

Supporting information

Using Sulfobutylated and Sulfomethylated Lignin as Dispersant for Kaolin Suspension

Derya Yesim Hopa,^[a, b] Pedram Fatehi^{*[a,c]}

[a] Department of Chemical Engineering, Lakehead University, Thunder Bay, ON, Canada P7B 5E1.

[b] Department of Chemical Engineering, Afyon Kocatepe University, Afyonkarahisar, 03200, Turkey.

[c] State Key Laboratory of Biobased Material and Green Papermaking, Qilu University of Technology (Shandong Academy of Sciences), Jinan, China, 250353

The mineralogical analysis of a 0.5 g of air-dried kaolin sample was carried out via X-Ray Diffraction (XRD) analysis using Pananalytical Expert Pro Diffractometer (Malvern, UK) with a Cu K α ($\lambda = 1.5405 \text{ \AA}$) radiation source. The sample was scanned between 6°-95° with a step of 0.02° per second. A phase identification software (Match, Crystal Impact, Bonn, Germany) was used to determine the mineralogical composition of kaolin.

Table S1. The results of mineralogical analysis and calculated oxide composition for kaolin.

| Mineral | Weight (%) | Calculated oxides | Weight (%) |
|------------|------------|--------------------------------|------------|
| Kaolinite | 72.22 | Al ₂ O ₃ | 43.70 |
| Dickite | 18.53 | SiO ₂ | 55.17 |
| Quartz | 3.24 | Fe ₂ O ₃ | 0.62 |
| Cordierite | 6.02 | MgO | 0.50 |

Table S2. Charge density values of sulfomethylated lignin with different ratios of formaldehyde to lignin (F/L) and sodium metabisulfite to lignin (S/L).

| F/L = 1:1 | S/L | Charge density (meq/g) |
|-----------|--------|------------------------|
| | 0.5:1 | -2.00 |
| | 0.75:1 | -2.06 |
| | 1.25:1 | -2.15 |
| F/L = 2:1 | 1:1 | -2.31 |

