

Article

# Bioinspired Design and Computational Prediction of SCS Nickel Pincer Complexes for Hydrogenation of Carbon Dioxide

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**Abstract:** Inspired by the structures of the active site of lactate racemase and H<sub>2</sub> activation mechanism of mono-iron hydrogenase, we proposed a series of sulphur–carbon–sulphur (SCS) nickel complexes and computationally predicted their potentials for catalytic hydrogenation of CO<sub>2</sub>. Density functional theory calculations reveal a metal–ligand cooperated mechanism with the participation of a sulfur atom in the SCS pincer ligand as a proton receiver for the heterolytic cleavage of H<sub>2</sub>. For all newly proposed complexes containing functional groups with different electron-donating and withdrawing abilities in the SCS ligand, the predicted free energy barriers for the hydrogenation of CO<sub>2</sub> to formic acid are in a range of 22.2–25.5 kcal/mol in water. Such a small difference in energy barriers indicates limited contributions of those functional groups to the charge density of the metal center. We further explored the catalytic mechanism of the simplest model complex for hydrogenation of formic acid to formaldehyde and obtained a total free energy barrier of 34.6 kcal/mol for the hydrogenation of CO<sub>2</sub> to methanol.

**Keywords:** lactate racemase; density functional theory; nickel; hydrogenation of CO<sub>2</sub>

## 1. Introduction

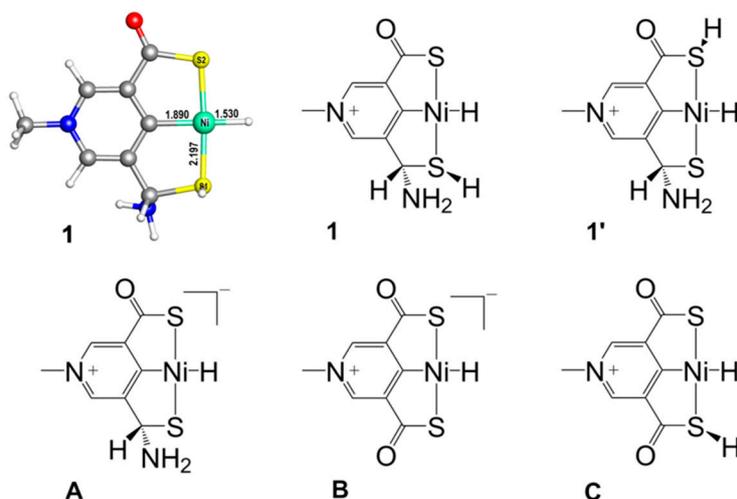
Increasingly severe climate change is driving people to look for effective ways to reduce the concentration of greenhouse gases, especially carbon dioxide, in the atmosphere [1–3]. Transition metal-catalyzed CO<sub>2</sub> reduction has attracted increasing attention because it provides a promising way to use CO<sub>2</sub> as an abundant and non-toxic carbon source for the synthesis of valuable chemicals and fuels [4–6]. With the proposal of the “methanol economy” by Olah and co-workers [7–9], catalytic hydrogenation of CO<sub>2</sub> to methanol (CO<sub>2</sub> + 3H<sub>2</sub> → CH<sub>3</sub>OH + H<sub>2</sub>O) has become one of the most attractive strategies for the utilization of CO<sub>2</sub> as C1 building block and potential hydrogen storage material [10–14]. The above reaction usually contains three cascade catalytic cycles, CO<sub>2</sub> + H<sub>2</sub> → HCOOH, HCOOH + H<sub>2</sub> → CH<sub>2</sub>O + H<sub>2</sub>O, and CH<sub>2</sub>O + H<sub>2</sub> → CH<sub>3</sub>OH.

Although people have achieved some progresses in homogeneous catalytic hydrogenation of CO<sub>2</sub> with the development of base metal iron [15–17], cobalt [18], and manganese [19] catalysts in recent years, most of the experimentally reported efficient CO<sub>2</sub> hydrogenation catalysts contain noble metals and air- and moisture-sensitive phosphine ligands [6,20–32]. For the catalytic hydrogenation of CO<sub>2</sub> to methanol reaction, only a few base metal catalysts are reported so far. For example, Beller and co-workers [18] recently reported a homogeneous cobalt/triphos-based catalyst for the reduction of CO<sub>2</sub> at 70 bar of H<sub>2</sub> and 10 bar of CO<sub>2</sub> under 100 °C and production of methanol with turnover numbers

(TONs) up to 78. Pombeiro and co-workers [16] reported direct synthesis of methanol from CO<sub>2</sub> and H<sub>2</sub> catalyzed by an Fe(II) scorpionate complex achieved 44% yield of methanol with TONs and turnover frequencies (TOFs) up to 2387 and 167 h<sup>-1</sup>, respectively, at 80 °C and 75 pressure. Bernskoetter and co-workers [17] reported an Fe(II) pincer complex for homogeneously catalytic conversion of H<sub>2</sub> and CO<sub>2</sub> to methanol with 250 psi of CO<sub>2</sub> and 1150 psi of H<sub>2</sub> at 100 °C with a TON of 590. Prakash and co-workers [19] investigated phosphorus–nitrogen–phosphorus (PNP) manganese pincer complexes for homogeneous hydrogenation of CO<sub>2</sub> to methanol at 80 bar and 150 °C with TONs up to 36. We can see those experimentally reported base metal catalysts relied on rigid reaction conditions and have rather low activities or methanol selectivity. The rational design of cost-effective non-noble metal catalysts for efficient conversion of H<sub>2</sub> and CO<sub>2</sub> to methanol under mild conditions (<100 °C) is still highly desirable and challenging. In addition to the above experimental studies, Yang and co-workers [33–37] computationally designed several Mn, Fe, and Co complexes as potential catalysts for the production of methanol from H<sub>2</sub> and CO<sub>2</sub>. Those bio-inspired design and computational predictions indicate that metal–ligand cooperation (MLC) is essential for the formation of metal hydride complexes by heterolytic cleavage of H<sub>2</sub> for the reductions of CO<sub>2</sub> and formic acid. However, the application of nickel complexes for catalytic hydrogenation of CO<sub>2</sub> to methanol is still insufficiently investigated.

