

Article

Heating Changes Bio-Schwertmannite Microstructure and Arsenic(III) Removal Efficiency

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Abstract: Schwertmannite (Sch) is an efficient adsorbent for arsenic(III) removal from arsenic(III)-contaminated groundwater. In this study, bio-schertmannite was synthesized in the presence of dissolved ferrous ions and *Acidithiobacillus ferrooxidans* LX5 in a culture media. Bio-synthesized Sch characteristics, such as total organic carbon (TOC), morphology, chemical functional groups, mineral phase, specific surface area, and pore volume were systematically studied after it was dried at 105 °C and then heated at 250–550 °C. Differences in arsenic(III) removal efficiency between 105 °C dried-sch and 250–550 °C heated-sch also were investigated. The results showed that total organic carbon content in Sch and Sch weight gradually decreased when temperature increased from 105 °C to 350 °C. Sch partly transformed to another nanocrystalline or amorphous phase above 350 °C. The specific surface area of 250 °C heated-sch was 110.06 m²/g compared to 5.14 m²/g for the 105 °C dried-sch. Total pore volume of 105 °C dried-sch was 0.025 cm³/g with 32.0% mesopore and 68.0% macropore. However, total pore volume of 250 °C heated-mineral was 0.106 cm³/g with 23.6% micropore, 33.0% mesopore, and 43.4% macropore. The arsenic(III) removal efficiency from an initial 1 mg/L arsenic(III) solution (pH 7.5) was 25.1% when 0.25 g/L of 105 °C dried-sch was used as adsorbent. However, this efficiency increased to 93.0% when using 250 °C heated-sch as adsorbent. Finally, the highest efficiency for arsenic(III) removal was obtained with sch-250 °C due to high amounts of sorption sites in agreement with the high specific surface area (SSA) obtained for this sample.

Keywords: bio-schwertmannite; specific surface area; pore volume; arsenic(III); adsorption

1. Introduction

Groundwater is generally the primary resource for drinking water and irrigation [1], and a large amount of groundwater is used without any pre-treatment. Arsenic contamination of groundwater is an important environmental problem that affects human health and quality of life [2,3]. Arsenic is a highly toxic element that has been categorized as a “Group 1 carcinogen” by the International Agency for Research on Cancer [4]. Long-term exposure to arsenic-contaminated drinking water can cause serious health problems, such as skin damage, bladder cancer, liver cancer, kidney cancer, nasal cancer, and gastric cancer [5]. Because arsenic pollution in groundwater can dramatically raise

the human risk for cancer and other serious diseases, the World Health Organization (WHO) has admonished countries to limit the As concentration in drinking water to 10 µg/L [6]. However, high concentrations of arsenic ($\text{As} > 10 \text{ µg/L}$) in groundwater are frequently found in China, India, Nepal, Bangladesh, and other countries [3,7–9]. Chakraborti et al. [10] investigated the arsenic contamination of groundwater in Patna district (Bihar, India); the study detected a maximum arsenic concentration of 1466 µg/L and found that 61% and 44% of tube-wells contained arsenic at concentrations that exceeded 10 µg/L and 50 µg/L, respectively. The concentration of arsenic in groundwater in the Datong Basin (Shanxi, China) usually exceeds the WHO-recommended value (10 µg/L) for drinking water and can reach 1820 µg/L [11]. In the environment, arsenic(III) and arsenic(V) are the most common in water environments. As(V) is the dominant species in an oxidative environment, whereas arsenic(III) is dominant in a reducing environment [12]. Generally, groundwater is characterized by a relatively high pH of 7–8.5 [13]. Furthermore, it has also been reported that arsenic(III) is 25–60 times more toxic than arsenic(V) [14]. Therefore, removal of arsenic(III) from groundwater is an important problem for the environmental engineering discipline.

Adsorption technology is one of the most popular methods for the treatment of arsenic(III)-containing groundwater. Schwertmannite [15], activated alumina [16], FeCl_3 -based materials [17], chitosan-functionalized graphene oxide [18], bismuth impregnated biochar [19], calix [4] pyrrole [20], and other adsorbents have been used for arsenic(III) removal. Schwertmannite, which is a poorly crystalline Fe(III)-oxyhydroxy-sulfate mineral whose chemical formula is $\text{Fe}_8\text{O}_8(\text{OH})_{8-2x}(\text{SO}_4)_x$ ($1 < x < 1.75$) [21,22] that acts as a good adsorbent for arsenic(III), has gained increasing attention in recent years [23]. In nature, schwertmannite is commonly found in sulfate-rich environments such as acid sulfate soil or acid mine drainage systems [24–26], and is typified by a “hedgehog-like” or “spherical” morphological shape and porous structure [26,27]. In addition, researchers generally accept that the pores in schwertmannite are connected and form tunnels, and that nearly all OH^- groups and one-third of the SO_4^{2-} are adsorbed on the material’s surface, with the remaining SO_4^{2-} located in the tunnel structures [28]. In general, the OH^- and SO_4^{2-} groups on the schwertmannite surface or in tunnel structures can ligand-exchange with arsenic(III) species, a process that determines the arsenic(III) removal efficiency [13,29,30]. Besides occurring naturally, schwertmannite can be produced by different synthesis techniques through biotic and abiotic methods; these are becoming essential for the treatment of arsenic(III)-contaminated drinking water [15,27]. Among these synthesis techniques, schwertmannite biosynthesis by *Acidithiobacillus ferrooxidans* oxidizing FeSO_4 solution has been proved to be an attractive approach for the preparation of the material within a short reaction time (48–72 h) under eco-friendly reaction conditions without a strong oxidizer and high temperature, which makes the removal of arsenic(III) from groundwater using bio-schwertmannite shows the great application prospect, especially in developing or less developed countries, because of its low producing cost. Moreover, Liao et al. [13] found that both Freundlich and Langmuir isotherms could be used to describe the adsorption behavior of arsenic(III) on biogenic schwertmannite; a maximum arsenic(III) adsorption capacity of 113.9 mg/g bio-schwertmannite can be achieved in the pH range 7–10. The maximum arsenic(III) adsorption capacity of bio-schwertmannite was higher than that of other adsorption materials under similar pH conditions, such as 1–4 mg/g for activated alumina [16], 27.1 mg/g for FeCl_3 -based materials [17], 64.2 mg/g for chitosan-functionalized graphene oxide [18], and 15.28 mg/g for calix pyrrole [4,20]. Therefore, the removal arsenic(III) from groundwater or other contaminated water sources using bio-schwertmannite has great commercial value.