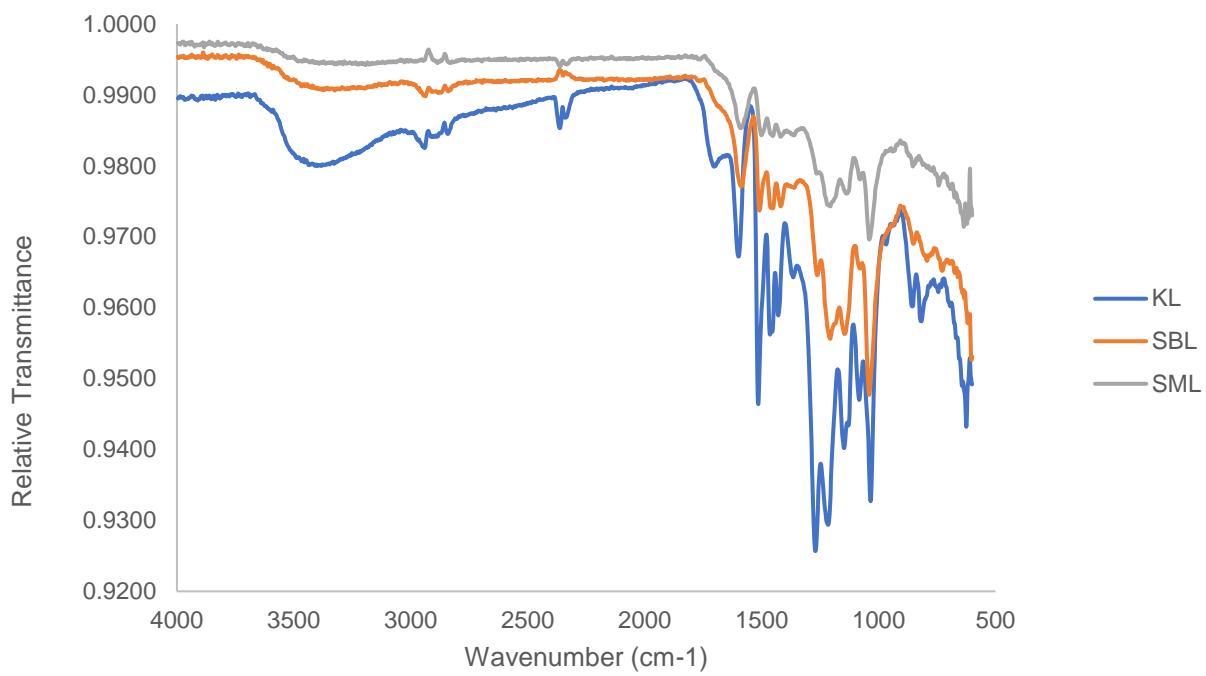


Figure S1. FTIR spectra of modified lignin (SBL and SML) samples compared with KL.

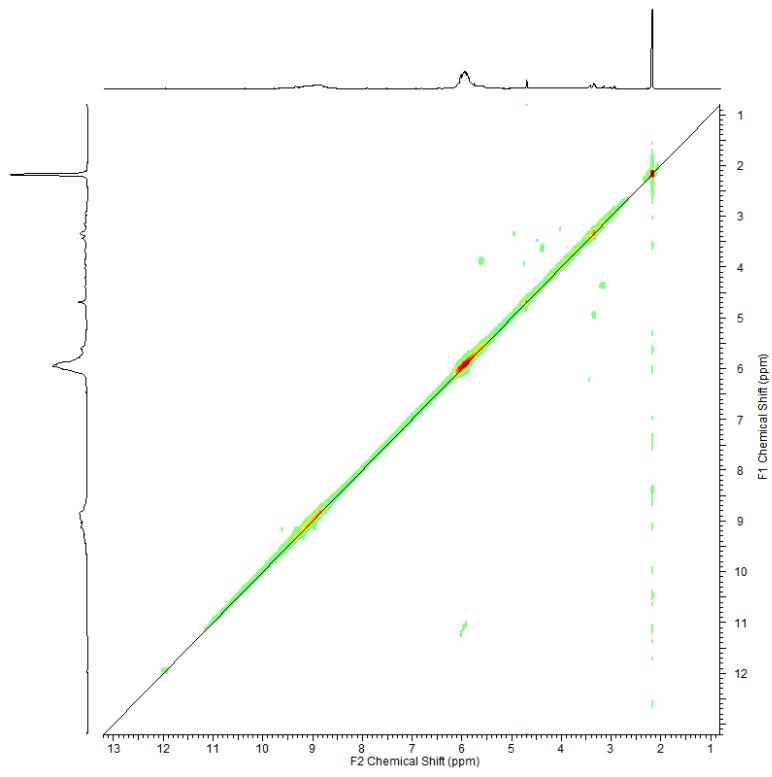


Figure S2. ^1H - ^1H 2D COSY map of KL.