In our recent study, we computationally predicted a series of scorpion-like sulphur–carbon–sulphur (SCS) nickel complexes as a mimic of the active site of lactate racemase [38] for lactate racemization [39] and dehydrocoupling of ammonia–borane for transfer hydrogenation of ketones and imines [40]. However, those scorpion-like (SCS)Ni complexes are not good candidates for hydrogenation reactions because of their high barriers for H<sub>2</sub> activation. In Yang and Hall’s computational study of monoiron hydrogenase catalysis [41], they found the heterolytic cleavage of H<sub>2</sub> by Fe and sulfide ligand has a rather low free energy barrier of 6.6 kcal/mol. Such results indicate that the metal center and sulfide ligand may cooperate as an intramolecular frustrated Lewis pair (FLP) for H<sub>2</sub> activation. Inspired by the above findings, we would like to explore the potentials of SCS nickel pincer structures for catalytic H<sub>2</sub> activation and (de)hydrogenation reactions.

As shown in Figure 1, we proposed an SCS nickel hydride complex (**1**) by removing the imidazole tail, replacing the substituent group coordinated to Ni with a hydride, and adding a proton to the sulfur atom (S1) in the amine side arm of the SCS ligand. **1'** is an isomer of **1** with the proton on the other sulfur atom (S2) in the SCS ligand. **A** is an anionic complex with deprotonated sulfur. Our density functional theory (DFT) calculations indicate that **A** and **1'** are 10.3 and 13.2 kcal/mol less stable than **1**, respectively, in water.



**Figure 1.** Proposed Ni pincer complex in this work (**1**) and its isomer with the proton on the other sulfur atom (**1'**), deprotonated structure of **1** (**A**), Ni hydride complex with Hu and co-workers’ symmetric SCS pincer ligand (**B**), and **B** with a proton on sulfur (**C**).

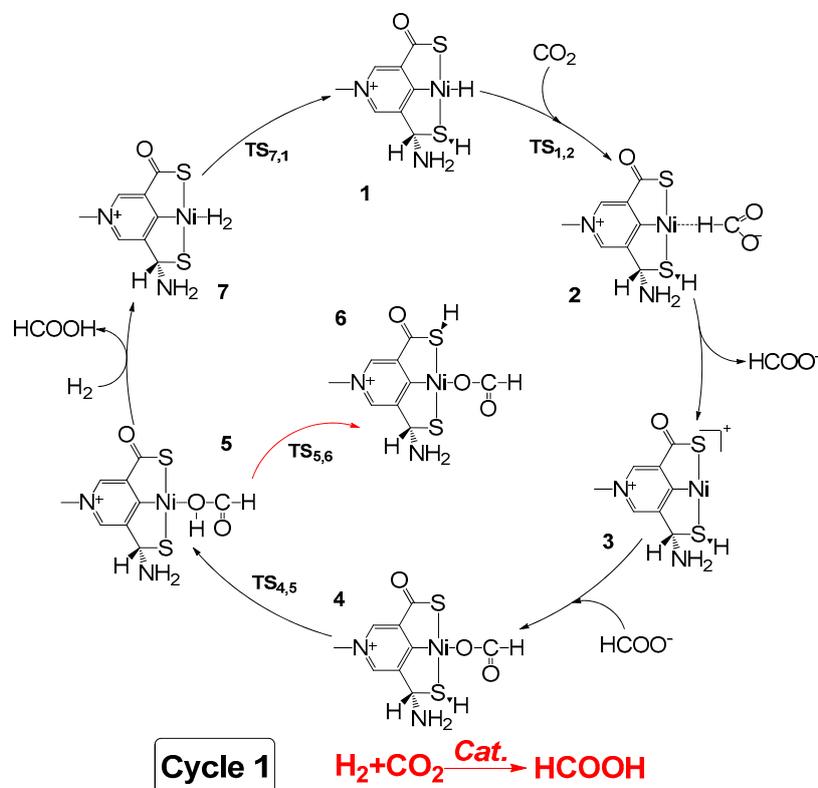
Hu and co-workers [42] recently reported three SCS nickel pincer complexes with symmetric C=O group in the arms of the pincer ligand as functional models of lactate racemase and studied their catalytic properties. Their DFT calculations suggest that the pyridinium carbon atom coordinated to Ni act as a hydride acceptor for lactate racemization, and the  $\beta$ -hydrogen elimination for the formation of Ni–H structure is energetically prohibitive. Figure 1 also lists a Ni hydride complex **B** with symmetric SCS pincer ligand and its protonated structure **C**, which is 7.3 kcal/mol less stable than **B** in water. Such results indicate that the C(H)(NH<sub>2</sub>) group in the SCS ligand significantly increased the nucleophilicity of its connecting sulfur atom and is likely to facility H<sub>2</sub> activation. We also calculated the transition state for **C** to CO<sub>2</sub> hydride transfer and obtained a free energy barrier of 36.3 kcal/mol. Therefore, the symmetric Ni complexes with two C=O groups in the SCS ligand are unlikely to catalyze the hydrogenation of CO<sub>2</sub>.

