

Article

Flotation Performance, Structure-Activity Relationship and Adsorption Mechanism of O-Isopropyl-N-Ethyl Thionocarbamate Collector for Elemental Sulfur in a High-Sulfur Residue

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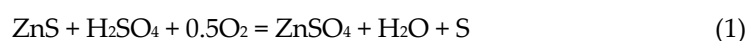
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Abstract: O-isopropyl-N-ethyl thionocarbamate (IPETC) collector was used to selectively recover elemental sulfur from a high-sulfur residue, and its flotation performance, structure–property relationship and adsorption mechanism to elemental sulfur were studied. The raw ore flotation test showed that IPETC displayed superior flotation performance to the elemental sulfur compared with sodium ethyl xanthate (SEX) and ammonium dibutyl dithiophosphate (ADDTP) collectors. Pure mineral flotation and adsorption experiments further demonstrated that among the three collectors, IPETC had the strongest collecting power and the optimum selectivity towards elemental sulfur. The structure–property relationship research based on density functional theory (DFT) calculation supported the above conclusion. The adsorption mechanism analysis manifested that IPETC adsorption on elemental sulfur surface was a chemical process by separately generating normal covalent bond between carbonyl S atom and S atom and a backdonation covalent bond between O atom and S atom, which was confirmed by the FTIR spectrum analysis result. IPETC exhibits excellent collecting ability and selectivity for elemental sulfur and therefore it has bright application prospects.

Keywords: O-isopropyl-N-ethyl thionocarbamate (IPETC); elemental sulfur; flotation performance; structure–property relationship; adsorption mechanism

1. Introduction

Elemental sulfur is an important chemical material and has been widely used for producing sulfuric acid and other products such as fertilizers, matches, food preservation agents, carbon disulfide, cement, gun powder, surfactants, detergents, pharmaceuticals, pesticides and vulcanized rubbers [1–7]. Elemental sulfur is usually the by-product of hydrometallurgical leaching of nonferrous sulfide ores. For example, oxygen pressure acid leaching of zinc concentrate (sphalerite) can yield large amounts of elemental sulfur [8–14], and the overall reaction is shown in Equation (1).



As indicated, zinc in sphalerite is transformed into zinc sulfate which can be used to produce zinc metal through electrowinning and sulfur remains in the residue in the form

of elemental sulfur [15–18]. The residue is considered to be acid-producing since its sulfur component can react with oxygen and water to generate thiosalts and sulfuric acid [19]. At present, the produced high-sulfur residue is usually subjected to safe storage for most zinc refineries in China, which not only causes serious resource waste but also poses huge threat to the environment and human health [20,21]. Therefore, it is necessary to recover elemental sulfur from the residue with effective technology.

Elemental sulfur has good natural hydrophobicity, and therefore without the use of any additional agent, it can be efficiently enriched by froth flotation, an extensively used technology in mineral processing for separating target mineral from gangue mineral by using their natural hydrophobicity difference [22–25]. However, the nonferrous metal sulfides (ZnS, FeS₂, PbS and Ag₂S) that are not completely oxidized during pressure leaching also remain in the residue [26], and inevitably they will float upward together with elemental sulfur and enter the concentrate because of their similar floatability and intimate association relationship [27,28]. This leads to the decrease of elemental sulfur grade in the concentrate product. To enhance the grade, a hot filtration procedure is usually implemented for the concentrate, which can realize the efficient separation of elemental sulfur from the concentrate by melting and filtration at 115–155 °C [19]. Nevertheless, a significant portion of the elemental sulfur is still hard to separate from the concentrate with hot filtration since this process generally needs raw material whose elemental sulfur content is more than 70% to ensure a satisfactory performance [26]. The elemental sulfur content in most of the high-sulfur residue is usually in 40–60% [15], and thus it is difficult to obtain a high elemental sulfur recovery. A feasible method to solve the above problem is to add suitable collector during the flotation procedure to realize the selective flotation of elemental sulfur.

The study aims to find an efficient collector to selectively separate elemental sulfur from a high-sulfur pressure acid leaching residue of zinc sulfide concentrate. First, the flotation performance of O-Isopropyl-N-Ethyl thionocarbamate (IPETC), ammonium dibutyl dithiophosphate (ADDTP) and sodium ethyl xanthate (SEX) for elemental sulfur were compared through a raw ore flotation experiment. Then, pure mineral flotation and adsorption tests were performed to further compare their flotation properties. After that, the structure-activity relationships of three collectors were researched using density functional theory (DFT) calculation through comparing their geometric configurations, electronic structures and corresponding flotation performances. Based on the above results, IPETC was selected as the optimum elemental sulfur collector, and its adsorption mechanism on the elemental sulfur surface was investigated using adsorption simulation calculation and Fourier-transform infrared spectroscopy (FTIR) analysis.

2. Experimental

2.1. Material and Reagents

The pressure acid leaching residue of zinc sulfide concentrate used in this study was provided by Hulun Buir Chihong Mining Industry Co., Inner Mongolia, China, and it has been used in our previous work [15]. The chemical composition analysis showed that the sulfur content arrived at 46.21%, and 81.97% of the sulfur occurred in the form of elemental sulfur. Therefore, it was a high-sulfur residue and used for closed-circuit flotation experiment to recover elemental sulfur.

Previous literature showed that natural minerals and minerals after leaching have similar surface properties [29,30], and thus natural minerals are usually adopted as the raw materials for mineral flotation study. According to the result of the X-ray diffraction (XRD) spectra analysis [15], elemental sulfur, sphalerite, pyrite, chalcopyrite, albite and anglesite were the main minerals in the residue, and therefore these pure mineral samples obtained from Guangdong Province in China were chosen for the pure mineral flotation experiment. The high-purity minerals were first collected by handpicking, and then crushed and ground in a porcelain ball mill. Afterwards, the samples were dry-sieved to

obtain the $-74 + 38 \mu\text{m}$ fractions for the flotation experiment, and the fractions of less than $5 \mu\text{m}$ were used for FTIR spectrum measurement. The X-ray fluorescence (XRF) analysis results for the samples in Tables 1–3 showed that their purities were over 96%, which could meet the desirable requirement in this study.

In order to investigate the difference of flotation property of pure minerals and pure minerals after oxygen pressure acid leaching, a comparative experiment was conducted, and the result is shown in Table 4. The method of the flotation test was the same as that in Section 2.2.1.2 Pure Mineral Flotation, and the collector (IPETC) dosage and pulp pH value are $3 \times 10^{-5} \text{ mol/L}$ and 8.0, respectively. The recoveries of pure minerals of elemental sulfur, sphalerite, pyrite, chalcopyrite, albite and anglesite were 99.2%, 51.6%, 78.3%, 82.5%, 23.6% and 13.3%, respectively, while the recoveries of the above pure minerals after leaching were 99.0%, 51.3%, 78.2%, 82.1%, 23.4% and 13.5%, respectively. Thus, a very close flotation recovery was obtained for each pure mineral and pure mineral after leaching, which can be ascribed to their close floatability. Therefore, it can be deduced that the physical and chemical properties of the surfaces of the pure minerals and pure minerals after leaching are very close.

The analytically pure O-Isopropyl-N-Ethyl thionocarbamate (IPETC), ammonium dibutyl dithiophosphate (ADDTP) and sodium ethyl xanthate (SEX) were bought from a commercial company of Hunan Province, China, and used as the collector. Methyl isobutyl carbinol (MIBC) from the same company was used as the foaming agent. The other reagents used are also analytically pure. Ultrapure water was used throughout all experiments.

Table 1. Chemical compositions of pure minerals of elemental sulfur, sphalerite, pyrite and chalcopyrite (wt %).

Mineral	Cu	Pb	Zn	Fe	S	Purity
Elemental sulfur	----	----	----	0.02	98.68	98.68
Sphalerite	----	0.04	66.20	0.52	32.9	98.65
Pyrite	----	0.21	0.07	45.25	53.26	97.21
Chalcopyrite	34.13	0.01	0.27	30.10	34.22	98.76

Note: “----” indicates that the element was not detected.

Table 2. Chemical composition of pure albite mineral (wt %).

Mineral	Na ₂ O	Al ₂ O ₃	SiO ₂	Purity
Albite	11.36	19.19	67.86	96.27

Table 3. Chemical composition of pure anglesite mineral (wt %).

Mineral	PbO	SO ₃	Fe ₂ O ₃	Purity
Anglesite	69.32	28.23	0.052	96.27

Table 4. The flotation results of pure minerals and pure minerals after oxygen pressure acid leaching whose conditions are the following: initial sulfuric acid concentration, 60 g/L, temperature, 150 °C, oxygen partial pressure, 1.0 MPa and time, 2.5 h.

Raw Material	Recovery (%)					
	Elemental Sulfur	Sphalerite	Pyrite	Chalcopyrite	Anglesite	Albite
Pure minerals	99.2	51.6	78.3	82.5	23.6	13.3
Pure minerals after oxygen pressure acid leaching	99.0	51.3	78.2	82.1	23.4	13.5

2.2. Experimental Methods

2.2.1. Flotation Experiment

2.2.1.1. Raw Ore Flotation

The flotation experiment of pressure acid leaching residue was carried out in a self-aeration XFD-63 flotation machine from Jilin Prospecting Machinery Factory, China [31]. The residue and ultrapure water were first put into the cell to form the pulp with 25% density, and then slime was added to adjust the pulp pH to about 8.0. Afterwards, the inhibitor ($\text{Na}_2\text{S} + \text{Na}_2\text{SO}_3 + \text{ZnSO}_4$), collector and foaming agent (MIBC) were successively added into the pulp which was agitated at 1650 rpm for 3 min after adding each reagent. The flotation was conducted for 5 min, and elemental sulfur concentrate and tailing were collected, filtrated, dried, weighed and assayed to calculate elemental sulfur recovery.

