


Communication

# Pyrrolylquinoline-BF<sub>2</sub> and BPh<sub>2</sub> BODIPY-Type Analogues: Synthesis, Structural Analysis and Photophysical Properties

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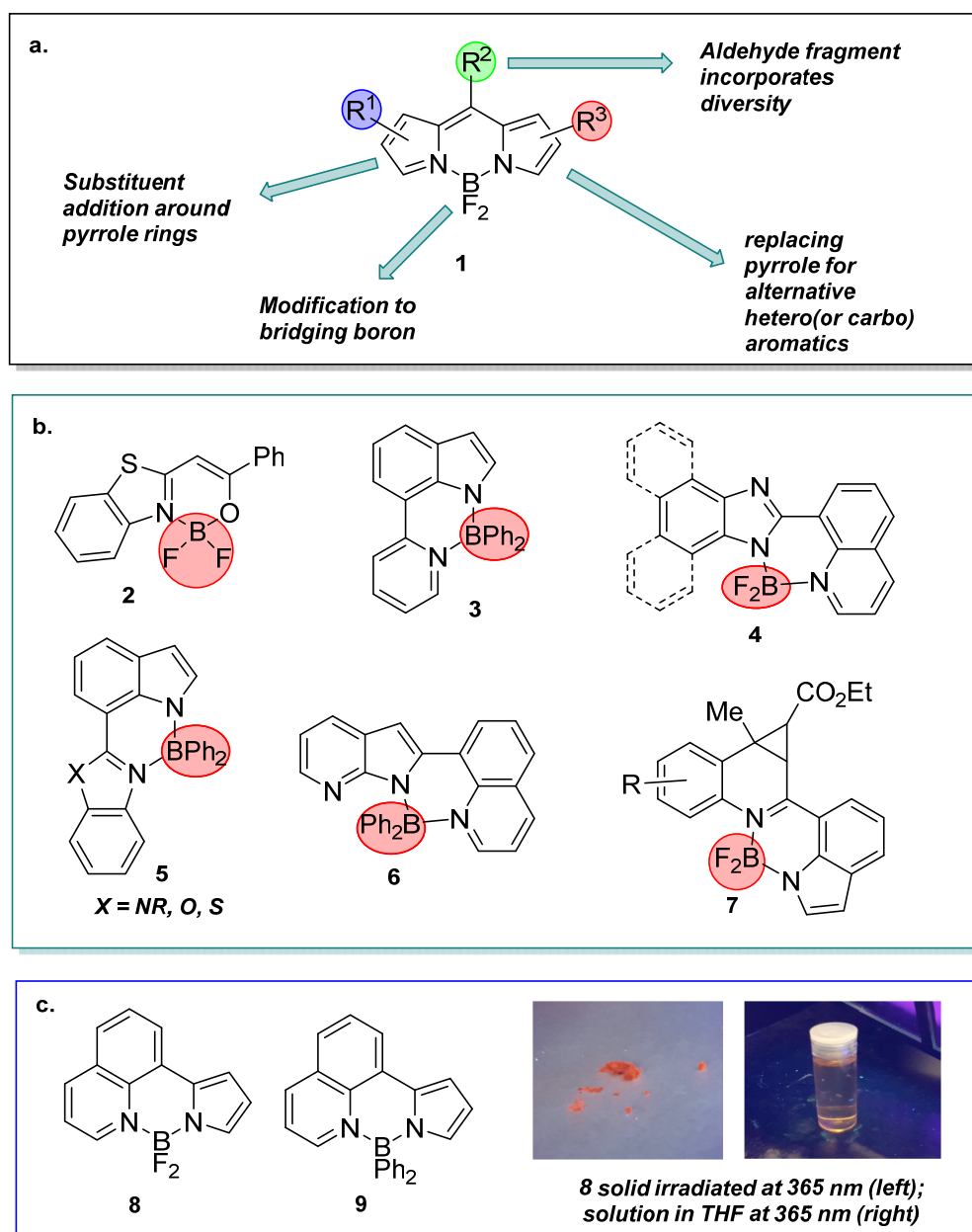
**Abstract:** Two new pyrrolylquinoline-substituted heteroaromatic-containing compounds bearing a central boron bridge have been prepared by a short, high-yielding sequence consisting of Suzuki-coupling of 8-bromoquinoline and *N*-Boc 2-pyrroleboronic acid, thermolytic *tert*-butyloxycarbonyl deprotection, and subsequent boron chelation (either using boron trifluoride or triphenylborane). Both derivatives display longer wavelength absorption maxima ( $\lambda_{\text{max}}^{\text{abs}}$ ) than a previously reported indolopyridine-BPh<sub>2</sub> analogue, in agreement with the smaller HOMO-LUMO energy gap predicted by DFT quantum chemical calculations. Both of the pyrrolylquinoline-boron chelates display weak emission (quantum yields 0.3–0.9%) and the BPh<sub>2</sub> complex displays a very broad, long-wavelength emission ( $\lambda_{\text{max}}^{\text{em}} = 715 \text{ nm}$ , MeCN), which may be due to dimer emission and results in a large pseudo-Stokes' shift ( $7753 \text{ cm}^{-1}$ ) for this compound.

