

Article

Structural Characterization of Titanium–Silica Oxide Using Synchrotron Radiation X-ray Absorption Spectroscopy

Arpaporn Teamsinsungvon ^{1,2,3}, Chaiwat Ruksakulpiwat ^{1,2,3}, Penphitcha Amonpattaratkit ⁴
and Yupaporn Ruksakulpiwat ^{1,2,3,*}

¹ School of Polymer Engineering, Institute of Engineering, Suranaree University of Technology, Nakhon Ratchasima 30000, Thailand; arpaporn.te@gmail.com (A.T.); charuk@sut.ac.th (C.R.)

² Center of Excellence on Petrochemical and Materials Technology, Chulalongkorn University, Bangkok 10330, Thailand

³ Research Center for Biocomposite Materials for Medical Industry and Agricultural and Food Industry, Nakhon Ratchasima 30000, Thailand

⁴ Synchrotron Light Research Institute (SLRI), 111 University Avenue, Muang District, Nakhon Ratchasima 30000, Thailand; penphitcha@slri.or.th

* Correspondence: yupa@sut.ac.th; Tel.: +66-44-22-3033

Abstract: In this study, titania–silica oxides (Ti_xSi_y oxides) were successfully prepared via the sol–gel technique. The Ti and Si precursors were titanium (IV), isopropoxide (TTIP), and tetraethylorthosilicate (TEOS), respectively. In this work, the effects of pH and the Ti/Si atomic ratio of titanium–silicon binary oxide (Ti_xSi_y) on the structural characteristics of Ti_xSi_y oxide are reported. ^{29}Si solid-state NMR and FTIR were used to validate the chemical structure of Ti_xSi_y oxide. The structural characteristics of Ti_xSi_y oxide were investigated using X-ray diffraction, XRF, Fe-SEM, diffraction particle size analysis, and nitrogen adsorption measurements. By applying X-ray absorption spectroscopy (XAS) obtained from synchrotron light sources, the qualitative characterization of the Ti–O–Si and Ti–O–Ti bonds in Ti–Si oxides was proposed. Some Si atoms in the SiO_2 network were replaced by Ti atoms, suggesting that Si–O–Ti bonds were formed as a result of the synthesis accomplished using the sol–gel technique described in this article. Upon increasing the pH to alkaline conditions (pH 9.0 and 10.0), the nanoparticles acquired a more spherical shape, and their size distribution became more uniform, resulting in an acceptable nanostructure. Ti_xSi_y oxide nanoparticles were largely spherical in shape, and agglomeration was minimized. However, the $\text{Ti}_{50}\text{Si}_{50}$ oxide particles at pH 10.0 become nano-sized and agglomerated. The presence of a significant pre-edge feature in the spectra of $\text{Ti}_{50}\text{Si}_{50}$ oxide samples implied that a higher fraction of Ti atoms occupied tetrahedral symmetry locations, as predicted in samples where Ti directly substituted Si. The proportion of Ti atoms in a tetrahedral environment agreed with the value of 1.83 given for the Ti–O bond distance in Ti_xSi_y oxides produced at pH 9.0 using extended X-ray absorption fine structure (EXAFS) analysis. Photocatalysis was improved by adding 3% wt TiO_2 , SiO_2 , and Ti_xSi_y oxide to the PLA film matrix. TiO_2 was more effective than $\text{Ti}_{50}\text{Si}_{50}$ pH 9.0, $\text{Ti}_{50}\text{Si}_{50}$ pH 10.0, $\text{Ti}_{50}\text{Si}_{50}$ pH 8.0, and SiO_2 in degrading methylene blue (MB). The most effective method to degrade MB was $\text{TiO}_2 > \text{Ti}_{70}\text{Si}_{30} > \text{Ti}_{50}\text{Si}_{50} > \text{Ti}_{40}\text{Si}_{60} > \text{SiO}_2$. Under these conditions, PLA/ $\text{Ti}_{70}\text{Si}_{30}$ improved the effectiveness of the photocatalytic activity of PLA.

Keywords: mixed oxide; sol–gel method; titania; silica; XAS

Citation: Teamsinsungvon A.; Ruksakulpiwat C.; Amonpattaratkit P.; Ruksakulpiwat Y. Structural characterization of titanium–silica oxide using Synchrotron radiation X-ray absorption spectroscopy. *Polymers* **2022**, *14*, 2729. <https://doi.org/10.3390/polym14132729>

Academic Editors: Wei Jiang and Yubing Hu

Received: 8 June 2022

Accepted: 30 June 2022

Published: 3 July 2022

Publisher’s Note: MDPI stays neutral with regard to jurisdictional claims in published maps and institutional affiliations.



Copyright: © 2022 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<https://creativecommons.org/licenses/by/4.0/>).

1. Introduction

In recent years, binary oxides of Ti and Si have gained a great deal of attention in a wide range of applications, including TiO_2 – SiO_2 [1], Ag– TiO_2 – SiO_2 [2], and the ordered mesoporous TiO_2 /SBA-15 matrix [3]. Sol–gel processing is an adjustable method that has attained commercial success because it allows for the control of the desired properties of

a material during the preparation process. As a result, the sol–gel method has become one of the most preferred methods to produce oxide materials. Titanium dioxide (TiO_2), often known as titania, is an important substance. This multifunctional material may be used in coatings [4,5], biomaterials [6], photocatalysts [7], food packaging [8], and chemical sensors [9], to mention a few applications. It possesses a number of desirable characteristics, including chemical resistance, chemical stability [10], photocatalytic activity, outstanding photostability, biocompatibility, and antibacterial activity [11]. Furthermore, titanium dioxide (TiO_2) is widely recognized as a large-bandgap semiconductor with photocatalytic activity, whereas silicon dioxide (SiO_2), often known as silica, has a well-defined ordered structure, a large surface area, is cost-effective to produce, and is easy to modify [12].

Adding SiO_2 to the formulation of a polymer is also known to increase its modulus of elasticity, strength, heat and fire resistance, wear resistance, insulating properties, and other properties [13]. In addition to drug delivery, bioimaging, gene transport, and engineering, SiO_2 is employed as a food ingredient. Moreover, the FDA has classified silica as a "generally recognized as safe" (GRAS) substance, making it an attractive candidate for biological activities [14].

Titanium–silicon oxide (Ti_xSi_y) is a fascinating material family that has gained a great deal of attention in recent years. The oxide has been widely employed as a catalyst and support for a variety of processes. It is constructed of TiO_2 and SiO_2 , and it not only possesses the advantages of both TiO_2 (photocatalytic properties and antimicrobial biomaterials) and SiO_2 (high thermal stability and excellent mechanical strength), but it also expands their applications by developing additional catalytic active sites, owing to the chemical link between the two materials [15]. There are several techniques for manufacturing Ti_xSi_y oxide, but sol–gel approaches appear to be the most cost-effective. These methods do not necessitate the use of costly solvents. Titanium alkoxide (e.g., titanium isopropoxide) is frequently utilized in sol–gel techniques.

One of the most prevalent methods for producing oxide materials is the sol–gel method. This method allows for the management of textural characteristics, as well as a high degree of mixing in bulk mixed oxides while retaining the oxide material anion-free. Furthermore, the materials produced using this method are not microstructurally ordered, have a large surface area, and show pore size variation [16,17].

The powerful technique of synchrotron X-ray absorption spectroscopy (XAS) using tunable, very intense X-rays from a high-energy electron storage ring has been applied to investigate the structural properties of materials. XAS is particularly attractive because of its ability to deliver electronic structure as well as geometric information. A typical *K*-edge absorption spectrum is divided into two sections: (i) the X-ray absorption near-edge structure (XANES) (<50 eV) contains bond lengths and angles, as well as information about the three-dimensional structure around the absorbing atom; (ii) the extended X-ray absorption fine structure (EXAFS) region (typically > 50 eV) provides information on the initial coordination shell around the absorbing atom, including coordination numbers and bond lengths [18].

In most of the previous reports, XAS has been employed to determine information on the coordination environment of tetravalent $\text{Ti}[\text{Ti}(\text{IV})]$ in structurally complex oxide materials and bond distances between Ti–O and Ti–Si atoms. Niltharach et al. [19] used XANES methods to investigate the structural details of sol–gel-produced TiO_2 samples with and without the inclusion of Ce. The XANES results also showed that the sample produced under the low hydrolysis condition had a significant number of Ti atoms in forms other than anatase and rutile TiO_2 . Won Bae Kim et al. [20] used the linear combination of two reference XANES spectra to estimate the pre-edge of the Ti *K*-edge in order to quantitatively analyze the percentages of Ti–O–Si and Ti–O–Ti bonds. The findings of pre-edge fitting in conjunction with XRD and XPS suggested that monolayer coverage was attained at around 7–10 wt.% Ti loading, where the concentration of Ti in Ti–O–Si was saturated to 0.56 mmol-Ti/g material. Shuji Matsuo et al. [21] determined the local Ti

environments in the sol, gel, and xerogels of titanium oxide prepared by a sol–gel method using titanium *K*-edge XANES. All of the samples could be divided into three groups: the anatase group, the anatase-like structure group, and the weak Ti–Ti interaction group.

A variety of materials were produced by changing the composition of the primary sol and/or the thermal treatment. The nature of the precursor, the molar ratio between silicon and titanium alkoxide, the type of solvent, and the use of modifying agents all had an influence on the final material's microstructure and hence its properties. Furthermore, while manufacturing Ti_xSi_y oxide via sol–gel, the pH of the solution is one of the most essential factors determining Ti_xSi_y oxide's properties. The hydrolysis and condensation behavior of a solution during gel formation, as well as the structure of the binary oxide, is influenced by its pH [22]. In order to maintain the sample in a powder morphology with a large surface area, the gel must be stabilized. The goal of this study was to evaluate the effects of pH and Ti/Si atomic ratio on Ti_xSi_y oxide using the sol–gel technique (Stöber method). XANES and EXAFS were used to evaluate Ti–O–Si and Ti–O–Ti connectivity in Ti_xSi_y oxides, the local atomic structure, the bond distances between Ti–O and Ti–Si atoms, the coordination number, and the valence state of titanium atoms TiO_2 and all mixed oxide samples. In addition, the photocatalytic behavior of all samples was investigated by comparing their degradation of methylene blue (MB) solutions exposed to UV radiation.

2. Materials and Methods

2.1. Materials

Tetraethylorthosilicate (TEOS, 98%, AR-grade) and titanium (IV) isopropoxide (TTIP, 98%, AR-grade) were purchased from Acros (Geel, Belgium). Absolute ethanol ($\text{C}_2\text{H}_5\text{OH}$, AR-grade), hydrochloric acid (HCl, AR-grade), and ammonium hydroxide (NH_4OH , AR-grade) were supplied by Carlo Erba Reagents (Emmendingen, Germany). PLA4043D was provided by NatureWorks LLC (Minnetonka, MN, USA).

2.2. Synthesis of titania–silica binary oxide (Ti_xSi_y oxide), silica, and titanium dioxide

Sol–gel techniques were used to prepare titanium–silicon oxide (Ti_xSi_y) utilizing a modified Stöber procedure, including the simultaneous hydrolysis and condensation of TEOS [20]. Figure 1 illustrates the experimental technique for producing Ti_xSi_y oxide particles. TEOS and TTIP were used without purification.

The total moles of Ti-alkoxide and Si-alkoxide were determined to be 0.12 mol. For the pre-hydrolysis step, the mixture of two alkoxides was added into a five-neck round-bottom flask containing 150 cm³ of ethanol for 1 h. A condenser was used in the condensation step to prevent the solvent from vaporizing out of the process. Stirring raised the temperature of the reactor to 75–80°C. After an hour of mixing, a yellowish translucent sol was observed after adding TTIP and 25 ml of ethanol to the reactor drop by drop for 1 hour. DI water was added to the sol, which was then aged at room temperature for 2 hours at 80°C. The aged gel was then refined in a centrifugal separator for 10 minutes at 6000 rpm. Finally, the solid particles were calcined for 3 hours at 450°C. A reference pure SiO_2 sample was prepared in NH_4OH in the same manner as described above but without the addition of TTIP [23]. In addition, reference pure TiO_2 was generated in the same manner as Ti_xSi_y oxide in acid [24]. Tables 1 and 2 reveal the composition of the sol–gel liquid solution used to produce Ti_xSi_y oxide.

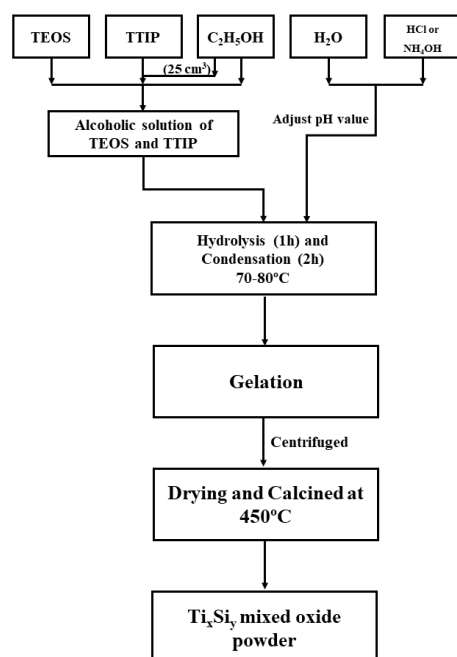


Figure 1. Experimental procedures for preparing Ti_xSi_y oxide particles.

Table 1. Moles of components in the preparation of $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at various pH values.

Samples	TTIP (mol)	TEOS (mol)	$\text{C}_2\text{H}_5\text{OH}$ (mol)	HCl/ NH_4OH (mol)	H_2O (mol)
SiO_2	-	0.120	1.889	0.060 ^b	0.440
$\text{Ti}_{50}\text{Si}_{50}$ pH 8.0	0.011	0.109	1.889	0.025 ^b	0.440
$\text{Ti}_{50}\text{Si}_{50}$ pH 9.0	0.011	0.109	1.889	0.060 ^b	0.440
$\text{Ti}_{50}\text{Si}_{50}$ pH 10.0	0.011	0.109	1.889	0.075 ^b	0.440
TiO_2	0.120	-	4.293	0.081 ^a	0.702

^a moles of HCl, ^b moles of NH_4OH .

Table 2. Moles of components in Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

Samples ^u	TTIP (mol)	TEOS (mol)	$\text{C}_2\text{H}_5\text{OH}$ (mol)	$\text{NH}_4\text{OH}/\text{HCl}$ (mol)	H_2O (mol)
SiO_2	-	0.120	1.889	0.060 ^b	0.440
$\text{Ti}_{70}\text{Si}_{30}$	0.020	0.100	1.889	0.060 ^b	0.440
$\text{Ti}_{50}\text{Si}_{50}$	0.011	0.109	1.889	0.060 ^b	0.440
$\text{Ti}_{40}\text{Si}_{60}$	0.008	0.112	1.889	0.060 ^b	0.440
TiO_2	0.120	-	4.293	0.081 ^a	0.702

^a moles of HCl, ^b moles of NH_4OH

2.3. Characterization of titanium–silicon oxide

2.3.1. ^{29}Si Solid-state nuclear magnetic resonance spectroscopy (^{29}Si solid-state NMR)

The ^{29}Si solid-state NMR spectra were used to evaluate the structure of the silicate phase in SiO_2 and Ti_xSi_y oxide. ^{29}Si NMR spectra were recorded on a Bruker Avance III HD 500 MHz spectrometer (Billerica, MA, USA).

2.3.2. Fourier transform infrared spectroscopy (FT-IR)

Infrared spectra of Ti_xSi_y oxide, TiO_2 , and SiO_2 were recorded by Fourier transform infrared spectroscopy (FTIR) (Bruker, Tensor 27, Billerica, MA, USA) using attenuated-total reflectance (ATR) equipped with a platinum diamond crystal (TYPE A225/QL). The spectra were recorded at wavenumbers from 400 to 4000 cm^{-1} with 4 cm^{-1} resolution

through the accumulation of 64 scans. All samples of Ti_xSi_y oxide, TiO_2 , and SiO_2 were dried in an oven at 70°C for 4 hours before testing.

2.3.3. Field emission scanning electron microscopy (FE-SEM)

FE-SEM images of Ti_xSi_y oxide, TiO_2 , and SiO_2 were examined by a field-emission scanning electron microscope (FESEM-EDS, Carl Zeiss Auriga, Oberkochen, Germany). The film specimens were coated with carbon prior to the investigation. An acceleration voltage of 3 kV was used to collect SEM images of the samples.

2.3.4. X-ray diffraction (XRD)

The diffractograms of Ti_xSi_y oxide, TiO_2 , and SiO_2 were recorded via powder X-ray diffraction (Bruker, Model D2 phaser, Billerica, MA, USA), with $\text{CuK}\alpha$ radiation, scanning from 10° to 90° at a rate of $0.05^\circ/\text{s}$, with a current of 35 mA and 35 mA.