With the above initial analysis, we believe **1** is a relatively stable nickel hydride complex and could potentially be a catalyst or intermediate for H<sub>2</sub> activation and (de)hydrogenation reactions. We further investigated the detailed reaction mechanism of **1** catalyzed hydrogenation of CO<sub>2</sub> to formic acid, analyzed the influence of various substitutes in the SCS pincer ligand to energy barriers, and examined the catalytic activity of **1** for hydrogenation of formic acid to methanol.

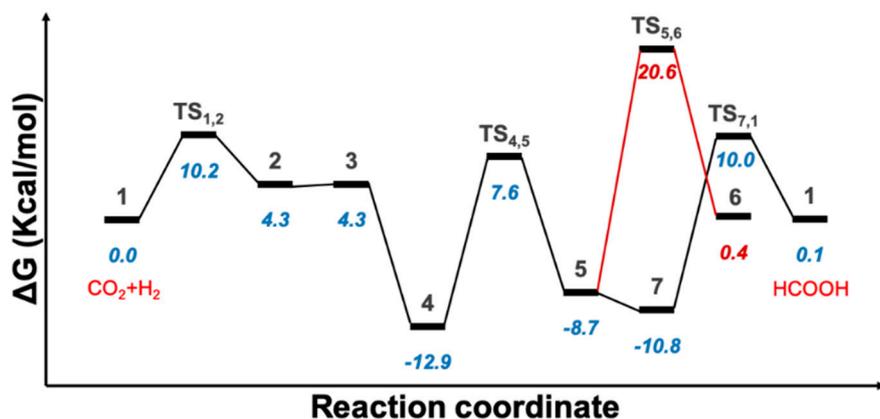
## 2. Results and Discussion

### 2.1. Predicted Catalytic Cycles for the Hydrogenation of CO<sub>2</sub> to Formic Acid

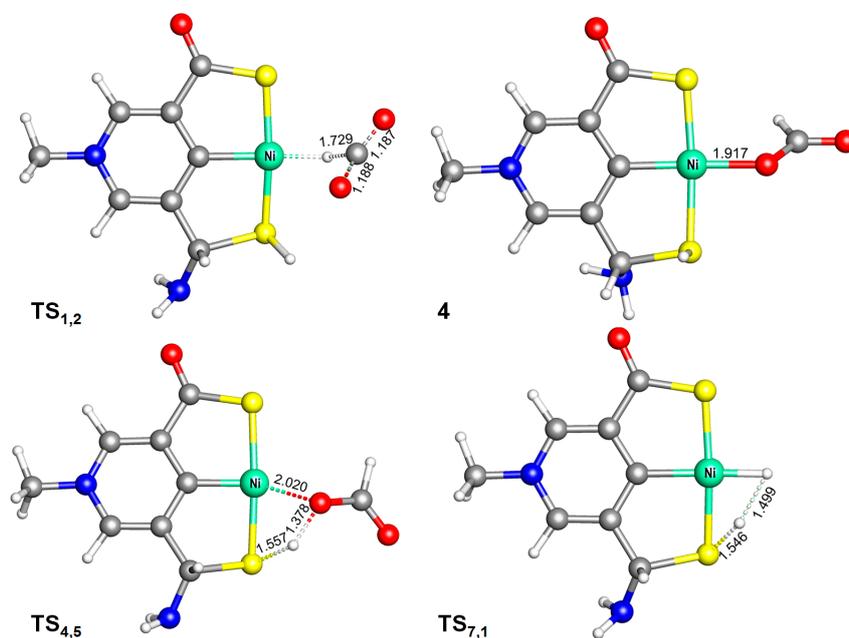
Using **1** as the starting point of the reaction, we proposed a plausible catalytic cycle (Cycle 1, Scheme 1) for the hydrogenation of CO<sub>2</sub> to formic acid based on DFT calculations. The corresponding relative free energies in the reaction pathway and optimized key structures in Cycle 1 are displayed in Figures 2 and 3, respectively.



**Scheme 1.** Predicted catalytic cycle for hydrogenation of CO<sub>2</sub> to formic acid (Cycle 1).



**Figure 2.** Free energy profile for the hydrogenation of CO<sub>2</sub> to formic acid catalyzed by **1**. Solvent corrected relative free energies are shown in italic with blue color.



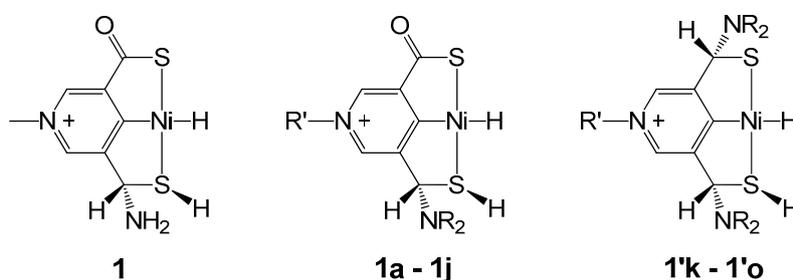
**Figure 3.** Optimized structures of TS<sub>1,2</sub> (673.2i cm<sup>-1</sup>), **4**, TS<sub>4,5</sub> (1614.2i cm<sup>-1</sup>), and TS<sub>7,1</sub> (1013.6i cm<sup>-1</sup>) in Cycle 1. Bond lengths are in Å.