**Keywords:** fluorescent molecules; structural analysis; boron bridged complex; DFT analysis; X-ray crystal analysis

## 1. Introduction

Since the initial report by Treibs and Kreuzer, who detailed the first examples of BODIPY derivatives [1], the field of chemistry associated with developing new fluorescent materials involving  $\pi$ -conjugated systems has drawn considerable attention across scientific boundaries [2–6]. The need for developing new fluorescent molecules is supported by the growing number of applications that they can be found in, for example, as metal or ion sensors [7–10], dyes used in laser spectroscopy [11–15], photo-labels in biological studies [16–20], incorporated into polymers [21,22], molecular rotors [23,24], amongst others. Over the past two decades, notable modifications to the core BODIPY structure have involved modification of the bridging boron atom (and replacement with other atoms) [25–27], addition of substituents around the pyrrole rings [28–30], substitution of the pyrrole rings for other hetero(carbo)aromatics to highlight just a few (Figure 1a) [31–33]. In some cases, a complete overhaul of the dye structure, yet keeping a tetrahedral boron(III) centre, has resulted in a number of novel dyes exhibiting unusual spectroscopic responses [34]. The fusion of quinoline derived units with either indoles [35], pyridines [36] and imidazoles [37], and subsequently formed as boron(III) complexes, has been reported in the literature (and their fluorescent properties studied). These findings built on the seminal work dating back to the late sixties, in which Hohaus reported the first quinoline-boron(III) complexes that

exerted photophysical properties [38]. Furthermore, important alternative heteroaromatics such as thiazole [39] or benzimidazole [37] have also been studied, along with the inclusion of quinoline carboxaldehydes to access numerous quinoline-containing BODIPY analogues (Figure 1b) [40].



**Figure 1.** (a) Selected sites of modification to a classical BODIPY core **1**; (b) selected examples of BODIPY analogues bearing heteroaromatic moieties **2–7**; (c) structures of BF<sub>2</sub> and BPh<sub>2</sub> pyrrolylquinoline derivatives **8** and **9** included in this study, with images showing appearance under 365 nm light irradiation.

Despite these studies, to the best of our knowledge, the boron(III) chelated pyrrolylquinoline type fluorescent dyes shown in Figure 1c have not been reported. These are analogues of the simple pyridine-indole chelate **3** [41] in which the benzo-ring fusion has been transposed between the heteroaromatic cores. In this report, we detail the synthesis of such compounds using Suzuki coupling (bromo-quinoline-pyrrole boronic acid cross-coupling) as the initial step to build the core structure, followed by *N*-*tert*-butoxycarbonyl removal

and subsequent boron complexation reactions to access the desired fluorophores. Along with full characterization of each compound (multi-nuclear NMR, X-ray, DFT, absorption and fluorescence analysis), we herein detail our findings.

## 2. Materials and Methods

**General Methods:**  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra were recorded directly with a Jeol Lambda 500 MHz, Jeol ECS-400 MHz or Bruker Avance 300 MHz. HRMS data were provided by the EPSRC National Mass Spectrometry Service (University of Swansea, UK). X-ray diffraction data were obtained on an Oxford Diffraction Gemini (Xcalibur, Atlas, Gemini ultra-diffractometer) and, using synchrotron radiation, on a Crystal Logics diffractometer fitted with a Rigaku Saturn detector (Xcalibur, Atlas, Gemini ultra-diffractometer equipped with an fine-focus sealed X-ray tube ( $\lambda_{\text{CuK}\alpha} = 1.54184 \text{ \AA}$ ) and an Oxford Cryosystems CryostreamPlus open-flow  $\text{N}_2$  cooling device. Rigaku Oxford Diffraction, 2015). IR spectra were obtained as neat samples using a Perkin Elmer L1600300, Spectrum Two LiTa FTIR spectrometer (Llantrisant, UK) scanning from  $4000\text{--}600 \text{ cm}^{-1}$ . THF and  $\text{Et}_2\text{O}$  were distilled from sodium/benzophenone and used directly. DCM was distilled from calcium hydride and used directly. Quantum chemical calculations were performed using the ORCA ab initio program [42]. Geometry optimizations and TD-DFT calculations were performed using the B3LYP functional, the def2-TZVPP basis set, RIJCOSX algorithm and the Conductor-like Polarizable Continuum Model (CPCM, acetonitrile).

