

## Article

# Basic Properties of MgAl-Mixed Oxides in CO<sub>2</sub> Adsorption at High Temperature

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**Abstract:** The increase of consciousness towards global warming and the need to reduce greenhouse gas emissions lead to the necessity of finding alternative applications based on easy-to-use materials in order to control and reduce global CO<sub>2</sub> emissions. Layered Double Hydroxides (LDHs) and LDH-derived materials are potentially good adsorbents for CO<sub>2</sub>, thanks to their low cost, easy synthesis, high sorption capacity, and surface basicity. They have been intensively studied in CO<sub>2</sub> capture at high temperature, presenting variable sorption capacities for MgAl LDHs with the same composition, but prepared under different synthesis conditions. The ambient temperature coprecipitation synthesis method is an attractive one-step procedure to synthesize LDHs under mild conditions, with low energy consumption and short synthesis time. The present study is based on the synthesis and characterization of hydrotalcites by a mild-conditions coprecipitation process and the production of derived mixed oxides to be used as CO<sub>2</sub> adsorbents. A critical comparison to similar materials is reported. Moreover, the effect of the surface basicity of the derived mixed oxides (measured by adsorption calorimetry) and the CO<sub>2</sub> sorption capacity are discussed, showing a linear correlation between the amount of weak and very strong basic sites and the CO<sub>2</sub> adsorption behavior.

**Keywords:** Layered Double Hydroxides; mixed oxides; basicity; CO<sub>2</sub> adsorption; adsorption calorimetry



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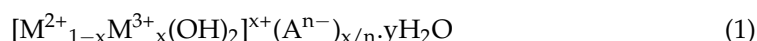
## 1. Introduction

Global warming, pollution, and health concerns often derive from the environmental impacts caused from industrial activities. The need to decrease the actual massive use of fossil fuels boosts the research on alternative solutions for limiting the emissions of CO<sub>2</sub>. CO<sub>2</sub> is among the main greenhouse gases, and its contribution has been estimated to more than 60% of the total global warming. The Kyoto Protocol and international conventions intend to reduce global emissions by 50% with respect to those measured in 2006 by 2050. Consequently, the necessity of finding easy-to-use materials for CO<sub>2</sub> adsorption becomes more than urgent.

Several methods for the safe control and disposal of CO<sub>2</sub> emissions have been widely studied [1]. For example, steam reforming of hydrocarbons is the most suitable process for hydrogen production, but it releases high amounts of carbon dioxide [2–4] that must be adsorbed. An appropriate CO<sub>2</sub> adsorbent should satisfy the following criteria: a low cost, a fast kinetic, a high adsorption capacity and selectivity, and a high thermal and chemical stability for several adsorption cycles [1,5–7]. Clay minerals are potentially good adsorbents, but they are generally stable only up to 200 °C [8–10]. Due to the loss of interlayer water (dehydration), the irreversible degradation of the structure takes part at higher temperatures than 200 °C. Among all the materials tested for CO<sub>2</sub> capture and

storage, zeolites [11,12] and LDHs [7,13–15] have been widely studied thanks to their high surface area, developed pore structure, and high charge density [5]. Zeolites present high charge density and tunable pore size [16,17], while LDH-based materials represent an interesting alternative material thanks to their easy synthesis and the tunability of their chemical composition. Various preparation routes of LDH-based materials have been considered for synthesizing materials for CO<sub>2</sub> capture. In this frame, the impregnation of commercial hydrotalcites [18], the synthesis of tunable Mg/Al LDHs [14,15,19], and the formation of mixed oxides derived from their calcination [6,20–22] have been reported.

Layered Double Hydroxides are lamellar materials made of stacked octahedral sheets containing a mixture of divalent and trivalent cations, according to the following structural formula:



where  $M^{2+}$  is a divalent metal cation ( $Mg^{2+}$ ,  $Ca^{2+}$ , or  $Zn^{2+}$ , for example),  $M^{3+}$  is a trivalent metal cation ( $Al^{3+}$ ,  $Fe^{3+}$ , or  $Co^{3+}$ , for example), and  $A^{n-}$  is a compensating anion (such as  $Cl^-$ ,  $CO_3^{2-}$ , or  $NO_3^-$ ).

LDHs are generally compared to conventional clay-like crystalline structures and are named sometimes “anionic clay” due to the positive charge deficit that is compensated by anions in the interlayer space. Hydrotalcite is a particular type of LDH made of magnesium and aluminum cations with a fixed Mg/Al molar ratio equal to 3.