2.3.5. X-ray absorption near-edge structure spectroscopy (XANES) and extended X-ray absorption fine structure spectroscopy (EXAFS)

All TiO_2 and SiO_2 standard compounds and Ti_xSi_y oxide samples were analyzed by XANES and EXAFS spectroscopy at the Ti *K*-edge at beamline 8 of the electron storage ring (using an electron energy of 1.2 GeV, a bending magnet, beam current of 80–150 mA, and 1.1 to 1.7×10^{11} photons s^{-1}) at the Synchrotron Light Research Institute (SLRI), Nakhon Ratchasima, Thailand [25]. Finely ground, homogenized powder of each sample was spread as a thin film (area $2.0\text{ cm} \times 0.5\text{ cm}$) and carefully dispersed with a spatula to yield a homogeneous particle distribution and to avoid hole effects on Kapton tape (Lanmar Inc., Northbrook, IL, USA) mounted on a sample holder. At least five spectra were obtained for each standard compound and oxide sample. All XANES and EXAFS spectra were measured in the transmission mode with ionization chamber detectors. For the acquisition of all spectra, a Ge (220) double crystal monochromator with an energy resolution ($\Delta E/E$) of 2×10^{-4} was used. An energy range of 4936–5824 eV with an energy step of 2, 0.2, 0.05 k eV was used for Ti *K*-edge spectra. The photon energy was calibrated against the *K*-edge of Ti foil at $4966 \pm 0.2\text{ eV}$. Finally, the normalized XANES and EXAFS data were processed and analyzed after background subtraction in the pre-edge and post-edge regions using a software package including (i) ATHENA(Chicago, Illinois, USA) for XAS data processing, (ii) ARTEMIS(Chicago, Illinois, USA) for EXAFS data analysis using theoretical standards from FEFF, and (iii) HEPHAESTUS(Chicago, Illinois, USA) software for a collection of beamline utilities based on tables of atomic absorption data. This package is based on the IFEFFIT library of numerical and XAS algorithms and is written in the Perl programming language using the Perl/Tk graphics toolkit [26].

2.3.6. X-ray fluorescence (XRF)

The chemical compositions of Ti_xSi_y oxide, TiO_2 , and SiO_2 were investigated by energy-dispersive XRF (EDS Oxford Instrument ED 2000, Abingdon, UK) with a Rh X-ray tube with a vacuum medium.

2.3.7. Particle size distribution

The particle size distribution of Ti_xSi_y oxide, TiO_2 , and SiO_2 was evaluated by a diffraction particle size analyzer (DPSA) (Malvern Instruments, model Mastersizer 2000, Great Malvern, England). The Ti_xSi_y oxide, TiO_2 , and SiO_2 were dispersed in absolute ethanol and analyzed by a He-Ne laser. The average particle size distribution was investigated from the standard volume percentiles at 10, 50, and 90%. The average volume-weighted diameter was used to define the average particle size.

2.3.8. Specific surface area (BET)

The specific surface areas of Ti_xSi_y oxide, TiO_2 , and SiO_2 were investigated by a nitrogen adsorption analyzer (BELSORP MINI II, model Bel-Japan, Osaka, Japan). The samples of Ti_xSi_y oxide, TiO_2 , and SiO_2 were degassed at 300°C for 3 h before measurement. The Brunauer, Emmett, Teller (BET) and Barrett, Joyner, Halenda (BJH) methods were used to calculate the specific surface area and pore size distribution, respectively.

2.3.9. Photocatalytic activity characterization

PLA, PLA/ TiO_2 , PLA/ SiO_2 , and PLA/ Ti_xSi_y oxide composites with 3 wt.% of Ti_xSi_y oxide were prepared by the solvent-casting method. The polymer was dissolved in solvent (10% w/v) by stirring. Firstly, Ti_xSi_y oxide was dispersed in chloroform with ultrasonic treatment for 1 day. After this, PLA was added to the Ti_xSi_y oxide and strictly stirred for 4 days. The dispersions of PLA composites were additionally ultrasonically treated for 1 h using a frequency of 42 kHz, four times per day. The treated dispersions were gently poured onto Petri dishes, and the solvent was evaporated at room temperature. The films were dried to a constant mass, dried at room temperature for 24 h, and kept in the oven at 40°C for 4 h. Film thickness was measured by using a micrometer in eight replicates for each sample, from which an average value was attained. Pure PLA and nanocomposite films had a uniform thickness of 200–250 μm .

The photocatalytic activity of PLA composite films was evaluated in terms of the degradation of methylene blue (MB) under UV light according to Chinese standard GB/T 23762-2009. PLA composite films were placed in a flask and then 200 mL of MB solution (10 mg/L) was added. The flask was placed on a mechanical shaker at 50 rpm in a UV chamber with 4 UV lamps (LP Hg lamps, 8 watts, main light emission at 245 nm). The schematic design of the investigational setup for the photocatalytic process is presented in Figure 2. Then, 4 mL MB solution was collected every 60 min and analyzed using a UV–vis spectrophotometer (Cary300, Agilent Technology, CA, USA). In order to maintain the volume of MB solution in the flask, the samples were returned after each measurement. The maximum absorbance of MB occurred at 664 nm (Figure 3). The spectrometer was calibrated with a solution of MB at 1mg/L, 3mg/L, 5mg/L, 7mg/L, and 10mg/L concentrations, respectively. The calibration curve of methylene blue aqueous solutions is shown in Figure 4. In order to precisely predict the decomposition of MB under UV light, the PLA and composite films, as well as the solution, were stored in a black box without any photocatalyst. The concentration of MB was also collected every 60 min to estimate the absorption of MB.

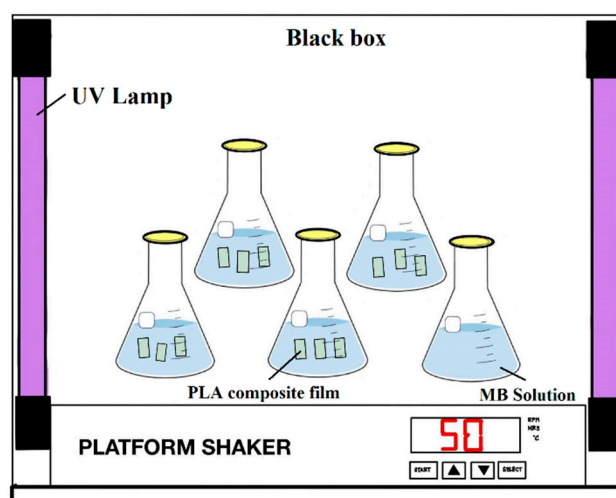


Figure 2. Schematic illustration of the experimental setup for photocatalytic process.

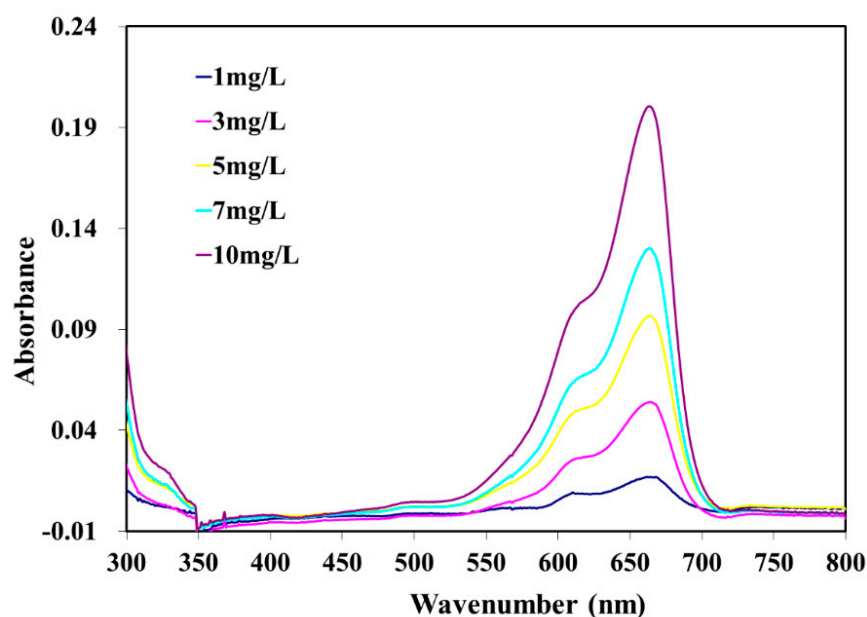


Figure 3. Wavelength-absorbance curves of methylene blue (MB) aqueous solutions.

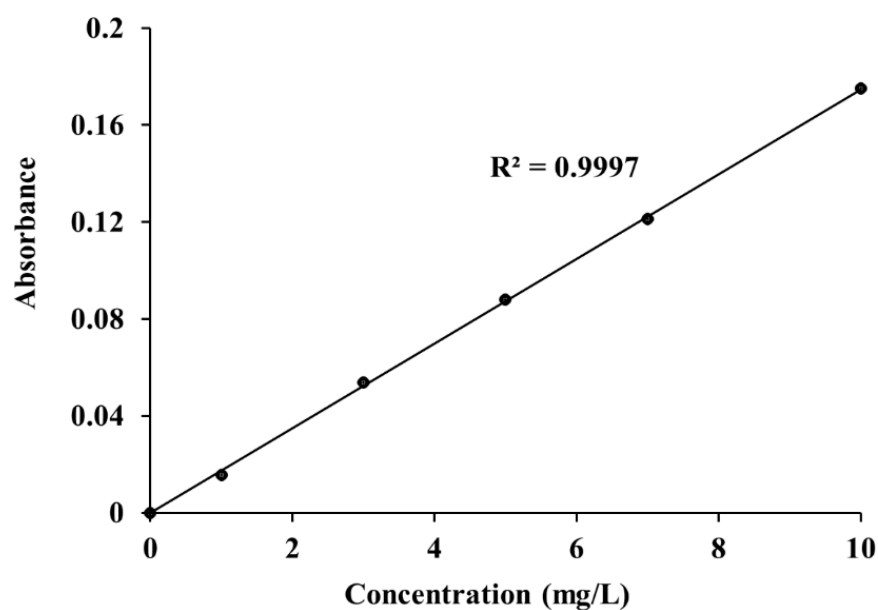


Figure 4. Calibration curve of methylene blue (MB) aqueous solutions.

3. Results

3.1. Effect of pH

^{29}Si Solid-state NMR spectroscopy was used to investigate the structures of Ti_xSi_y oxide (particularly for 1:1 atomic ratio) and SiO_2 in order to determine the development of the framework of silicon atoms in these materials. The structural analysis of ^{29}Si NMR is illustrated in Figure 5, which presents the spectra of the silica and the $\text{Ti}_{50}\text{Si}_{50}$ oxide. In general, the ^{29}Si NMR spectra of oxides show three peaks for different environments for Si: isolated silanol groups $(\text{SiO})_3\text{Si}-\text{OH}$, Q^3 ; geminal silanols $(\text{SiO})_2\text{Si}-(\text{OH})_2$, Q^2 ; and silicon in the siloxane binding environment without hydroxyl groups $(\text{SiO})_4\text{Si}$, Q^4 (Q^n represents a SiO_4 unit, with n being the number of bridging oxygen atoms) [27,28]. According to related references, the three resonance signals of pure SiO_2 (synthesized by sol-gel

method) appearing at -93.9, -103.3, and -110.2 ppm can be attributed to the configurations of Q^2 , Q^3 , and Q^4 , respectively. The spectrum of silica (Figure 5a) exhibits a chemical shift of -110 ppm, which is attributed to the Q^4 units, and a minor peak appears at -103 ppm that would conform to the Q^3 units corresponding to the siloxane (Si-O-Si) and silanol ($\equiv\text{Si-OH}$) groups of silica. The strength of Q^4 structural units $[\text{Si}(\text{SiO})_4]$ represents a three-dimensional network of silica, as evidenced by the intensity of the peaks in Figure 5a. There would also be a very small proportion of Si nuclei directly bonded to the hydroxyl groups, influenced by the chemical shift of the Q^3 structural unit $[\text{Si}(\text{SiO})_3\text{OH}]$ [17]. Additionally, Figure 5b indicates that silica was chemically combined with titanium oxide for the formation of the $\text{Ti}_{50}\text{Si}_{50}$ oxide, causing Ti atoms to be disrupted in the three-dimensional silica network and the microstructure surrounding the silicon atoms. When comparing the peak at approximately -103 ppm to the quantity of siloxane (Si-O-Si) and silanol ($\equiv\text{Si-OH}$) groups in pure SiO_2 , a very substantial increase in Q^3 units (mono-substitution) is detected, which might be due to the existence of significant Si-O-Ti bonds in this sample, signaling atomic mixing. The presence of Ti-O-Si-O-Ti-O groups may have caused the appearance of Q^2 units, which might be linked to disubstituted silica in tetrahedra [29]. A similar observation was reported by Pabón, Retuert, and Quijada (2007) [17]. Obviously, Ti atoms replaced some Si atoms in the SiO_2 network, indicating that, as a consequence of the synthesis performed by the sol-gel method defined in this work, a large number of Si-O-Ti bonds were created, in agreement with the results of the FTIR and XRD.

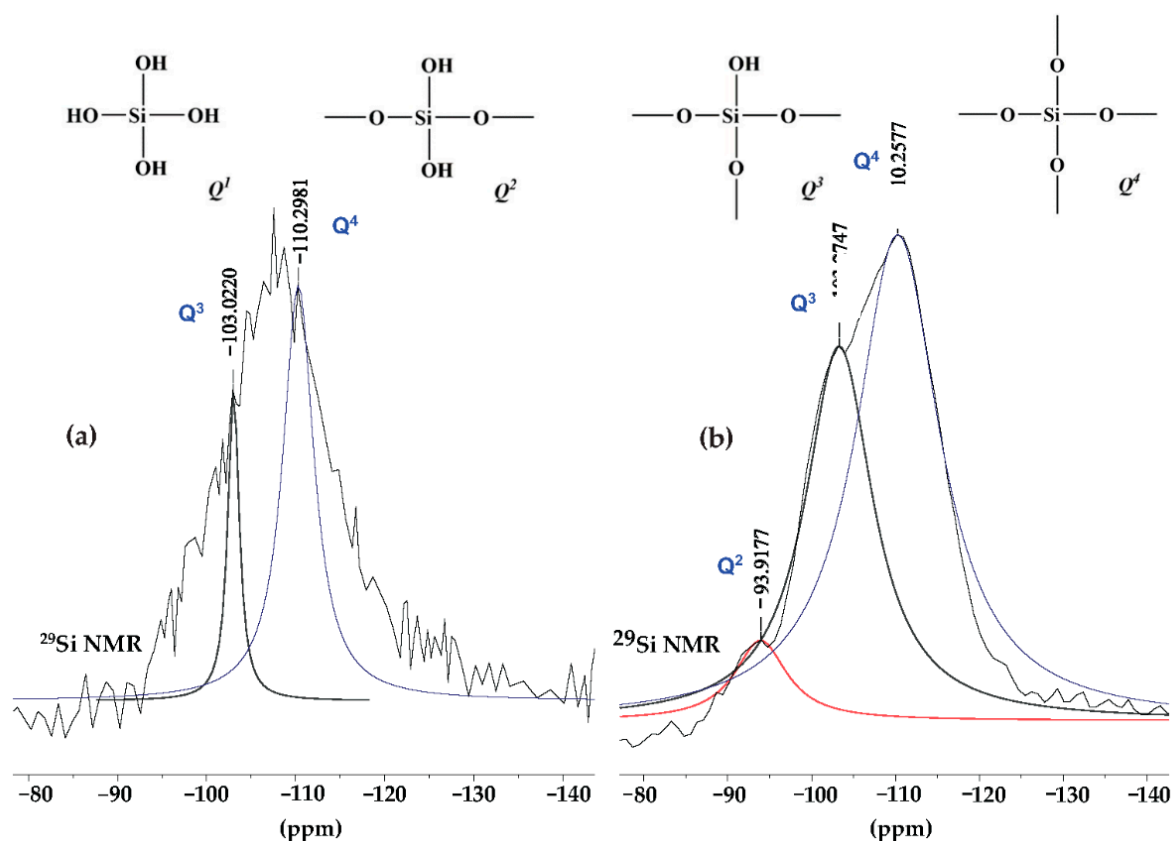


Figure 5. ^{29}Si -NMR patterns of (a) SiO_2 and (b) $\text{Ti}_{50}\text{Si}_{50}$ oxide.

FTIR spectra of TiO_2 , SiO_2 , and $\text{Ti}_{50}\text{Si}_{50}$ oxides synthesized at various pH values are shown in Figure 6. TiO_2 and all $\text{Ti}_{50}\text{Si}_{50}$ oxides displayed a prominent band at 400 to 700 cm^{-1} in their FTIR spectra, which corresponded to the bending and stretching mode of Ti-O-Ti and was characteristic of well-ordered TiO_6 octahedrons [30].

