

Supplementary information

Voltammetric Investigation of Ferulic Acid at Disposable Pencil Graphite Electrode

Iulia Gabriela David ^{1,*}, Dana Elena Popa ^{1,*}, Mihaela Buleandra ¹, Silvia Nicoleta Codreanu ¹, Lorelei Croitoru ¹, Laura Andreea Iordache ¹ and Hassan Noor ^{2,3}

¹ Department of Analytical Chemistry and Physical Chemistry, Faculty of Chemistry, University of Bucharest, Panduri Av. 90-92, District 5, 050663 Bucharest, Romania; mihaela.buleandra@g.unibuc.ro (M.B.); codreanu.silvia233@gmail.com (S.N.C.); croitoru.lorelei@gmail.com (L.C.); lauraiordache24@gmail.com (L.A.I.)

² Department of Surgery, Faculty of Medicine, "Lucian Blaga" University Sibiu, Lucian Blaga Street 25, 550169 Sibiu, Romania; hassan.noor@ulbsibiu.ro

³ European Hospital Medlife-Polisano, Strada Izvorului 1A, 550169 Sibiu, Romania

* Correspondence: gabrielaiulia.david@g.unibuc.ro (I.G.D.); elena.popa@chimie.unibuc.ro (D.E.P.)

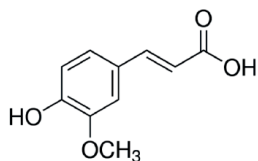


Figure S1. FA chemical structure

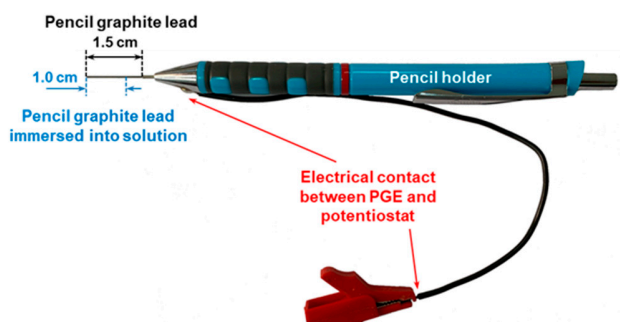


Figure S2. The PGE.

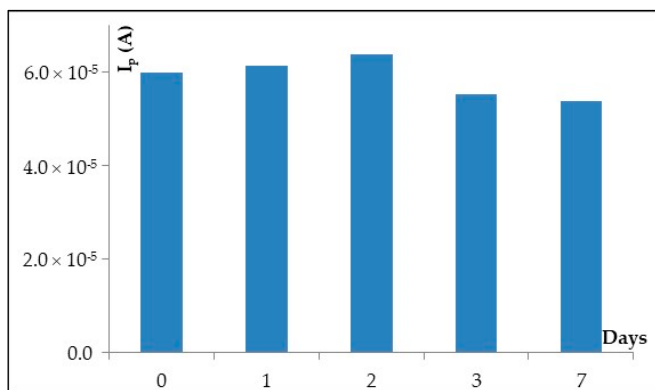


Figure S3. The variation of the DPV anodic peak current recorded at HB_PGE for 1.00×10^{-3} mol/L FA in ABS pH 4.00 prepared at different days from the same ethanolic 1.00×10^{-2} mol/L FA stock solution.

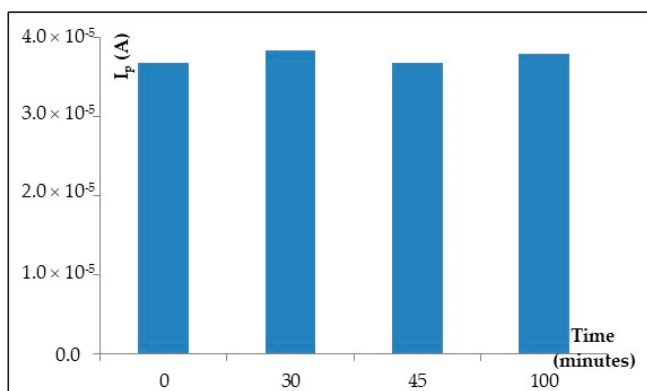


Figure S4. The variation of the DPV anodic peak current recorded at different time intervals at HB_PGE for a 5.00×10^{-4} mol/L FA in ABS pH 4.00.

