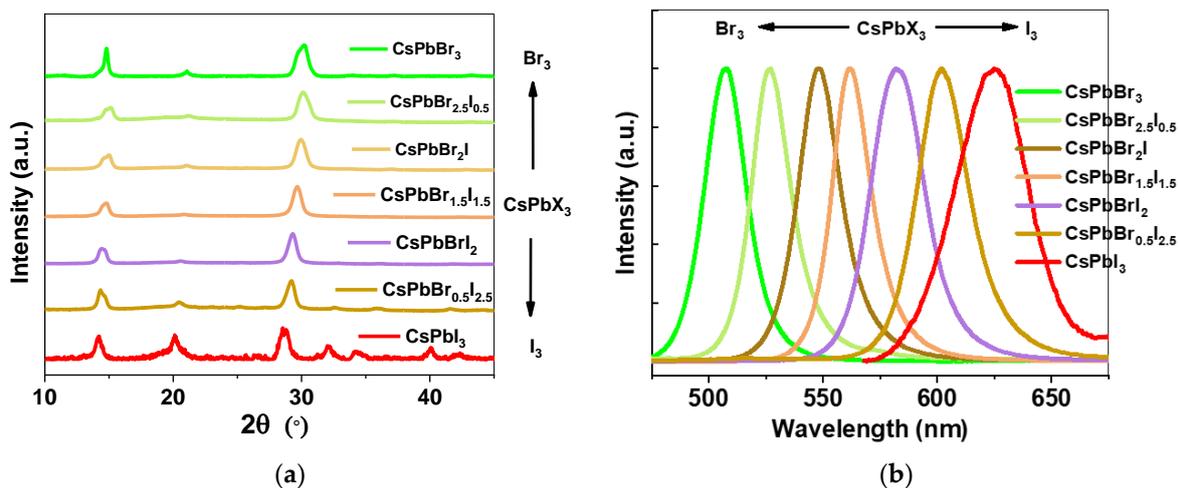
We then studied the influence of mixed halogen cation based on  $\text{CsPbCl}_y\text{Br}_{3-y}$  ( $y = 0, 0.5, 1, 1.5, 2, 2.5, \text{ and } 3$ ). From the XRD patterns (Figure 2a), the diffraction peaks exhibited successive shift, which was linearly dependent on the halide composition. The composition modulation of halide anions does not affect the cationic sublattice; meanwhile, the cubic structure is still maintained [21]. Moreover, it means the precursor solutions should be uniform during the film fabrication procedure. From the PL spectra shown in Figure 2b,

we also found the successively red shift emission peaks, which were 434 nm for CsPbCl<sub>3</sub>, 446 nm for CsPbCl<sub>2.5</sub>Br<sub>0.5</sub>, 453 nm for CsPbCl<sub>2</sub>Br, 471 nm for CsPbCl<sub>1.5</sub>Br<sub>1.5</sub>, 490 nm for CsPbClBr<sub>2</sub>, 499 nm for CsPbCl<sub>0.5</sub>Br<sub>2.5</sub>, and 507 nm for CsPbBr<sub>3</sub> perovskites, respectively. The redshift of the PL peaks indicated that the Cl anion was gradually replaced by Br to form CsPbCl<sub>y</sub>Br<sub>3-y</sub> ( $y = 0, 0.5, 1, 1.5, 2, 2.5, \text{ and } 3$ ).



**Figure 2.** (a) X-ray diffraction patterns and (b) photoluminescence spectra for mixed-halide CsPbCl<sub>y</sub>Br<sub>3-y</sub> ( $y = 0, 0.5, 1, 1.5, 2, 2.5, \text{ and } 3$ )

We further studied the influence of mixed halogen cation based on CsPbBr<sub>y</sub>I<sub>3-y</sub> ( $y = 0, 0.5, 1, 1.5, 2, 2.5, \text{ and } 3$ ). They exhibited similar trends as observed in CsPb(Cl/Br)<sub>3</sub> perovskites (as shown in Figure 3). The XRD peaks showed a successive shift, which was linearly dependent on the composition. The successive redshift of the photoluminescence peak was also observed, indicating that the Br anion was gradually mixed with I anion and formed CsPbBr<sub>y</sub>I<sub>3-y</sub> ( $y = 0, 0.5, 1, 1.5, 2, 2.5, \text{ and } 3$ ) perovskite.