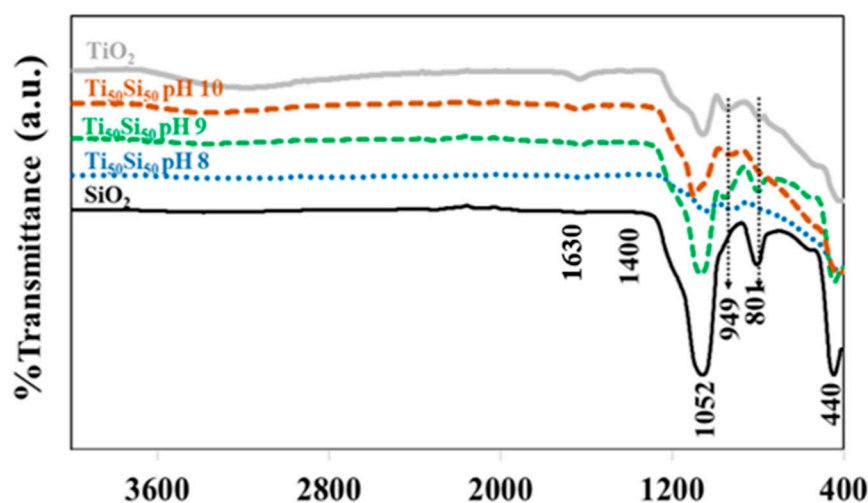


Figure 6. FTIR spectra of TiO_2 , SiO_2 , and Ti_xSi_y oxides synthesized at pH 8.0, 9.0, and 10.0.

Strong bands near 3440 and 1630 cm^{-1} can be seen in Figure 6, which are assigned to the stretching and deformation vibration of the hydroxyl groups present ($\text{TiO}_2\text{-OH}$) on the surfaces of TiO_2 and all mixed oxide samples. The intensive band of OH-group asymmetrical and symmetrical stretching vibrations at 3440 cm^{-1} and O-H deformation vibration at 1630 cm^{-1} could confirm the large amount of water molecules [31]. However, the OH-group vibration bands were substantially weaker as the calcination temperatures increased [32]. The peaks at 440, 801, and 1052 cm^{-1} corresponded to the rocking, symmetric, and asymmetric stretching vibrations of Si-O-Si of silica, respectively [3]. However, TiO_2 also presents a peak at 1000–1200 cm^{-1} , which is assigned to the deformation vibration of Ti-O-Ti of TiO_2 [33]. The peak at 949 cm^{-1} was commonly attributed to Ti-O-Si vibrations by various authors, indicating the existence of Ti-O-Si linkages [34]. This band can also be attributed to the Si-OH bond, which should be attributed to Si-OH bonding. However, this absorption band disappeared after the SiO_2 particles were calcined at 450°C. On the other hand, the 949 cm^{-1} absorption band was still observed for all Ti_xSi_y oxide samples that had been calcined at a temperature of 450°C. Accordingly, the 949 cm^{-1} peak in the FTIR spectra for Ti_xSi_y oxide particles should not be attributed to Si-OH bonds but rather be associated with Si-O-Ti bonding [35,36]. This indicated titanium being combined into the framework of silica. Therefore, the influence of pH on the FTIR of the $\text{Ti}_{50}\text{Si}_{50}$ oxide was not observed. Based on the evidence from NMR and FTIR, the probable mechanism for the synthesized Ti_xSi_y oxides based on results of earlier studies [37] is suggested in Figure 7.

Figure 8 shows FE-SEM images of SiO_2 and $\text{Ti}_{50}\text{Si}_{50}$ oxide powder nanoparticles synthesized at various pH values. SiO_2 particles were mostly spherical in shape, with low agglomeration (Figure 8a). Meanwhile, large bulk particles with high agglomeration of $\text{Ti}_{50}\text{Si}_{50}$ oxides are shown in Figure 8b. This agglomeration with an irregular shape was a result of the Ti and Si precursors in the sol solution during the processing, synthesized due to the absence of appropriate hydroxide ions (OH^- ions) at pH 8.0 [38]. Figure 8c and 8d show that the nanoparticles were spherical in shape and the size distribution of particles became more uniform, with a respectable nanostructure, upon increasing the pH to alkaline conditions (pH 9.0 and 10.0). In this work, Ti_xSi_y oxide nanoparticles were essentially spherical in form, with low agglomeration. However, the size of the $\text{Ti}_{50}\text{Si}_{50}$ oxide particles at pH 10.0 was nano-sized, with the appearance of agglomeration (Figure 8d). Wu, Wu, and Lü, (2006) claimed that the colloidal sol-gel technique can generate large particles composed of agglomerated nanoparticles and either condensed or porous polycrystalline microparticles [39]. Therefore, pH 9.0 was selected to achieve the compositionally controlled synthesis of Ti_xSi_y oxide particles that are spherical in shape.

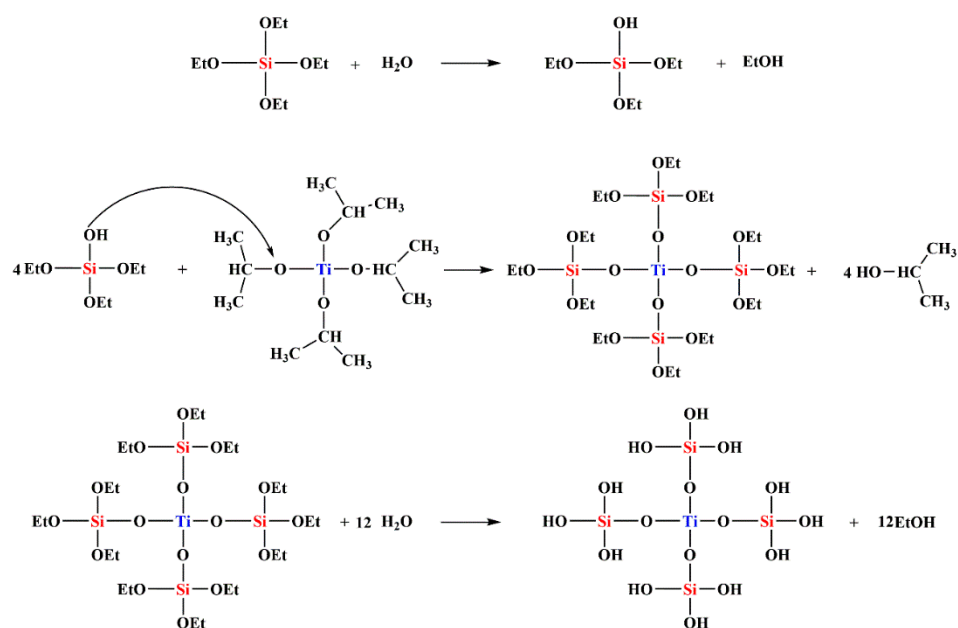


Figure 7. Proposed probable mechanism for synthesized Ti_xSi_y oxides.

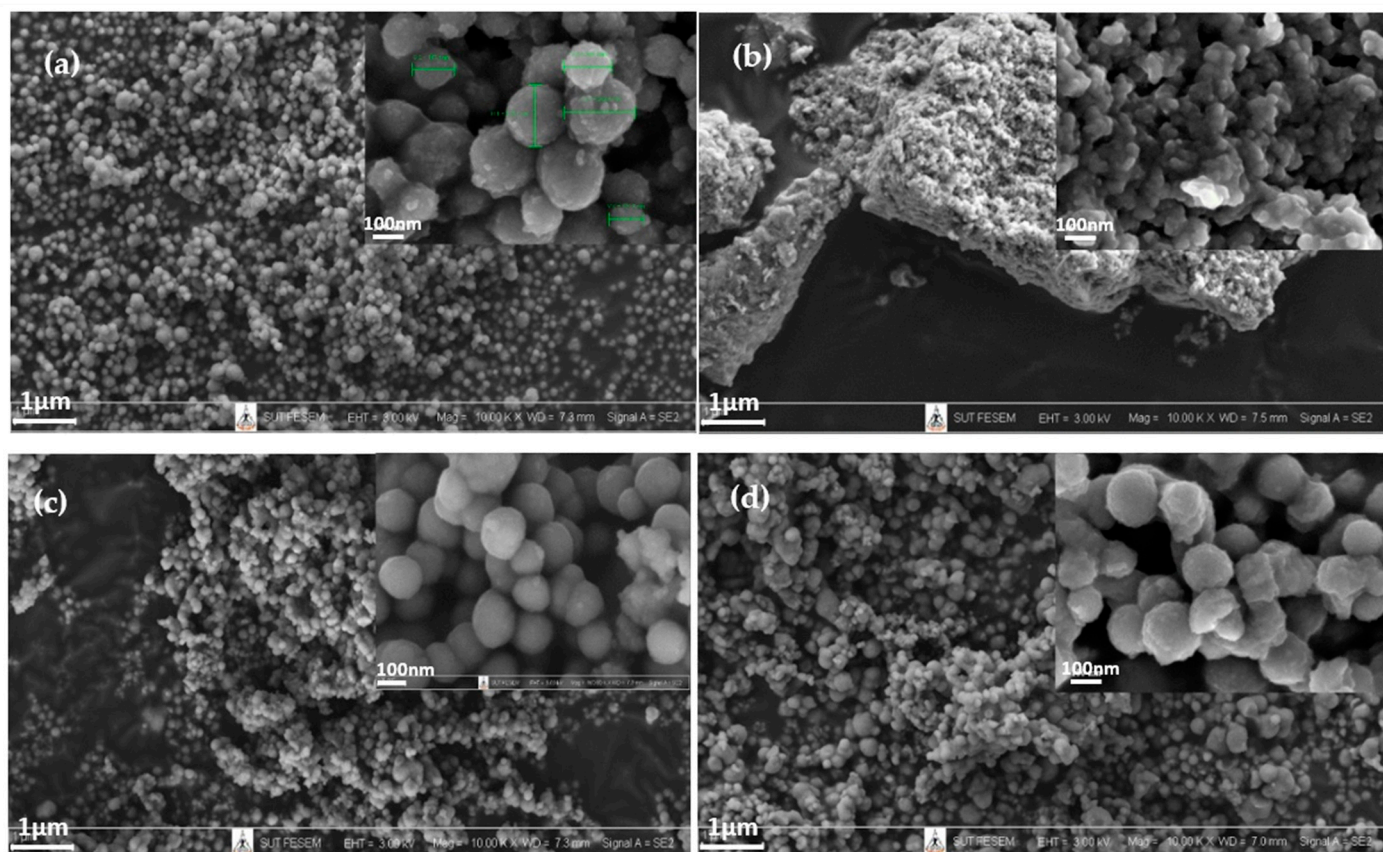


Figure 8. Low ($\times 10\text{k}$ WD = 7.0–7.5 mm EHT = 3.00 kV) and high ($\times 100\text{k}$ WD = 6.9–7.3 mm EHT = 3.00 kV) magnification SEM images of $\text{Ti}_{50}\text{Si}_{50}$ oxides synthesized at various pH values: (a) SiO_2 , (b) $\text{Ti}_{50}\text{Si}_{50}$ pH 8.0, (c) $\text{Ti}_{50}\text{Si}_{50}$ pH 9.0, (d) $\text{Ti}_{50}\text{Si}_{50}$ pH 10.0.

XRD patterns of $\text{Ti}_{50}\text{Si}_{50}$ oxide powders synthesized at various pH values (8.0, 9.0, and 10.0) are shown in Figure 9. Silica showed one amorphous peak located at 2θ angles at 23° . In most of the $\text{Ti}_{50}\text{Si}_{50}$ oxide powders prepared at pH 8.0 and 9.0 (Figure 9), XRD patterns exhibited a broad amorphous peak because of the existence of silica. Furthermore, the formation of the $\text{Ti}_{50}\text{Si}_{50}$ oxides suggests that titanium oxide was combined with the silica by atomic mixing, and consequently the distribution of Ti atoms occurred in the microstructure around the silicon atoms and the three-dimensional silica network. However, the $\text{Ti}_{50}\text{Si}_{50}$ oxide powders prepared at pH 10.0 exhibited an intense TiO_2 anatase peak, suggesting that the formation of higher crystallites of anatase occurred in $\text{Ti}_{50}\text{Si}_{50}$ oxides when the high pH of the sol–gel solution was applied during hydrolysis [38]. It can be clearly seen that the crystallinity of anatase in $\text{Ti}_{50}\text{Si}_{50}$ oxide samples is enhanced by adjusting the pH value of the sol.

Furthermore, the XRD peaks for $\text{Ti}_{50}\text{Si}_{50}$ oxide powder with $\text{pH} > 9$ present a crystalline nature with 2θ angles at 25° (101) and 48° (200), which correspond to the crystalline phase (anatase) of TiO_2 [3,34], due to the presence of a sufficient amount of OH^- to form TiO_2 . It is possible that the pH influences the particle size and crystallinity of $\text{Ti}_{50}\text{Si}_{50}$ oxides. The crystal structure of TiO_2 was tetragonal; $a = b = 3.782 \text{ \AA}$, $c = 9.502 \text{ \AA}$ [40]. However, no rutile phase was detected in any case, according to the absence of (110) diffraction peak at 27.4° .

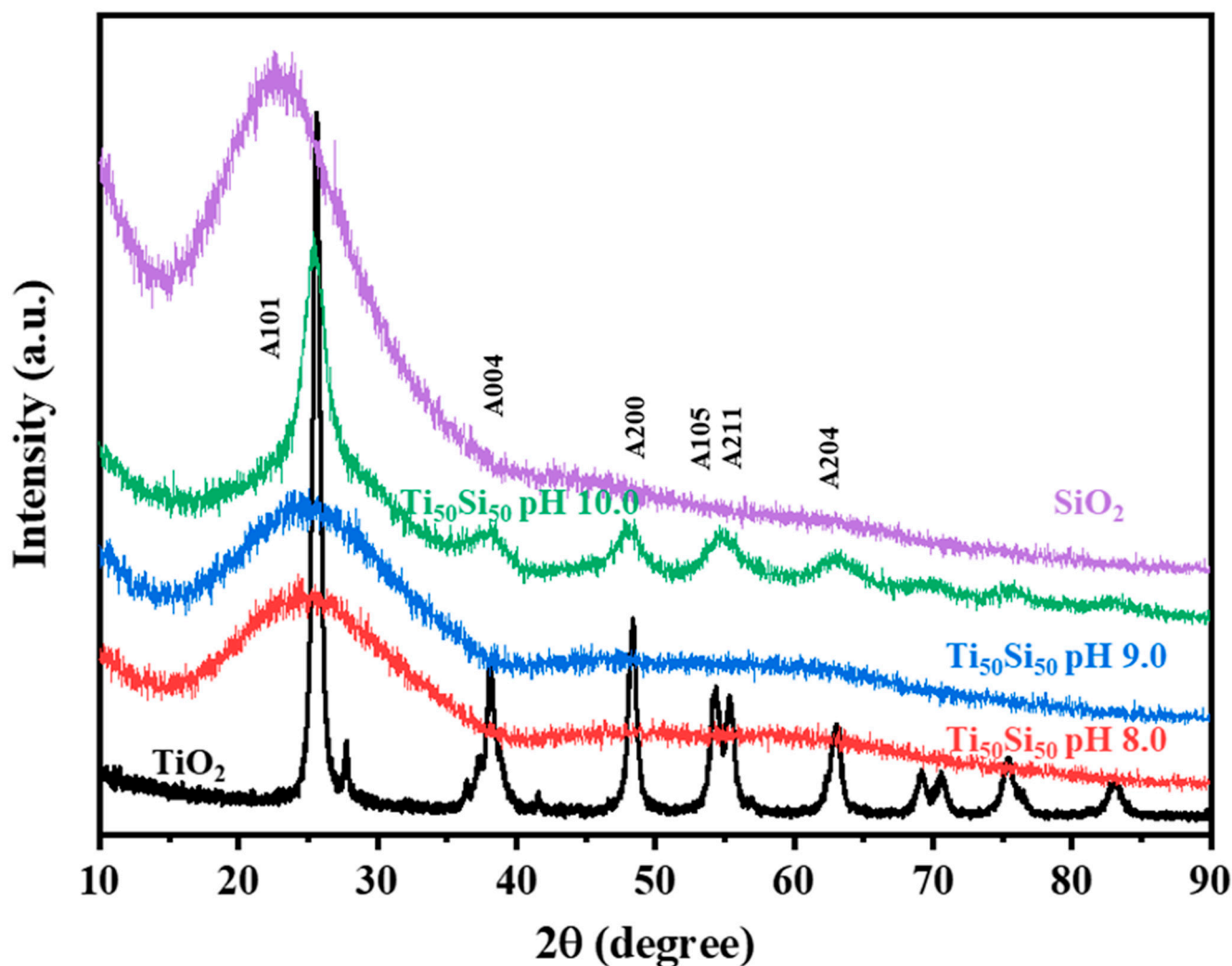


Figure 9. XRD patterns of $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized with pH 8.0, 9.0, and 10.0.