Regarding their widest application, CO<sub>2</sub> adsorption is strongly related to the number and the strength of surface basic sites [6,23,24]. Even if they present a lower CO<sub>2</sub> adsorption capacity, when compared to other more conventional sorbents [25–28], the high presence of water molecules can increase their CO<sub>2</sub> adsorption capacity [29]. However, these materials are significantly more active for CO<sub>2</sub> adsorption after thermal decomposition, to form basic mixed oxides [30,31]. Thanks to the presence of  $Al^{3+}$  cations in the lattice, hydrotalcite-derived materials exhibit good performances for the adsorption of CO<sub>2</sub> [32,33]. The parameters that play a role in improving the surface acidity and basicity in LDH-derived materials are the presence of compensating anions in the interlayer space [34–36], the synthesis method [37,38], and the temperature of the treatment [34]. The tunability of the surface acid/basic properties is an important parameter, depending on the target application. As an example, the removal of  $Cl^-$  by anion exchange leads to an increase in the surface basicity, as reported by Tichit et al. [31]. The authors concluded that the  $Cl^-$  anions block the basic sites of hydrotalcites. The surface basicity of these materials has been improved by adding alkali salts and by increasing the Mg/Al molar ratio [39]. In addition, the presence of transition metals in the structure of the hydrotalcites tends to increase the surface acidity of the corresponding hydrotalcite-derived mixed oxides, as reported by Pavel et al. [40]. Moreover, the possibility to be regenerated [41,42] and the adsorption reversibility [43] of the hydrotalcite materials are also important points to enhance the efficiency of these sorbents in industrial adsorption units.

Coprecipitation in mild conditions is a more recent synthesis route of choice to prepare hydrotalcites in a shorter time than that employed in the conventional coprecipitation method [35,44,45]. This coprecipitation method consists of dissolving the inorganic salts containing the divalent and trivalent metal cations in a solvent, generally ethanol or water, and increasing the solution pH by adding a basic solution. The base addition promotes the condensation reaction. This procedure allows for obtaining compounds without secondary phases in a shorter time and with lower energy consumption than the conventional coprecipitation method, but with lower crystallinity [46–48].

In the present study, a series of hydrotalcite precursors (LDHs), prepared by the mild-conditions coprecipitation method, was firstly characterized by X-ray Diffraction, solid-state <sup>27</sup>Al Nuclear Magnetic Resonance spectroscopy, X-ray fluorescence spectroscopy, BET surface area measurements, and Thermogravimetric Analyses. Due to the difficulty of having insights on the condensation mechanism, it is still not clear why differences in the acid/base properties (even if very small) can be measured. Further investigation is then strongly needed on this point. The present manuscript aims to give some interesting insight to contribute to this debate. At different

preparation pH values, the condensation reaction to form the hydrotalcites might take part following different pathways (not yet elucidated up to now) to form a final product with slightly different ratios among the Mg- and Al-containing phases. Segregation of Mg and Al hydroxide can also take over, and, even if not detectable by the techniques employed, drive the acid-basic properties of the final materials. In the present paper, differently to that reported in the most part of the already published articles, the surface basicity (the key property for CO<sub>2</sub> adsorption) of the calcined materials (mixed oxides) was probed by SO<sub>2</sub> adsorption calorimetry. In a second time, the CO<sub>2</sub> sorption capacities of the mixed oxides have been measured in order to compare their performances with similar materials reported in the literature and find correlations between the physico-chemical characteristic and the surface basicity. CO<sub>2</sub> adsorption was here carried on at the temperature of 200 °C to verify the compatibility of the prepared material with industrial processes where a high adsorption temperature is required (among others in the so-called “carbon capture and utilization process”, where CO<sub>2</sub>, once captured, is utilized as a feedstock and converted catalytically into methane or methanol).