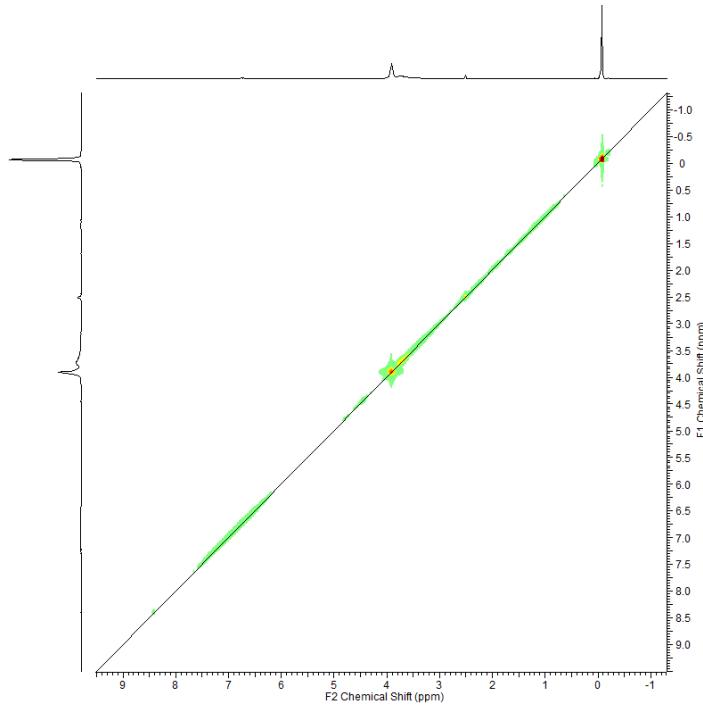


Figure S3. ^1H - ^1H 2D COSY map of SML.

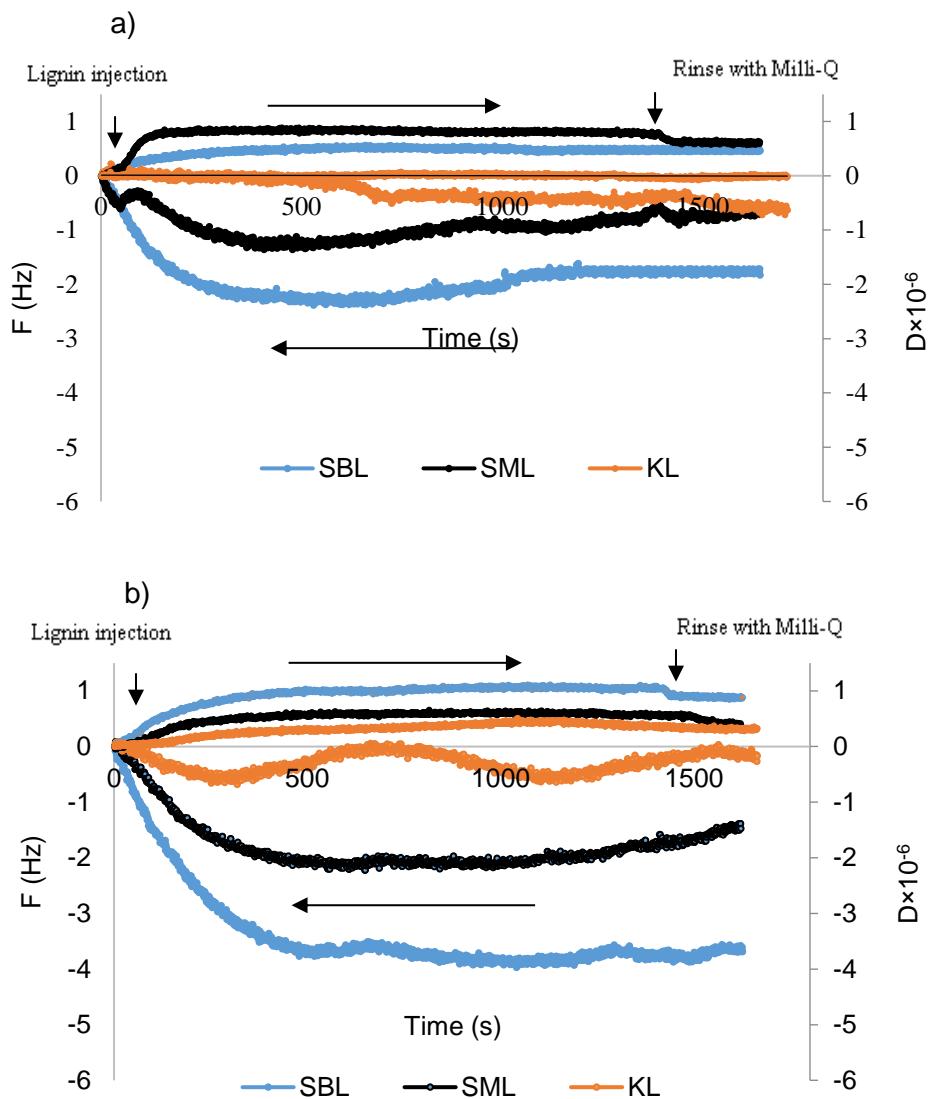


Figure S4. Adsorption of lignin derivatives on Al_2O_3 coated quartz sensors for concentrations **a)** 100 mg/L **b)** 200 mg/L at pH 7.8.