2.2.1.2. Pure Mineral Flotation

The pure mineral flotation experiment was performed in an XFG5-35 flotation machine (Jilin Prospecting Machinery Factory, Jilin, china) with a 40 mL plexiglass cell [32]. The agitating speed was kept at 1650 r min^{-1} with a mechanical impeller. In each test, 2.0 g pure mineral sample was placed into an ultrasonic bath for 5 min of cleaning, and then transferred into the flotation cell. After that, the pulp pH was adjusted to a desired value with dilute NaOH or H_2SO_4 solution. Finally, the collector and the foaming agent were sequentially added into the pulp which was agitated for 2 min. The flotation was carried out for 5 min, and the obtained concentrate and tailing were weighted and analyzed after filtration and drying to calculate mineral recovery.

2.2.2. FTIR Spectrum

IPETC solution ($1 \times 10^{-2} \text{ mol/L}$) was first mixed with elemental sulfur sample (0.5 g, $-38 \mu\text{m}$). After magnetically stirring for 60 min, the solution was filtered, and the obtained solid product was dried at 35 °C under vacuum for 24 h for subsequent infrared detection. The Fourier-transform infrared (FTIR) spectra were recorded by a G510PFTIR infrared spectrometer (Nicolet Company, Madison, USA) in the wavenumber range from 400 cm^{-1} to 4000 cm^{-1} using the KBr technique.

2.2.3. DFT Calculation

The initial molecular/ion model of collector was constructed and then optimized by MM2 and PM3 methods in Chemoffice 2008 program. The obtained geometry was further optimized by DFT method at the B3LYP/6-31G (d, p) level in the Gaussian 03 program [31]. Finally, periodic structure optimization was completed in the CASTEP module of Materials Studio 4.4 and then the optimized structure was subjected to DFT calculations.

The calculations of mineral plane and crystal were completed by CASTEP and Dmol3 modules in Materials Studio 4.4 program, in which CASTEP module was used for the establishment of plane and crystal models as well as geometric structure optimization, state density analysis and Mulliken population analysis, and Dmol3 module was used to analyze frontier orbital energy. The plane wave basis set and DND basis set were separately implemented in the CASTEP and Dmol3 modules for the DFT calculations. The establishment and optimization of adsorption model of IPETC on elemental sulfur surface

and related calculation were both performed in the CASTEP module of Materials Studio 4.4.

2.3. Analytical Methods

The elemental sulfur content in the high-sulfur residue and flotation concentrate was detected by carbon tetrachloride dissolution followed by gravimetric analysis [33]. The concentrations of sulfhydryl collectors were determined by UV-Vis spectroscopy [34].

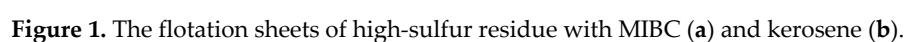
3. Results and Discussion

3.1. Flotation Performance of Collectors

Before the use of three collectors IPETC, ADDTP and SEX, single foaming agent MIBC and non-polar collector kerosene were separately used for recovering elemental sulfur from the high-sulfur residue. The flotation sheets are displayed in Figure 1a,b, and the flotation results are shown in Table 5. It can be seen that a low and close elemental sulfur recovery was obtained for MIBC and kerosene. Furthermore, the elemental sulfur grades in the concentrate were low, which is because large amounts of sulfide minerals such as sphalerite, pyrite, chalcopyrite in the residue entered the concentrate with elemental sulfur. The above results indicated that the selective and effective separation of elemental sulfur from the high-sulfur residue cannot be realized using MIBC and kerosene.

The flotation performances of IPETC, ADDTP and SEX collectors for elemental sulfur in the high-sulfur residue were compared. After detailed conditional experiment in our laboratory, the closed-circuit flotation flowchart in Figure 2 was ascertained and adopted, and the obtained results are listed in Table 6. Despite the high pulp pH, the collectors will also adsorb a small amount of sphalerite. Thus, zinc sulfate is added to increase the inhibitory effect on sphalerite, even though some zinc ions are present in the pulp itself. It is not necessary to inhibit the gangue minerals in the residue, such as albite and anglesite, because they have poor natural floatability and the sulfhydryl collector (i.e., IPETC) hardly adsorbs on their surfaces.

From Table 6, the elemental sulfur recoveries for SEX, ADDTP and IPETC were 93.67%, 97.83% and 99.87% while their elemental sulfur grades in the concentrate were 81.53%, 82.25% and 84.47%, respectively. Thus, both the maximum elemental sulfur recovery and grade were achieved using IPETC collector. The zinc flotation result was also shown in Table 6. The zinc grades in the elemental sulfur concentrate were 1.33%, 1.15% and 0.88% for SEX, ADDTP and IPETC, and their recoveries were 13.33%, 12.08% and 9.06%, respectively. Thus, compared with elemental sulfur, only a small amount of zinc was floated and entered the concentrate. Moreover, the zinc grade in the concentrate and its recovery were the lowest for IPETC. The above results indicated that in comparison with SEX and ADDTP, IPETC exhibited more excellent collecting ability and selectivity for elemental sulfur in the high-sulfur residue, and thus it has good industrial application potential.



Collector	Product	Yield (%)	Elemental Sulfur Grade (%)	Elemental Sulfur Recovery (%)
MIBC	Elemental sulfur concentrate	57.31	71.17	88.83
	Tailing	42.69	12.02	11.17
	Feed residue	100.00	45.92	100.00
Kerosene	Elemental sulfur concentrate	56.70	71.24	88.00
	Tailing	43.30	12.72	12.00
	Feed residue	100.00	45.90	100.00

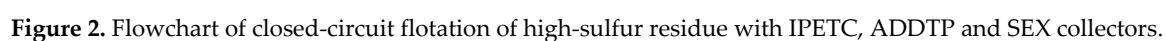


Table 6. The result of closed-circuit flotation of the high-sulfur residue with IPETC, ADDTP and SEX collectors.

Collector	Product	Yield (%)	Grade (%)		Recovery (%)	
			Elemental Sulfur	Zinc	Elemental Sulfur	Zinc
SEX	Elemental sulfur concentrate	43.55	81.53	1.33	93.67	13.33
	Tailing	56.45	4.25	6.67	6.33	86.67
	Feed residue	100.00	37.91	4.34	100.00	100.00
ADDTP	Elemental sulfur concentrate	45.47	82.25	1.15	97.83	12.08
	Tailing	54.53	1.52	6.98	2.17	87.92
	Feed residue	100.00	38.23	4.33	100.00	100.00
IPETC	Elemental sulfur concentrate	44.78	84.47	0.88	99.87	9.06
	Tailing	55.22	0.09	7.16	0.13	90.94
	Feed residue	100.00	37.88	4.35	100.00	100.00

3.2. Pure Mineral Flotation and Adsorption Experiments

3.2.1. Pure Mineral Flotation Experiment

The collecting performances of IPETC, ADDTP and SEX for pure minerals of elemental sulfur, sphalerite, pyrite, chalcopyrite, albite and anglesite were compared, and the results of effects of collector dosage and pulp pH on pure mineral recovery are displayed in Figures 3a–c and 4a–c.

As shown in Figure 3a–c, the recoveries of elemental sulfur, sphalerite, pyrite and chalcopyrite went up with the increase of IPETC, ADDTP and SEX dosages in their initial ranges of $(2-3) \times 10^{-5}$, $(0.4-1.2) \times 10^{-4}$ and $(0.4-1.2) \times 10^{-4}$ mol/L. Afterwards, no obvious increases of the recoveries of four pure minerals were observed with the further increase of collector dosage. Therefore, the optimal IPETC, ADDTP and SEX dosages were separately 3×10^{-5} , 1.2×10^{-4} and 1.2×10^{-4} mol/L. The recoveries of albite and anglesite were low in the whole abscissa range, and collector dosage had little effect on them, which could be ascribed to their poor natural floatability.

As indicated in Figure 4a–c, the recoveries of elemental sulfur and three sulfides were all augmented as the pulp pH increased in the initial range of 2–8. Nevertheless, further increase of pulp pH from 8 to 12 resulted in the decrease of the recoveries, but the decrease degrees of recoveries of three sulfides were larger than that of elemental sulfur, suggesting that increased pH had less negative impact on elemental sulfur flotation in this pH range. Thus, the optimal pulp pH for the four minerals was 8. The pulp pH exerted similar effect on anglesite recovery, but its optimum pH was 6 where the maximum recovery was achieved. The albite recovery gradually dropped in the pulp pH range of 2–8 and then basically remained steady.

In addition, for each collector, both the maximum elemental sulfur recovery and the differences between the maximum recoveries of elemental sulfur and the other five pure minerals followed the order of IPETC > ADDTP > SDD. Therefore, among the three collectors, IPETC presented the optimal collecting ability and selectivity to elemental sulfur.