## 2. Materials and Methods

### 2.1. Synthesis of the Layered Double Hydroxides

The synthesis of the MgAl LDHs precursors involves magnesium chloride and aluminum chloride as sources of the atoms constituting the material framework. Magnesium chloride hexahydrate MgCl<sub>2</sub>·6H<sub>2</sub>O (Sigma-Aldrich, BioXtra, Burlington, MA, USA, ≥99.0 wt%), aluminum acetylacetonate Al(C<sub>5</sub>H<sub>7</sub>O<sub>2</sub>)<sub>3</sub> (Sigma-Aldrich, ReagentPlus®, 99 wt%), aluminum chloride hexahydrate AlCl<sub>3</sub> · 6H<sub>2</sub>O (Fluka Analytical, ≥99.0 wt%), ethanol CH<sub>3</sub>CH<sub>2</sub>OH (Carlo Erba Reagents, Emmendingen, Germany, ≥97 wt%), sodium hydroxide NaOH (Carlo Erba Reagents, 99.9 wt%), and deionized water (18.2 MΩ·cm) were used.

At first, the required amounts of magnesium and aluminum precursors (to fix the Mg/Al molar ratio at 3 according to the general structural formula of the hydrotalcites: [Mg<sub>6</sub>Al<sub>2</sub>(OH)<sub>16</sub>](A<sup>n−</sup>)<sub>2</sub>·yH<sub>2</sub>O), where A<sup>n−</sup> are the Cl<sup>−</sup> and C<sub>5</sub>H<sub>7</sub>O<sub>2</sub><sup>−</sup> that were dissolved into 50 mL of ethanol under stirring. The Mg/Al molar ratios were calculated according to the amounts of each reactant and their molar masses. Then, an aqueous solution of sodium hydroxide (NaOH) 1M was added to adjust the pH at 11 for the HTCL-1 and HTCL-3 samples and 10 for the HTCL-5 sample (see Table 1). The mixture was left under stirring at room temperature for 24 h. The solid was recovered by centrifugation at 11,000 rpm for 10 min, washed three times with 30 mL of ethanol, and finally dried for 24 h in an oven at 60 °C. The samples were then labeled “HTLC-x”, where x is the number of the sample. Two different precursors (AlCl<sub>3</sub>·6H<sub>2</sub>O and Al(C<sub>5</sub>H<sub>7</sub>O<sub>2</sub>)<sub>3</sub>) were used in order to verify if the substitution of Cl<sup>−</sup> with C<sub>5</sub>H<sub>7</sub>O<sub>2</sub><sup>−</sup> would lead to an increase of the surface basicity; indeed, in Ref. [31], the authors concluded that the Cl<sup>−</sup> anions block the basic sites of hydrotalcites. Moreover, two samples obtained using AlCl<sub>3</sub>·6H<sub>2</sub>O as the Al source were prepared by adding different amounts of NaOH solution (28 mL for HTLC-1 and 16 mL for HTLC-5) in order to verify if the different conditions of condensation (at pH 11 and 10, respectively) will lead to different CO<sub>2</sub> adsorption properties of the derived oxides.

**Table 1.** Coprecipitation synthesis parameters of the MgAl LDHs precursors synthesized with magnesium chloride and two different aluminum sources, and different amounts of OH<sup>−</sup>.

Sample	Al Source	Mg/Al Ratio	NaOH Amount (mL)
HTCL-1	AlCl <sub>3</sub> ·6H <sub>2</sub> O	3.02	28
HTCL-3	Al(C <sub>5</sub> H <sub>7</sub> O <sub>2</sub> ) <sub>3</sub>	2.99	22
HTCL-5	AlCl <sub>3</sub> ·6H <sub>2</sub> O	3.00	16

Finally, the powders of hydrotalcite precursors were calcined at 450 °C for 5 h in an oven. The calcined samples (hydrotalcite-derived mixed oxides) were labeled “HTLC-x

CAL", x corresponding to the number of the sample. The resultant oxides were used for CO<sub>2</sub> adsorption in this study.

## 2.2. Physico-Chemical Characterization Techniques

The structural properties of the synthesized samples were investigated by X-ray Diffraction using a Panalytical X'Pert PRO MPD diffractometer with Cu K $\alpha$  radiation ( $\lambda = 1.5418 \text{ \AA}$ ), and performed from 2 to 70° 2 $\theta$ , with a step of 0.017° 2 $\theta$  and a time per step of 218 s, with a total time for acquisition of 1 h 15, on randomly oriented powder samples. Each reflection is associated with a distance between planes according to the Bragg's Law:  $2d_{hkl} \cdot \sin(\theta) = n \cdot \lambda$  (where  $d_{hkl}$  corresponds to the distance of the plane (hkl),  $\theta$  the diffraction angle,  $\lambda$  the K $\alpha$  radiation used (1.5418  $\text{\AA}$ ), and  $n$  the periodicity index).