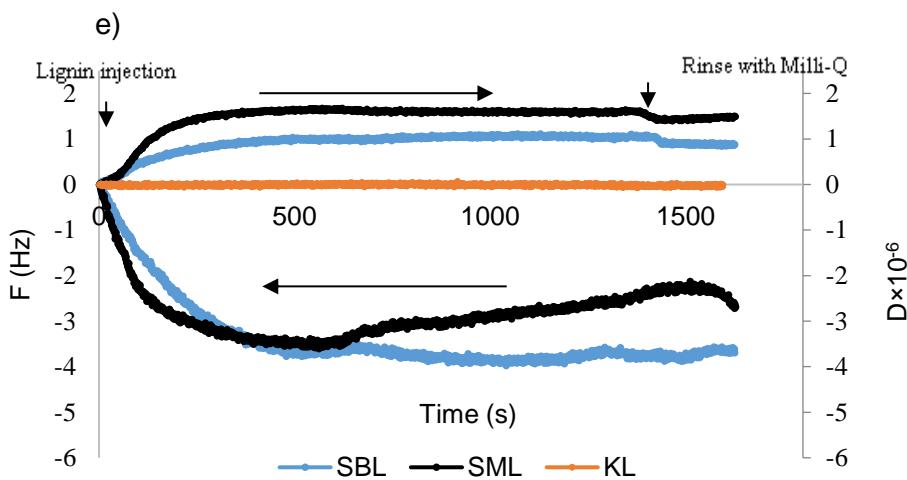
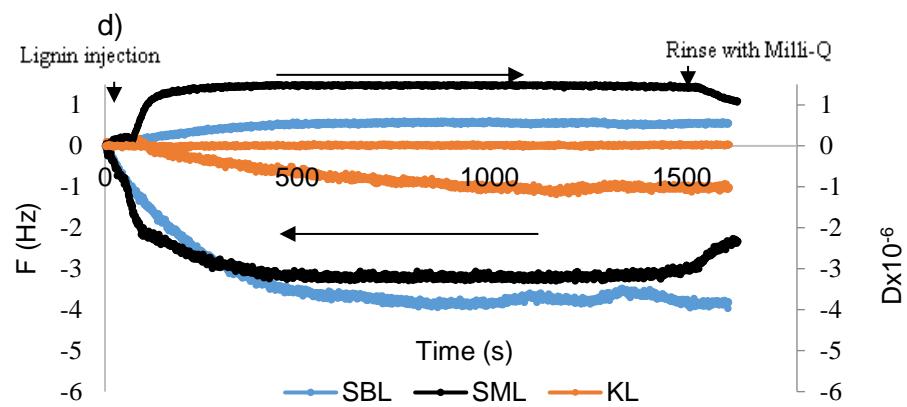
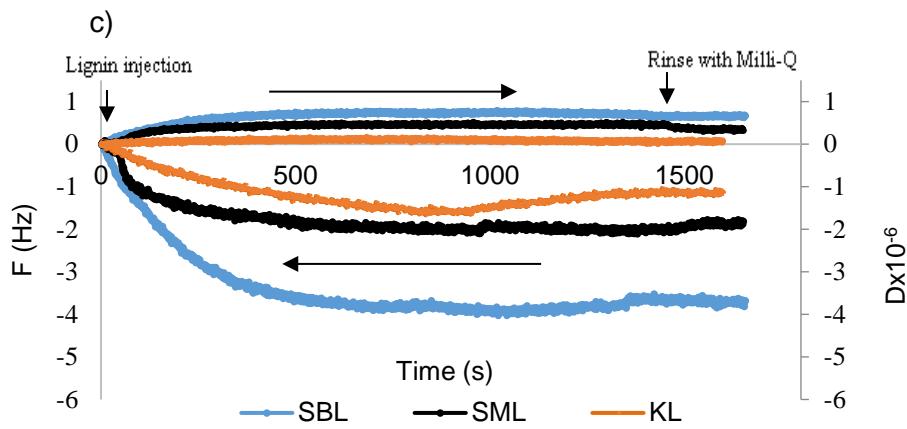