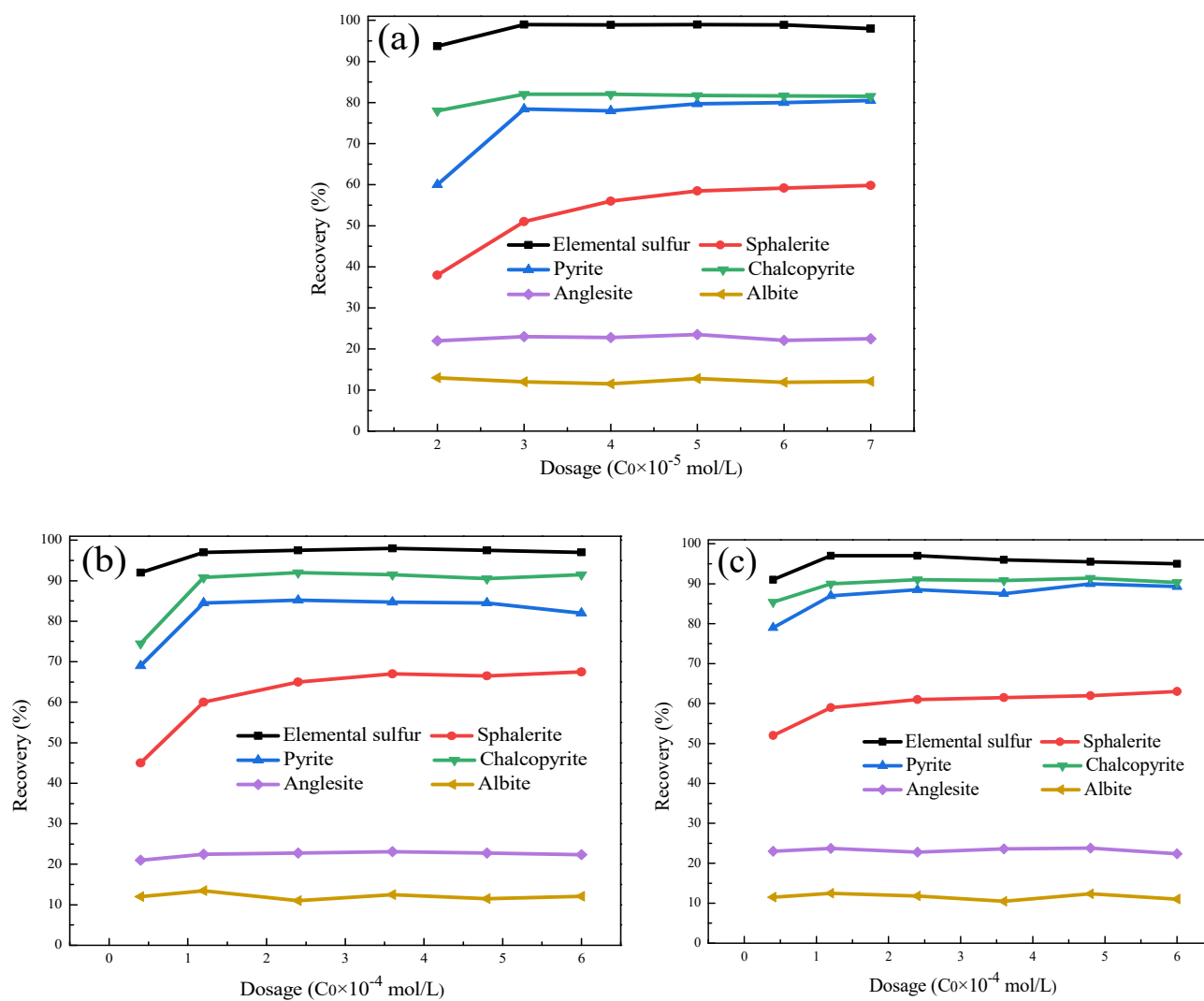


Figure 3. Effect of collector dosage on pure mineral recovery at pulp pH value 8 ((a) IPETC; (b) ADDTP; (c) SEX).

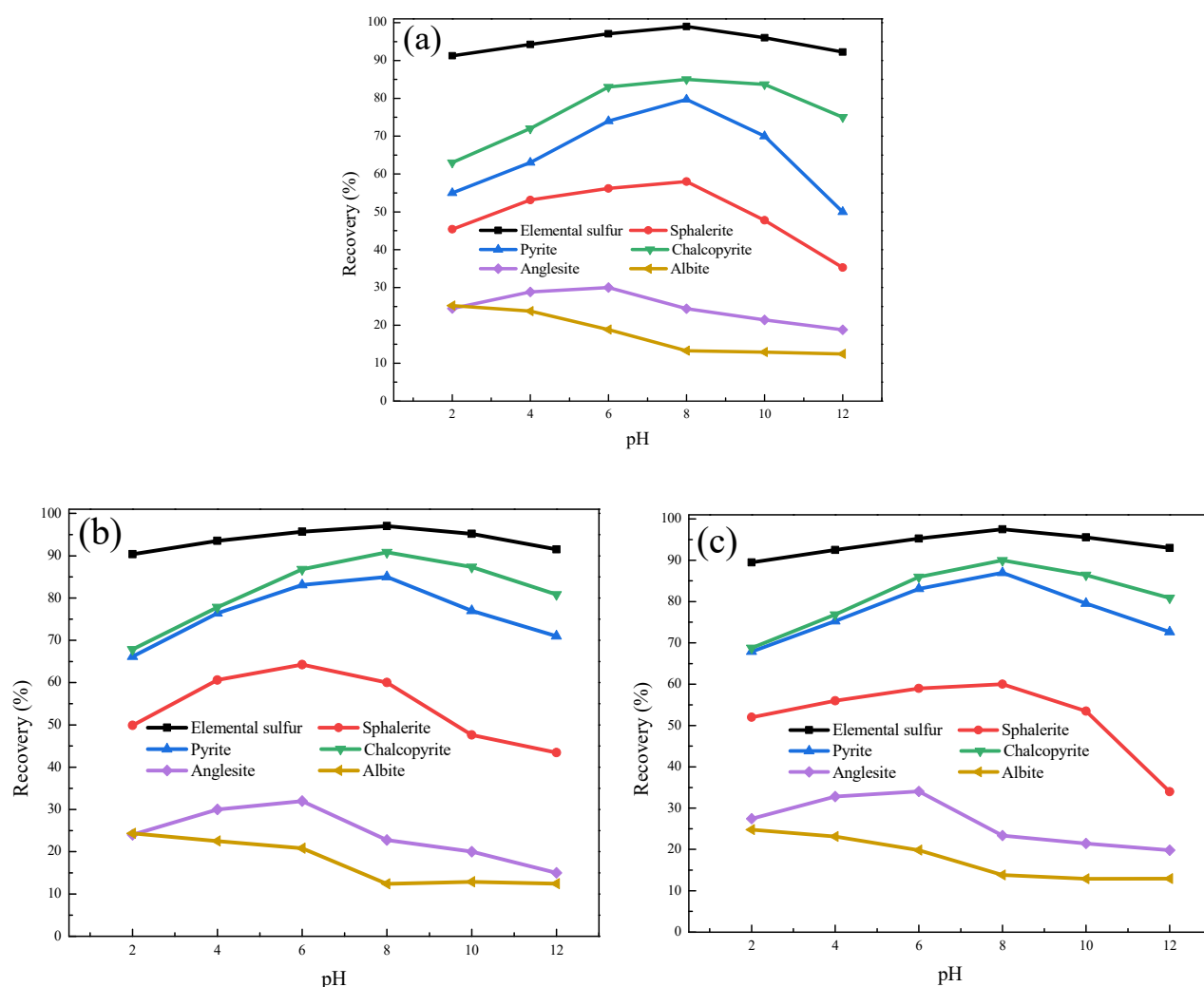


Figure 4. Effect of pulp pH on pure mineral recovery at optimum collector dosage ((a) IPETC; (b) ADDTP; (c) SEX).

3.2.2. Pure Mineral Adsorption Experiment

The adsorption amounts of IPETC, ADDTP and SEX on pure mineral surfaces were also compared, and the effect of initial collector concentration is shown in Figure 5a–c. With the increase of concentrations of three collectors, their adsorption amounts on the surfaces of elemental sulfur, albite and anglesite basically remained steady, but the adsorption amount on elemental sulfur surface kept at a high level while those on the surfaces of albite and anglesite were very low. For sphalerite, pyrite and chalcopyrite, the adsorption amounts of three collectors on their surfaces first increased and then were roughly unchanged.

However, among the three collectors, the adsorption amount of IPETC on elemental sulfur surface was the biggest under the same collector concentration, suggesting that IPETC possessed the strongest interaction with elemental sulfur surface. Moreover, IPETC exhibited the greatest difference between adsorption quantities on elemental sulfur surface and the surfaces of other pure minerals. The above results indicated that IPETC not only possessed the biggest adsorption amount on elemental sulfur surface but also displayed the best adsorption selectivity for elemental sulfur, which is consistent with the result of the pure mineral flotation test in Section 3.2.1.