At the beginning of the reaction, a formate anion was formed with a CO<sub>2</sub> molecule attacking **1** and taking the hydride on Ni through TS<sub>1,2</sub> (Figure 3). The free energy barrier for hydride transfer from Ni to C was 10.2 kcal/mol. The formate anion in **2** could easily dissociate and come back for the formation of a much more stable intermediates **4**, which was 12.9 kcal/mol more stable than **1** because of the formation of a strong Ni–O bond (1.917Å) in it. Then the proton on S1 transferred to the O atom bonding to Ni through TS<sub>4,5</sub> and formed a formic acid molecule with a free energy barrier of 20.5 kcal/mol. The Ni···O distance in **5** was slightly elongated to 2.04 Å, which indicates a weaker interaction between Ni and O after the formation of the O–H bond. The HCOOH/H<sub>2</sub> exchange in **5** happened quickly and formed a 2.1 kcal/mol more stable dihydrogen complex **7**. The intramolecular H<sub>2</sub> cleavage in **7** for the regeneration of **1** with the assistance of a S1 atom in the SCS ligand had a free energy barrier of 20.8 kcal/mol (**7** → TS<sub>7,1</sub>). We also examined the stability of an isomer of **4** with a proton on S2, which could be formed by transferring the hydroxyl proton nearby the Ni in **5** to S2 through TS<sub>5,6</sub> with a free energy barrier of 29.3 kcal/mol, and found **6** was 9.1 less stable than **4**. The relative energies in Figure 2 show **4** and TS<sub>1,2</sub> were the rate-determining states in Cycle 1 with a total

free energy barrier of 23.2 kcal/mol ( $4 \rightarrow \text{TS}_{1,2}$ ), which indicates the reaction could easily happen at room temperature for a quick formation of formic acid.

## 2.2. Influence of Substituents in the SCS Ligand

To understand the influence of various functional groups in the SCS pincer ligand to the catalytic activity and find out potential nickel complexes with better catalytic performance, we built ten SCS nickel complexes (**1a–1j**, Figure 4) by replacing the amino hydrogens (R) and the methyl group in the pyridinium ring (R') in **1** with various substituents. Considering the difficulties in synthesizing asymmetric SCS pincer ligand, we also proposed five complexes with symmetrical amine groups in the SCS pincer ligand (**1'k–1'o**, Figure 4). Similar nickel complexes with symmetrical amine groups in SCS ligand were recently reported by Hu and co-workers [43]. Since the free energy profile in Figure 2 indicates that **4** and  $\text{TS}_{1,2}$  were the rate-determining states in Cycle 1, and  $\text{TS}_{7,1}$  was very close to  $\text{TS}_{1,2}$  in relative energy, we calculated the relative free energies of  $4 \rightarrow \text{TS}_{1,2}$  and  $4 \rightarrow \text{TS}_{7,1}$  with different functional groups (Table 1).



**Figure 4.** Proposed SCS nickel pincer complexes with different substituents at R and R'.

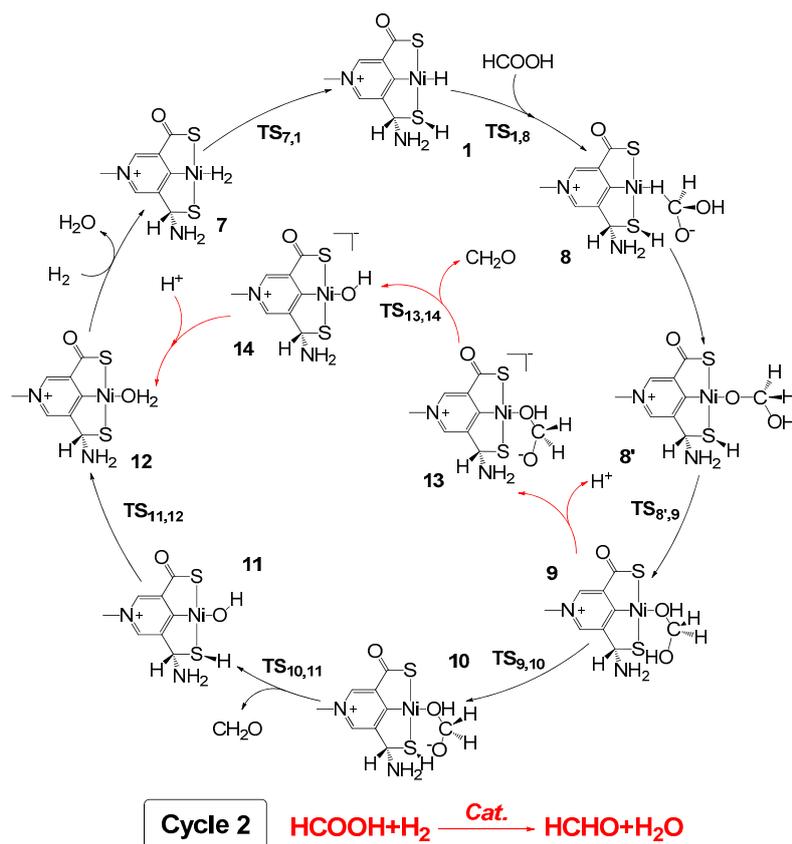
**Table 1.** Relative free energies of  $4 \rightarrow \text{TS}_{1,2}$  and  $4 \rightarrow \text{TS}_{7,1}$  with different functional groups at R and R' in the SCS nickel pincer complexes.