### 2.1. 8-(1-tert-Butoxycarbonylpyrrol-2-yl)quinoline 12

*N*-Boc-2-pyrroleboronic acid (0.456 g, 2.16 mmol, 1.5 eq), 8-bromoquinoline (0.300 g, 1.44 mmol, 1 eq),  $\text{K}_3\text{PO}_4$  (0.612 g, 2.88 mmol, 2 eq), and (2-dicyclohexylphosphino-2',4',6'-diisopropyl-1,1'-biphenyl)[2-(2'-amino-1,1'-biphenyl)]palladium(II) methanesulfonate (LPd(XPhos)OMs) (0.049 g, 0.057 mmol 0.04 eq) were dissolved in dry THF (15 mL) and water (5 mL). The reaction mixture was heated under reflux for 6 h, then allowed to cool to room temperature. The mixture was diluted with  $\text{CH}_2\text{Cl}_2$  (30 mL) and washed with water ( $2 \times 10 \text{ mL}$ ). The organic solvent was dried over  $\text{MgSO}_4$  and filtered. The organic solvent was removed under reduced pressure to yield an oil, which was purified by column chromatography ( $\text{CH}_2\text{Cl}_2$ ) to give the title compound **12** as a dark oil (0.350 g, 1.19 mmol, 82%).  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ):  $\delta$  8.73 (dd,  $J = 4.2, 1.8 \text{ Hz}$ , 1H), 7.96 (dd,  $J = 8.3, 1.8 \text{ Hz}$ , 1H), 7.62 (dd,  $J = 8.2, 1.5 \text{ Hz}$ , 1H), 7.57 (dd,  $J = 7.1, 1.5 \text{ Hz}$ , 1H), 7.40 (dd,  $J = 3.1, 2.0 \text{ Hz}$ , 1H), 7.37 (dd,  $J = 8.2, 7.2 \text{ Hz}$ , 1H), 7.18 (dd,  $J = 8.2, 4.2 \text{ Hz}$ , 1H), 6.26–6.11 (m, 2H), 0.80 (s, 9H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ )  $\delta$  150.12, 149.60, 147.67, 136.00, 134.66, 132.16, 129.59, 128.06, 127.78, 126.14, 122.43, 121.02, 114.56, 110.60, 82.43, 27.20. IR (neat):  $\nu_{\text{max}}/\text{cm}^{-1}$  2980, 1733, 1593. HRMS-ES Calcd for  $\text{C}_{18}\text{H}_{18}\text{N}_2\text{O}_2 + \text{H}^+$ : 295.1440. Found: 295.1447.

### 2.2. 8-(Pyrrol-2-yl)quinoline 13

A round bottomed flask containing 8-(1-tert-butoxycarbonylpyrrol-2-yl)quinoline **12** (0.300 g, 1.02 mmol), was heated for 30 min at  $200^\circ\text{C}$  under nitrogen resulting in the evolution of gas (presumed to be  $\text{CO}_2$ , isobutene). The flask was then allowed to cool to room temperature. The crude product, a greenish solid, was directly used for the next reaction without purification (0.198 g, quant.).  $^1\text{H}$  NMR (400 MHz,  $\text{CDCl}_3$ )  $\delta$  12.68 (br s, 1H), 8.93 (dd,  $J = 4.3, 2.0 \text{ Hz}$ , 1H), 8.18 (dd,  $J = 8.2, 1.8 \text{ Hz}$ , 1H), 8.16 (dd,  $J = 8.2, 1.5 \text{ Hz}$ , 1H), 7.63–7.53 (m, 2H), 7.42 (dd,  $J = 6.2, 3.1 \text{ Hz}$ , 1H), 7.09–7.05 (m, 1H), 6.96 (ddd,  $J = 6.3, 4.2, 3.0 \text{ Hz}$ , 1H), 6.38–6.42 (m, 1H).  $^{13}\text{C}$  NMR (101 MHz,  $\text{CDCl}_3$ )  $\delta$  148.56, 144.61, 137.08, 131.63, 129.17, 129.09, 126.86, 125.04, 125.01, 120.77, 119.10, 108.97, 107.04.

### 2.3. 8,8-. Difluoro-8H-7l4,8l4-pyrrolo[1',2':3,4][1,3,2]diazaborinino[5,6,1-ij]quinolone 8

In a Schlenk tube, 8-(pyrrol-2-yl)quinoline (0.300 g, 1.54 mmol, 1 eq) was dissolved in  $\text{CH}_2\text{Cl}_2$  (7 mL), diisopropylethylamine (1.16 g, 8.96 mmol, 5.8 eq) was added, followed by the addition of  $\text{BF}_3 \cdot \text{Et}_2\text{O}$  (1.75 g, 12.4 mmol, 8 eq). The reaction mixture was stirred for 20 h. The solvent was then removed under reduced pressure to yield a solid, which was