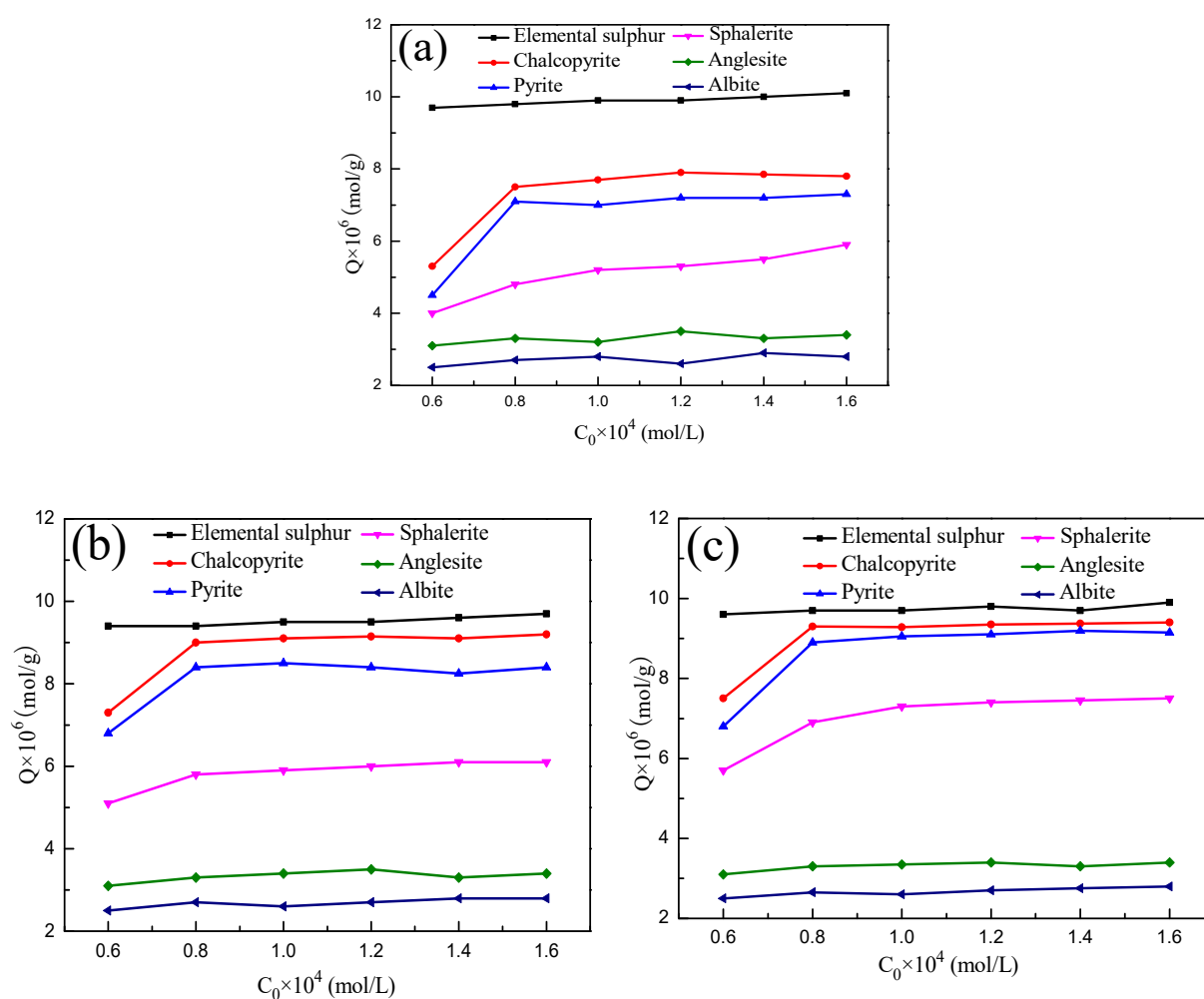


Figure 5. Effect of initial collector concentration on their adsorption amounts on pure mineral surfaces at pulp pH 8 ((a) IPETC; (b) ADDTP; (c) SEX).

3.3. Structure–Property Relationships of Collectors

3.3.1. Geometry Configuration

The simulation calculations of collector molecules/ion were carried to study their structure–property relationships. The optimized geometry configurations of ADDTP, SEX and IPETC are presented in Figure 6a–c, and the bond lengths and dihedral angles are displayed in Table 7.

Table 7. Bond lengths and dihedral angles of collectors.

Collector	Bond Length (Å)		Dihedral Angle (°)	
	1C-10S	1C-11S	9O-1C-10S-11S	
SEX	1.707	1.699	−179.992	
	29S-31P	30S-31P	14C-27O-31P-29S	9C-26O-31P-30S
	2.006	2.006	−59.008	−59.003
IPETC	8C-13S	13S-8C-11N-2C	11N-8C-14O-9C	
	1.679	−178.654	177.847	

The bond length of carbonyl S atom in SEX is longer than that in IPETC, indicating that carbonyl S in IPETC is less likely to lose electrons. As a result of this, normal covalent bond is easier to be generated in SEX, which is not unbeneficial to its selectivity. Therefore, IPETC presents a better collecting selectivity for elemental sulfur in terms of bond length

of carbonyl S atom. The carbonyl S atom in ADDTP bonds with the P atom to form a P=S bond which is different with C=S bond in SEX and IPETC, and thus their bond lengths are not comparable.

Judging from the dihedral angle of three collectors, the atoms of -S-C=N-C- and -N-C=O-C- groups in IPETC and -O-C(=S)-S- group in SEX are nearly located in the same plane, respectively. Furthermore, the p orbitals that are not involved in hybridization exist in the valence electron layers of these atoms. Thus, conjugated big π -bond is easy to be formed in the three groups.

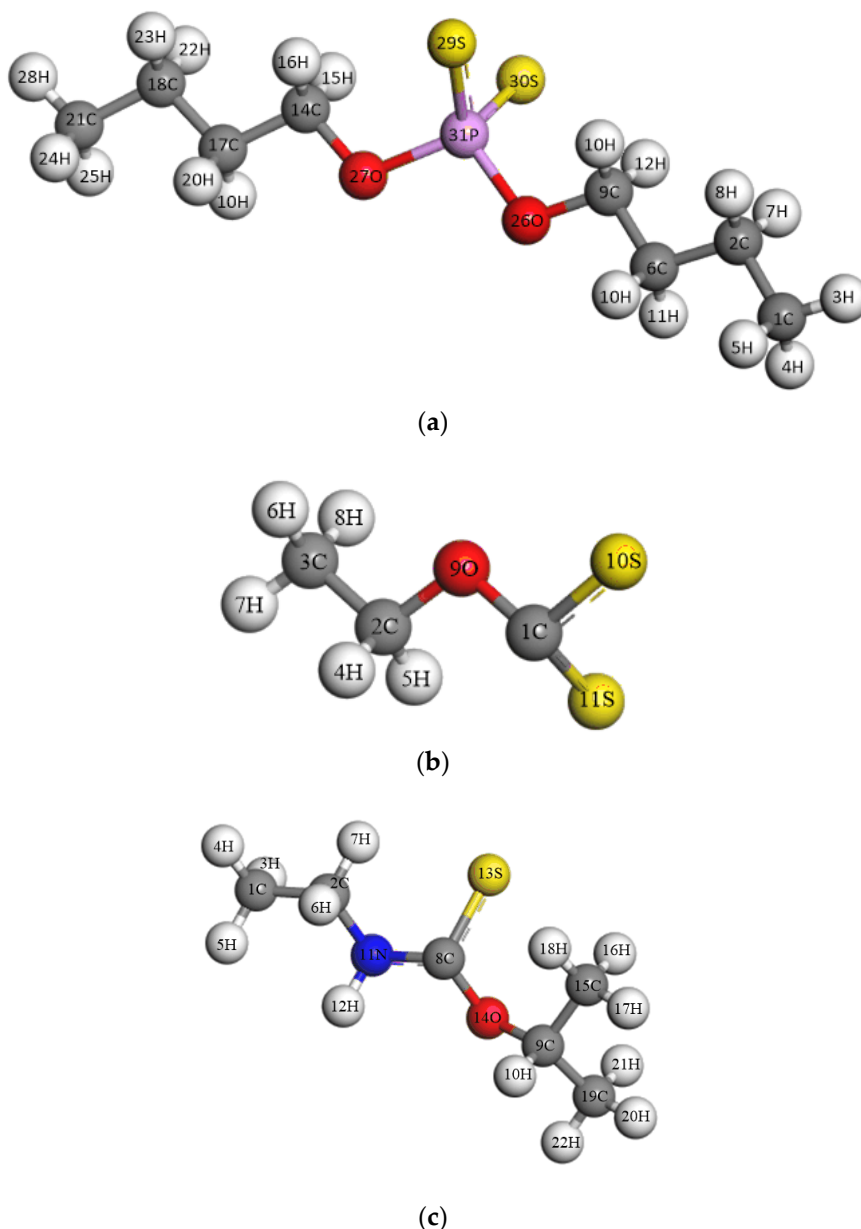


Figure 6. Optimized geometry configurations of collector molecules or ions ((a) IPETC; (b) ADDTP; (c) SEX).

3.3.2. Electronic Structure

3.3.2.1. Mulliken Population Analysis

The Mulliken populations of bond and atom charge of SEX, ADDTP and IPETC are indicated in Table 8. The sequence of bond populations of C=S or P=S bonds in the three collectors is IPETC > ADDTP > SEX. A larger population means stronger covalency of the

covalent bond. Thus, in comparison with IPETC, it is more likely for the carbonyl S atoms in ADDTP and SEX to offer electrons, leading to the easier generation of a normal covalent bond between collector and mineral surface. Therefore, in terms of the bond population, IPETC shows better selectivity to elemental sulfur.

The absolute values of Mulliken charge populations of carbonyl S atoms follow the order of ADDTP > SEX > IPETC. A larger absolute value means stronger electrostatic attraction between carbonyl S atom in collector and S atom in elemental sulfur surface. Nevertheless, electrostatic interaction has no directivity, and therefore the larger absolute value is unfavorable to the collector selectivity. Thus, in terms of atom charge population, IPETC also exhibits better selectivity for elemental sulfur than ADDTP and SEX.

Table 8. Mulliken populations of C=S or P=S bond and atom charge of collectors.