Titanium K-edge X-ray-absorption near-edge structure (XANES) spectroscopy is widely used to derive information on the coordination environment of tetravalent Ti[Ti(IV)] in structurally complex oxide materials. Normalized Ti K-edge XANES spectra of anatase TiO₂ standards, synthesized TiO₂ materials, and Ti₅₀Si₅₀ oxides synthesized at various pH values were collected and are presented in Figure 10. The spectrum is mostly separated into two regions: (i) the pre-edge region (4940 to 4990 eV) and (ii) the post-edge region (4990 to 5010 eV) [41]. The anatase is composed of pre-edge features A₁–A₃ and a characteristic shoulder B [42,43]. The origin of these features was described by Farges et al. The first pre-edge (A₁) is mainly attributed to quadrupolar transitions to t_{2g} levels of the TiO₆ octahedron. The second pre-edge (A₂) and the third pre-edge (A₃) are attributed to 1s to 3d dipolar transitions [43]. The B features may be attributed to the interactions of the central Ti 4p orbital hybridized with the near Ti or O atom (Wu et al., 1997). The intensive absorbance of the B feature may be due the local structures of the Ti orbital hybridized for the TiO₂–SiO₂ photocatalyst [41]. Both TiO₂ anatase and synthesized TiO₂ display three minor pre-edge peaks, which are attributed to transitions from the 1s level of Ti to 1t_{1g}, 2t_{2g}, and 3e_g molecular orbitals (in sequence of increasing energy) [44]. The three minor pre-edge peaks for TiO₂ anatase are outstanding fingerprints for the crystalline phase of titanium dioxide materials. Thus, the formation of anatase is specified by the near-edge region of TiO₂.

As silicon merged into the Ti₅₀Si₅₀ oxide, a large single pre-edge feature was observed to dominate the other weaker feature. From Figure 10, the significant pre-edge feature in Ti₅₀Si₅₀ oxide samples indicates that a larger proportion of Ti atoms occupy tetrahedral symmetry positions, as expected for samples where Ti directly replaces Si in the oxide sample. Moreover, the pre-edge area of Ti₅₀Si₅₀ oxides was not affected by the various pH values applied during synthesis. Furthermore, the XANES spectra of the Ti₅₀Si₅₀ oxides have a higher intensity of pre-edge peaks than TiO₂ anatase, indicating that Ti is no longer entirely confined to octahedral sites when combined with Si. Therefore, the results of the XANES spectra of mixed oxide samples are in reasonable agreement with the XRD results.

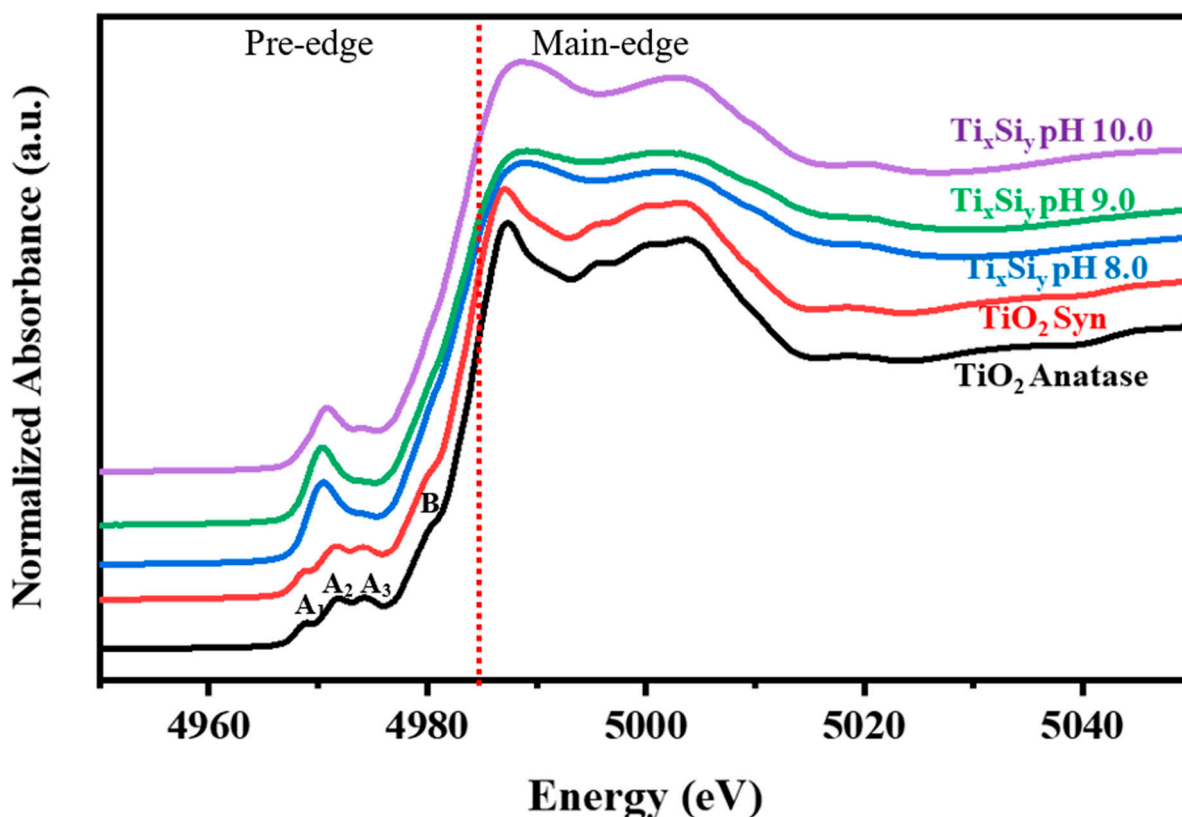
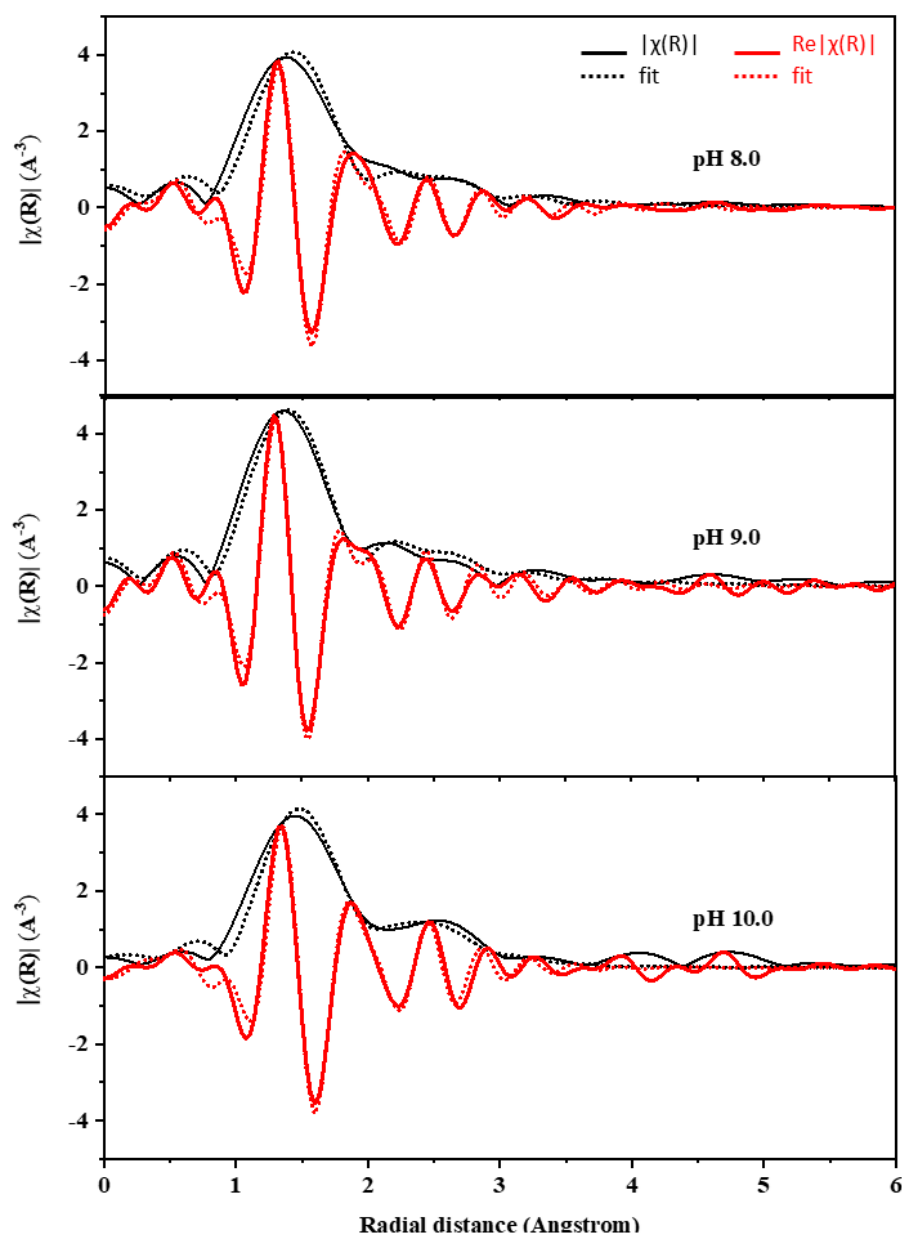


Figure 10. Ti K-edge XANES spectra of $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized with pH 8.0, 9.0, and 10.0.

The EXAFS analysis at the Ti K-edge was used to investigate the coordination number of Ti atoms and bond distances between Ti–O and Ti–Si atoms. The program package of ATHENA–ARTEMIS was used to perform the background noise correction and the normalization of the raw data [26]. The EXAFS data were FT to the R-space to define the local atom structure and relative bond length with respect to the absorbing atom. The ATOM and FEFF codes were used to simulate the FT data (in R-space) by creating a systematic theoretical structure of SiO_2 [45]. The range of data reserved for the transformation was $3\text{--}9\text{ \AA}^{-1}$ in k-space. Structural parameters were attained, without the phase corrections, by fitting the data in R-space, within the intermission of $1.1\text{--}3\text{ \AA}$ (Figure 11). The shell parameters, such as Ti–O and Ti–Si bond distances (R), coordination number (CN), Debye–Waller factor (σ^2), and R-factor, are presented in Table 3. Note that the amplitude reduction factor (S_0^2) was preliminarily determined to be 0.52 by fitting the EXAFS spectrum of SiO_2 .

**Figure 11.** Ti K-edge XANES spectra of $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at pH 8.0, 9.0, and 10.0.

The first and second shells in the FTs arise from the single scattering paths of cation–oxygen and cation–cation, respectively [46]. In this study, we first focus on the Ti–O and Ti–Si shells of all Ti_xSi_y oxide samples. TiO_2 anatase has an average Ti–O bond distance of 1.95 Å. The results from curve fitting of the EXAFS provide an interatomic distance of 1.99, 1.82, and 1.94 Å for Ti atom dilute silica of $\text{Ti}_{50}\text{Si}_{50}$ oxides synthesized at various pH levels of 8.0, 9.0, and 10.0, respectively. This result can be compared favorably with the bond distance of Ti–O in the titanium silicalite molecular sieve of 1.80 Å, corresponding to the cation existing in a tetrahedral environment [47,48]. Neurock and Manzer (1996) calculated the Ti–O bond distance using density functional methods, obtaining a value of 1.81 Å, suggesting Ti in a tetrahedral cluster [49]. On the other hand, we found that there is evidence of a longer Ti–O distance at 1.99 at pH 8.0, indicating a six-fold Ti site (octahedral), which agrees with Loshmanov et al. [47], who investigated TiO_2 in SiO_2 with neutron diffraction and attempted to isolate the Ti–O bond by subtracting their radial distribution functions from that of pure SiO_2 . They assigned a value of 1.95 Å. By analogy with the Ti–O distance in its octahedrally bonded oxides (1.91–2.01 Å), they inferred an octahedral coordination for Ti in SiO_2 . This may be due to the effect of pH on the reaction of the Ti_xSi_y oxides synthesized. Obviously, the value of 1.83 Å described for the Ti–O bond distance in Ti_xSi_y oxides synthesized at pH 9.0 is agreeable with the absolute value of the Ti atoms occupied in a tetrahedral environment, which supports the comparable assumption from XANES. Furthermore, all of the $\text{Ti}_{50}\text{Si}_{50}$ oxides presented a small value for the Debye–Waller factor, corresponding to the structural disturbance in the materials [45]. The Debye–Waller factor is used to describe the influence of static and thermal disorder on the EXAFS spectrum. A large value for the σ^2 can be caused by a variance in the ligand distances (static disorder), whereas small Debye–Waller factors artificially increase the ligand contribution [50,51].

Table 3. Structural parameters (obtained from fittings of Ti K-edge EXAFS data) of first two coordination shells around Ti atom of Ti and Si in $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at various pH values and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

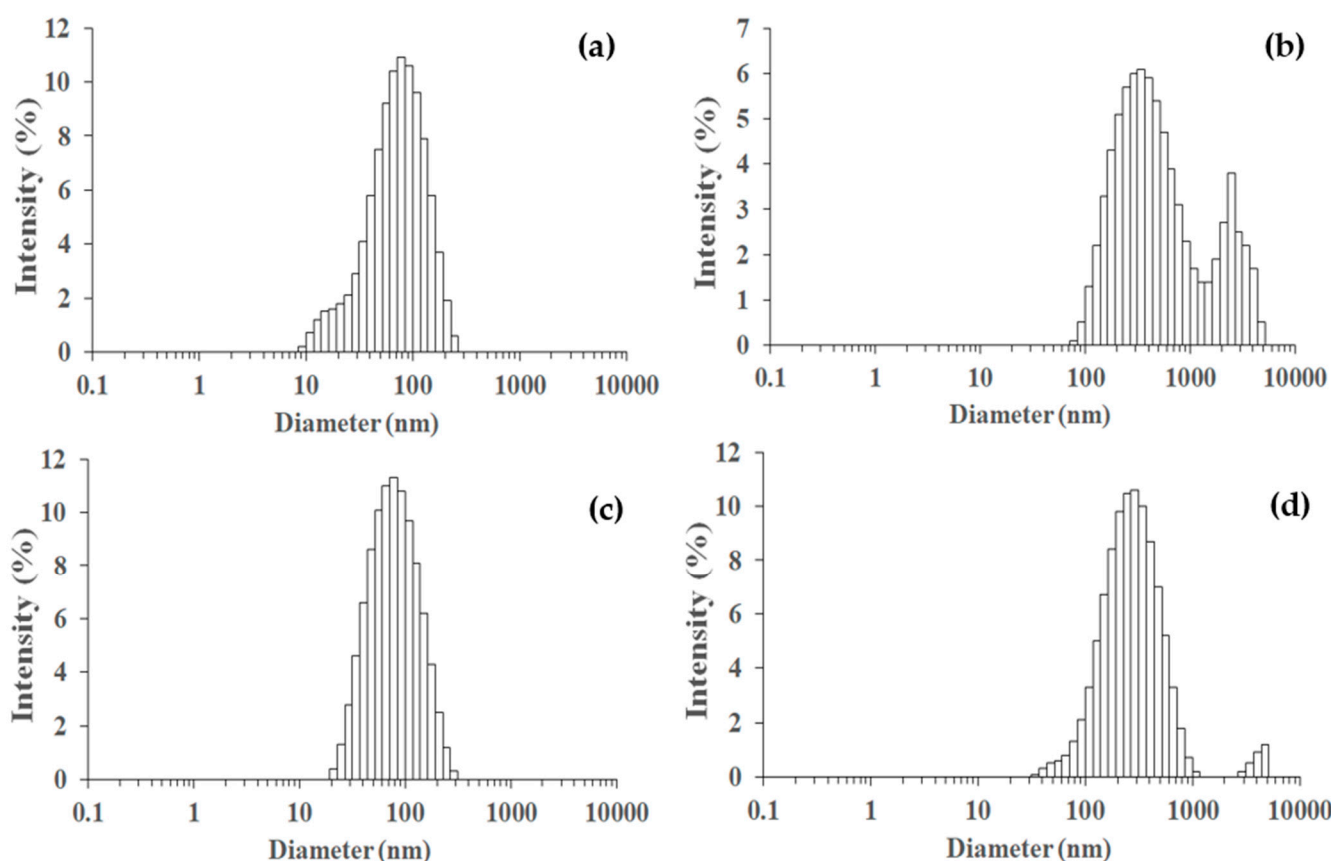
Sample	Shell	Bond	$\Delta E(\text{eV})$	CN	$R(\text{\AA})$	σ^2	R-factor
$\text{Ti}_{50}\text{Si}_{50}$ pH 8.0	1	Ti–O	4.11±3.34	4.14 ±0.18	1.99±0.02	0.0053±0.0013	0.0176
	2	Ti–Si		1.27±0.44	2.79±0.03	0.0018±0.0019	
$\text{Ti}_{50}\text{Si}_{50}$ pH 9.0	1	Ti–O	2.72±2.03	4.07±0.12	1.83±0.01	0.0042±0.0008	0.0082
	2	Ti–Si		1.00±0.13	2.77±0.02	0.0020±0.0020	
$\text{Ti}_{50}\text{Si}_{50}$ pH 10.0	1	Ti–O	6.70±3.29	4.08±0.19	1.94±0.02	0.0049±0.0024	0.0158
	2	Ti–Si		1.52±0.22	2.81±0.03	0.0013±0.0031	
$\text{Ti}_{70}\text{Si}_{30}$	1	Ti–O	6.68±3.23	4.15±0.19	1.90±0.02	0.0051±0.0014	0.0198
	2	Ti–Si		0.97±0.21	2.800±0.04	0.0006±0.0050	
$\text{Ti}_{50}\text{Si}_{50}$	1	Ti–O	2.72±2.03	4.07±0.12	1.83±0.01	0.0042±0.0008	0.0082
	2	Ti–Si		1.00±0.13	2.77±0.02	0.0020±0.0020	
$\text{Ti}_{40}\text{Si}_{60}$	1	Ti–O	6.54±2.27	4.09 ±0.13	1.82±0.01	0.0048±0.0010	0.0103
	2	Ti–Si		1.19±0.03	2.79±0.03	0.0023±0.0017	

The chemical compositions of TiO_2 , SiO_2 , and $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at various pH values by XRF are shown in Table 4. The XRF analysis showed that $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at pH 8.0 showed a higher percentage of Si (54.39%) than those synthesized at pH 9.0 (47.87%) and pH 10.0 (9.82%) because TTIP can be hydrolyzed more quickly than TEOS in a solution at a high pH. In contrast, the atomic content of Ti in $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at various pH values increased with the increasing pH value of the sol (43.72%).

