

Phase evolution and textural changes during the direct conversion and storage of CO₂ to produce calcium carbonate from calcium hydroxide

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The contents of the supporting information are listed below:

- Figure S1. Determination of the weight change in calcium hydroxide using thermogravimetric analyses (TGA).
- Figure S2. Changes in the characteristic peaks of Ca(OH)₂ ($d = 4.90 \text{ \AA}$, $q = 1.28 \text{ \AA}^{-1}$, h k l: (0 0 1)) and the integrated peak intensity are represented in (a) and (b) respectively [40]. The relative integrated intensity of calcium hydroxide represented in (b) is the integrated intensity of the characteristic peak at a given temperature normalized to the integrated intensity at 30 °C. Vertical bars in (b) represent estimated 5% standard deviation uncertainties.

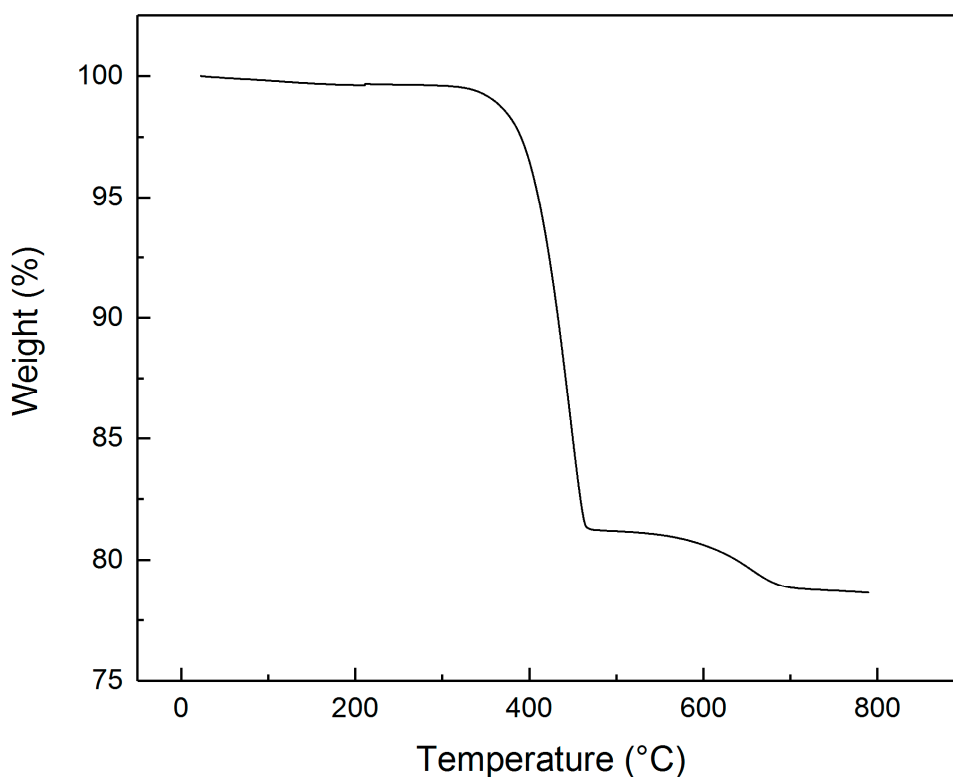


Figure S1. Determination of the weight change in calcium hydroxide using thermogravimetric analyses (TGA).

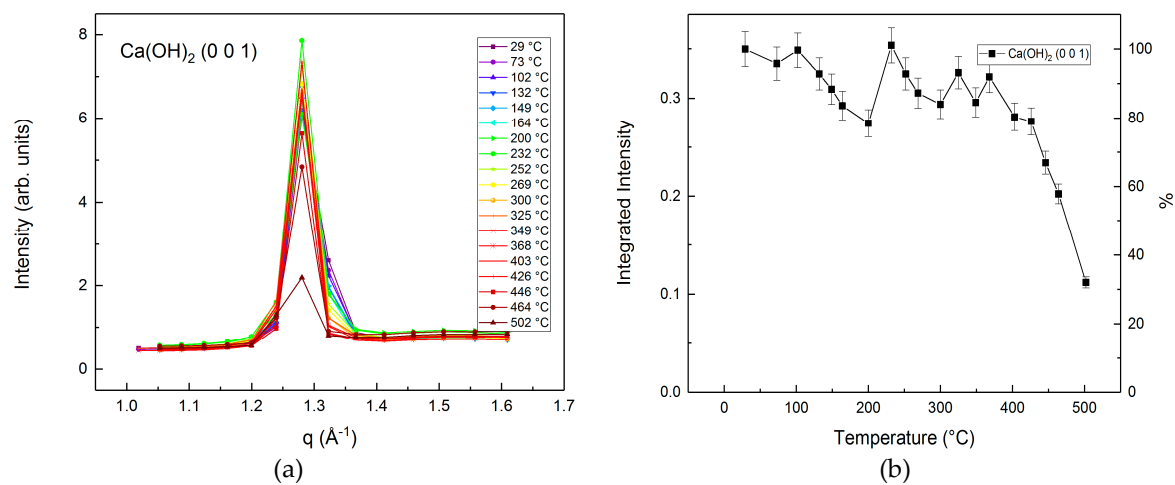


Figure S2. Changes in the characteristic peaks of Ca(OH)_2 ($d = 4.90 \text{ \AA}$, $q = 1.28 \text{ \AA}^{-1}$, h k l: (0 0 1)) and the integrated peak intensity are represented in (a) and (b) respectively [40]. The relative integrated intensity of calcium hydroxide represented in (b) is the integrated intensity of the characteristic peak at a given temperature normalized to the integrated intensity at 30 °C. Vertical bars in (b) represent estimated 5% standard deviation uncertainties.