Unfortunately, bio-synthesized schwertmannite particles aggregate too closely because of the extracellular polymeric substances (EPS) of *A. ferrooxidans* that exist in the biosynthesis systems [27,31,32], thereby reducing the schwertmannite specific surface area (SSA). Previous studies have shown that the SSA of bio-synthesized schwertmannite can vary from ~3 to ~50 m^2/g [15,27,31,33]. Further improving the SSA of bio-synthesized schwertmannite is an important challenge to the wider use of this material as an adsorbent for arsenic(III) removal from groundwater. The EPS secreted by *A. ferrooxidans* consists mainly of extracellular lipopolysaccharide and outer membrane

proteins [34]. In order to remove the EPS from bio-synthesized schwertmannite for increasing the SSA of bio-synthesized schwertmannite, preliminary experiments have attempted to extract the EPS using organic solvents such as ethanol and acetone (the ratio of bio-synthesized schwertmannite weight and organic solvent volume was 1:100). It was found that the SSA of bio-synthesized schwertmannite only increased from 2.98 to 3.85 or 4.73 m²/g after ethanol or acetone treatment, respectively. It is well known that organic carbon, such as the EPS of microorganisms, can be effectively removed under high temperature conditions. However, whether heating can improve the SSA of bio-schwertmannite, or result in changes to the bio-schwertmannite microstructure, and changes to the arsenic(III) removal capacity of bio-schwertmannite have not been studied.

The aims of the present study were (1) to investigate bio-synthesized schwertmannite characteristics, such as total organic carbon, morphology, chemical functional groups, mineral phase, specific surface area, and pore volume distribution, as a function of heat treatment at 250–550 °C; and (2) to examine the difference of arsenic(III) removal capacity between 105 °C dried bio-synthesized schwertmannite and 250–550 °C heated bio-synthesized schwertmannite.

2. Materials and Methods

2.1. *A. ferrooxidans* LX5 and Its Culture

A strain of iron-oxidizing bacteria, *A. ferrooxidans* LX5 (CGMCC No. 0727), obtained from the China General Microbiological Culture Collection Center was used in this study. The *A. ferrooxidans* LX5 strain was inoculated in modified 9K liquid culture medium with an initial pH of ~2.50 [35]. The medium was composed of FeSO₄·7H₂O (44.2 g/L), (NH₄)₂SO₄ (3.5 g/L), Ca(NO₃)₂·4H₂O (0.0168 g/L), K₂HPO₄ (0.058 g/L), KCl (0.119 g/L), and MgSO₄·7H₂O (0.583 g/L). Then, the culture was incubated at 28 °C in a rotatory shaker (SHZ-A model, Changzhou Henglong instrument Co., Ltd, Changzhou, China) at 180 r/min until the ferrous ions in the system were completely oxidized, and after 60 h was filtered through quantitative filter paper (φ12.5 cm, Jiejie brand, Fushun city Dongyang Industrial and Trading Co., Ltd., Fushun, China) to remove iron-based precipitate. The filtered culture with pH of ~2.20 containing ~3 × 10⁷ cells/mL of *A. ferrooxidans* LX5 cells was used for bio-synthesized schwertmannite preparation. All solutions were prepared with deionized water using analytical grade reagents.

2.2. Preparation of Bio-Synthesized Schwertmannite

Bio-schwertmannite was prepared using *A. ferrooxidans* LX5 to bio-oxidize a ferrous ions solution in a custom-designed 25-L bioreactor. In brief, 13.5 L of solution containing 0.16 mol/L ferrous ions (added as FeSO₄·7H₂O) was dispensed into a bioreactor fitted with an air compressor and temperature controller. The aeration rate and temperature in the bioreactor were controlled at 0.055 m³/min and ~28 °C, respectively. A 1.5 L quantity of *A. ferrooxidans* LX5 culture prepared as described in Section 2.1 was added into the bioreactor resulting in a bacterial density of ~3 × 10⁶ cells/mL in the bioreactor. During the schwertmannite bio-synthesis process, the 1 mL of solution was withdrawn from the system at 12 h intervals, filtered through a 0.22 μm membrane filter, and analyzed to determine ferrous ions concentration in filtrate using 1, 10 phenanthroline method [27]. After 60 h, the ferrous ions were bio-oxidized completely. Then, the cultures were filtered through quantitative filter paper to obtain a suspended reddish-brown precipitate. The precipitate was rinsed twice with distilled water before being dried in a digital vacuum drying oven (GZX-9246MBE model, Shanghai Boxun industry Co., Ltd., Shanghai, China) at 105 °C to constant weight, after which the residue was ground and passed through a 100-mesh screen (Zhang Xingsha Sieve Factory, Shaoxin, China). The dried residue was characterized using X-ray diffraction (XRD) and field-emission scanning electron microscopy (SEM) to confirm that the collected precipitate was schwertmannite.