Figure S5. Adsorption of lignin derivatives on Al_2O_3 coated quartz sensors for concentrations **c)** 300 mg/L **d)** 400 mg/L **e)** 600 mg/L at pH 7.8.

Table S3. Pseudo-first-order fitting parameters for the mass uptake of KL

| Concentration (mg/L) | R ² | k (s ⁻¹) | Γ _e (ng/cm ²) | Γ _{exp.} (ng/cm ²) |
|----------------------|----------------|----------------------|---|--|
| 300 | 0.99 | 0.0021 | 32.4 | 26.8 |
| 400 | 0.93 | 0.0004 | 53.7 | 16.2 |

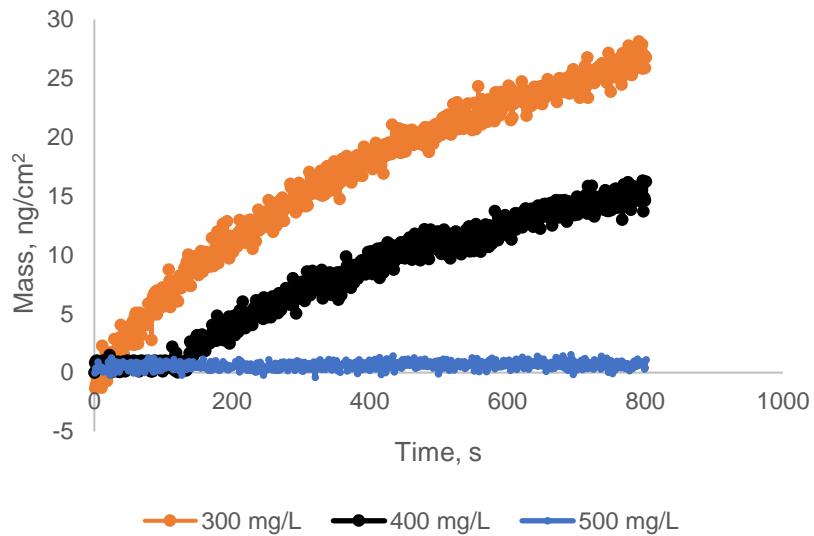


Figure S6. Kinetics of the mass deposition on Al_2O_3 surface for the adsorption of KL