Collector	Bond Population			Atom Charge Population		
SEX	10S-1C	11S-1C	1C	9O	10S	11S
	0.41	0.31	−0.099	−0.174	−0.379	−0.405
ADDTP	29S-31P	30S-31P	27O	29S	30S	31P
	0.72	0.74	−0.729	−0.759	−0.811	1.521
IPETC	8C-13S		8C	11N	13S	14O
	1.00		0.067	0.022	−0.206	−0.238

3.3.2.2. Density of States Analysis

The densities of states of three collectors are indicated in Figure 7a–c. The valence band tops of SEX, ADDTP and IPETC are mainly comprised of 3p orbitals of 10S and 11S, 29S and 30S, and 13S atoms, respectively. Their conduction band bottoms are primarily comprised of 2p orbitals of 1C and 9O atoms and 3p orbitals of 10S and 11S atoms, 2p orbitals of 26O, 27O and 31P atoms and 3p orbitals of 29S and 30S atoms, and 2p orbitals of 8C, 11N and 14O atoms and 3p orbital of 13S atom, respectively. According to the band theory, the valence band top is the highest occupied molecular orbital (HOMO), and the conduction band bottom is the lowest unoccupied molecular orbital (LUMO), and they have the highest chemical activity and are separately apt to lose and gain electrons. Therefore, it can be concluded that the oxidation reaction centers of SEX, ADDTP and IPETC are separately 10S and 11S, 29S and 30S, and 13S atoms. Accordingly, the reduction reactions can occur on 1C, 9O, 10S and 11S atoms for SEX, 26O, 27O, 29S, 30S and 31P atoms for ADDTP, and 8C, 11N, 13S and 14O atoms for IPETC, respectively.

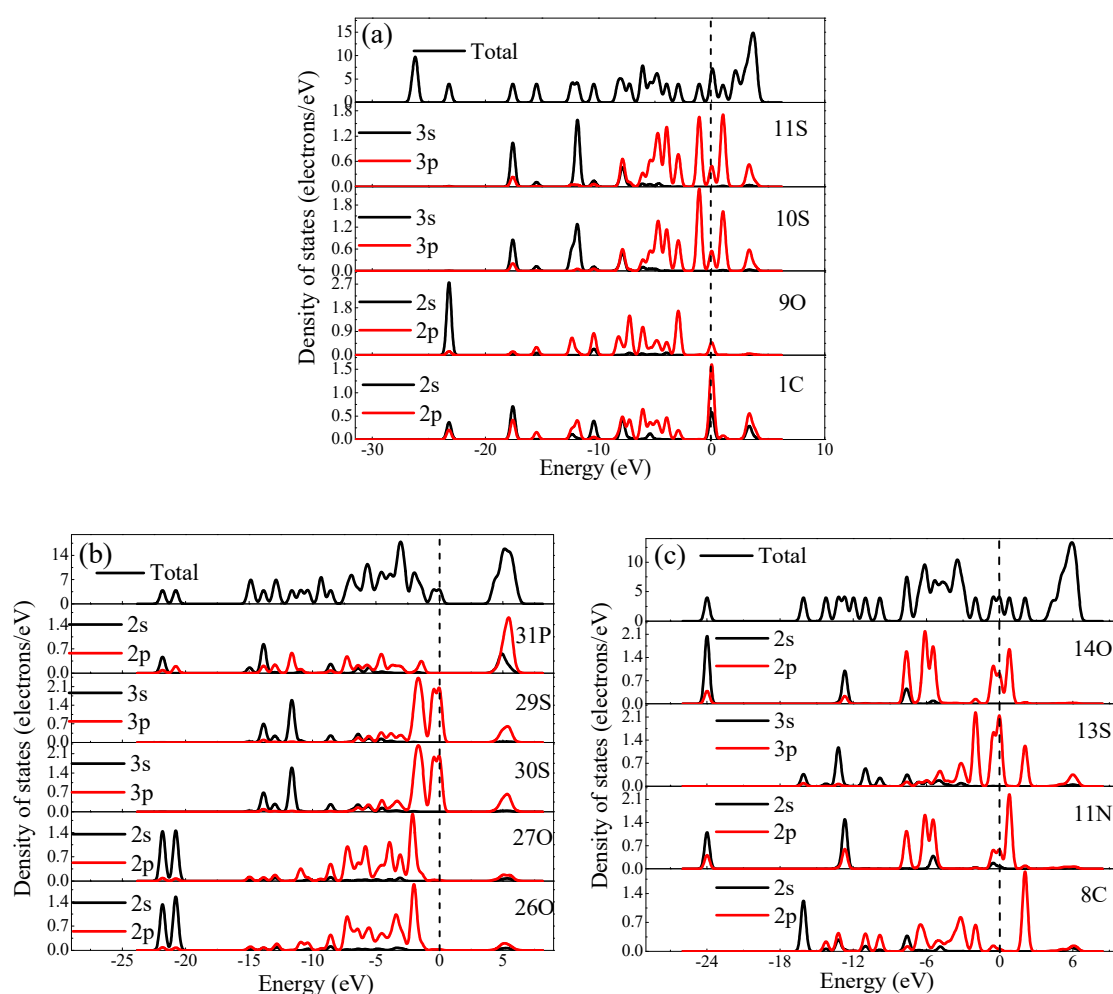


Figure 7. Densities of states of collectors ((a) IPETC; (b) ADDTP; (c) SEX).

3.3.2.3. Frontier Orbital Analysis

The electronic cloud pictures of frontier molecular orbitals of three optimized SEX, ADDTP and IPETC models are given in Figure 8a–f. From these subgraphs, the HOMOs of SEX, ADDTP and IPETC are mainly composed of electron orbitals of 10S and 11S, 29S and 30S, and 13S atoms, respectively. The LUMOs are primarily constituted by electron orbitals of 1C, 9O, 10S and 11S atoms for SEX, 26O, 27O, 29S, 30S and 31P atoms for ADDTP, and 8C, 11N, 13S and 14O atoms for IPETC, respectively. The above results are consistent with the results of density of states analysis in section 3.3.2.2. Furthermore, the atoms of -S-C=N-C- and -N-C=O-C- groups in IPETC and -O-C(=S)-S- group in SEX are all almost in the same plane based on the dihedral angle values in Table 7. Moreover, it can be known from the result of density of states analysis in section 3.3.2.2 that the LUMO orbitals of IPETC and SEX all constitute the p orbital of each atom in these functional groups. Therefore, it can be inferred that the LUMO orbitals of IPETC and SEX both are conjugated big π -bonds formed by electron orbitals of these functional groups.

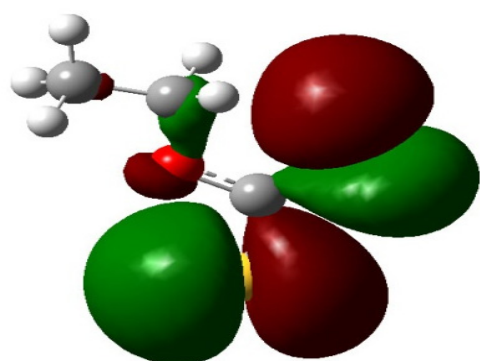
The frontier orbital energies of three collectors and energy differences between collectors and minerals (ΔE) are shown in Table 9. According to frontier molecular orbital theory, a smaller absolute value of energy difference between the HOMO of one reactant and the LUMO of another one means that they are prone to react with each other. ΔE_1 are all evidently smaller than ΔE_2 , suggesting that the chemical reaction between collectors and minerals is led by the electron transfer from collectors to minerals. Thus, the normal covalent bond predominates in the interaction. ΔE_1 between each collector and elemental

sulfur is smaller than that between each collector and chalcopyrite/sphalerite/pyrite/anglesite/albite, indicating that the normal covalent bond between each collector and elemental sulfur is stronger than that between each collector and five other minerals. In comparison with SEX and ADDTP, both ΔE_1 between IPETC and elemental sulfur and ΔE_1 elemental sulfur/ ΔE_1 five other minerals of IPETC are the minimum. Thus, IPETC displays the optimal collecting power and selectivity for elemental sulfur among the three collectors.

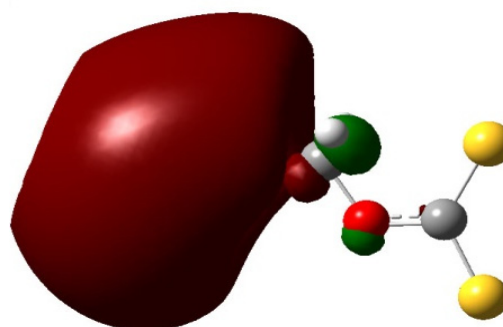
Table 9. Frontier orbital energies of collectors and energy differences between collectors and minerals.