2.3. Bio-Schwertmannite Heating

Bio-schwertmannite heating was accomplished in a box-type electric resistance furnace (SX-8-10 model, Yingshan State Experimental Equipment Factory, Huanggang, China). Twelve 18 mL porcelain crucibles were triple-rinsed with deionized water and heated at 550 °C for 2 h; the crucibles were then cooled in a desiccator to room temperature before use. Then, approximately 2.0000 g bio-schwertmannite was placed in each crucible, and was weighed and numbered. Three crucibles were assigned randomly to 2 h heat treatments at 250, 350, 450, or 550 °C. Each treatment was performed in triplicate. After treatment, the heated bio-schwertmannite was cooled and weighed to determine the weight loss of bio-schwertmannite during heating. Samples of untreated bio-schwertmannite and heated bio-schwertmannite were collected for determination of total organic carbon, morphology, chemical functional groups, mineral phase, specific surface area, and total pore volume. The distribution of specific surface area and pore volume among micropores (<2 nm diameter), mesopores (2–50 nm), and macropores (>50 nm) also was determined. It should be pointed out that the weight loss and total organic carbon concentration of bio-schwertmannite have been determined for each of the triplicates. The morphology, chemical functional groups, mineral phase, specific surface area, and total pore volume have been determined for the mixture of the triplicates.

2.4. Arsenic(III) Adsorption Efficiency of Heated Bio-Schwertmannite

The As₂O₃ was dried in a digital vacuum drying oven for 2 h. A 1000 mg/L arsenic(III) stock solution was prepared by dissolving dried As₂O₃ in NaOH solution, and was stored at 4 °C before use. Before the arsenic(III) adsorption experiment, a solution containing 1.0 mg/L arsenic(III) was prepared by diluting the stock solution. Each solution was adjusted to pH ~7.50 by dropwise addition of 0.1 mol/L HCl or NaOH. Arsenic(III) adsorption was determined by adding 10 mg of untreated, 250 °C, or 550 °C heated bio-schwertmannite to 100 mL capped plastic bottles (inner diameter 46 mm), each containing 40 mL solution (initial 1.0 mg/L arsenic(III)). The untreated, 250 °C, or 550 °C heated bio-schwertmannite concentration in arsenic(III) adsorption systems was 0.25 g/L. All plastic bottles were shaken in a rotary shaker (HZQ-F160 model, Harbin, China) at 28 °C and 180 r/min. It is noted that 4 h is adequate for arsenic(III) to reach adsorption equilibrium on bio-schwertmannite [13]. After 4 h adsorption, the solution from different treatments was filtered through a 0.45 µm membrane and analyzed for arsenic(III). Each treatment was conducted in triplicate.

It is noted that the main aims during arsenic(III) removal experiment in here was to examine the difference of arsenic(III) removal capacity between 105 °C dried bio-schwertmannite and 250–550 °C heated bio-schwertmannite. To find the maximum arsenic(III) adsorption capacity of 105 °C dried bio-schwertmannite or 250–550 °C heated bio-schwertmannite was not the objective. According to the previous reports from our research group [13,27], 0.25 g/L was chosen as the only bio-schwertmannite concentration. This does not mean that dried bio-schwertmannite and heated bio-schwertmannite can adsorb arsenic(III) from groundwater under these low concentrations.

2.5. Analytical Procedures

The weight of untreated and heated bio-schwertmannite was measured using an electronic balance (Sartorius BS 124S, Beijing, China) with a precision of 0.0001 g. The surface morphology of untreated or heated bio-schwertmannite was investigated using field-emission scanning electron microscopy (SEM, JSM-7001F, Tokyo, Japan) with an accelerating voltage of 10 kV [35]. The bio-schwertmannite phase was determined using X-ray diffraction (XRD, MiniFlex II, Tokyo, Japan) with a Cu Kα X-ray source [35]. Fourier transform infrared (FTIR) spectroscopy (Bruker, Model Tensor 27, Karlsruhe, Germany) was used to characterize various chemical functional groups in untreated and heated bio-schwertmannite [23]. Total pore volume and the specific surface area of bio-schwertmannite were determined by the N₂-BET method using a surface area and porosimetry analyzer (Quantachrome, Boynton Beach, USA) [36]. Total organic carbon (TOC) of bio-schwertmannite was analyzed by the

combustion/non-dispersive infrared (NDIR) method (Multi N/C-3100 model, Analytikjena, Jena, Germany) [37]. Arsenic(III) concentration was determined using a dual-channel atomic fluorescence photometer (AFS-2100, Beijing, China) [38] with a detection limit of 0.01 µg/L.

3. Results and Discussion

3.1. Color, Total Organic Carbon (TOC) Content, and Weight Loss of Bio-Schwertmannite

The color of untreated bio-schwertmannite (bio-schwertmannite dried at 105 °C) was reddish-brown, as also found in previous studies [23,39]. However, the color changed from reddish-brown to black when the heating temperature increased and reached 550 °C. The color of bio-schwertmannite began turning black at 250 °C and was completely black at 350–550 °C. The total organic carbon (TOC) content and weight loss of bio-schwertmannite changed noticeably as the heating temperature varied (Figure 1). This showed that the EPS secreted by *A. ferrooxidans* and dead *A. ferrooxidans* cells was removed from bio-schwertmannite during the heating process. The relationship between the heating temperature and the loss of *A. ferrooxidans* cells and the secreted EPS was measured by determining the amount of TOC in bio-schwertmannite after heating treatment (Figure 1a). The initial TOC content in untreated bio-schwertmannite was 1.54 mg/g. The TOC content decreased by 55.2% in 250 °C heated bio-schwertmannite and by 89.0%–89.6% in bio-schwertmannite heated at 350–550 °C. Thus, TOC was gradually removed as the heating temperature increased from 105 to 550 °C.