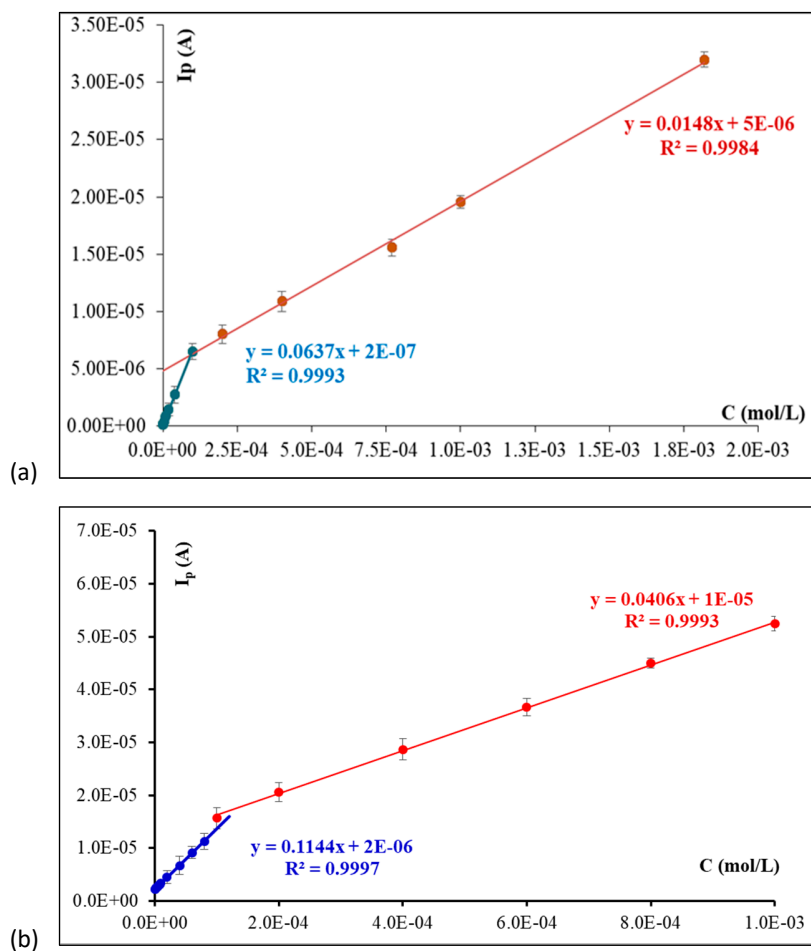


Figure S5. The $I_p=f(C)$ dependencies obtained for the analysis of FA in BRB pH 4.56 solutions at HB_PGE by (a) DPV and (b) SWV.

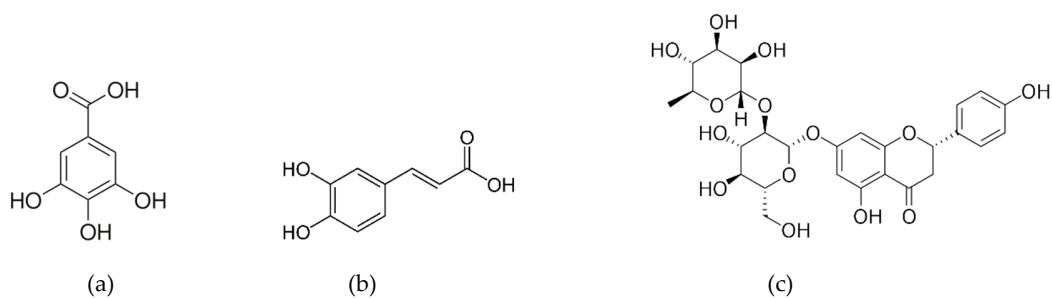


Figure S6. Chemical structures of the polyphenols tested as possible interferents in FA DPV analysis at HB_PGE (a) gallic acid; (b) caffeic acid and (c) naringin.

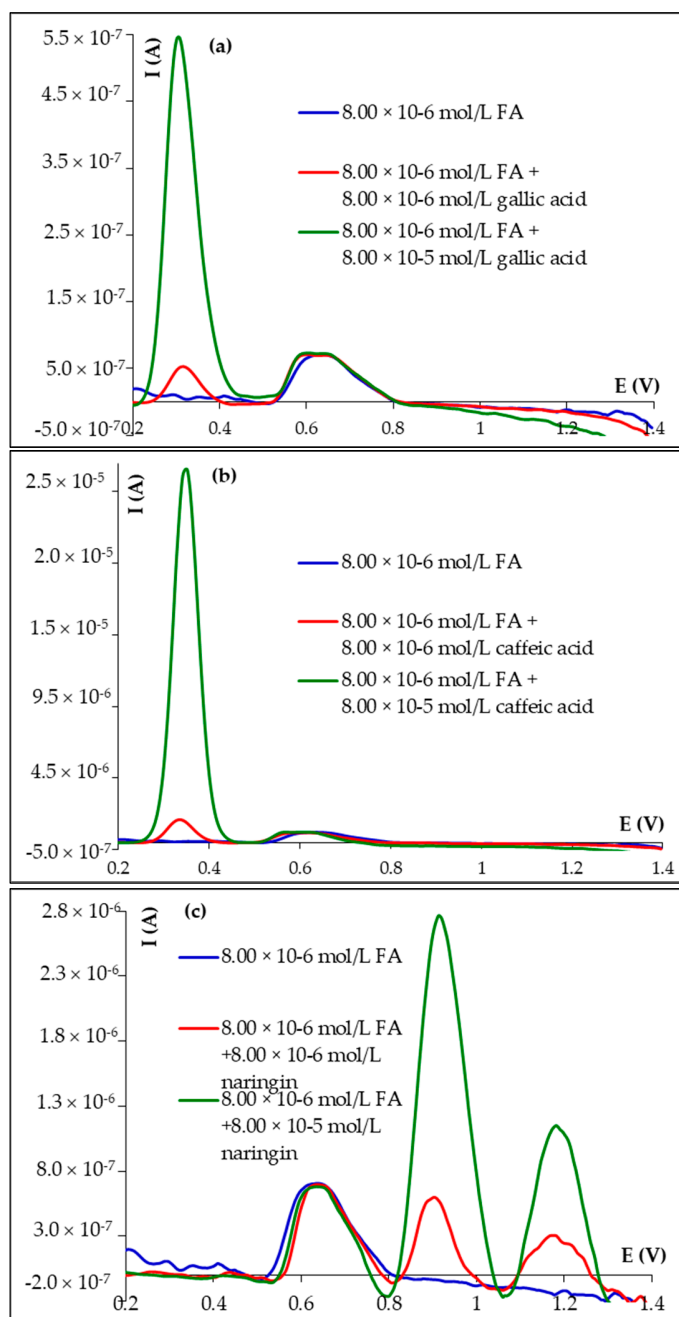


Figure S7. Differential pulse voltammograms recorded at HB_PGE for 8.00×10^{-6} mol/L FA in BRB pH 4.56 solutions in the presence of (a) gallic acid; (b) caffeic acid and (c) naringin.