purified by column chromatography ( $\text{CH}_2\text{Cl}_2$ ) to give the title compound **8** as an orange solid (0.270 g, 1.11 g, 72%), mp = 163–165 °C.  $R_f$  = 0.6 ( $\text{CH}_2\text{Cl}_2$ ).  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ )  $\delta$  9.16 (d,  $J$  = 5.4, 1.6 Hz, 1H), 8.62 (dd,  $J$  = 8.3, 1.6 Hz, 1H), 8.17 (dd,  $J$  = 7.1, 1.8 Hz, 1H), 7.82–7.62 (m, 3H), 7.30 (dd,  $J$  = 2.6, 1.4 Hz, 1H), 6.91 (d, 1H), 6.45 (dd,  $J$  = 3.5, 2.5 Hz, 1H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ )  $\delta$  145.10, 144.36, 129.70, 129.43, 127.67, 125.93, 125.41, 124.28, 123.60, 121.04, 111.75, 109.97, 108.05.  $^{11}\text{B}$  NMR (96 MHz,  $\text{CDCl}_3$ )  $\delta$  1.85 (t,  $J$  = 29.3 Hz).  $^{19}\text{F}$  NMR (282 MHz,  $\text{CDCl}_3$ )  $\delta$  -140.05 (q,  $J$  = 29.3 Hz). IR (neat):  $\nu_{\text{max}}/\text{cm}^{-1}$ ; 3062, 1588, 1510, 1378, 1299, 1259, 1209, 1138, 1107, 1065, 1039, 979, 879, 832, 738, 519. HRMS-ES Calcd for  $\text{C}_{13}\text{H}_9\text{BF}_2\text{N}_2 + \text{H}^+$ : 243.0905. Found: 243.0908.

#### 2.4. 8,8-. Diphenyl-8H-7l4,8l4-pyrrolo[1',2':3,4][1,3,2]diazaborinino[5,6,1-ij]quinolone **9**

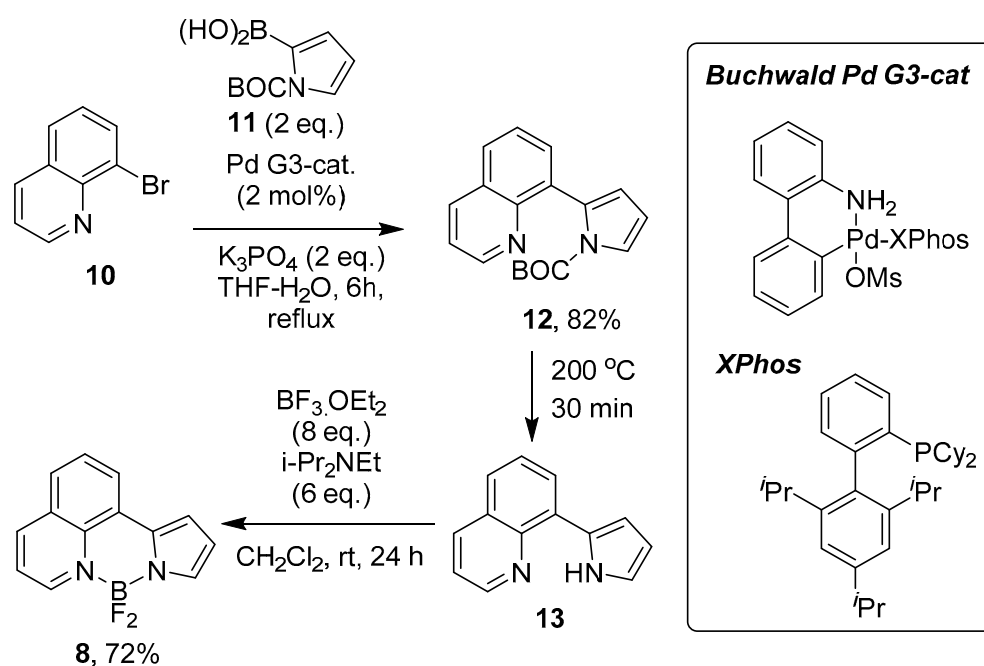
In a Schlenk tube, 8-(pyrrol-2-yl)quinoline (0.280 g, 1.44 mmol, 1.0 eq) **13** and triphenylborane (0.698 g, 2.88 mmol, 2.0 eq) were dissolved in toluene (10 mL). The reaction mixture was heated under reflux for 24 h and then allowed to cool to room temperature. The solvent was removed under reduced pressure to yield a solid, which was purified by column chromatography ( $\text{CH}_2\text{Cl}_2$ ) to give the title compound **9** as a red solid (0.350 g, 0.97 mmol, 68%), mp 260–263 °C.  $R_f$  = 0.7 ( $\text{CH}_2\text{Cl}_2$ ).  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ )  $\delta$  8.76 (dd,  $J$  = 5.5, 1.6 Hz, 1H), 8.42 (dd,  $J$  = 8.3, 1.6 Hz, 1H), 8.12 (dd,  $J$  = 7.4, 1.5 Hz, 1H), 7.69–7.51 (m, 4H), 7.22–7.17 (m, 5H), 7.12–7.06 (m, 4H), 6.95 (dd,  $J$  = 3.5, 1.4 Hz, 1H), 6.68–6.66 (m, 1H), 6.37 (dd,  $J$  = 3.5, 2.4 Hz, 1H).  $^{13}\text{C}$  NMR (75 MHz,  $\text{CDCl}_3$ )  $\delta$  147.01, 142.27, 136.58, 133.29, 129.95, 129.41, 128.93, 128.75, 128.31, 127.75, 127.47, 127.25, 127.17, 126.53, 124.06, 123.00, 120.32, 110.22, 107.60.  $^{11}\text{B}$  NMR (96 MHz,  $\text{CDCl}_3$ )  $\delta$  4.05 (s). IR (neat):  $\nu_{\text{max}}/\text{cm}^{-1}$ ; 3068, 3045, 1587, 1538, 1507, 1429, 1374, 1362, 1230, 1230, 1200, 1177, 1171, 1071, 872, 827, 725, 640, 619. HRMS-ES Calcd for  $\text{C}_{25}\text{H}_{19}\text{BF}_2\text{N}_2 + \text{H}^+$ : 359.1720. Found: 359.1726.