Table 4. Percentage of atomic and atomic ratio of Ti and Si in TiO₂, SiO₂, and Ti₅₀Si₅₀ oxide synthesized at various pH values and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

Sample	Atomic (%)		Atomic ratio of Ti/Si
	Si	Ti	
SiO ₂	100	0	-
Ti ₅₀ Si ₅₀ pH 8.0	54.39	45.61	0.84
Ti ₅₀ Si ₅₀ pH 9.0	47.87	52.13	1.09
Ti ₅₀ Si ₅₀ pH 10.0	9.82	93.18	9.49
Ti ₇₀ Si ₃₀	29.79	70.21	2.36
Ti ₅₀ Si ₅₀	47.87	52.13	1.09
Ti ₄₀ Si ₆₀	60.19	39.81	0.66
TiO ₂	0	100	-

The particle sizes of TiO₂, SiO₂, and Ti₅₀Si₅₀ oxides synthesized at various pH values as determined by the Zetasizer Nano ZS and SEM micrographs are shown in Table 5 and Figure 12. The mean particle size distributions of TiO₂ and SiO₂ were 40.3 and 147.8 nm, while Ti₅₀Si₅₀ oxide synthesized at pH 9.0 presented a mean particle size distribution of 136.2 nm. However, the mean particle size distribution of Ti₅₀Si₅₀ oxide synthesized at pH 8.0 and 10.0 covered a range of 350–650 nm due to the tendency towards agglomeration.

**Figure 12.** Particle size distributions measured on a Zetasizer Nano ZS of SiO₂ and Ti₅₀Si₅₀ oxide synthesized at various pH values: (a) SiO₂, (b) Ti₅₀Si₅₀ pH 8.0, (c) Ti₅₀Si₅₀ pH 9.0, (d) Ti₅₀Si₅₀ pH 10.0.

To determine the correct diameter of TiO_2 , SiO_2 , and Ti_xSi_y oxide, SEM micrographs (Figure 8) were used for measurement with ImageJ. The number average diameter, d_n , was calculated from a minimum of 200 particles according to the following equations [52]:

$$d_n = \frac{\sum_i n_i d_i}{\sum_i n_i} \quad (1)$$

$$SD = \sqrt{\frac{\sum_i n_i (d_i - d_n)^2}{N}} \quad (2)$$

where d_i is the diameter of a particle and n_i is the total number of particles having diameter, d_i .

It can be observed that the number average diameter of the sample increased with the increasing pH value of the synthesis.

The textural properties of the SiO_2 , TiO_2 , and $\text{Ti}_{50}\text{Si}_{50}$ oxides synthesized at various pH values are shown in Table 6. It can be observed that the surface areas (S_{BET}) of the oxides were 177.02, 225.68, and $62.75 \text{ m}^2\text{g}^{-1}$ for $\text{Ti}_{50}\text{Si}_{50}$ synthesized at pH 8.0, 9.0, and 10.0, respectively. In addition, the mean pore diameter of the $\text{Ti}_{50}\text{Si}_{50}$ oxides was 2.56, 5.85, and 27.31 nm. Meanwhile, the S_{BET} of SiO_2 and TiO_2 was 116.90 and $74.07 \text{ m}^2\text{g}^{-1}$, with an average pore size of 37.08 and 12.75 nm.

Table 5. Particle sizes of SiO_2 , TiO_2 , and $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at various pH values, and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

Sample	Zetasizer Nano ZS		SEM
	Particle size, d (nm)	Polydispersity index, PDI	Value of average diameter, d_n (nm)
SiO_2	147.8	0.248	144.2 ± 11.3
$\text{Ti}_{50}\text{Si}_{50}$ pH 8.0	639.3	0.627	41.5 ± 16.7
$\text{Ti}_{50}\text{Si}_{50}$ pH 9.0	136.2	0.239	135.4 ± 12.3
$\text{Ti}_{50}\text{Si}_{50}$ pH 10.0	397.0	0.322	114.2 ± 24.2
$\text{Ti}_{70}\text{Si}_{30}$	573.1	0.630	149.3 ± 15.5
$\text{Ti}_{50}\text{Si}_{50}$	136.2	0.239	135.4 ± 12.3
$\text{Ti}_{40}\text{Si}_{60}$	825.5	0.907	131.8 ± 13.3
TiO_2	40.3	0.490	27.8 ± 6.3

Table 6. Specific surface area, pore volume, and mean pore size of TiO_2 , SiO_2 , and $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at various pH values, and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

Sample	Specific Surface Area, S_{BET} (m^2g^{-1})	Pore Volume, V_p (cm^3g^{-1})	Mean pore diameter, d_p (nm)
SiO_2	116.90	1.06	51.08
Ti_xSi_y pH 8.0	177.02	0.11	2.56
Ti_xSi_y pH 9.0	225.68	0.33	5.85
Ti_xSi_y pH 10.0	62.75	0.42	27.31
$\text{Ti}_{70}\text{Si}_{30}$	569.07	1.42	10.96
$\text{Ti}_{50}\text{Si}_{50}$	225.68	0.33	5.85
$\text{Ti}_{40}\text{Si}_{60}$	68.34	0.36	21.97
TiO_2	74.07	0.23	12.75

A typical nitrogen adsorption–desorption isotherm of the SiO_2 , TiO_2 , and $\text{Ti}_{50}\text{Si}_{50}$ oxides synthesized at various pH values is presented in Figure 13. The isotherm of SiO_2 is a type II curve according to the IUPAC isotherm classification, indicating that the obtained SiO_2 particles contain macropores. An inflection point occurs near the completion of the first adsorbed monolayer, which is typical for non-porous or macroporous materials with a pore size of >50 nm. In addition, all $\text{Ti}_{50}\text{Si}_{50}$ oxides exhibited type II isotherms with pore diameters of 2.56, 5.85, and 27.31 nm. However, the isotherm of TiO_2 is a type IV curve

according to the IUPAC isotherm classification. An abrupt hysteretic loop is detected from this curve, which is typical for mesoporous materials with pores in the range of 2 to 50 nm that show capillary condensation and evaporation [53,54].

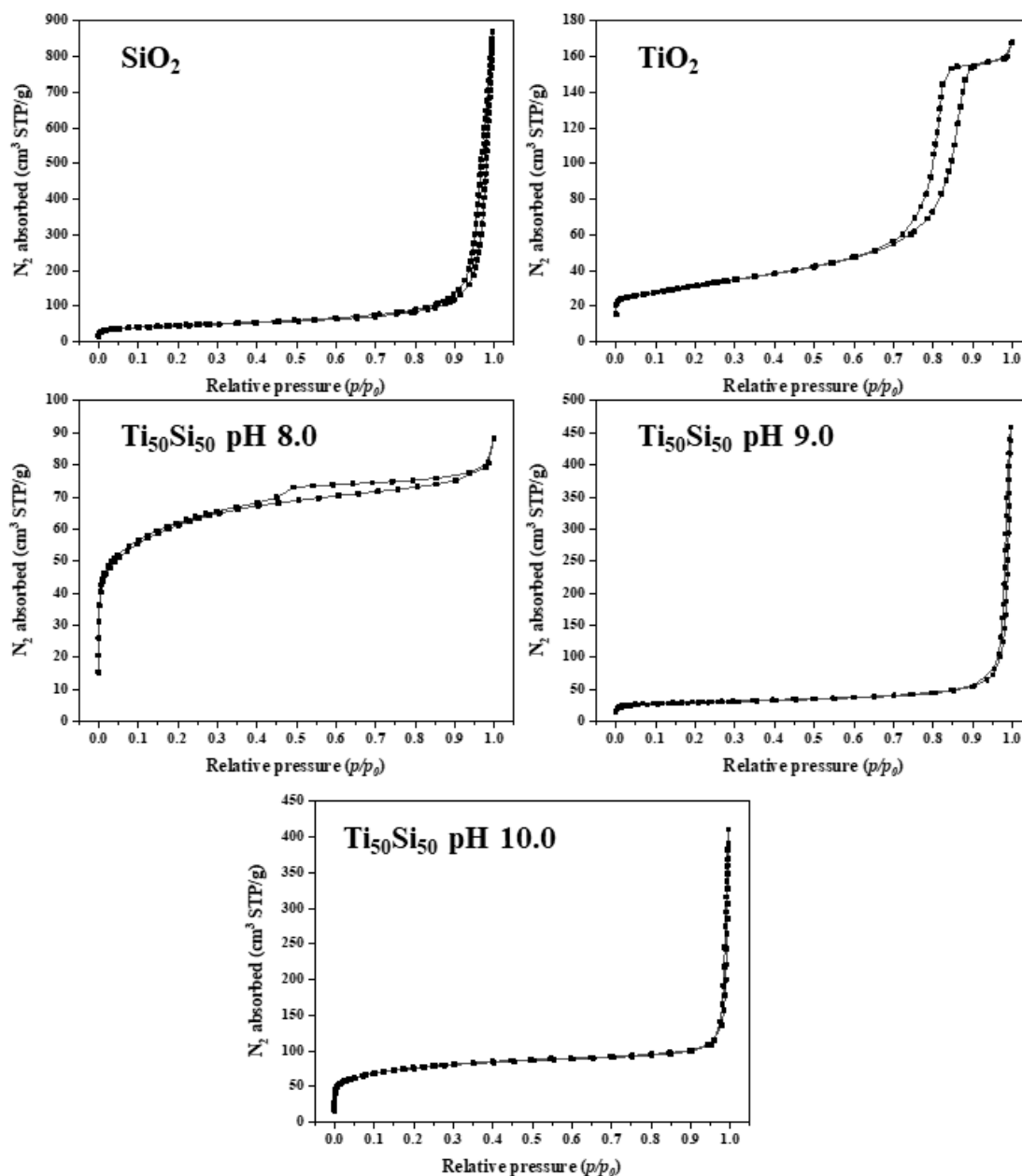


Figure 13. Nitrogen adsorption–desorption isotherms of SiO_2 , TiO_2 , and $\text{Ti}_{50}\text{Si}_{50}$ oxide synthesized at various pH values.

3.2. Effect of Ti/Si ratio

FTIR spectra of TiO_2 , SiO_2 , and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios are shown in Figure 14. All Ti_xSi_y oxides have three major peaks in their FTIR spectra, with wavenumbers of 801, 949, and 1052 cm^{-1} . The peaks at 801 and 1052 cm^{-1} corresponded to symmetric and asymmetric stretching vibrations of Si-O-Si of silica, while the peak at 949 cm^{-1} was commonly attributed to Ti-O-Si vibrations, indicating the existence of Ti-O-Si linkages [34]. This suggested that titanium had been incorporated into the silica framework. Therefore, there was no effect of the Ti/Si atomic ratio on the FTIR of Ti_xSi_y oxide.

The FE-SEM image of TiO_2 exhibits a large majority of particles with high agglomeration, as shown in Figure 15a. $\text{Ti}_{70}\text{Si}_{30}$ and $\text{Ti}_{50}\text{Si}_{50}$ oxides show particles that are homogeneous with a good nanostructure and mostly spherical in shape when Ti/Si is more than 1 (Figure 15b,c). However, the size of $\text{Ti}_{40}\text{Si}_{60}$ oxide particles becomes gradually aggregated with increasing moles of Si (Figure 15d).

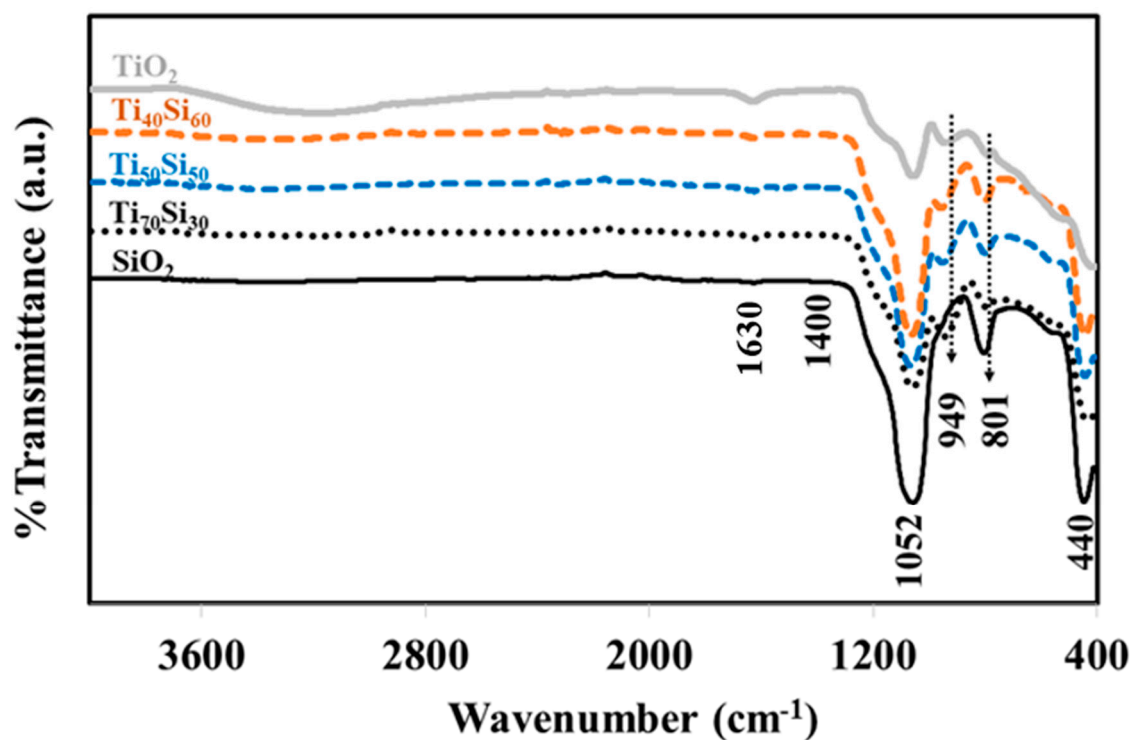


Figure 14. FTIR spectra of TiO_2 , SiO_2 , and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

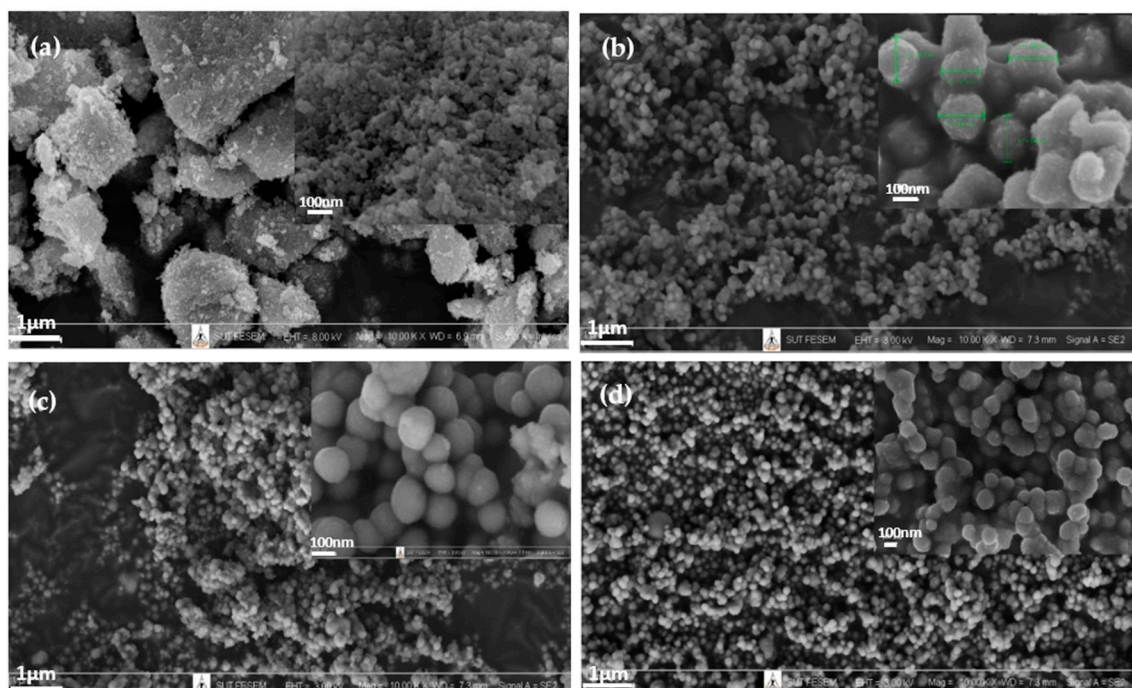


Figure 15. Low ($\times 10k$ WD = 7.3 mm EHT = 3.00 kV) and high ($\times 100k$ WD = 6.9–7.3 mm EHT = 3.00 kV) magnification SEM images of TiO_2 and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios: (a) TiO_2 , (b) $Ti_{70}Si_{30}$, (c) $Ti_{50}Si_{50}$, (d) $Ti_{40}Si_{60}$.