Solid-state <sup>27</sup>Al MAS NMR spectroscopy spectra were obtained using a Bruker AVANCE II 400 MHz spectrometer at 104.3 MHz with magic angle spinning (MAS). The samples were packed in a 2.5 mm diameter cylindrical rotor, spun at a spinning frequency of 25 kHz, and recorded for 8 h. A short delay time of 1 s and a 4  $\mu$ s single pulse were used.

Thermogravimetric analysis (TGA) of the synthesized samples was performed using a Mettler-Toledo TGA/DSC1 LF1100 apparatus, in alumina sample holders, under argon, with a flow rate of about 100 mL·min<sup>−1</sup> from 25 to 600 °C and a heating rate of 5 °C·min<sup>−1</sup>. An empty sample holder was recorded as reference to correct the baseline deviation.

N<sub>2</sub> adsorption–desorption isotherms were obtained on a Micromeritics ASAP 2420 apparatus at −196.15 °C. The samples were degassed at 100 °C for 15 h before the measurements. The data were analyzed by means of the software MicroActive 5.02.

## 2.3. Basicity Measurement by SO<sub>2</sub> Adsorption Calorimetry

The surface basicity of the samples was determined by adsorption microcalorimetry of SO<sub>2</sub> at 150 °C. The experimental set-up was constituted of a Setaram C80 heat-flow calorimeter linked to a homemade volumetric apparatus equipped with a Barocel Capacitance manometer, for pressure measurements. Approximately 80 mg of the samples was preheated in a quartz cell by heating overnight under vacuum at 350 °C. The differential heat of adsorption was measured as a function of the coverage by repeatedly sending small doses of the gas probe onto the sample until an equilibrium pressure of 0.5 Torr was reached. The sample was then outgassed for 40 min at the same temperature, and a second adsorption run was performed at 150 °C on each sample, until an equilibrium pressure of approximately 0.2 Torr. The difference between the amount of the probe gas adsorbed during the first and the second runs represents the irreversibly adsorbed amount ( $V_{irr}$ ) of the probe gas, which provided an estimation of the number of strong basic sites.

## 2.4. CO<sub>2</sub> Adsorption Tests

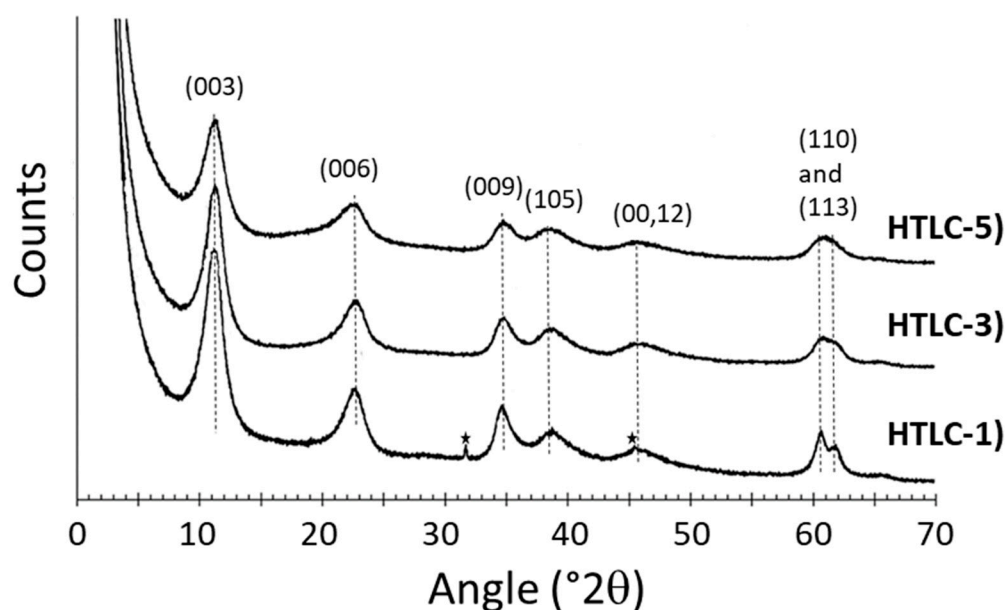
A Setaram Sensys thermogravimetry–differential scanning calorimetry (TG-DSC) apparatus has been used to measure the adsorption of CO<sub>2</sub>. Samples were pretreated in situ at 300 °C under pure N<sub>2</sub> flowing at 20 mL·min<sup>−1</sup>. Then, the temperature was decreased and stabilized to the adsorption temperature (200 °C) always under N<sub>2</sub> flow. Once the mass stabilized, the gas was switched to CO<sub>2</sub> with a constant flow of 20 mL·min<sup>−1</sup> and maintained during the CO<sub>2</sub> sorption experiments for 4 h until complete saturation of the samples. The mass losses and uptakes were measured in order to respectively evaluate the materials CO<sub>2</sub> adsorption capacity.