**Figure 3.** (a) X-ray diffraction patterns and (b) photoluminescence spectra for mixed-halide CsPbBr<sub>y</sub>I<sub>3-y</sub> ( $y = 0, 0.5, 1, 1.5, 2, 2.5, \text{ and } 3$ ).

#### 4. Conclusions

In summary, we synthesized and characterized CsPbX<sub>3</sub> by adjusting the halogen anions (Cl, Br, I, mixed Cl/Br, and mixed Br/I). With controllable hot-injection methods, these all-inorganic perovskites exhibited a high-symmetry cubic phase. Meanwhile, by halogen anion engineering, they showed a successive tunable optical bandgap with luminescence covering the entire visible range (434–625 nm). Moreover, the successive bandgap

energy of CsPbX<sub>3</sub> were found to have a linear relationship with the halogen content. We are convinced that the present findings will be helpful for the development of promising related application, such as all-inorganic perovskite solar cells, perovskite light-emitting diodes, and perovskite photodetectors.

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## References

1. Abdi-Jalebi, M.; Andaji-Garmaroudi, Z.; Cacovich, S.; Stavrakas, C.; Philippe, B.; Richter, J.M.; Alsari, M.; Booker, E.P.; Hutter, E.M.; Pearson, A.J.; et al. Maximizing and stabilizing luminescence from halide perovskites with potassium passivation. *Nature* **2018**, *555*, 497–501. [[CrossRef](#)] [[PubMed](#)]
2. Kim, G.; Min, H.; Lee, K.S.; Lee, D.Y.; Yoon, S.M.; Seok, S.I. Impact of strain relaxation on performance of  $\alpha$ -formamidinium lead iodide perovskite solar cells. *Science* **2020**, *370*, 108–112. [[CrossRef](#)] [[PubMed](#)]
3. Kojima, A.; Teshima, K.; Shirai, Y.; Miyasaka, T. Organometal halide perovskites as visible-light sensitizers for photovoltaic cells. *J. Am. Chem. Soc.* **2009**, *131*, 6050–6051. [[CrossRef](#)] [[PubMed](#)]
4. Chart, Best Research-Cell Efficiency Chart. Available online: <https://www.nrel.gov/pv/cell-efficiency.html> (accessed on 20 December 2020).
5. Tan, Z.-K.; Moghaddam, R.S.; Lai, M.L.; Docampo, P.; Higler, R.; Deschler, F.; Price, M.; Sadhanala, A.; Pazos, L.M.; Credgington, D. Bright light-emitting diodes based on organometal halide perovskite. *Nat. Nanotechnol.* **2014**, *9*, 687–692. [[CrossRef](#)] [[PubMed](#)]
6. Kim, Y.-H.; Kim, S.; Kakekhani, A.; Park, J.; Park, J.; Lee, Y.-H.; Xu, H.; Nagane, S.; Wexler, R.B.; Kim, D.-H.; et al. Comprehensive defect suppression in perovskite nanocrystals for high-efficiency light-emitting diodes. *Nat. Photonics* **2021**, *15*, 148–155. [[CrossRef](#)]
7. Lee, B.; He, J.; Chang, R.P.; Kanatzidis, M.G. All-solid-state dye-sensitized solar cells with high efficiency. *Nature* **2012**, *485*, 486.
8. Li, J.; Huang, J.; Zhao, A.; Li, Y.; Wei, M. An inorganic stable Sn-based perovskite film with regulated nucleation for solar cell application. *J. Mater. Chem. C* **2020**, *8*, 8840–8845. [[CrossRef](#)]
9. Shockley, W.; Queisser, H.J. Detailed balance limit of efficiency of p-n junction solar cells. *J. Appl. Phys.* **1961**, *32*, 510–519. [[CrossRef](#)]
10. Rühle, S. Tabulated values of the Shockley–Queisser limit for single junction solar cells. *J. Sol. Energy* **2016**, *130*, 139–147. [[CrossRef](#)]
11. Li, J.; Hu, P.; Chen, Y.; Li, Y.; Wei, M. Enhanced performance of Sn-based perovskite solar cells by two-dimensional perovskite doping. *ACS Sustain. Chem. Eng.* **2020**, *8*, 8624–8628. [[CrossRef](#)]
12. Zhao, D.; Ding, L. All-perovskite tandem structures shed light on thin-film photovoltaics. *Sci. Bull.* **2020**, *65*, 1144–1146. [[CrossRef](#)]
13. Borriello, I.; Cantele, G.; Ninno, D. Ab initio investigation of hybrid organic-inorganic perovskites based on tin halides. *Phys. Rev. B* **2008**, *77*, 235214. [[CrossRef](#)]
14. Noh, J.H.; Im, S.H.; Heo, J.H.; Mandal, T.N.; Seok, S.I. Chemical management for colorful, efficient, and stable inorganic–organic hybrid nanostructured solar cells. *Nano Lett.* **2013**, *13*, 1764–1769. [[CrossRef](#)] [[PubMed](#)]
15. Liashenko, T.G.; Cherotchenko, E.D.; Pushkarev, A.P.; Pakštas, V.; Naujokaitis, A.; Khubezhov, S.A.; Polozkov, R.G.; Agapev, K.B.; Zakhidov, A.A.; Shelykh, I.A.; et al. Electronic structure of CsPbBr<sub>3–x</sub>Cl<sub>x</sub> perovskites: Synthesis, experimental characterization, and DFT simulations. *PCCP* **2019**, *21*, 18930–18938. [[CrossRef](#)] [[PubMed](#)]
16. Møller, C.K. Crystal structure and photoconductivity of caesium plumbahalides. *Nature* **1958**, *182*, 1436. [[CrossRef](#)]
17. Protesescu, L.; Yakunin, S.; Bodnarchuk, M.I.; Krieg, F.; Caputo, R.; Hendon, C.H.; Yang, R.X.; Walsh, A.; Kovalenko, M.V. Nanocrystals of cesium lead halide perovskites (CsPbX<sub>3</sub>, X = Cl, Br, and I): Novel optoelectronic materials showing bright emission with wide color gamut. *Nano Lett.* **2015**, *15*, 3692–3696. [[CrossRef](#)] [[PubMed](#)]