Collector	Orbital Energy (eV)	ΔE^a	Elemental Sulfur	Chalcopyrite	Pyrite	Sphalerite	Anglesite	Albite
SEX	HOMO	−4.57	ΔE_1	0.39	0.64	0.78	1.54	1.69
	LUMO	−3.07	ΔE_2	2.55	2.60	2.21	2.38	2.44
ADDTP	HOMO	−5.56	ΔE_1	0.32	0.51	0.53	0.73	0.95
	LUMO	−4.67	ΔE_2	2.42	1.00	1.37	2.04	2.31
IPETC	HOMO	−5.18	ΔE_1	0.15	0.76	1.04	1.43	1.51
	LUMO	−3.76	ΔE_2	1.91	1.86	1.77	1.91	1.97

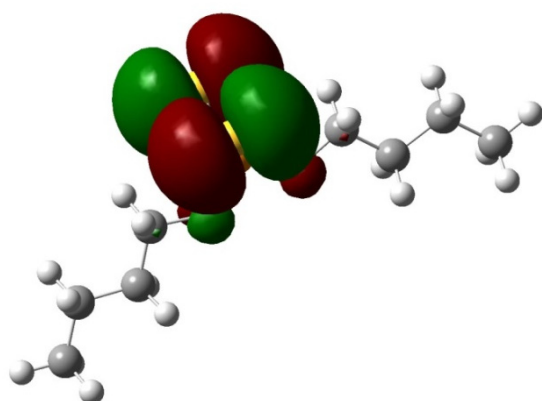
^a $\Delta E_1 = |E(\text{HOMO, collector}) - E(\text{LUMO, mineral})|$ and $\Delta E_2 = |E(\text{HOMO, mineral}) - E(\text{LUMO, collector})|$.



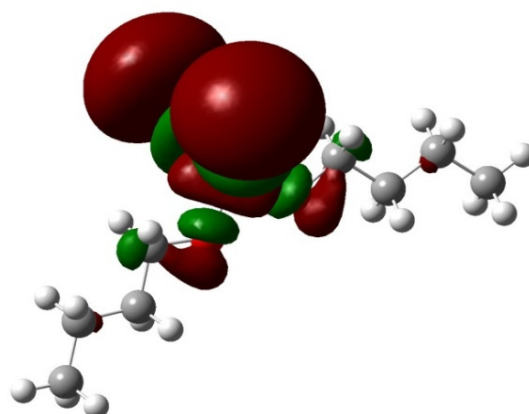
(a)



(b)



(c)



(d)

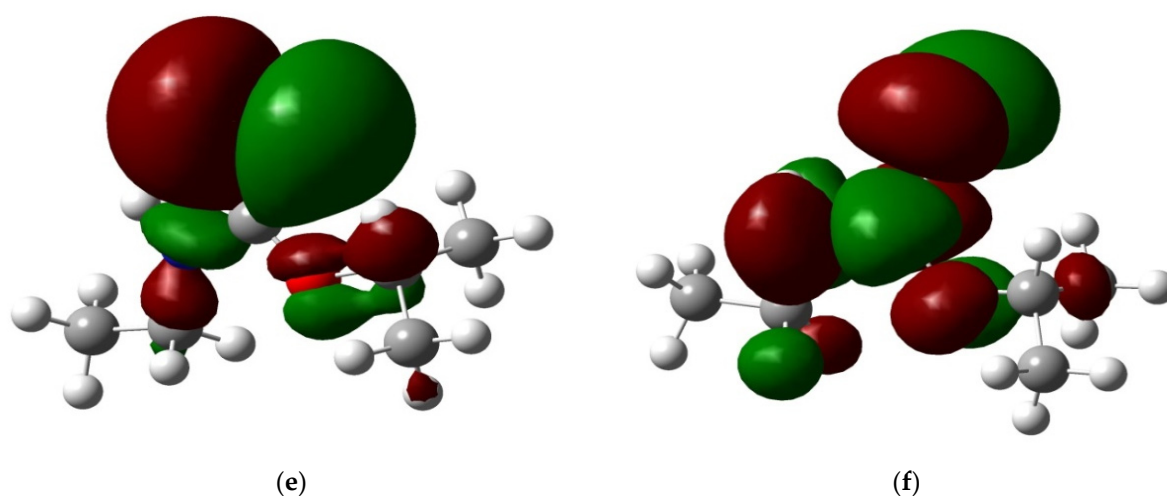


Figure 8. Electron cloud pictures of frontier molecular orbitals of optimized collector models ((a) HOMO of SEX; (b) LUMO of SEX; (c) HOMO of ADDTP; (d) LUMO of ADDTP; (e) HOMO of IPETC; (f) LUMO of IPETC).

3.4. Mechanism of IPETC Adsorption on Elemental Sulfur Surface

3.4.1. Adsorption Configuration and Adsorption Energy

After many adsorption position tests, six possible configurations of IPETC adsorption on perfect elemental sulfur (110) plane are indicated in Figure 9a–f, and the adsorption energies are shown in Table 10. The adsorption energies of six configurations are all negative, and therefore these adsorption reactions can happen spontaneously. Nevertheless, the adsorption energy is the lowest for the configuration of simultaneous adsorption of carbonyl S together with O, suggesting that this configuration is the most stable.

The adsorption energies of six configurations of IPETC adsorption on the perfect crystal planes of chalcopyrite (112), sphalerite (110), pyrite (100), anglesite (001) and albite (001) are displayed in Table 11. Obviously, the most stable adsorption configuration of IPETC on the five crystal planes is also the simultaneous adsorption of carbonyl S together with O. Nevertheless, the adsorption energies for the first four minerals are negative and obey the sequence of chalcopyrite < pyrite < sphalerite < anglesite, which are all larger than that for elemental sulfur in Table 9. Thus, it is easier for IPETC adsorption on elemental sulfur surface to occur. The adsorption energies for albite are all positive, and therefore IPETC adsorption on albite surface is difficult, which supported the results of pure mineral flotation and adsorption experiments in Section 3.2 that both the albite recovery and its adsorption amount on albite surface are very low.

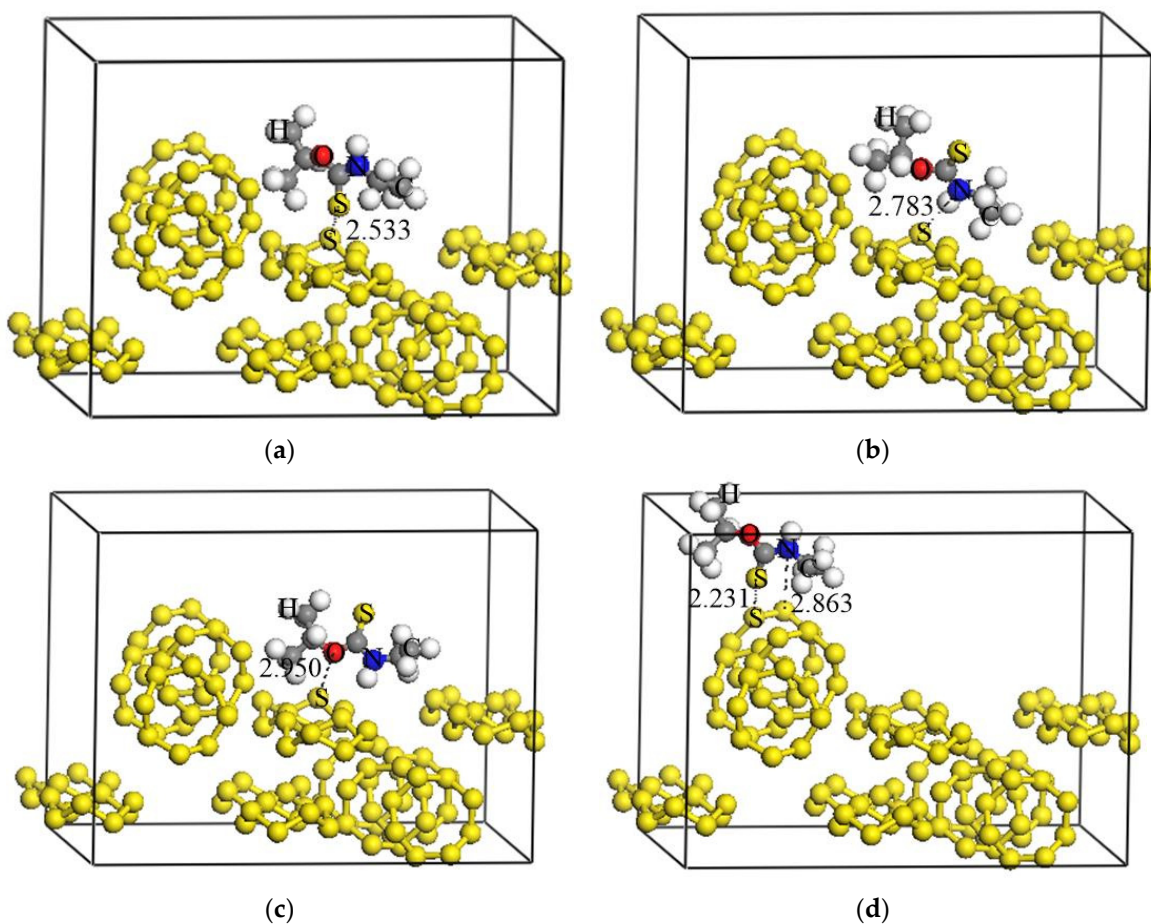
The steadiest adsorption configurations of IPETC on perfect planes of chalcopyrite, pyrite, sphalerite and anglesite are presented in Figure 10a–d, and the Mulliken populations of corresponding adsorption bonds are shown in Table 12. For each configuration, the population of carbonyl S adsorption bond is much bigger than that of O adsorption bond. Thus, the interaction between carbonyl S and bonding atom in mineral surface is stronger.

Table 10. Adsorption energy of each adsorption configuration of IPETC on elemental sulfur (110) crystal plane.

Adsorption Configuration	Adsorption Energy (kJ/mol)
Adsorption of carbonyl S	−19.79
Adsorption of N	−10.23
Adsorption of O	−6.16
Simultaneous adsorption of carbonyl S together with N	−21.82
Simultaneous adsorption of carbonyl S together with O	−26.98
Simultaneous adsorption of N and O	−16.53

Table 11. Adsorption energy of each adsorption configuration of IPETC on each crystal plane.