### 3. Results

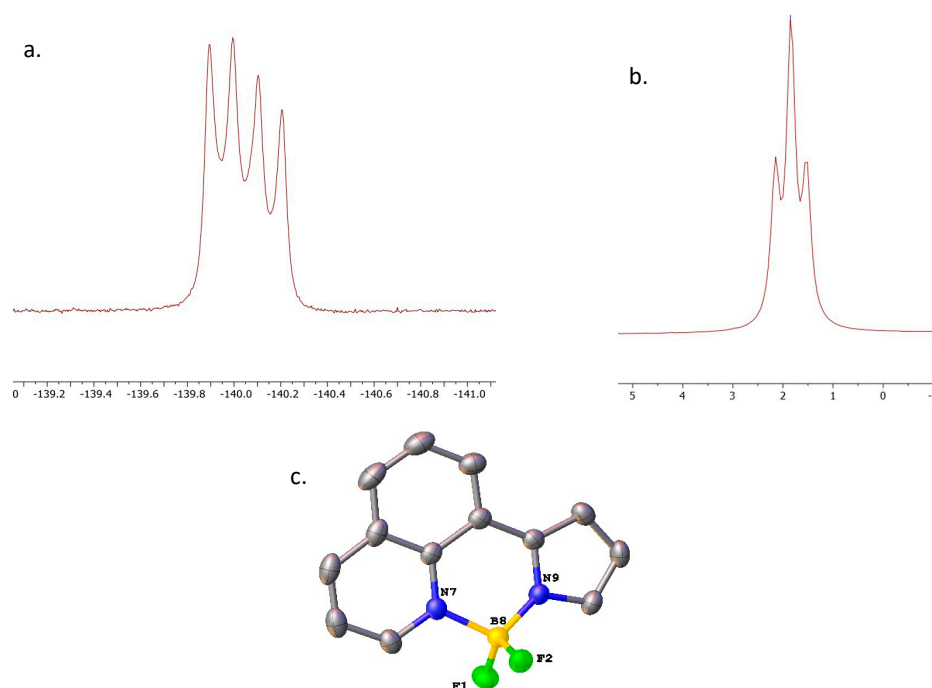
#### 3.1. Synthesis

The synthesis of both pyrrolylquinoline boron bridged analogues **8** and **9** requires access to intermediate **13** before boron complexation. We were pleased to find that Suzuki cross-coupling reaction between commercially available 8-bromoquinoline **10** and *N*-Boc pyrrole-2-boronic acid **11** employing Buchwald's Pd G3 pre-catalyst in combination with XPhos ligand (in 2:1 THF:H<sub>2</sub>O) afforded cross-coupled product **12** in an excellent 82% yield (Scheme 1) [43]. A facile *tert*-butoxycarbonyl deprotection could be realized by solvent-free heating at 200 °C under inert atmosphere to yield the carbamate **12**. Thermolytic cleavage of the *tert*-butyl group, followed by decarboxylation, affords the NH-pyrrole **13** in an essentially quantitative yield (along with isobutene and carbon dioxide as additional products). The desired BF<sub>2</sub>-chelated pyrrolylquinoline derivative **8** was formed from the pyrrole **13** in 72% yield by reaction with boron trifluoride etherate (8 equivalents) in the presence of Hünig's base (6 equivalents).

To confirm the structure of BF<sub>2</sub>-chelated pyrrolylquinoline compound **8**, we initially investigated multi-nuclear NMR analysis for this product (see Supplementary Materials for details of  $^1\text{H}$ ,  $^{13}\text{C}$ ,  $^{19}\text{F}$  and  $^{11}\text{B}$  NMR spectroscopy). Characteristically,  $^{19}\text{F}$  NMR indicated a solitary signal at −140 ppm, with the expected quartet observed due to F-B coupling (Figure 2a). Likewise, the  $^{11}\text{B}$  NMR spectrum showed a triplet at  $\delta$  1.85 ppm for the BF<sub>2</sub>-group (Figure 2b). The boron chelate **8** was crystallized from a solvent system of chloroform and diethyl ether (1:3) and the molecular structure was confirmed by single crystal X-ray diffraction. The bond angle 107.72° (N9-B8-N7) for **8** is close to tetrahedral and similar to the corresponding angle 106.6° (N2-B4-N1) in a typical unsubstituted BODIPY **1** (see Figure 1a) [44]. This suggests that there is no induced strain on the aromatic system of **8** (Figure 2).