From the XRD patterns, only broad-range signals due to the uniform dispersion of titanium atoms onto the SiO_2 network can be seen in the XRD patterns of $Ti_{70}Si_{30}$, $Ti_{50}Si_{50}$, and $Ti_{40}Si_{60}$ oxides (Figure 16). As a result, the absence of XRD diffraction signals suggests that anatase crystallization is still hindered by the titania aggregate's segregation from the silica network [55].

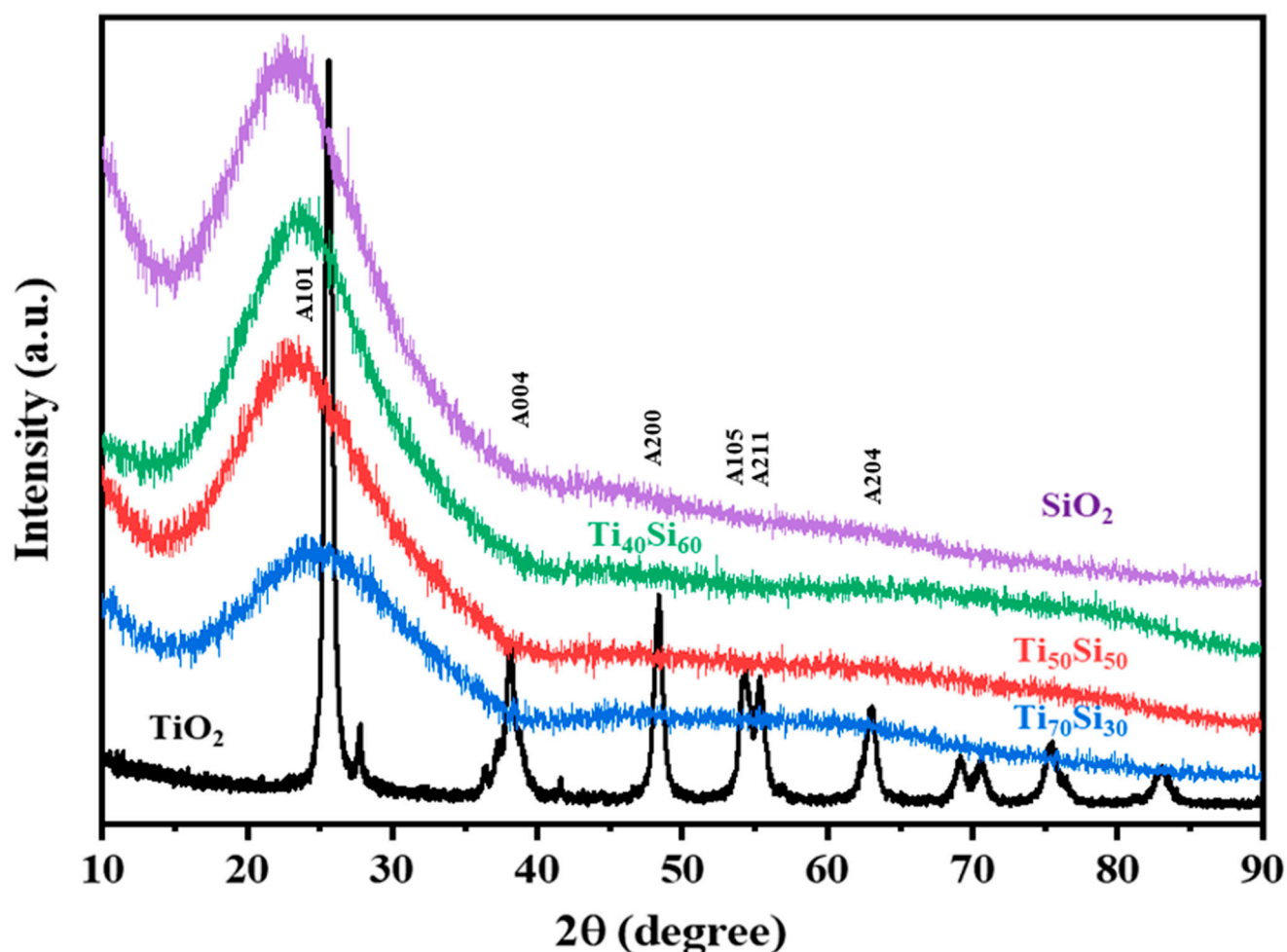


Figure 16. XRD patterns of SiO_2 , TiO_2 , and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

Ti K-edge XANES spectra of standard TiO_2 anatase, synthesized TiO_2 materials, and Ti_xSi_y oxides synthesized with different Ti/Si ratios were collected and are presented in Figure 17. Ti is expected to readily replace Si in the tetrahedral SiO_2 framework at low Ti concentrations in $Ti_{40}Si_{60}$. The features and position in the rising edge of the B shoulders are clearer and the positions are lower. This suggests that the Ti–O–Ti linkages in Ti_xSi_y oxides increase and are incomplete in anatase [21].

The structural parameters of Ti_xSi_y oxide synthesized at pH 9.0 with varying Ti/Si ratios were attained without phase corrections by fitting the data in R-space within the intermission of 1.1–3 Å (Figure 18). The shell parameters, such as Ti–O and Ti–Si bond distances (R), coordination number (CN), Debye–Waller factor (σ^2), and R-factor, are presented in Table 3. The amplitude reduction factor (S_0^2) was estimated to be 0.52 based on the EXAFS spectra of SiO_2 . According to the EXAFS curve fitting findings, the interatomic

distances for Ti atom dilute silica of $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxide are 1.90, 1.83, and 1.82, respectively. The average Ti–O bond distance dropped from 1.90 to 1.82 when Ti in Ti_xSi_y oxide increased. Furthermore, the 1.83 and 1.82 Ti–O bond lengths found for $\text{Ti}_{50}\text{Si}_{50}$ and $\text{Ti}_{40}\text{Si}_{60}$ oxide correspond with the absolute value of the Ti atoms occupied in a tetrahedral environment, validating the comparable assumption from XANES.

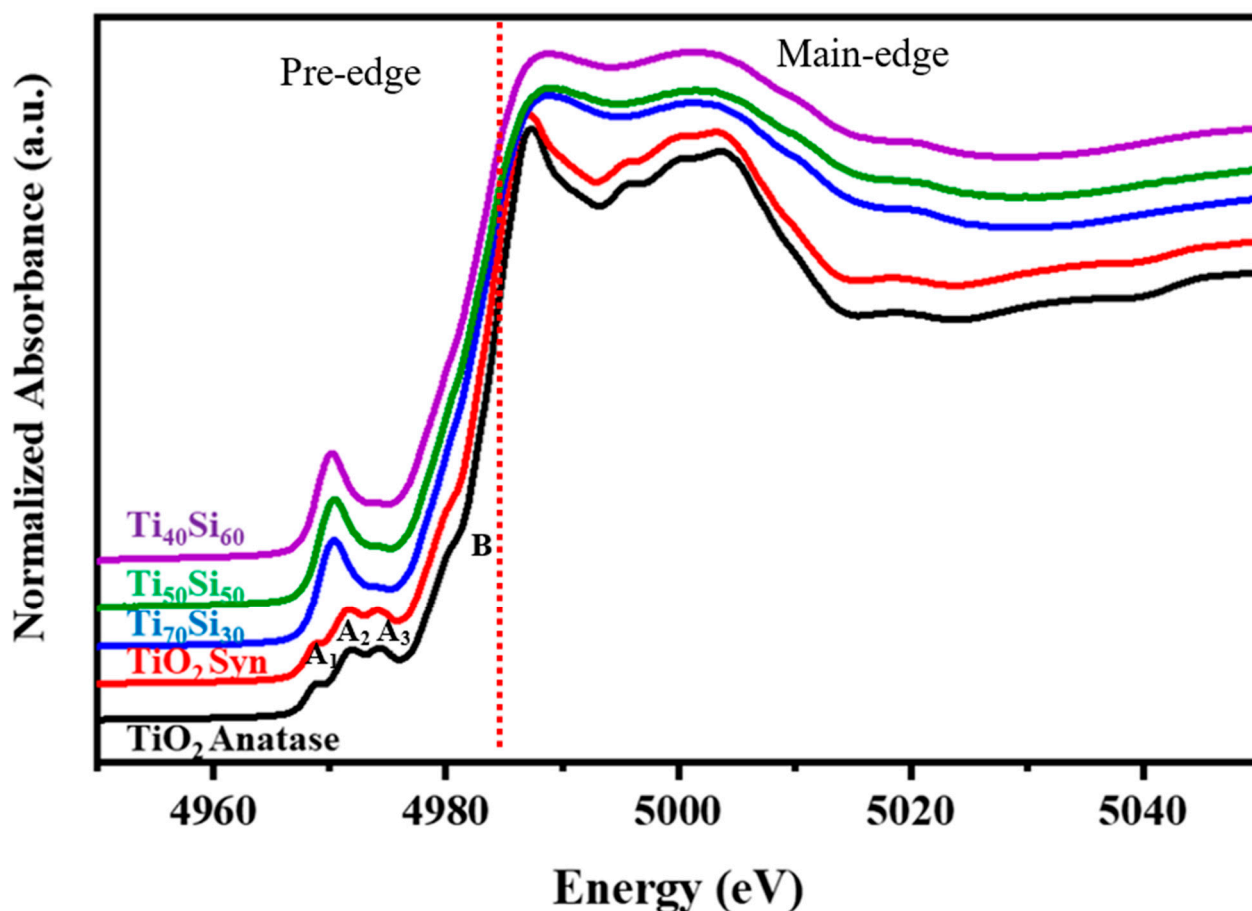


Figure 17. Ti K-edge XANES spectra of TiO_2 and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

The proportion of Si atoms in $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxide was verified by XRF analysis to be 29.79, 47.87, and 60.19%, respectively. Ti content was 70.21, 52.13, and 39.81% in $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxide, respectively (Table 4).

Table 5 and Figure 19 display the particle sizes of TiO_2 , SiO_2 , and Ti_xSi_y oxides synthesized at various Ti/Si ratios, as measured by the Zetasizer Nano ZS and SEM micrographs. The mean particle size distribution of $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxide was 573.1, 136.1, and 825.5 nm, respectively. In addition, SEM micrographs (Figure 15) were used to measure the correct diameter of TiO_2 , SiO_2 , and Ti_xSi_y oxide by ImageJ. The number average diameter, d_n , was estimated using Equations (1) and (2) from the minimum value given in [52]. $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxide had an average diameter of 149.3, 135.4, and 131.8 nm, respectively, as reported in Table 6. As the amount of Ti incorporated into the silica network increased, the number average diameter decreased.

The $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxides had surface areas (S_{BET}) of 569.07, 225.68, and 68.34 m^2g^{-1} , respectively. The S_{BET} displayed an upward trend as the Ti content increased. According to a prior study, the S_{BET} of the Ti_xSi_y oxide was enhanced as the Ti concentration in the silica network increased [56].

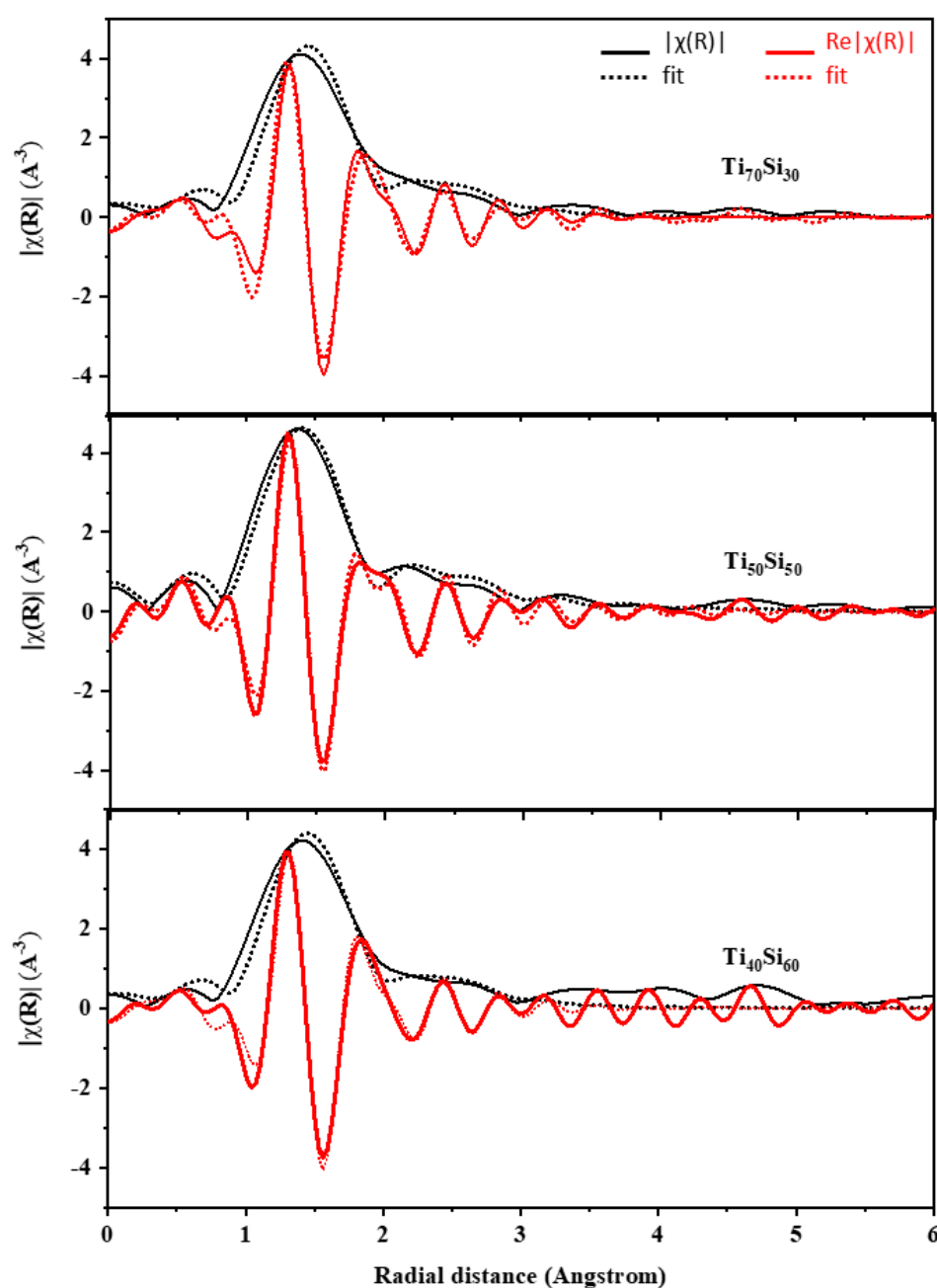


Figure 18. The experimental EXAFS data of Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios in R-space within the intermission of 1.1–3 Å.