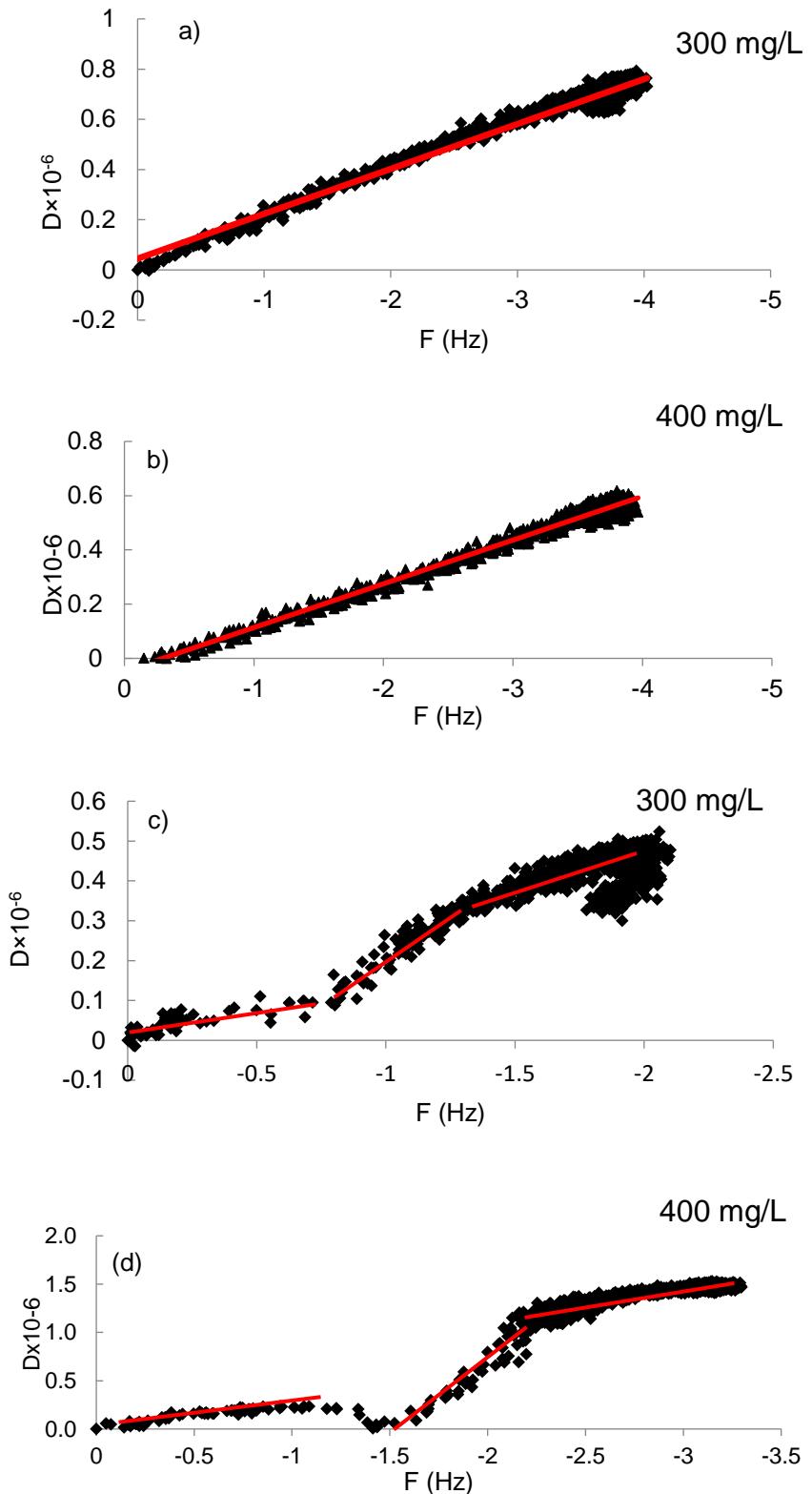


Figure S7. Dissipation change of the sensors as a function of frequency change for **a)** 300 mg/L SBL **b)** 400 mg/L SBL **c)** 300 mg/L SML **d)** 400 mg/L SML.

Table S4. Normalized dissipations together with correlation coefficients for lignin samples

| Concentration (mg/L) | SBL | | SML | |
|----------------------|---------------------------------|---------------------------------|---------------------------------|---------------------------------|
| | K | K ₁ | K ₂ | K ₃ |
| 300 | 0.18 (R ² =0.97) | 0.11 (R ² =0.74) | 0.42 (R ² =0.82) | 0.18 (R ² =0.80) |
| 400 | 0.16 (R ² =0.98) | 0.22 (R ² =0.89) | 1.58 (R ² =0.92) | 0.29 (R ² =0.80) |
| 500 | 0.20 (R ² =0.92) | 0.22 (R ² =0.93) | 0.59 (R ² =0.90) | 0.40 (R ² =0.98) |