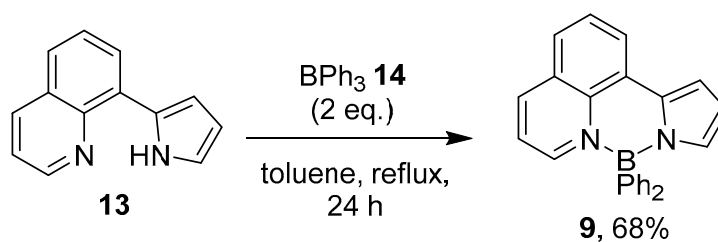
**Scheme 1.** Synthesis of  $\text{BF}_2$ -chelated pyrrolylquinoline **8**.



**Figure 2.** (a)  $^{19}\text{F}$  NMR spectrum (282 MHz,  $\text{CDCl}_3$ ) of the pyrrolylquinoline- $\text{BF}_2$  complex **8**. (b)  $^{11}\text{B}$  NMR spectrum (96 MHz,  $\text{CDCl}_3$ ) of the pyrrolylquinoline- $\text{BF}_2$  complex **8**. (c) X-ray single crystal analysis of pyrrolylquinoline- $\text{BF}_2$  complex **8** (hydrogen atoms have been omitted for clarity).

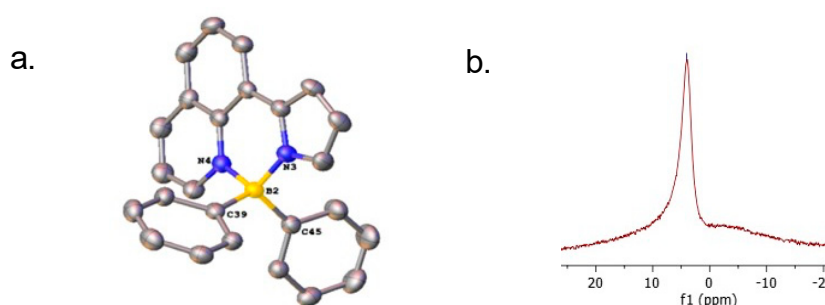
The corresponding pyrrolylquinoline  $\text{BPh}_2$  complex **9** was accessed according to the procedure employed by Curiel and co-workers for the synthesis of the analogous indole **3** [41]. The pyrrolylquinoline intermediate **13** was treated with two equivalents of triphenylborane **14** in toluene under reflux. The  $\text{BPh}_2$  chelate **9** was isolated in 68% yield (Scheme 2).





**Scheme 2.** BPh<sub>2</sub>-chelation of pyrrolylquinoline intermediate **13**.

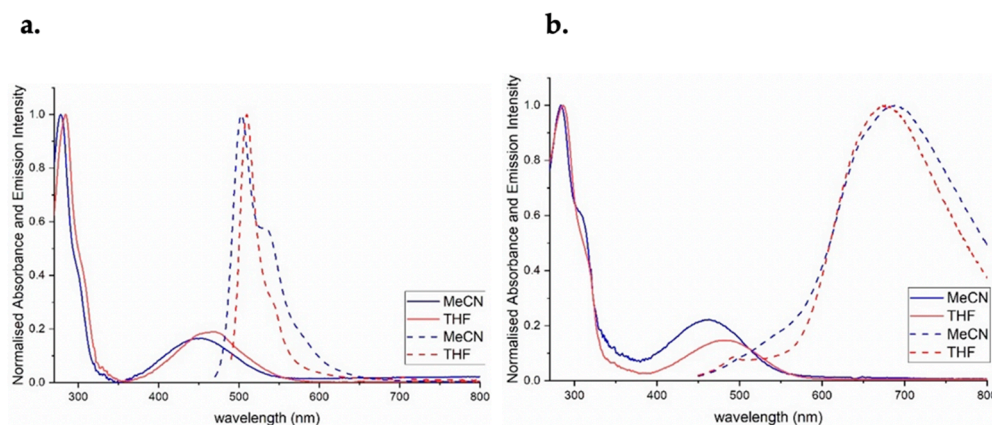
The successful formation of pyrrolylquinoline-BPh<sub>2</sub> **9** was confirmed by X-ray structure and by the <sup>11</sup>B NMR spectroscopy, showing the expected singlet peak at δ 4.05 ppm (Figure 3).



**Figure 3.** (a) X-ray crystal structure for BPh<sub>2</sub>-adduct **9**. (Hydrogen atoms have been omitted for clarity). (b) <sup>11</sup>B NMR spectrum for **9** (96 MHz, CDCl<sub>3</sub>).

### 3.2. Photophysical Properties and DFT Analysis

The UV-vis and emission spectra, measured in acetonitrile and tetrahydrofuran at room temperature, for the two quinoline-derived BODIPY-type analogues (BF<sub>2</sub>) **8** and (BPh<sub>2</sub>) **9** are shown in Figure 4 and the key parameters are summarized in Table 1 together with the previously reported [42] data for the ‘structurally transposed’ indolopyridine-BPh<sub>2</sub> analogue **3** for comparison. Fluorescein was used as the reference compound for the determination of luminescence quantum yield (fluorescein Φ<sub>f</sub> = 0.925, λ<sub>ex</sub> = 405 nm, in NaOH<sub>(aq)</sub> (0.1 M)) [45].