A typical nitrogen adsorption–desorption isotherm of the SiO_2 , TiO_2 , and $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxides is presented in Figure 20. Type II isotherms with pore diameters of 10.96, 5.85, and 21.97 nm were found in $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxides, corresponding to mesoporous materials with pore sizes ranging from 2 to 50 nm. However, the TiO_2 isotherm is classified as a type IV curve according to the IUPAC. This curve shows an abrupt hysteric loop, which is common in mesoporous materials with pores between 2 and 50 nm that exhibit capillary condensation and evaporation [53,54].

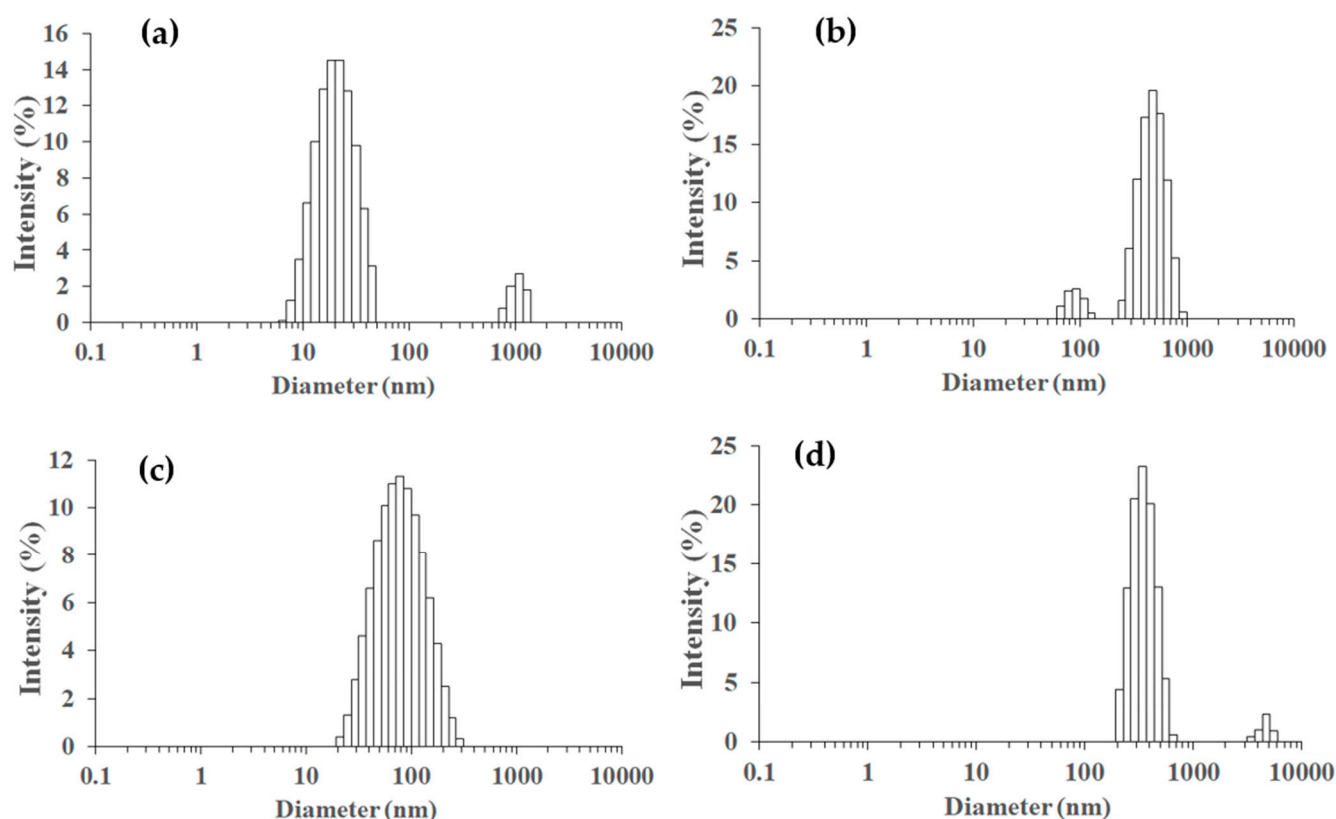
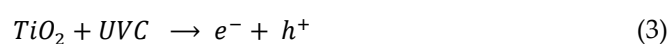


Figure 19. Particle size distributions measured on a Zetasizer Nano ZS of TiO_2 and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios: (a) TiO_2 , (b) $\text{Ti}_{70}\text{Si}_{30}$, (c) $\text{Ti}_{50}\text{Si}_{50}$, (d) $\text{Ti}_{40}\text{Si}_{60}$.

3.3. Photocatalytic degradation of methylene blue (MB)

The degradation of methylene blue (MB) was used to study the photocatalytic activity of PLA and PLA/ Ti_xSi_y oxide composite films. In addition, UV irradiation may also lead to the decomposition of MB without the presence of any photocatalyst. The degradation of MB on PLA and PLA/ Ti_xSi_y oxide composite films was caused by a change in MB concentration in aqueous solution under UV irradiation, as illustrated in Figures 21 and 22. The addition of 3%wt TiO_2 , SiO_2 , and Ti_xSi_y oxide to the PLA film matrix increased the effectiveness of photocatalysis in the degradation of MB. TiO_2 is more effective at degrading MB than $\text{Ti}_{50}\text{Si}_{50}$ pH 9.0, $\text{Ti}_{50}\text{Si}_{50}$ pH 10.0, $\text{Ti}_{50}\text{Si}_{50}$ pH 8.0, and SiO_2 . In the degradation of MB, the efficiency of the materials was of the following order: $\text{TiO}_2 > \text{Ti}_{70}\text{Si}_{30} > \text{Ti}_{50}\text{Si}_{50} > \text{Ti}_{40}\text{Si}_{60} > \text{SiO}_2$.

Photocatalytic activity is generally accepted to occur at the surface of a photocatalyst. As a result, the surface area of PLA nanocomposite film, which is dependent on the size of nanoparticles, film morphology, and thickness, influences photocatalytic reactivity [7]. The PLA nanocomposite film including TiO_2 degrades MB more efficiently than employing only photocatalysis because of the two mechanisms of degradation. The first is that in which UVC directly degrades MB, while the second occurs when TiO_2 receives light energy greater than the bandgap energy, and electrons in the valence band (VB) are excited into the conduction band (CB), resulting in the generation of a hole (h^+) (Equation (3)). This hole can oxidize MB (Equation (4)) or oxidize H_2O to produce OH^\cdot (Equation (5)). The e^- in CB can reduce O_2 at the surface of TiO_2 to generate O_2^- (Equation (6)). Radicals (OH^\cdot , O_2^-) and h^+ interact with MB to create peroxide derivatives and hydroxylate or completely degrade to CO_2 and H_2O [7]. The photodegradation pathway can be summarized by Equations (3)–(6).



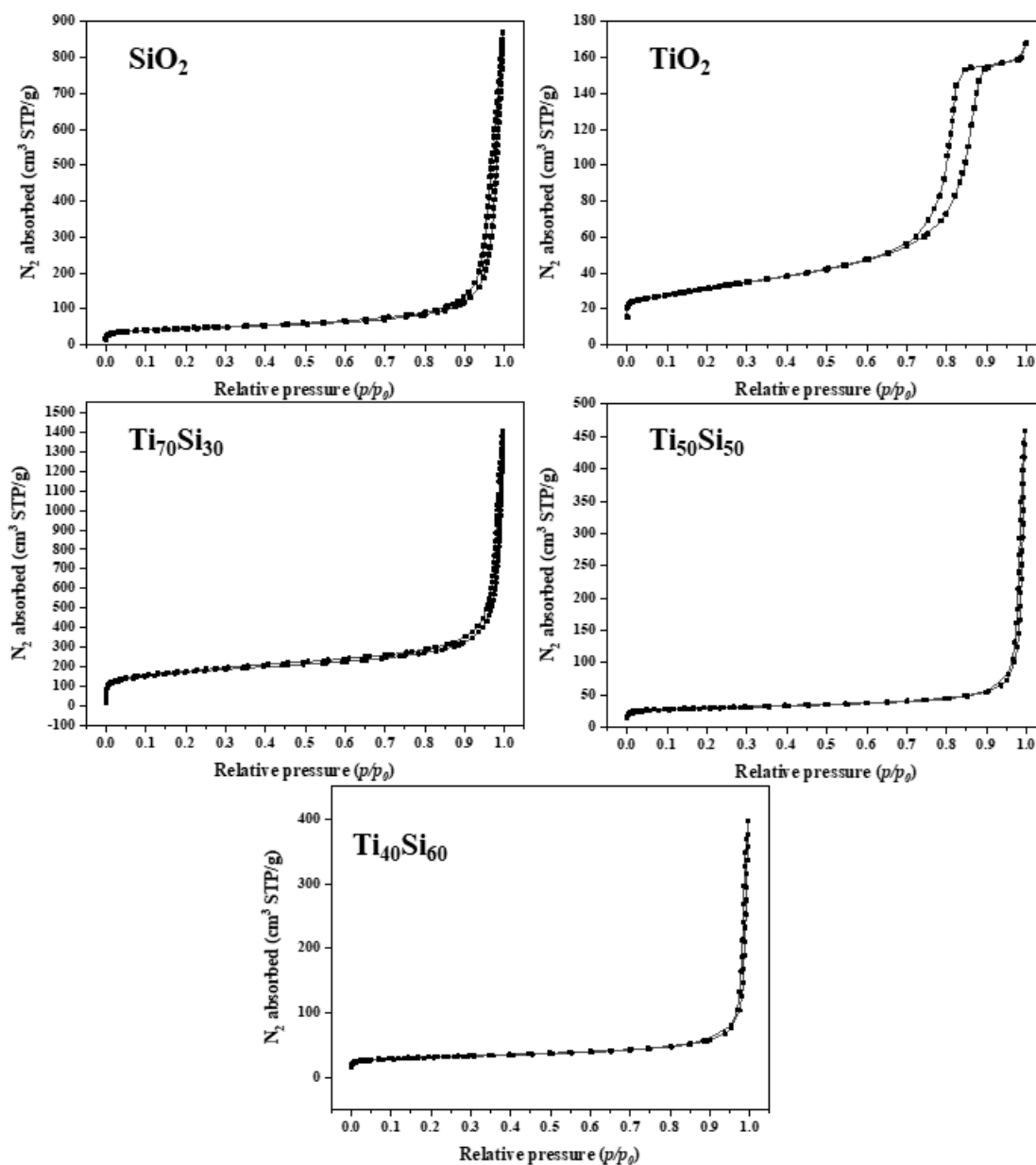


Figure 20. Nitrogen adsorption–desorption isotherms of SiO₂, TiO₂, and Ti_xSi_y oxide synthesized at pH 9.0 with different Ti/Si ratios.

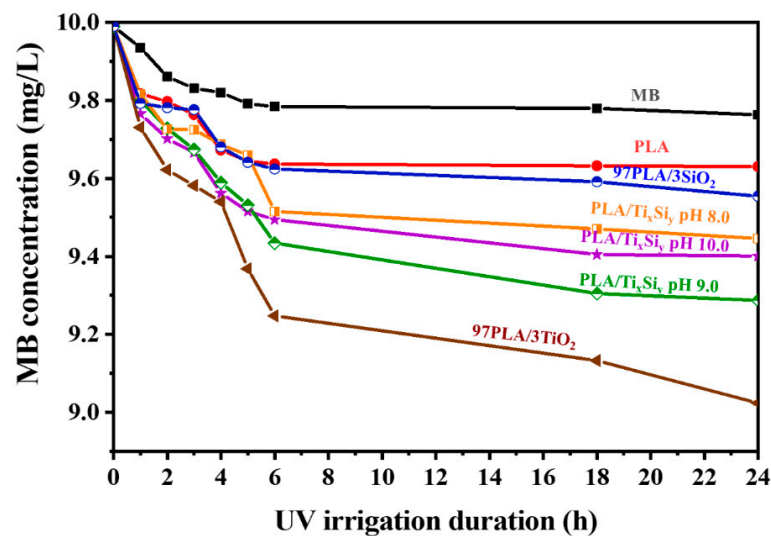


Figure 21. Concentration of methylene blue (MB) due to absorption of PLA film and PLA/ Ti₅₀Si₅₀ oxide (synthesized at pH 8.0, 9.0, and 10.0) films under UV irradiation.

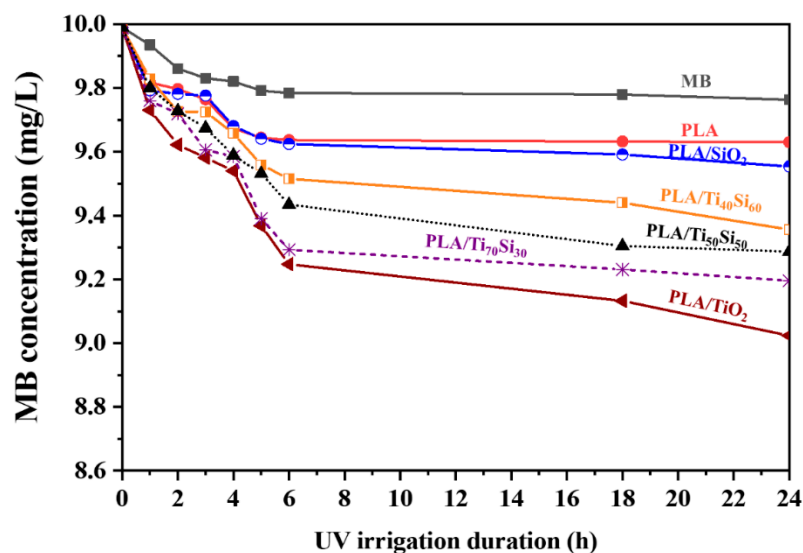


Figure 22. Concentration of methylene blue (MB) due to absorption of PLA film and PLA/ Ti_xSi_y oxide (synthesized at pH 9.0 with different Ti/Si ratios) films under UV irradiation.

4. Conclusions

Titania–silica oxides (Ti_xSi_y oxides) can be prepared by the sol–gel technique. The effect of pH and the Ti/Si atomic ratio of titanium–silicon binary oxide (Ti_xSi_y) on the structural characteristics of Ti_xSi_y oxide were investigated by using the sol–gel technique. ²⁹Si solid-state NMR and FTIR measurements indicated that certain Si atoms in the SiO₂ network had been replaced by Ti atoms, implying the formation of Si–O–Ti connections. As the pH was elevated to alkaline conditions (pH 9.0 and 10.0), the nanoparticles of Ti₅₀Si₅₀ oxide acquired a more spherical shape and their size distribution became more uniform, resulting in an acceptable nanostructure. Agglomeration was reduced in Ti₅₀Si₅₀ oxide nanoparticles, which were mostly spherical in form. However, the Ti₅₀Si₅₀ oxide particles at pH 10.0 become nano-sized and agglomerated. In Ti₅₀Si₅₀ oxide samples, the existence of a large pre-edge feature indicated that a higher percentage of Ti atoms occupied tetrahedral symmetry positions, as expected in samples where Ti directly substitutes Si. The result of 1.83 Å for the Ti–O bond distance in Ti_xSi_y oxides generated at pH 9.0 accords with

the fraction of Ti atoms occupied in a tetrahedral environment. The average diameter of the sample increased with the increasing pH of the sol. In addition, $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxide had an average diameter of 149.3, 135.4, and 131.8 nm, respectively. As the amount of Ti incorporated into the silica network increased, the average diameter decreased. The S_{BET} displayed an upward trend as the Ti content increased. Type II isotherms with pore diameters of 10.96, 5.85, and 21.97 nm were found in $\text{Ti}_{70}\text{Si}_{30}$, $\text{Ti}_{50}\text{Si}_{50}$, and $\text{Ti}_{40}\text{Si}_{60}$ oxides, corresponding to mesoporous materials with pore sizes ranging from 2 to 50 nm, which is common in mesoporous materials. However, the TiO_2 isotherm was classified as a type IV curve. The addition of 3%wt TiO_2 , SiO_2 , and Ti_xSi_y oxide to the PLA film matrix increased the effectiveness of photocatalysis in the degradation of MB. TiO_2 was more effective at degrading MB than $\text{Ti}_{50}\text{Si}_{50}$ pH 9.0, $\text{Ti}_{50}\text{Si}_{50}$ pH 10.0, $\text{Ti}_{50}\text{Si}_{50}$ pH 8.0, and SiO_2 . In the degradation of MB, the order of efficiency is $\text{TiO}_2 > \text{Ti}_{70}\text{Si}_{30} > \text{Ti}_{50}\text{Si}_{50} > \text{Ti}_{40}\text{Si}_{60} > \text{SiO}_2$. Moreover, PLA/ $\text{Ti}_{70}\text{Si}_{30}$ also increased the effectiveness of the photocatalytic activity of PLA. Finally, PLA/ $\text{Ti}_{70}\text{Si}_{30}$ improved the efficiency of the photocatalytic activity of PLA.