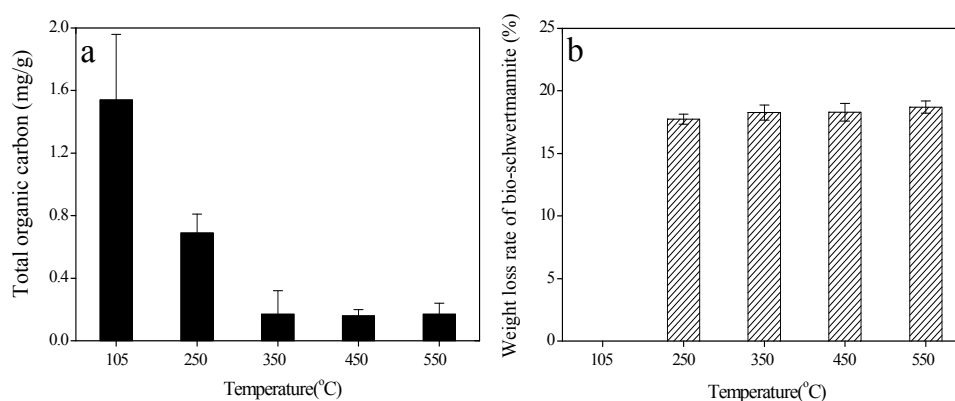


Figure 1. Total organic carbon content (a) and weight loss of bio-schwertmannite (b) before and after heating.

The weight loss of bio-schwertmannite after heating was calculated as a percentage of the initial weight of material dried at 105 °C (Figure 1b). As shown, the weight loss reached 17.75% at 250 °C. However, the weight loss increased just from 17.75% to 18.71% when the heating temperature increased from 250 to 550 °C. In other words, the weight loss of bio-schwertmannite increased significantly when the heating temperature increased in the range of 105–250 °C and increased slightly in the range of 250–550 °C. This result was similar to that reported by Sun et al. [40] who found that the weight loss of bio-schwertmannite collected from the Malan mine in Shanxi province, China was 18.12% when the bio-schwertmannite heating temperature increased from 99 to 535 °C. In addition, Gan et al. [23] synthesized schwertmannite using acidophilic bacterium *A. ferrooxidans* 23270 with a $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}:\text{AlPO}_4$ ratio of 200:1. They found that the mass of bio-schwertmannite decreased only slightly (~2%) when the heating temperature changed from 250 to 550 °C, as observed from the bio-schwertmannite thermal analyses curve. Figure 1 shows conclusively that the removal of EPS and *A. ferrooxidans* cells by heating contributes little to bio-schwertmannite weight loss because the initial TOC content in untreated bio-schwertmannite was negligible (1.54 mg/g). Therefore, the significant loss of bio-schwertmannite mass during heating may be related to the loss of chemical functional groups.

3.2. Chemical Functional Groups and Phase of Bio-Schwertmannite

The variation in chemical groups associated with bio-schwertmannite can be observed from the FTIR spectra (Figure 2).

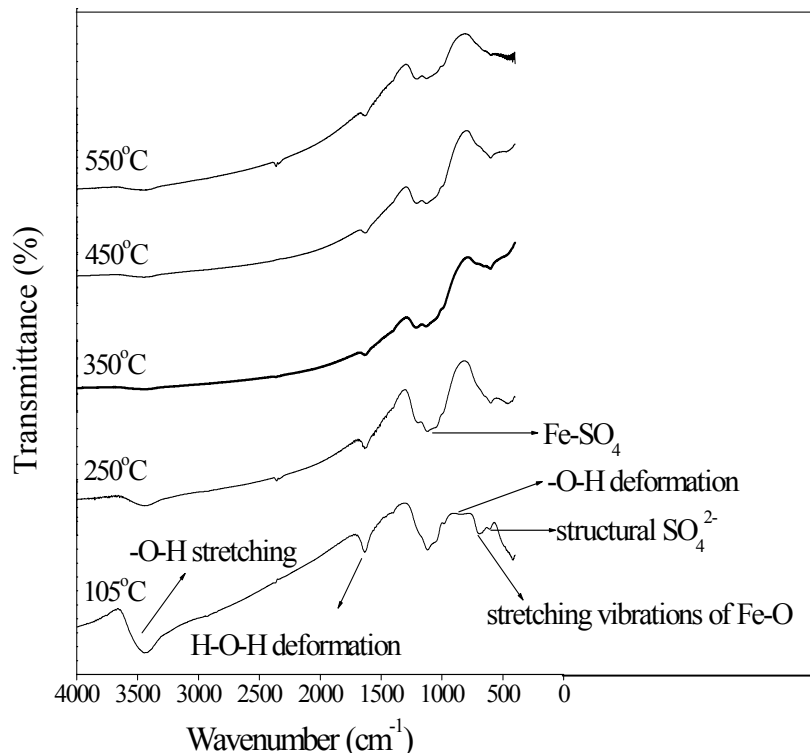


Figure 2. Fourier transform infrared (FTIR) spectroscopy of bio-schwertmannite before and after heating.