	R	R'	$4 \rightarrow \text{TS}_{1,2}$ (kcal/mol)	$4 \rightarrow \text{TS}_{7,1}$ (kcal/mol)
<b>1</b>	H	CH <sub>3</sub>	23.2	23.0
<b>1a</b>	F	CH <sub>3</sub>	23.9	23.7
<b>1b</b>	Cl	CH <sub>3</sub>	22.9	22.5
<b>1c</b>	Br	CH <sub>3</sub>	22.2	21.3
<b>1d</b>	CH <sub>3</sub>	CH <sub>3</sub>	22.6	23.0
<b>1e</b>	H	H	23.6	23.5
<b>1f</b>	H	COOH	23.4	21.0
<b>1g</b>	H	NH <sub>2</sub>	22.9	23.3
<b>1h</b>	H	F	23.3	22.5
<b>1i</b>	H	Cl	23.3	20.9
<b>1j</b>	H	CN	23.2	22.0
<b>1'k</b>	H	CH <sub>3</sub>	24.6	25.2
<b>1'l</b>	F	CH <sub>3</sub>	25.3	24.5
<b>1'm</b>	Cl	CH <sub>3</sub>	22.9	22.2
<b>1'n</b>	Br	CH <sub>3</sub>	22.5	20.8
<b>1'o</b>	CH <sub>3</sub>	CH <sub>3</sub>	24.4	25.5

Comparing the calculated relative free energies listed in Table 1, we can see complex **1c** with R = Br and R' = CH<sub>3</sub> had the lowest total free energy barrier of 22.2 kcal/mol, while **1'o** had the highest barrier of 25.5 kcal/mol. Such a 3.3 kcal/mol difference in free energy barriers indicates that the substituents at R and R' in the SCS ligand have moderate influences on catalytic activity because those functional groups are far away from the reaction center and have limited contributions to the electron density on nickel.

### 2.3. Hydrogenation of Formic Acid to Formaldehyde and Water

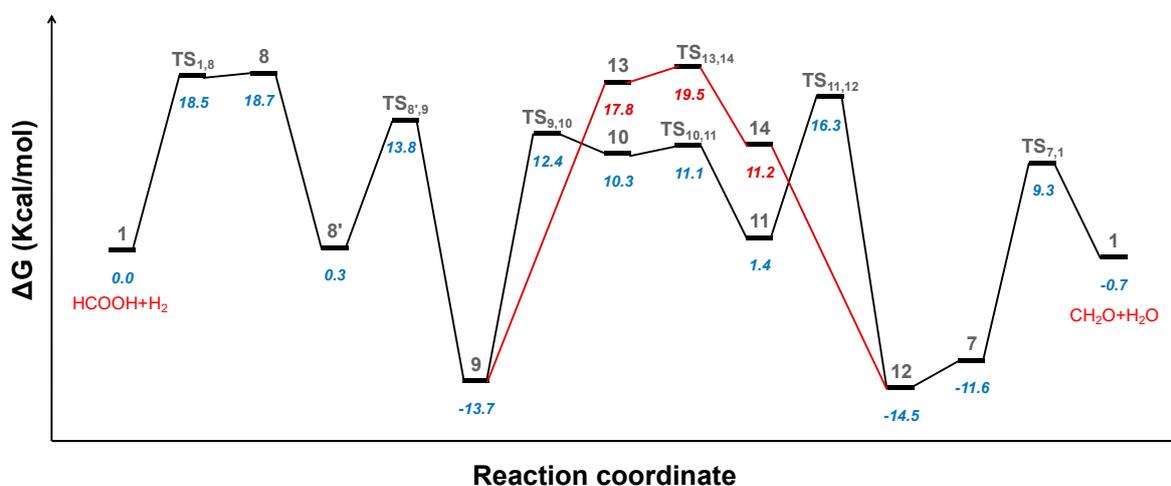
The relative free energies listed in Table 1 indicate **1** and its derivatives were promising candidates for catalytic hydrogenation of CO<sub>2</sub> to formic acid under mild conditions. We further investigated the catalytic activity of **1** for hydrogenation of formic acid. Scheme 2 and Figure 5 are a predicted mechanism for the hydrogenation of HCOOH to CH<sub>2</sub>O and H<sub>2</sub>O catalyzed by **1** (Cycle 2), and the corresponding reaction coordinate with relative free energies. The optimized structures of stable intermediates, **8'** and **12**, and transition states for hydride transfer (**TS**<sub>1,8</sub>), proton transfers (**TS**<sub>9,10</sub> and **TS**<sub>11,12</sub>), and C–O bond cleavages (**TS**<sub>10,11</sub> and **TS**<sub>13,14</sub>) are displayed in Figure 6.