Author Contributions: Conceptualization, Y.R.; methodology, A.T., C.R. P.A., and Y.R.; validation, A.T., C.R., P.A., and Y.R.; formal analysis, A.T. and P.A.; investigation, A.T., C.R., P.A., and Y.R.; resources, A.T., C.R., and Y.R.; data curation, A.T.; writing—original draft preparation, A.T.; writing—review and editing, A.T., C.R. P.A., and Y.R.; visualization, C.R. and Y.R.; supervision, C.R. and Y.R.; project administration, C.R. and Y.R.; funding acquisition, C.R. and Y.R. All authors have read and agreed to the published version of the manuscript.

Funding: This research was funded by Thailand Science Research and Innovation (TSRI), grant number 42853.

Institutional Review Board Statement: Not applicable

Informed Consent Statement: Not applicable

Data Availability Statement: Not applicable

Acknowledgments: The authors are grateful to Suranaree University of Technology (SUT); to the Center of Excellence on Petrochemical and Materials Technology (PETROMAT); to the Science, Research and Innovation Promotion Fund from Thailand Science Research and Innovation (TSRI); and to the Research Center for Biocomposite Materials for Medical Industry and Agricultural and Food Industry for the financial support. In addition, we would like to express our gratitude to the Synchrotron Light Research Institute (Public Organization) for the XAS experiment.

Conflicts of Interest: The authors declare no conflict of interest.

References

1. Rasalingam, S.; Peng, R.; Koodali, R.T. Review Article: Removal of Hazardous Pollutants from Wastewaters: Applications of TiO_2 - SiO_2 Mixed Oxide Materials. *Journal of Nanomaterials* **2014**, 1–42. <http://dx.doi.org/10.1155/2014/617405>
2. Wang, X.; Xue, J.; Wang, X.; Liu, X. Heterogeneous Ag- TiO_2 - SiO_2 composite materials as novel catalytic systems for selective epoxidation of cyclohexene by H_2O_2 . *PLOS ONE* **2017**, 12 (5), e0176332. <https://doi.org/10.1371/journal.pone.0176332>
3. Sahu, D.R.; Hong, L.Y.; Wang, S.-C.; Huang, J.-L. Synthesis, analysis and characterization of ordered mesoporous TiO_2 /SBA-15 matrix: Effect of calcination temperature. *Microporous and Mesoporous Materials* **2009**, 117, 640–649. <http://dx.doi.org/10.1016/j.micromeso.2008.08.013>
4. Petronella, F.; Pagliarulo, A.; Truppi, A.; Lettieri, M.; Masieri, M.; Calia, A.; Curri, M.L.; Comparelli, R. TiO_2 Nanocrystal Based Coatings for the Protection of Architectural Stone: The Effect of Solvents in the Spray-Coating Application for a Self-Cleaning Surfaces. *Coatings* **2018**, 8(10), 356. <https://doi.org/10.3390/coatings8100356>
5. Zhu, Y.; Buonocore, G.G.; Lavorgna, M. Photocatalytic activity of PLA/ TiO_2 nanocomposites and TiO_2 -active multilayered hybrid coatings. *Italian Journal of Food Science* **2012**, 24, 102–106.
6. Tiainen, H. D.; Wiedmer, H.; Haugen, J. Processing of highly porous TiO_2 bone scaffolds with improved compressive strength. *Journal of the European Ceramic Society* **2013**, 33, 15–24.
7. Chen, L.; Zheng, K.; Liu, Y. Geopolymer-supported photocatalytic TiO_2 film: Preparation and characterization. *Construction and Building Materials* **2017**, 151, 63–70.
8. Allodi, V.; Brutti, S.; Giarola, M.; Sgambetterra, M.; Navarra, M.A.; Panero, S.; Mariotto, G. Structural and Spectroscopic Characterization of A Nanosized Sulfated TiO_2 Filler and of Nanocomposite Nafion Membranes. *Polymers* **2016**, 8.
9. Bertuna, A.; Comini, E.; Poli, N.; Zappa, D.; Sberveglieri, G. Titanium Dioxide Nanostructures Chemical Sensor, *Procedia Engineering* **2016**, 168, 313–316.

10. Zapata, P.A.; Palza, H.; Cruz, L.S.; Lieberwirth, I.; Catalina, F.; Corrales, T.; Rabagliati, F.M. Polyethylene and poly(ethylene-co-1-octadecene) composites with TiO₂ based nanoparticles by metallocenic “in situ” polymerization. *Polymer* **2013**, *54*, 2690–2698.
11. Fonseca, C.; Ochoa, A.; Ulloa, M.T.; Alvarez, E.; Canales, D.; Zapata, P.A. Poly(lactic acid)/TiO₂ nanocomposites as alternative biocidal and antifungal materials. *Materials Science and Engineering: C* **2015**, *57*, 314–320.
12. Wu, F.; Lan, X.; Ji, D.; Liu, Z.; Yang, W.; Yang, M. Grafting polymerization of polylactic acid on the surface of nano-SiO₂ and properties of PLA/PLA-grafted-SiO₂ nanocomposites. *Journal of Applied Polymer Science* **2013**, *129*, 3019–3027.
13. Serenko, O.A.; Muzafarov, A.M. Polymer composites with surface modified SiO₂ nanoparticles: Structures, properties, and promising applications. *Polymer Science Series C* **2016**, *58*, 93–101.
14. S.-W. Ha, M.N. Weitzmann, G.R. Beck Jr, Chapter 4 - Dental and Skeletal Applications of Silica-Based Nanomaterials A2 - Subramani, Karthikeyan, in: W. Ahmed, J.K. Hartsfield (Eds.) *Nanobiomaterials in Clinical Dentistry*, William Andrew Publishing 2013, pp. 69–91.
15. Gao, X.; Wachs, I.E. Titania–silica as catalysts: molecular structural characteristics and physico-chemical properties, *Catalysis Today* **1999**, *51*, 233–254.
16. Pabón, E.; Retuert, J.; Quijada, R.; Zarate, A. TiO₂–SiO₂ mixed oxides prepared by a combined sol–gel and polymer inclusion method. *Microporous and Mesoporous Materials* **2004**, *67*, 195–203.
17. Pabón, E.; Retuert, J.; Quijada, R. Synthesis of mixed silica–titania by the sol–gel method using polyethylenimine: porosity and catalytic properties. *Journal of Porous Materials* **2007**, *14*, 151–158.
18. Milne, C.J.; Penfold, T.J.; Chergui, M. Recent experimental and theoretical developments in time-resolved X-ray spectroscopies. *Coordination Chemistry Reviews* **2014**, *277–278*, 44–68.
19. Niltharach, A.; Kityakarn, S.; Worayingyong, A.; Thienprasert, J.T.; Klysubun, W.; Songsiriritthigul, P.; Limpijumnong, S. Structural characterizations of sol–gel synthesized TiO₂ and Ce/TiO₂ nanostructures. *Physica B: Condensed Matter* **2012**, *407*, 2915–2918.
20. Kim, W.B.; Choi, S.H.; Lee, J.S. Quantitative Analysis of Ti–O–Si and Ti–O–Ti Bonds in Ti–Si Binary Oxides by the Linear Combination of XANES. *The Journal of Physical Chemistry B* **2000**, *104*, 8670–8678.
21. Matsuo, S.; Sakaguchi, N.; Wakita, H. Pre-edge features of Ti K-edge X-ray absorption near-edge structure for the local structure of sol-gel titanium oxides, *Anal Sci*, *21* (2005) 805–809.
22. Alias S.S.; Mohamad, A.A. ZnO: Effect of pH on the Sol–Gel Process., Synthesis of Zinc Oxide by Sol–Gel Method for Photoelectrochemical Cells., Springer, Singapore., SpringerBriefs in Materials., 2014, pp. 9–25.
23. Kim, T.G.; An, G.S.; Han, J.S.; Hur, J.U.; Park, B.G.; Choi S.-C., Synthesis of Size Controlled Spherical Silica Nanoparticles via Sol-Gel Process within Hydrophilic Solvent. *J. Korean Ceram. Soc* **2017**, *54*, 49–54.
24. Karkare, M.M. Choice of precursor not affecting the size of anatase TiO₂ nanoparticles but affecting morphology under broader view. *International Nano Letters* **2014**, *4*, 111.
25. Klysubun, W.; Tarawarakarn, P.; Thamsanong, N.; Amonpattaratkit, P.; Cholsuk, C.; Lapboonrueng, S.; Chaichuay, S.; Wongtepa, W. Upgrade of SLRI BL8 beamline for XAFS spectroscopy in a photon energy range of 1–13 keV, *Radiation Physics and Chemistry* **2020**, *175*, 108145.
26. Ravel, B. N. M.; Newville, M. Data analysis for X-ray absorption spectroscopy using IFEFFIT, 2005.
27. Schraml-Marth, M.; Walther, K.L.; Wokaun, A.; Handy, B.E.; Baiker, A. Porous silica gels and TiO₂/SiO₂ mixed oxides prepared via the sol-gel process: characterization by spectroscopic techniques. *Journal of Non-Crystalline Solids* **1992**, *143*, 93–111.
28. Degirmenci, V.; Erdem, Ö.F.; Ergun, O.; Yilmaz, A.; Michel, D.; Uner, D. Synthesis and NMR Characterization of Titanium and Zirconium Oxides Incorporated in SBA-15. *Topics in Catalysis* **2008**, *49*, 204–208.
29. Dirken, P.J.; Smith, M.E.; Whitfield, H.J. 17O and 29Si Solid State NMR Study of Atomic Scale Structure in Sol-Gel-Prepared TiO₂-SiO₂ Materials, *The Journal of Physical Chemistry* **1995**, *99*, 395–401.
30. Zeitler, V.A.; Brown, C.A. The Infrared Spectra of Some Ti–O–Si, Ti–O–Ti and Si–O–Si Compounds. *The Journal of Physical Chemistry* **1957**, *61*, 1174–1177.
31. Praveen, P.; Structural, Functional and Optical Characters of TiO₂ Nanocrystallites: Anatase and Rutile Phases, *St. Joseph’s Journal of Humanities and Science* **2019**, *6*, 43–53.
32. Chen, Y.-F.; Lee, C.-Y.; Yeng, M.-Y.; Chiu, H.-T. The effect of calcination temperature on the crystallinity of TiO₂ nanopowders. *Journal of Crystal Growth* **2003**, *247*, 363–370.
33. Vasquez, J.; Lozano, H.; Lavayen, V.; Lira-Cantú, M.; Gómez-Romero, P.; Santa Ana, M.A.; Benavente, E.; Gonzalez, G. High-yield preparation of titanium dioxide nanostructures by hydrothermal conditions. *J Nanosci Nanotechnol* **2009**, *9*, 1103–1107.
34. Yamashita, H.; Kawasaki, S.; Ichihashi, Y.; Harada, M.; Takeuchi, M.; Anpo, M.; Stewart, G.; Fox, M.A.; Louis, C.; Che, M. Characterization of Titanium–Silicon Binary Oxide Catalysts Prepared by the Sol–Gel Method and Their Photocatalytic Reactivity for the Liquid-Phase Oxidation of 1-Octanol. *The Journal of Physical Chemistry B* **1998**, *102*, 5870–5875.
35. Vives, S.; Meunier, C. Influence of the synthesis route on sol–gel SiO₂–TiO₂ (1:1) xerogels and powders, *Ceramics International* **2008**, *34*, 37–44.
36. Yu, H.-F.; Wang S.-M., Effects of water content and pH on gel-derived TiO₂–SiO₂. *Journal of Non-Crystalline Solids* **2000**, *261*, 260–267.
37. Yoldas, B.E. Formation of titania-silica glasses by low temperature chemical polymerization. *Journal of Non-Crystalline Solids* **1980**, *38*, 81–86.
38. Alias, S.S.; Ismail, A.B.; Mohamad, A.A. Effect of pH on ZnO nanoparticle properties synthesized by sol–gel centrifugation, *Journal of Alloys and Compounds* **2010**, *499*, 231–237.

39. Wu, L.; Wu, Y.; Lü Y. Self-assembly of small ZnO nanoparticles toward flake-like single crystals, *Materials Research Bulletin* **2006**, *41*, 128-133.
40. Etacheri, V.; Di Valentin, C.; Schneider, J.; Bahnemann, D.; Pillai, S.C. Visible-light activation of TiO₂ photocatalysts: Advances in theory and experiments. *Journal of Photochemistry and Photobiology C: Photochemistry Reviews* **2015**, *25*, 1-29.
41. Li Hsiung, T.; Paul Wang, H.; Wang, H.C. XANES studies of photocatalytic active species in nano TiO₂-SiO₂. *Radiation Physics and Chemistry* **2006**, *75*, 2042-2045.
42. Rossi, T.C.; Grolimund, D.; Nachtegaal, M.; Cannelli, O.; Mancini, G.F.; Bacellar, C.; Kinschel, D.; Rouxel, J.R.; Ohannessian, N.; Pergolesi, D.; Lippert, T.; Chergui, M. X-ray absorption linear dichroism at the Ti K-edge of anatase TiO₂ single crystals. *Physical Review B* **2019**, *100*, 245207.
43. Farges, F.; Brown, G.E.; Rehr, J.J. Ti K-edge XANES studies of Ti coordination and disorder in oxide compounds: Comparison between theory and experiment. *Physical Review B* **1997**, *56*, 1809-1819.
44. Davis, R.J.; Liu, Z. Titania-Silica: A Model Binary Oxide Catalyst System, *Chemistry of Materials* **1987**, *9*, 2311-2324.
45. Sharma, A.; Varshney, M.; Park, J.; Ha, T.-K.; Chae, K.-H.; Shin, H.-J. XANES, EXAFS and photocatalytic investigations on copper oxide nanoparticles and nanocomposites, *RSC Advances* **2015**, *5*, 21762-21771.
46. Khemthong, P.; Photai, P.; Grisdanurak, N. Structural properties of CuO/TiO₂ nanorod in relation to their catalytic activity for simultaneous hydrogen production under solar light. *International Journal of Hydrogen Energy* **2013**, *38*, 15992-16001.
47. Loshmanov, A.A.; Sigaev, V.N.; Khodakovskaya, R.Y.; Pavlushkin, N.M.; Yamzin, I.I. Small-angle neutron scattering on silica glasses containing titania. *Journal of Applied Crystallography* **1974**, *7*, 207-210.
48. Sandstrom, D.R.; Lytle, F.W.; Wei, P.S.P.; Greigor, R.B.; Wong, J.; Schultz, P. Coordination of Ti in TiO₂-SiO₂ glass by X-ray absorption spectroscopy. *Journal of Non-Crystalline Solids* **1980**, *41*, 201-207.
49. Neurock, M.; Manzer, L.E. Theoretical insights on the mechanism of alkene epoxidation by H₂O₂ with titanium silicalite, *Chemical Communications* **1996**, 1133-1134.
50. Scherk, C.G.; Ostermann, A.; Achterhold, K.; Iakovleva, O.; Nazikkol, C.; Krebs, B.; Knapp E.W.; Meyer-Klaucke, W.; Parak, F.G. The X-ray absorption spectroscopy Debye-Waller factors of an iron compound and of met-myoglobin as a function of temperature, *Eur Biophys J* **2001**, *30*, 393-403.
51. Wellenreuther, G.; Parthasarathy, V.; Meyer-Klaucke, W. Towards a black-box for biological EXAFS data analysis. II. Automatic BioXAS Refinement and Analysis (ABRA), *Journal of Synchrotron Radiation* **2010**, *17*, 25-35.
52. Wu, D.; Zhang, Y.; Yuan, L.; Zhang, M.; Zhou, W. Viscoelastic interfacial properties of compatibilized poly(ϵ -caprolactone)/polylactide blend, *Journal of Polymer Science Part B: Polymer Physics* **2010**, *48*, 756-765.
53. Gregg, S. J.; Sing, K.S.W. Adsorption, Surface Area and Porosity Auflage, 2ed., John Wiley & Sons, Ltd, Academic Press, London, 1982.
54. Zhang, X.; Zhang, F.; Chan, K.-Y. Synthesis of titania-silica mixed oxide mesoporous materials, characterization and photocatalytic properties, *Applied Catalysis A: General* **2005**, *284* (2005) 193-198.
55. Van Grieken, R.; Aguado, J.; López-Muñoz, M.J.; Marugán, J. Sol-Gel Titania and Titania-Silica Mixed Oxides Photocatalysts, *Solid State Phenomena* **2010**, *162*, 221-238.
56. Kosuge, K. Titanium-Containing Porous Silica Prepared by a Modified Sol-Gel Method, *The Journal of Physical Chemistry B* **1999**, *103*, 3563-3569.