In untreated bio-schwertmannite, the intense absorption peak observed in the region of $2900\text{--}3700\text{ cm}^{-1}$ can be attributed to --O--H stretching [23]. Furthermore, the band observed at 1635 cm^{-1} is the deformation of H--O--H [23], and the intense peaks shown at $1000\text{--}1250\text{ cm}^{-1}$ represent the strong SO_4 stretching vibration, which can be attributed to the formation of an Fe--SO_4 bridging complex [41]. Related features due to the presence of structural SO_4 include bands at 981 cm^{-1} and 607 cm^{-1} [42]. The peak at 692 cm^{-1} and broad absorption peak at 860 cm^{-1} can be attributed to Fe--O stretching and --O--H deformation [42].

Spectra shown in Figure 2 for heated bio-schwertmannite show that the peaks corresponding to the deformation of H--O--H (at 1635 cm^{-1}), --O--H stretching (at $2900\text{--}3700\text{ cm}^{-1}$), and --O--H deformation (at 860 cm^{-1}), gradually weakened (albeit only slightly) when the heating temperature increased to $350\text{ }^{\circ}\text{C}$. In addition, the band for structural SO_4^{2-} also showed a slight weakening trend. Adsorbed water molecules in a material are known to be removed upon heating at $105\text{ }^{\circ}\text{C}$. Sun et al. [40] reported that structural water in bio-schwertmannite can be removed when the heating temperature exceeds $120\text{ }^{\circ}\text{C}$. A comparison of Figures 1 and 2 leads to the conclusion that the weight loss of bio-schwertmannite during heating at $250\text{--}550\text{ }^{\circ}\text{C}$ was due mainly to the loss of structural water and --O--H groups.

X-ray diffraction was used to determine the mineral phase and identify the crystalline mineral and amorphous mineral. The XRD patterns of untreated and heat-treated bio-schwertmannite are shown in Figure 3. Because schwertmannite is a poorly crystalline $\text{Fe(III)-oxyhydroxy-sulfate}$ mineral [21], the characteristic peaks expressed in bio-schwertmannite XRD patterns are broad peaks [43,44]. The XRD pattern of schwertmannite has eight broad peaks (JCPDS, PDF 47-1775) and these occur

at $2\theta = 18.24^\circ, 26.27^\circ, 35.16^\circ, 39.49^\circ, 46.54^\circ, 55.30^\circ, 61.34^\circ$, and 63.68° . The typical broad peaks at $2\theta = 26.27^\circ, 35.16^\circ, 55.30^\circ$, and 61.34° are the most easily identified for schwertmannite.

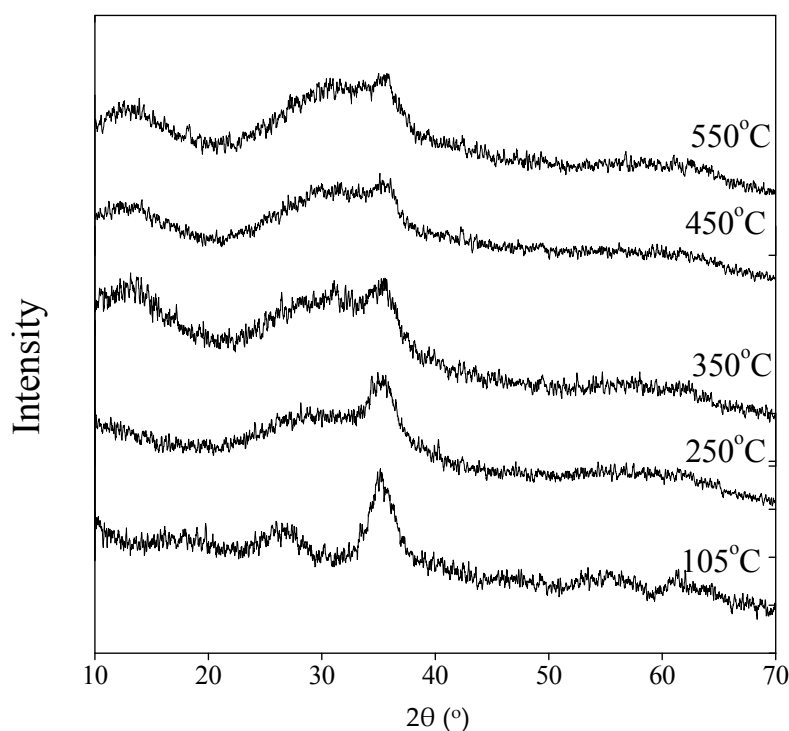


Figure 3. The X-ray diffraction (XRD) patterns of bio-schwertmannite before and after heating.

The results in Figure 3 show that the XRD patterns were similar for untreated bio-schwertmannite dried at 105°C and for material heated at 250°C . In other words, the mineral phase of bio-schwertmannite did not greatly change when the heating temperature increased to 250°C . However, the intensity of the typical broad peaks at $2\theta = 35.16^\circ$ in XRD patterns for bio-schwertmannite heated at 350 – 550°C were relatively weak compared to those for untreated bio-schwertmannite and for the material when heated to 250°C . Furthermore, a broad peak at $2\theta \approx 30^\circ$ – 35° appeared obviously in bio-schwertmannite heated at 350 – 550°C but was not observed in untreated bio-schwertmannite or weakly presented in that heated to 250°C . This additional peak indicated that some bio-schwertmannite underwent obvious phase transformation to another amorphous or nanocrystalline phase at 350 – 550°C even though schwertmannite remained in the main phase. In addition, it should be pointed out that this amorphous or nanocrystalline material may be the two-line ferrihydrite, which also can adsorb arsenic(III) and arsenic(V) [45,46]. Jiang et al. [45] reported that the chemical synthesized two-line ferrihydrite has a broad characteristic peak at $2\theta = 34^\circ$. Therefore, the bio-schwertmannite may have changed to two-line ferrihydrite under heating conditions, especially when the heating temperature reached above 350°C .