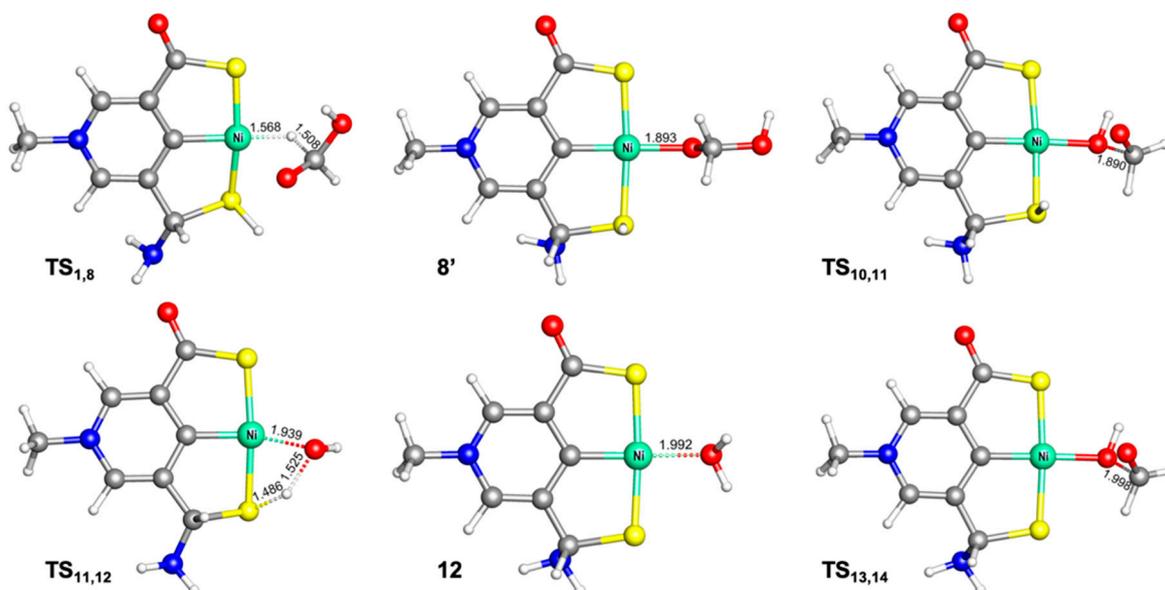
**Scheme 2.** Predicted mechanism for the hydrogenation of formic acid to formaldehyde and water catalyzed by **1** (Cycle 2). The deprotonation pathway is shown with red arrows.

When a formic acid molecule approaches **1**, it can take the hydride on Ni and forms an anionic group HOCH<sub>2</sub>O<sup>−</sup> through an 18.5 kcal/mol barrier transition state **TS**<sub>1,8</sub>. HOCH<sub>2</sub>O<sup>−</sup> in **8** could easily reorient and form an 18.4 kcal/mol more stable isomer **8'** with a strong Ni–O bond (1.893 Å). Then the proton on S1 transferred to the oxygen atom bonding to Ni and formed a methanediol molecule through **TS**<sub>8',9</sub>. After the formation of methanediol in **9**, there are two possible ways for the formation of formaldehyde, intramolecular proton transfer, and deprotonation to solvent. In the intramolecular proton transfer pathway, the hydroxyl proton far away from Ni transfers to S1 and forms intermediate **10**. Then a CH<sub>2</sub>O molecule can easily be released with C–O bond cleavage through **TS**<sub>10,11</sub>, which was only 0.8 kcal/mol higher than **10**. The hydroxyl group left on Ni then takes the proton on S1 and forms a water molecule in intermediate **12**. Instead of intramolecular proton transfer, the hydroxyl proton far away from Ni could be deprotonated in the solvent and form an anionic intermediate **13** with a free energy barrier of 31.5 kcal/mol. The following transition state (**TS**<sub>13,14</sub>) for the release of formaldehyde with a C–O bond cleavage was only 1.7 kcal/mol higher than **13**. The hydroxyl anion in

**14** can easily recapture a proton in the solvent and form a water molecule. The H<sub>2</sub>O/H<sub>2</sub> exchange in **12** for the formation of **7** was an only 2.9 kcal/mol uphill step.



**Figure 5.** Free energy profile for the hydrogenation of formic acid to formaldehyde and water catalyzed by **1** (Cycle 2). Solvent corrected relative free energies are shown in italic. The deprotonation pathway is shown in red.



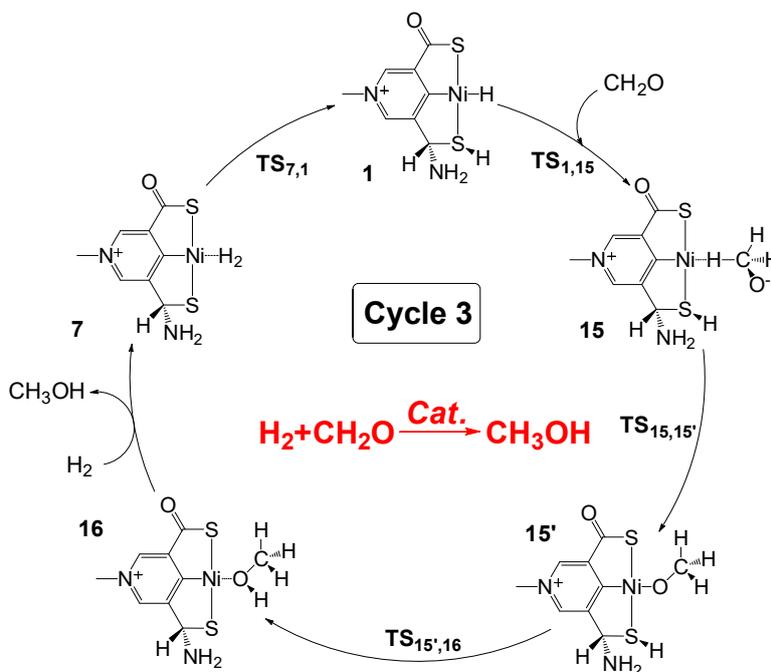
**Figure 6.** Optimized structures of TS<sub>1,8</sub> (508.7i cm<sup>-1</sup>), 8', TS<sub>10,11</sub> (262.4i cm<sup>-1</sup>), TS<sub>11,12</sub> (1352.3i cm<sup>-1</sup>), 12, and TS<sub>13,14</sub> (189.9i cm<sup>-1</sup>). Bond lengths are in Å.