**Figure 4.** Normalized absorption (solid lines) and emission spectra (dashed lines) of (a) pyrrolylquinoline-BF<sub>2</sub> **8** and (b) -BPh<sub>2</sub> **9** in MeCN and THF.

**Table 1.** Photophysical data for the selected BODIPY analogues.

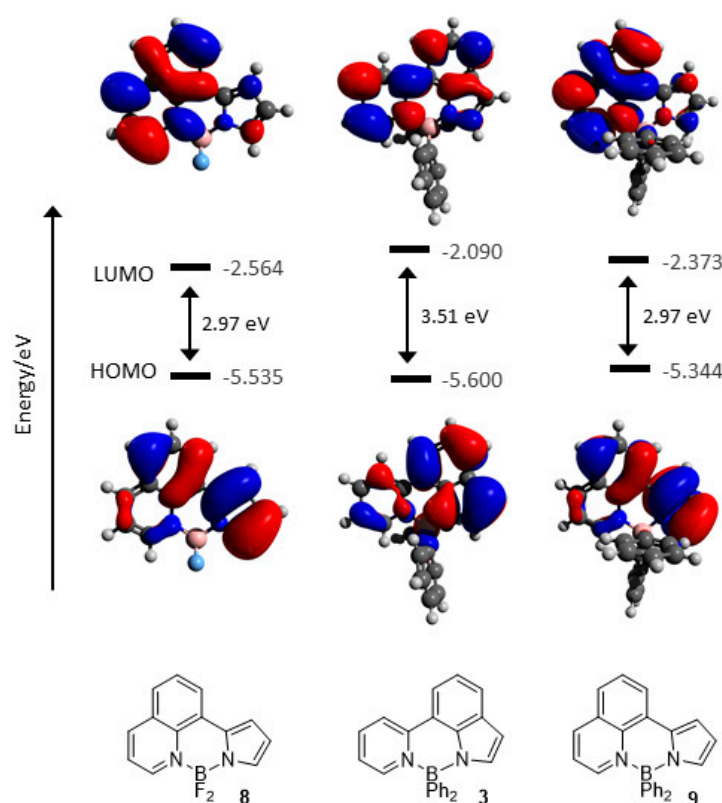
| BODIPY Analogue                                | Solvent | $\lambda_{\text{abs}} (\text{max})$<br>/nm | $\epsilon/10^3 \text{ M}^{-1}$<br>$\text{cm}^{-1}$ | $\lambda_{\text{em}} (\text{max})$<br>/nm | $\Delta\nu_{\text{SS}}/\text{cm}^{-1}$ <sup>c</sup> | $\Phi_f (\lambda_{\text{ex}})$ |
|------------------------------------------------|---------|--------------------------------------------|----------------------------------------------------|-------------------------------------------|-----------------------------------------------------|--------------------------------|
| 8                                              | MeCN    | 278, 450                                   | 15, 2.8                                            | 503                                       | 2342                                                | 0.007 (405)                    |
| 8                                              | THF     | 284, 470                                   | 32, 5.5                                            | 510                                       | 1669                                                | 0.003 (405)                    |
| 9                                              | MeCN    | 283, 460                                   | 13, 3.2                                            | 715                                       | 7753                                                | 0.009 (405)                    |
| 9                                              | THF     | 286, 480                                   | 22, 4.0                                            | 693                                       | 6403                                                | 0.005 (405)                    |
| Indolopyridine BPh <sub>2</sub> 3 <sup>a</sup> | MeCN    | 409                                        | 8.0                                                | 519                                       | 5182                                                | 0.12 <sup>b</sup>              |

<sup>a</sup> Data taken from Reference [41]; <sup>b</sup> reported as determined in ethanol [41]; <sup>c</sup> Stokes' shift.

The UV-vis absorption spectrum of the BF<sub>2</sub> dye **8** in THF solution shows a strong absorption band at 284 nm and a weaker absorption at 470 nm (Figure 4). A weak fluorescence emission ( $\phi_f = 0.3\%$ ) is observed at 510 nm, and the appearance of the emission spectrum strongly resembles that of a typical BODIPY-F<sub>2</sub> dye [2–6]. An increase in solvent polarity (from THF to MeCN) results in a hypsochromic shift in both the absorption and (to a lesser extent) emission  $\lambda_{\text{max}}$ ; consequently, the Stokes' shift increases from 1669 to 2342 cm<sup>−1</sup> (Table 1). The luminescence quantum yield also increases (to 0.7%) in the more polar solvent. The UV-vis absorption spectrum of BPh<sub>2</sub> pyrrolylquinoline chelate **9** is similar to that of the BF<sub>2</sub>-chelate with absorption peaks at 286 and 480 nm in THF, which undergo hypsochromic shifts in MeCN (Figure 4). The emission spectrum of BPh<sub>2</sub> pyrrolylquinoline chelate **9** consists of a very broad, featureless peak at  $\lambda_{\text{max}}$  693 nm in THF with a low quantum yield of emission (0.5%). In the more polar solvent MeCN, a bathochromic shift ( $\lambda_{\text{max}}$  715 nm) and increase in quantum yield (0.9%) is observed. The excitation spectra for both chelates **8** and **9** closely match the absorption spectra, indicating that the emission arises from a single excited state in each case.