3.3. Morphology, Specific Surface Area, and Total Pore Volume of Bio-Schwertmannite

Field-emission scanning electron microscopy (SEM) is the main method for observing mineral morphology [30,47]. The SEM-determined morphology, specific surface area (SSA), and total pore volume of bio-schwertmannite before and after heating are shown in Figure 4.

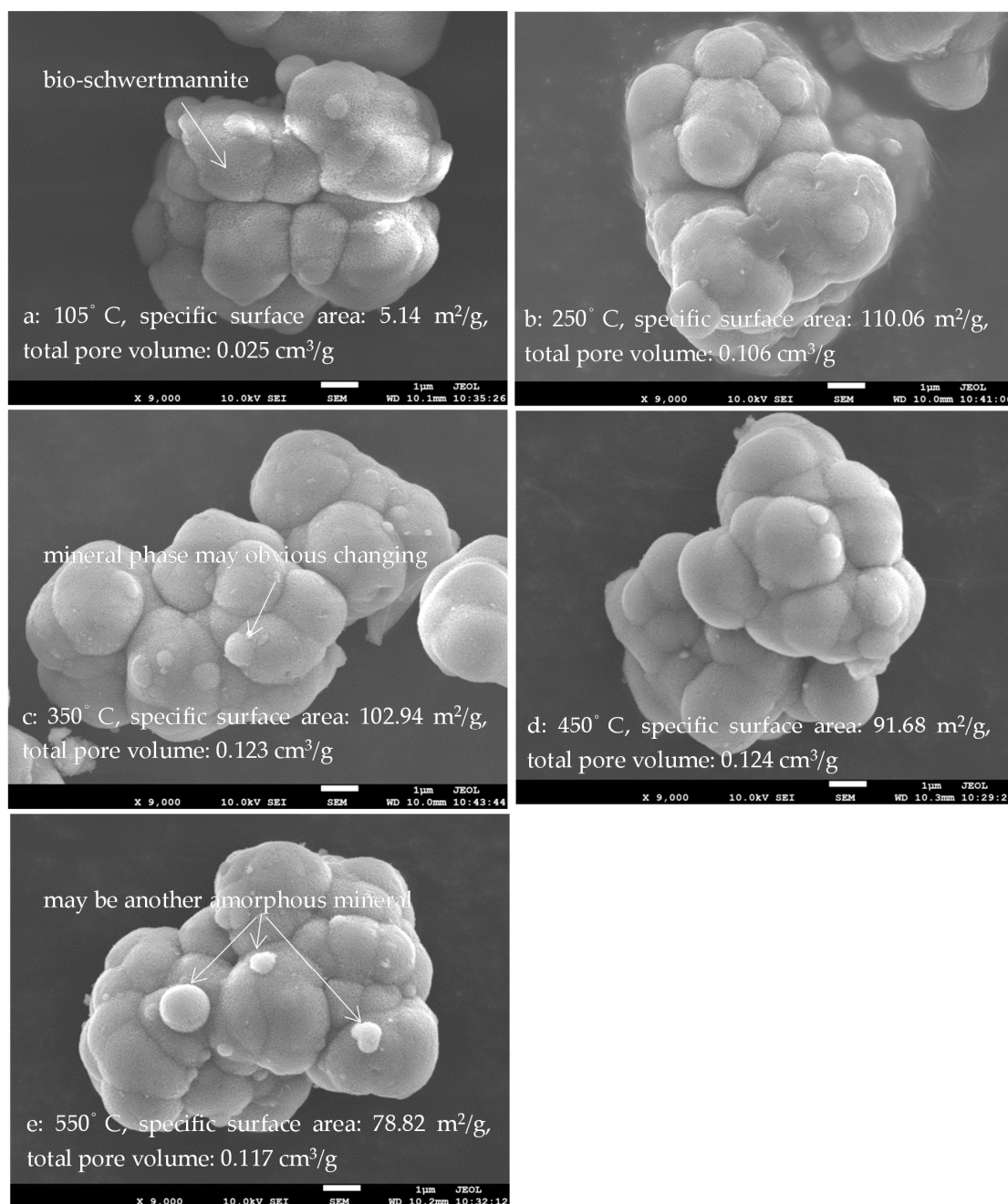


Figure 4. Morphology, specific surface area, and total pore volume of bio-schwertmannite before and after heating.

Untreated bio-schwertmannite clearly exhibited a structure comprised of aggregated homogeneous spherical particles (approximately 2–3 μm diameter) with coarse pores. These characteristics of bio-schwertmannite particles produced in this study were in agreement with the description by Gagliano et al. [48] of naturally occurring schwertmannite collected from a constructed wetland treating mine drainage. After heating at 250–550 $^{\circ}\text{C}$, the morphology of heated bio-schwertmannite showed no great difference compared to untreated bio-schwertmannite, but the bio-schwertmannite surface may have presented another mineral phase. This phenomenon can be confirmed from the results showed in Figure 3. The aggregation of individual bio-schwertmannite spherical particles showed no obvious change after heating at 250–550 $^{\circ}\text{C}$. These results indicate that the aggregation of

bio-schwertmannite particles was not only due to the EPS secreted by *A. ferrooxidans* or *A. ferrooxidans* cells. Chemical bonds between bio-schwertmannite particles also play a role in the aggregation, and these bonds are not easily broken by heating at 250–550 °C. In other words, the EPS of *A. ferrooxidans* or *A. ferrooxidans* cells promotes the production of chemical bonds between bio-schwertmannite particles, and is not the only aggregation effect.