18. Guhrenz, C.; Benad, A.; Ziegler, C.; Haubold, D.; Gaponik, N.; Eychmüller, A. Solid-state anion exchange reactions for color tuning of CsPbX<sub>3</sub> perovskite nanocrystals. *Chem. Mater.* **2016**, *28*, 9033–9040. [[CrossRef](#)]
19. Harada, J.; Sakata, M.; Hoshino, S.; Hirotsu, S. Neutron diffraction study of the structure in cubic CsPbCl<sub>3</sub>. *J. Phys. Soc. Jpn.* **1976**, *40*, 212–218. [[CrossRef](#)]
20. Sutton, R.J.; Filip, M.R.; Haghighirad, A.A.; Sakai, N.; Wenger, B.; Giustino, F.; Snaith, H.J. Cubic or orthorhombic? Revealing the crystal structure of metastable black-phase CsPbI<sub>3</sub> by theory and experiment. *ACS Energy Lett.* **2018**, *3*, 1787–1794. [[CrossRef](#)]
21. Nedelcu, G.; Protesescu, L.; Yakunin, S.; Bodnarchuk, M.I.; Grotevent, M.J.; Kovalenko, M.V. Fast anion-exchange in highly luminescent nanocrystals of cesium lead halide perovskites (CsPbX<sub>3</sub>, X= Cl, Br, I). *Nano Lett.* **2015**, *15*, 5635–5640. [[CrossRef](#)] [[PubMed](#)]