It is worth noting that the relative free energies of **13** and TS<sub>13,14</sub> were calculated based on an experimental value of  $-262.5$  kcal/mol for the solvent-free energy of the proton in water [44]. The solvent environment will strongly influence the free energy barriers of this reaction pathway and may lead to a lower total free energy barrier or even a different mechanism for hydrogenation of CO<sub>2</sub> and formic acid [45].

#### 2.4. Hydrogenation of Formaldehyde to Methanol

Scheme 3 is the mechanism for the hydrogenation of formaldehyde to methanol (Cycle 3). The corresponding free energy profile and optimized key structures in Cycle 3 are displayed in Figures 7 and 8, respectively. When a formaldehyde molecule approaches **1**, the hydride on Ni can easily be

transferred to its carbon atom for the formation of a methoxy anion through transition state  $TS_{1,15}$  (Figure 8), which was only 7.3 kcal/mol higher than **1** in free energy. The rotation of methoxy anion in **15** formed a 22.5 kcal/mol more stable isomer **15'** with a strong Ni-O bond (1.881 Å). Then a methanol molecule was formed with the transfer of the proton on S1 to the oxygen in methoxy through  $TS_{15',16}$ . The  $CH_3OH/H_2$  exchange in **16** for the regeneration of **7** was a 6.0 kcal/mol uphill step. The total free energy barrier of Cycle 3 was 25.8 kcal/mol ( $16 \rightarrow TS_{7,1}$ ).



Scheme 3. Predicted mechanism for hydrogenation of formaldehyde.

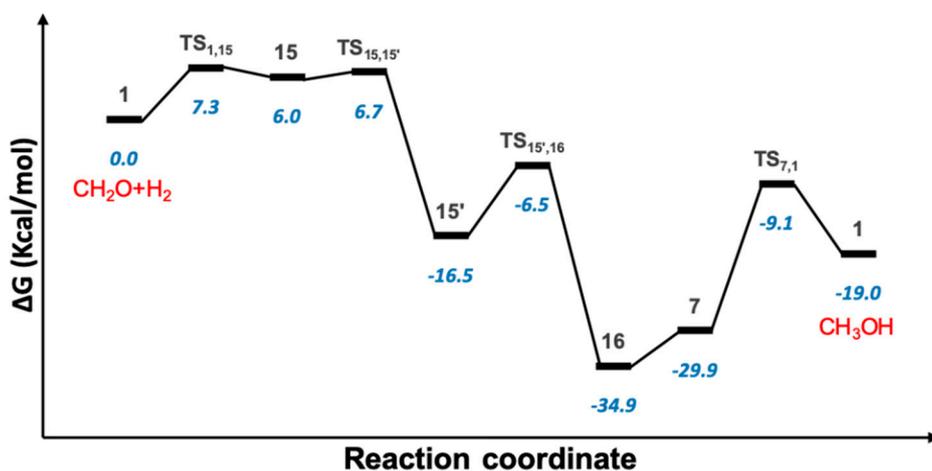
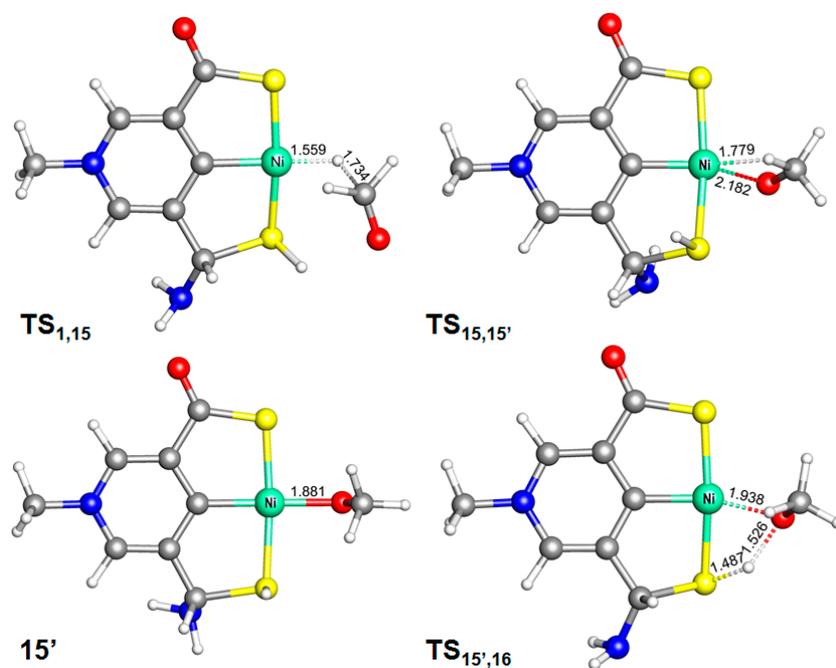


Figure 7. Free energy profile for the hydrogenation of formaldehyde to methanol catalyzed by **1** (Cycle 3). Solvent corrected relative free energies are shown in italic.