Although the morphology of heated bio-schwertmannite was not great changed by heating, the SSA and total pore volume were significantly increased by heating. The SSA and total pore volume of untreated bio-schwertmannite were 5.14 and 0.025 cm³/g, respectively. Previous researchers have found that the SSA of bio-synthesized schwertmannite varied from ~3 to ~50 m²/g [15,27,31,33]. However, the SSA and total pore volume sharply increased to 110.06 and 0.106 cm³/g, respectively, in bio-schwertmannite heated at 250 °C. The corresponding relationship between SSA and total pore volume for 250 °C-heated bio-schwertmannite in this study were similar to findings by Dou et al. [49] who chemically synthesized schwertmannite using FeCl₃ and Na₂SO₄; the material had an SSA and total pore volume of 206.1 and 0.22 cm³/g, respectively. Compared with that for bio-schwertmannite heated at 250 °C, the SSA of bio-schwertmannite heated at 350–550 °C sharply decreased while the total pore volume increased. For example, the SSA decreased to 78.82 m²/g and the total pore volume increased to 0.117 cm³/g in bio-schwertmannite heated at 550 °C. Thus, the SSA of heated bio-schwertmannite reached a maximum value in the 250 °C heat treatment.

For explicating the influence factor of specific surface area in 250 °C heated bio-schwertmannite, a chemically synthesized schwertmannite was prepared by chemically oxidizing FeSO₄ with H₂O₂ in this study. The SSA of the chemical synthesized schwertmannite was 3.45 m²/g. However, when the schwertmannite was heated at 250 °C the SSA remained almost constant (3.47 m²/g), although the weight loss reached 19.04% due to the loss of OH[−] groups and structural water. These results indicated that the loss of OH[−] groups and structural water has no influence on the SSA of schwertmannite during heating. In other words, the larger SSA of bio-schwertmannite heated at 250 °C in the current study was due to the removal of EPS secreted by *A. ferrooxidans* and dead *A. ferrooxidans* cells. Furthermore, the results shown in Figures 1 and 4 indicated that the TOC content has a significant negative correlation ($R^2 = 0.95$) with the total pore volume. The removal of EPS secreted by *A. ferrooxidans* and *A. ferrooxidans* cells could significantly increase the total pore volume of bio-schwertmannite. As noted, the SSA of heated bio-schwertmannite gradually decreased when the heating temperature increased from 350 to 550 °C, although there was no obvious change in total pore volume. This result may be regulated by the mineral phase changed when the heating temperature increased from 350 to 550 °C.

3.4. Specific Surface Area and Pore Volume Distribution among Micropores, Mesopores, and Macropores in Bio-Schwertmannite

The SSA of adsorbents results mainly from different types of pores [50]. In general, the pores in materials can be classified as micropores (diameter < 2 nm), mesopores (diameter 2–50 nm), and macropores (diameter > 50 nm) [51]. The distribution of SSA and pore volume among micropores, mesopores, and macropores in bio-schwertmannite is exhibited in Figure 5. In untreated bio-schwertmannite (dried at 105 °C), 58.2% and 41.2% of total SSA, 32.0% and 68.0% of total pore volume were contributed from mesopores and macropores, respectively. However, micropores, mesopores, and macropores contributed 35.1%, 26.2%, and 38.7%, respectively, of the total SSA, and 23.6%, 33.0%, and 43.4%, respectively, of the total pore volume in bio-schwertmannite heated at 250 °C. When heating increased to 350–550 °C, nearly all of the micropores disappeared from bio-schwertmannite. This result explains the SSA variations described in Section 3.3. In other words, because many micropores disappeared at high heating temperatures, the SSA of bio-schwertmannite decreased gradually from 110.06 to 78.82 m²/g while the corresponding total pore volume increased from 0.106 to 0.117 cm³/g as the heating temperature increased from 250 to 550 °C. Furthermore, the TOC content and weight loss of bio-schwertmannite nearly stabilized (Figure 1) when bio-schwertmannite heating exceeded 350 °C. Therefore, the gradual decrease in the SSA of

heated bio-schwertmannite as heating temperature increased from 350 to 550 °C (Figure 4) may have been caused by the mineral phase transformation (Figure 2). This phase transformation also could have caused the reduction of macropore volume and the amount of SSA contributed by macropores.