Adsorption Configuration	Adsorption Energy (kJ/mol)				
	Chalcopyrite (112) Plane	Pyrite (100) Plane	Sphalerite (110) Plane	Anglesite (001) Plane	Albite (001) Plane
Adsorption of carbonyl S	−12.53	−11.12	−10.02	−3.53	1.54
Adsorption of N	−6.88	−3.78	−2.12	−1.38	4.10
Adsorption of O	−9.12	−8.00	−7.79	−3.19	2.16
Simultaneous adsorption of carbonyl S together with N	−8.01	−4.56	−3.45	−2.41	2.33
Simultaneous adsorption of carbonyl S together with O	−16.89	−14.69	−12.21	−4.67	1.23
Simultaneous adsorption of N and O	−5.13	−6.69	−5.63	−0.43	3.49



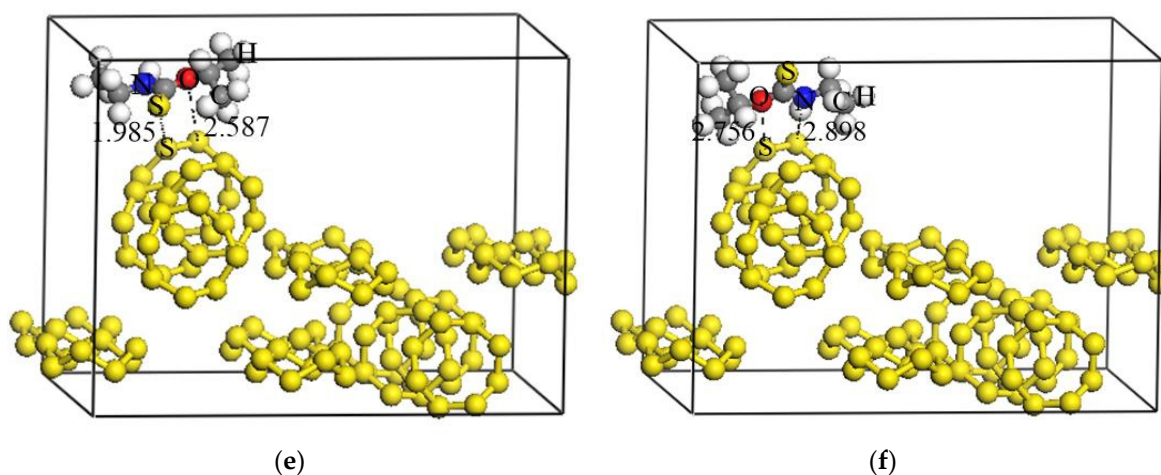


Figure 9. Adsorption configurations of IPETC on perfect elemental sulfur (100) crystal plane ((a) adsorption of carbonyl S atom; (b) adsorption of N atom; (c) adsorption of O atom; (d) adsorption of carbonyl S together with N atoms; (e) adsorption of carbonyl S together with O atoms; (f) adsorption of N together with O atoms).

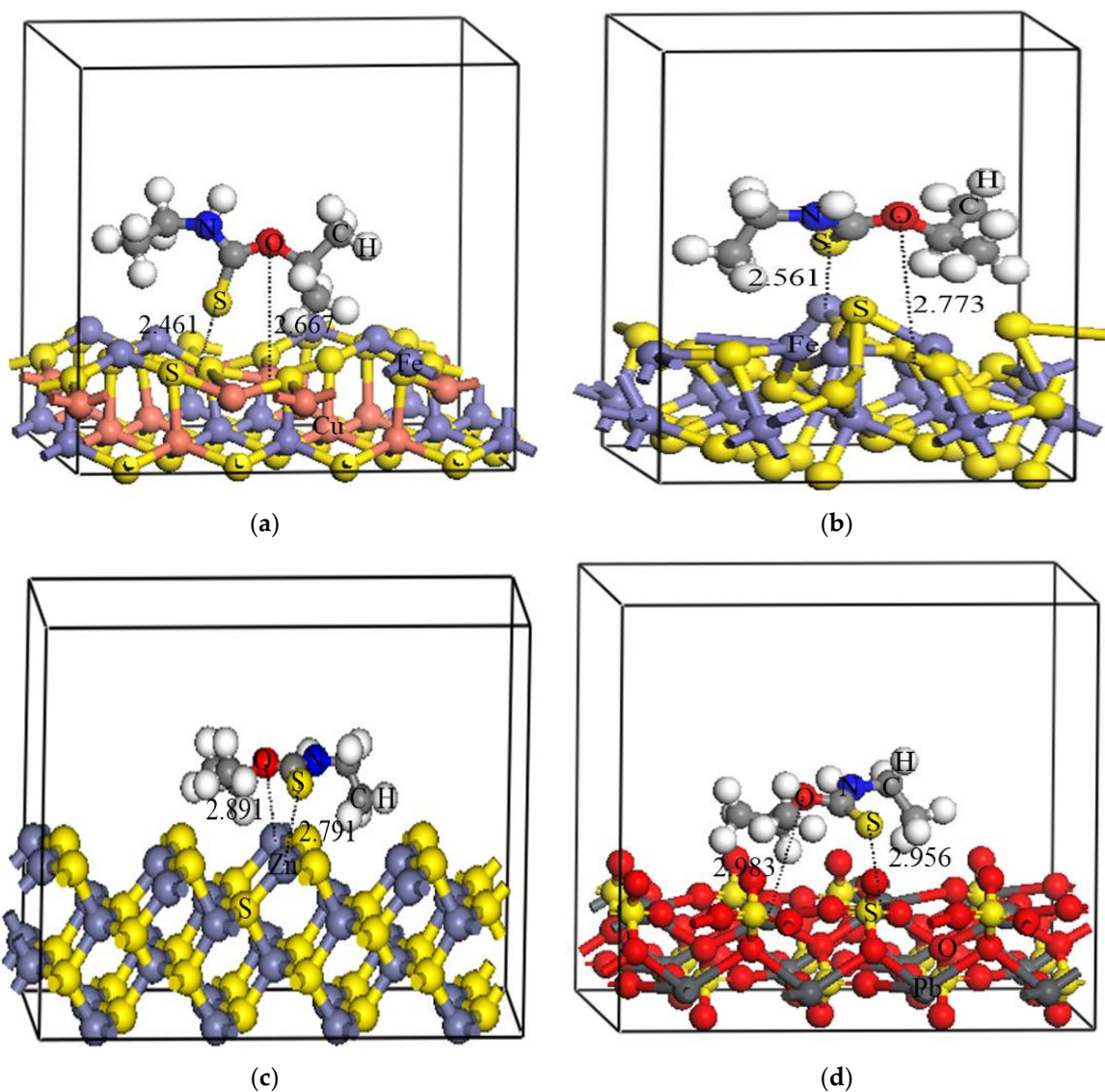


Figure 10. The steadiest adsorption configurations of IPETC on perfect crystal planes of chalcopryite (112) (a), pyrite (100) (b), sphalerite (110) (c) and anglesite (001) (d).

Table 12. Mulliken population of adsorption bonds of each steadiest adsorption configuration of IPETC on each crystal plane.

Crystal Plane	Adsorption Bond	Bond Length (Å)	Mulliken Population
Elemental sulfur (110)	S-S	1.985	0.42
	O-S	2.587	0.15
Chalcopyrite (112)	S-Cu	2.461	0.35
	O-Cu	2.667	0.12
Pyrite (100)	S-Fe	2.561	0.33
	O-Fe	2.773	0.08
Sphalerite (110)	S-Zn	2.791	0.23
	O-Zn	2.891	0.11
Anglesite (001)	S-Pb	2.956	0.19
	O-Pb	2.983	0.12

3.4.2. Mulliken Population and Density of States Analyses

The Mulliken populations and densities of states of bonding atoms before and after IPETC adsorption on perfect elemental sulfur (100) plane are presented in Table 13 and Figure 11, respectively. The 3p orbital of carbonyl S atom of IPETC donates some electrons to the 3p orbital of S atom on elemental sulfur surface to form a normal covalent bond, while the 2p orbital of O atom of IPETC accepts the electrons from the 3p orbital of S atom in elemental sulfur surface to form a backdonation covalent bond.

Table 13. Mulliken charge populations of bonding atoms before and after adsorption.

Bonding Atom	Adsorption State	Electron Charge (e)			Charge Population
		s	p	d	
O _(O-S)	Before adsorption	1.77	4.66	0.00	−0.44
	After adsorption	1.77	4.71	0.00	−0.46
S _(O-S)	Before adsorption	1.89	4.19	0.00	−0.01
	After adsorption	1.89	4.12	0.00	−0.07
Carbonyl S _(S-S)	Before adsorption	1.82	4.41	0.00	−0.23
	After adsorption	1.82	4.10	0.00	0.05
S _(S-S)	Before adsorption	1.89	4.10	0.00	0.01
	After adsorption	1.89	4.12	0.00	−0.01

The 2p state of O atom of IPETC slightly shifts to the higher energy direction, which means that it gets some electrons in the reaction. However, there is no evident variation in the 3p state of bonding S atom on elemental sulfur surface, indicating that the interaction between O and S atoms is weak. The 3p state of carbonyl S atom of IPETC moves towards the lower energy direction, which indicates that when the S atom of IPETC reacts with an S atom on elemental sulfur surface, it loses some electrons and its oxidation is strengthened. At the same time, the localization of 3p state of bonding S atom on elemental sulfur surface declines and evident hybridization happens on the 3p orbital of carbonyl S atom and 3p orbital of S atom from −5 eV to 0 eV. Thus, the interaction between carbonyl S atom in IPETC and S atom in elemental sulfur surface is strong.