The emission behaviour of the BPh<sub>2</sub> chelate **9** is in stark contrast to that of the BF<sub>2</sub> chelate **8**, and it may correspond to emissions from a dimer in solution as a result of poor solubility of the BPh<sub>2</sub> chelate in THF and MeCN, favouring dimerization in solution. We were unable to observe more typical 'monomer-like' emission even on attempted serial dilutions. As a consequence of this unexpected low energy emission, the apparent Stokes shift (Table 1) is much larger than that for either the BF<sub>2</sub> chelate **8** or the reported indolopyridine analogue **3**. It is interesting to note that the reported emission spectrum of the indolopyridine **3** in acetonitrile does feature a broad peak, typical of a BPh<sub>2</sub>-BODIPY [46], but this is centred at a much shorter wavelength (519 nm) than that observed for the isomeric quinoline chelate **9** [41].

Figure 5 shows the HOMO and LUMO orbitals and energies of these pyrrolylquinoline analogues **8** and **9** together with those for the indolopyridine isomer **3** for comparison, calculated using density functional theory (DFT) at the B3LYP/def2-TZVPP level, applying the conductor-like polarizable continuum model (CPCM, with solvent = MeCN) to account for solvent effects. As might be anticipated on the basis of the electronic nature of the ring systems, for both **8** and **9**, the HOMO is predominantly focused on the pyrrole and the carbocyclic ring of the quinoline, whereas the LUMO is predominantly focused on the quinoline rings. The HOMO-LUMO energy gaps for **8** and **9** are 2.97 eV in each case, which is substantially smaller than the HOMO-LUMO gap of 3.51 eV calculated for the indolopyridine **3** at the same level of theory and is in agreement with the significant red shift in the absorbance maximum of the pyrroloquinolines compared to the indole.



**Figure 5.** HOMO and LUMO molecular orbitals and energies (eV) for **8** and **9** together with those for the indolopyridine **3** for comparison. Calculated by DFT at the B3LYP/def2-TZVPP level using ORCA [42].

The principle absorption peak positions (those with  $f_{\text{osc}} > 0.02$ ) were predicted by time-dependent DFT calculations (ESI). Peaks at 608 ( $S_0 \rightarrow S_1$ ) and 278 nm ( $S_0 \rightarrow S_5$ ) were predicted for the  $\text{BF}_2$ -chelate **8**; at 594 ( $S_0 \rightarrow S_1$ ) and 300 nm ( $S_0 \rightarrow S_8$ ) for the  $\text{BPh}_2$ -chelate **9**; and at 454 ( $S_0 \rightarrow S_1$ ), 394 ( $S_0 \rightarrow S_2$ ), 330 ( $S_0 \rightarrow S_7$ ), and 272 nm ( $S_0 \rightarrow S_{11}$ ) for the indolopyridine- $\text{BPh}_2$  chelate **3**, included for comparison. For the indolopyridine- $\text{BPh}_2$  chelate **3**, the calculated absorption spectrum is in reasonable agreement with the reported experimental results, as reported previously [41]. For the new pyrrolylquinoline-chelates **8** and **9**,  $\lambda_{\text{max}}^{\text{abs}}$  is correctly predicted to occur at longer wavelengths than the indole **3**; however, the energy of this transition was significantly underestimated in both cases, leading to a longer wavelength prediction compared to the experimental result. Inspection of the molecular orbital contributions to all of these transitions confirms them to be  $\pi \rightarrow \pi^*$  in nature.

#### 4. Conclusions

Fluorescent boron-complexed dyes have become an area of intensive research due to their potential applications, which were discussed at the beginning of this article. Many approaches to the synthesis of boron(III)-complexed dyes have been reported in the literature, and we have adapted some of these in order to prepare a new core. The two new BODIPY-like pyrrolylquinoline boron-chelates **8** and **9** are prepared in good yields through a simple and short synthetic sequence, and their structures have been confirmed by means of single-crystal X-ray diffraction and multi-nuclear NMR spectroscopy. In comparison to the previously reported indolopyridine derivative **3**, which is isomeric to the pyrrole **9**, the quantum yields for both dyes were low. The emission spectrum of the difluoroboron compound **8** is typical for a heteroaromatic  $N,N,F,F$ -boron chelate but the  $\text{BPh}_2$ -chelate **9** displayed only a very broad, highly red-shifted emission, which may be due to emissions from a dimer in solution and results in a large apparent Stokes' shift. We envisage future



boron(III) analogues to be developed further by adding different substituents to the pyrrole moiety that may enhance the photophysical properties of these complexes. This could be achieved by brominating certain positions on the pyrrole ring, and then incorporating new aromatic or heteroatom moieties via a Suzuki cross-coupling. The apparent emission from an aggregate, in the BPh<sub>2</sub> chelate **9**, may warrant further investigation.

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