**Figure 8.** Optimized structures of  $15'$ ,  $TS_{1,15}$  ( $532.2i\text{ cm}^{-1}$ ),  $TS_{15,15'}$  ( $191.8i\text{ cm}^{-1}$ ), and  $TS_{15',16}$  ( $1347.0i\text{ cm}^{-1}$ ). Bond lengths are in Å.

### 3. Computational Details

Gaussian 09 suite of programs (Revision E.01, Gaussian, Inc., Wallingford, CT, USA, 2009) [46] were used for all DFT calculations with the M06 functional [47]. The all-electron 6-31++G(d,p) basis set [48–50] was used for H, C, N, O, and S atoms, while the Stuttgart relativistic effective core potential basis set (ECP10MDF) was used for Ni [51]. Numerical integrations were at the ultrafine grid level (99, 590). The reliability of M06 functional for this Ni catalytic system was evaluated by comparing the relative free energies between rate-determining states **4** and  $TS_{1,2}$  using different density functionals, which are listed in Table S1. The ground states of key intermediates were confirmed as singlets through comparison with their optimized high-spin analogs (Table S2). All structures were fully optimized with the solvent effect corrections using the integral equation formalism polarizable continuum model (IEFPCM) [52] and the solvation model based on density (SMD) radii [53] for water, which could be an environmentally benign, low-cost, and universal solvent for future experimental study. For thermal corrections on optimized structures, 298.15 K temperature, 1 atm pressure, and harmonic potential approximation were used. The experimental value of  $-262.5\text{ kcal/mol}$  for the free energy of the proton in water [44] was used to calculate deprotonation free energies. Intermediates and transition states were validated through the number of imaginary vibrational modes shown in frequency calculation results. Intrinsic reaction coordinate (IRC) calculations confirmed all transition states were connecting proper reactants and products. The JIMP2 molecular visualizing and manipulating program (version 0.091, Texas A&M University, College Station, TX, USA, 2006) [54] was used to draw 3D molecular structure figures displayed in the text.

### 4. Conclusions

In summary, our computational study predicted a series of (SCS)Ni(II) pincer complexes as promising candidates for catalytic hydrogenation of  $CO_2$  to formic acid with free energy barriers in a range of 22.2–25.5 kcal/mol in an aqueous solvent. Our DFT calculations reveal a metal–ligand cooperative mechanism with the participation of a sulfur atom in the SCS pincer ligand for the heterolytic cleavage of  $H_2$ . In general, the complexes with electron-donating groups have lower barriers for  $CO_2$  reduction. The free energy barrier differences for the formation of formic acid

catalyzed by proposed nickel complexes with various functional groups at R and R' positions were less than 3.3 kcal/mol, which indicates a moderate substituent effect for the catalytic activities. The above findings not only provide well-defined prototypical base metal complexes as promising candidates for catalytic hydrogenation of CO<sub>2</sub> to formic acid but also point to a way to design efficient catalysts for hydrogenation and dehydrogenation reaction, in which the ligand sulfide group may play an essential role for H<sub>2</sub> activation or formation through MLC.

Our further investigation of the catalytic activities of model complex **1** for hydrogenation of formic acid obtained a free energy barrier of 32.5 kcal/mol (**12** → **8**) for the formation of formaldehyde. The free energy barrier for hydrogenation of formaldehyde to methanol catalyzed by **1** was 25.8 kcal/mol (**16** → **TS<sub>7,1</sub>**). It is worth noting that although **1** was described as the starting point of the catalytic cycle, it was actually an unstable intermediate in the reaction. The dihydrogen complex **7**, which was a 10.9 kcal/mol more stable isomer of **1**, should be considered as the real catalyst. By comparing all relative free energies in the above three catalytic cycles, we can conclude that the formation of HOCH<sub>2</sub>O<sup>−</sup> anion groups via hydride transfer from Ni to formic acid is the rate-determining step in the whole catalytic hydrogenation of CO<sub>2</sub> to methanol reaction with a total free energy barrier of 34.6 kcal/mol (**16** → **8**). Although such a barrier is slightly too high for a reaction under mild conditions, it is a good starting point for further design of SCS nickel pincer complexes with improved catalytic activities for CO<sub>2</sub> reduction and various (de)hydrogenation reactions.

Moreover, it is worth noting that the pH value of the solvent could impact the catalytic activities and the stability of proposed SCS nickel pincer complexes. Some CO<sub>2</sub> will form hydrogen carbonate in water and make the solvent slightly acidic with the equilibria among dissolved CO<sub>2</sub> gas, carbonic acid, and hydrogen carbonate. In a weak acidic environment, the solvent free energy of proton is slightly higher, which will lead to higher barriers for the deprotonation pathways in Cycle 1 and 2, and make the proposed Ni complexes less likely to be deprotonated, but will not strongly impact the overall barriers of the catalytic cycles.

**Supplementary Materials:** The following are available online at <http://www.mdpi.com/2073-4344/10/3/319/s1>, Table S1: Absolute and relative free energies of rate-determining states **4** and **TS<sub>1,2</sub>** calculated by using different density functionals, Table S2: Absolute and relative electronic energies of singlet and triplet states.

**Author Contributions:** Conceptualization, X.L. and X.Y.; Methodology, X.L. and B.Q.; Writing – original draft, X.L.; Writing—review and editing, B.Q. and X.Y. All authors have read and agreed to the published version of the manuscript.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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