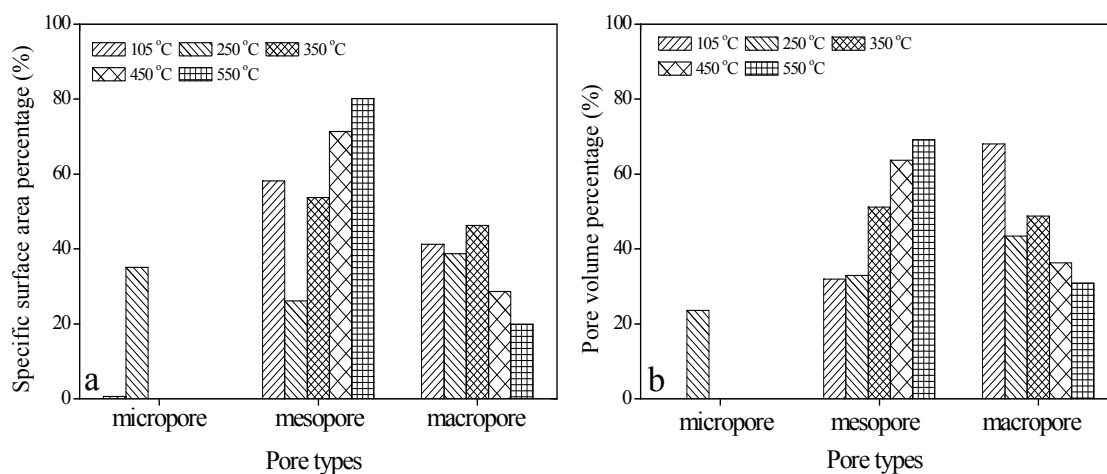


Figure 5. Specific surface area (a) and pore volume distribution of micropore, mesopore, and macropore in bio-schwertmannite (b) before and after heating.

3.5. Arsenic(III) Removal Efficiency of Bio-Schwertmannite

Representative samples of untreated bio-schwertmannite and of material heated at 250 and 550 °C were used to study the arsenic(III) removal efficiency of the bio-schwertmannite before and after heating (Figure 6).

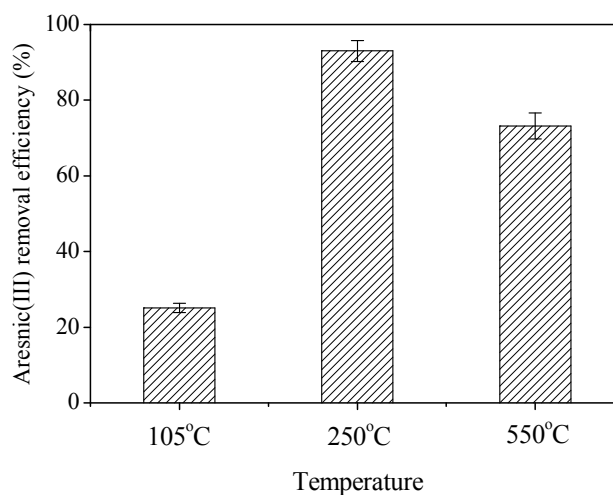


Figure 6. Arsenic(III) removal efficiency of bio-schwertmannite before and after heating.

The arsenic(III) removal efficiency of untreated bio-schwertmannite (0.25 g/L in a solution initially containing 1 mg/L arsenic(III) at pH ~7.5) was 25.1% (4 h reaction duration). However, this removal efficiency increased to 93.0% for bio-schwertmannite heated at 250 °C, but was only 73.2% for bio-schwertmannite heated at 550 °C. In other words, 13.4, 49.7, and 39.1 μmol of arsenic(III) can be removed per g of untreated 250 °C, or 550 °C heated bio-schwertmannite, respectively. These results highlighted that the SSA of bio-schwertmannite determined the arsenic(III) removal capacity. As shown in Section 3.2, some OH^- groups of bio-schwertmannite were removed when the heating temperature reached 250–550 °C. However, OH^- groups on the schwertmannite surface can exchange their

ligands with arsenic(III) species and improve the arsenic(III) removal efficiency [13]. Therefore, it is inferred that the arsenic(III) removal efficiency of heated bio-schwertmannite is mainly due to ligand exchange between SO_4^{2-} and arsenic(III) because the SO_4^{2-} groups remained intact in heated bio-schwertmannite at heating temperatures between 250 and 550 °C (Figure 2). It should be noted that part of arsenic(III) may adsorb on the remainder organic matter through complexing action when 250 °C heated bio-schwertmannite is used as an adsorbent for arsenic(III) removal. In addition, adsorbent regeneration is a quite important issue from an adsorbent life cycle perspective. After arsenic(III) adsorption, arsenic(III)-bearing bio-schwertmannite could be regenerated by 1.0 mol/L NaOH solution, and more than 80% of arsenic(III) adsorbed on the surface of bio-schwertmannite could be released after three successive regenerations [52].

4. Conclusions

Bio-schwertmannite can be bio-synthesized by *A. ferrooxidans* LX5 in 60 h, and the SSA of the resulting material is relatively small (5.14 m²/g). However, the SSA of bio-schwertmannite can be increased significantly (20-fold) by heating the material at 250 °C. Heating results in the removal of EPS secreted by *A. ferrooxidans* and *A. ferrooxidans* cells. Correspondingly, heating bio-schwertmannite at 250 °C increases the material's arsenic(III) removal efficiency (by 2.68 times in this study) due in part to increased SSA. If heating temperature increases from 350 to 550 °C, an obvious transformation in the mineral phase of bio-schwertmannite occurs. In this study, the SSA of bio-schwertmannite heated at 350 and 550 °C decreased to 102.94 and 78.82 m²/g, respectively, from a maximum of 110.06 m²/g achieved by heating at 250 °C. Thus, while heating bio-schwertmannite at 250 °C improves the material's properties as an adsorbent for arsenic(III) removal, heating at 350–550 °C causes these properties to deteriorate. The outcomes of this study provide guidance for the preparation and effective use of bio-schwertmannite as an adsorbent for removing arsenic(III) from contaminated groundwater. The details of mineral phase changes and whether the arsenic(III) can be partly oxidized to As(V) by dried and heated bio-schwertmannite will be the focus of further study.

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