Based on the above results, it can be concluded that when IPETC interacts with elemental sulfur surface, carbonyl S of IPETC offers some electrons to the S atom on mineral surface to form a normal covalent bond, while O of IPETC accepts some electrons from the S atom to generate a backdonation covalent bond. The backdonation covalent bond is relatively weak, and thus a normal covalent bond plays a leading role.

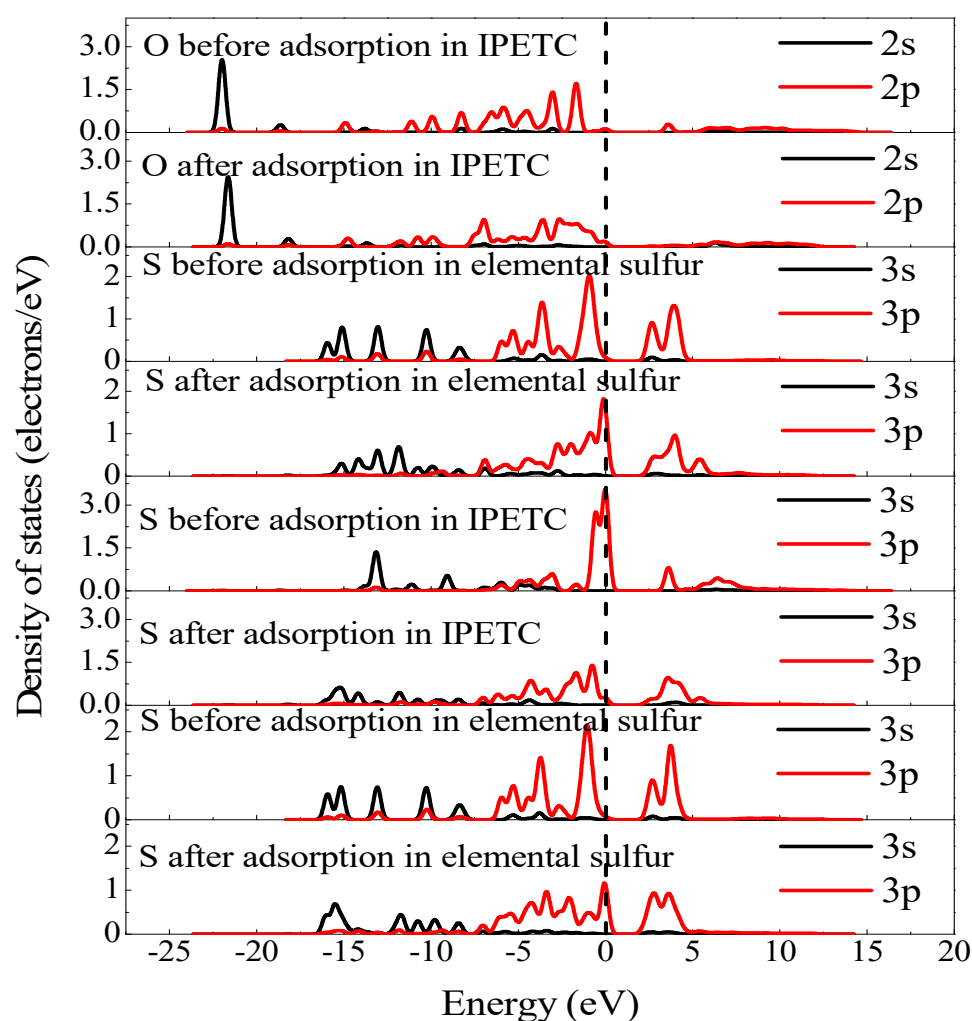


Figure 11. Densities of states of bonding atoms before and after IPETC adsorption on elemental sulfur (100) plane.

3.4.3. FTIR Spectrum Analysis

The infrared spectra of IPETC, elemental sulfur and IPETC-adsorbed elemental sulfur are presented in Figure 12. After interaction with IPETC, the infrared spectrum of elemental sulfur changed significantly, suggesting that IPETC chemisorbed on elemental sulfur surface. Two new strong absorption peaks occurred at 2318.44 cm^{-1} and 503.42 cm^{-1} , which are separately attributed to C=S-S and C-O-S coupled vibrations. The possible reason for this may be that when IPETC reacted with elemental sulfur, the S and O atoms of C=S and C-O bonds separately lost and obtained some electrons. As a result of this, a normal covalent bond and a backdonation covalent bond were formed between C=S and C-O bonds in IPETC and S atoms in elemental sulfur surface, resulting in the generation of two new coupled vibration peaks.

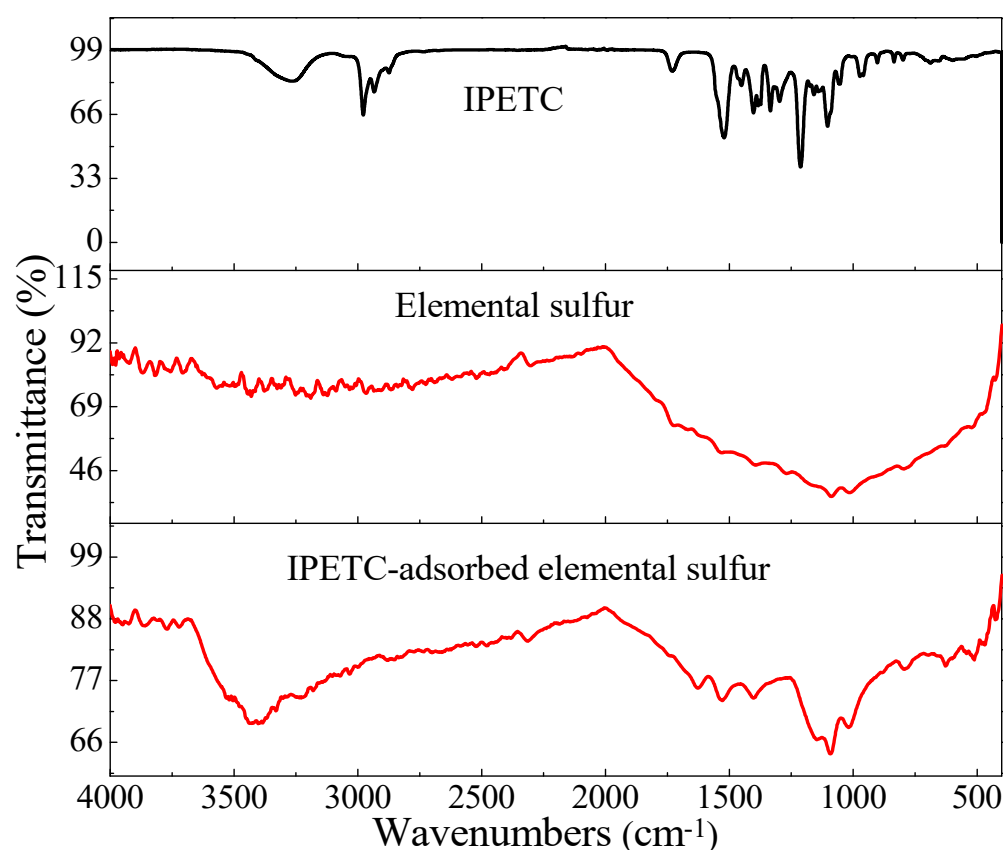


Figure 12. FTIR spectra of IPETC, elemental sulfur and IPETC-adsorbed elemental sulfur.

4. Conclusions

In this research, O-Isopropyl-N-Ethyl thionocarbamate (IPETC) collector was used to selectively recover elemental sulfur from a high-sulfur residue. The raw ore flotation test showed that in comparison with sodium ethyl xanthate (SEX) and ammonium dibutyl dithiophosphate (ADDTP) collectors, IPETC exhibited a superior collecting ability and selectivity to the elemental sulfur. Pure mineral flotation and adsorption experiments further proved that the collecting power and selectivity of IPETC to elemental sulfur are the optimum among the three collectors.

The bond length of C=S/P=S covalent bond, Mulliken population and frontier orbital energy difference between collectors and minerals manifested that the selectivity of three collectors to elemental sulfur is in the order of IPETC > ADDTP > SEX. The density of states and frontier orbital analyses showed that the HOMO of each collector is mainly constituted by electron orbitals of carbonyl S atom, the LUMOs of IPETC and SEX are comprised of conjugated big π -bond formed by electron orbitals of the coplanar functional groups, and the LUMO of ADDTP is composed of the electron orbitals of O, S and P atoms.

The adsorption configuration analysis indicated that the steadiest adsorption configuration of IPETC on the surfaces of elemental sulfur, chalcopyrite, sphalerite, pyrite, anglesite and albite is the simultaneous adsorption of carbonyl S together with O, and IPETC adsorption on elemental sulfur surface is the most stable. The Mulliken population and density of state analyses of bonding atoms indicated that when IPETC reacts with the elemental sulfur surface, the 3p orbital of carbonyl S atom of IPETC donates electrons to the 3p orbital of S atom in elemental sulfur surface to form normal covalent bond while the 2p orbital of O atom of IPETC obtains the electrons from the 3p orbital of S atom in elemental sulfur surface to generate a backdonation covalent bond, and the stronger normal covalent bond plays a dominant role. The FTIR spectrum analysis supported the generation of a normal covalent bond and a backdonation covalent bond during IPETC adsorption on elemental sulfur surface.

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