# 3. Results and Discussion

## 3.1. Characterization of the Hydrotalcite Precursors

Due to their 2-dimensional structure, the X-ray diffraction patterns of LDHs generally show (00 $\ell$ ) reflections that allow an estimation of the basal spacing ( $d_{003}$ ) and the c-cell parameter  $c = 3 \times d_{003}$ , corresponding to the interlayer distance plus the thickness of a single layer three times in case of rhombohedral symmetry [49]. Thus, (003) and (110) reflections are respectively related to the size of the unit-cell,  $c = 3 \times d_{003}$ , and the metal–

metal interatomic distance in a sheet,  $a = 2 \times d_{110}$ . Figure 1 displays the X-ray diffraction patterns of the precursors synthesized in this study. Seven main reflections are observed, 11.2, 22.7, 34.6, 38.7, 45.5, 60.6, and 61.8 °2θ, which correspond to distances of 7.9 ( $d_{003}$ ), 3.9 ( $d_{006}$ ), 2.6 ( $d_{009}$ ), 2.3 ( $d_{105}$ ), 2 ( $d_{00,12}$ ), 1.52 ( $d_{110}$ ), and 1.49 ( $d_{113}$ ) Å, respectively, that, according to the Bragg's law, were detected on the sample HTLC-1 that was the only sample presenting NaCl impurities (not completely removed during the post-preparation washing procedure). These positions correspond to MgAl LDHs, also called hydrotalcites, despite the broadness of the peaks related to a lower structural organization (crystallinity) due to the coprecipitation synthesis process (short synthesis time without heating). This phenomenon is observed in the samples HTLC-3 and HTLC-5 by the overlapping of the  $d_{110}$  and  $d_{113}$  planes, which is related to the reactants used.

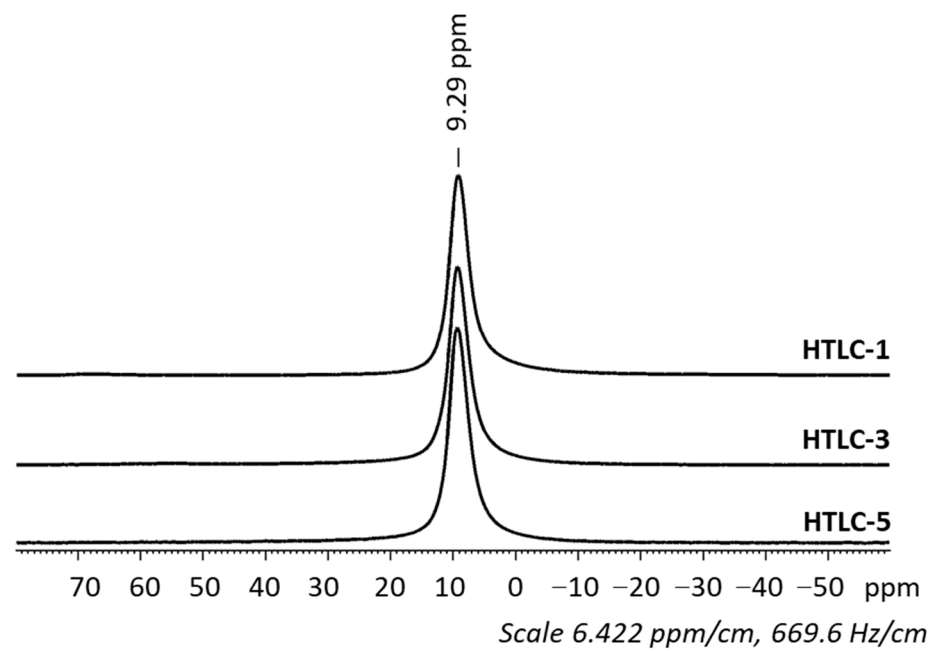


**Figure 1.** X-ray Diffractograms of the hydrotalcite samples. The star symbols correspond to NaCl impurities.

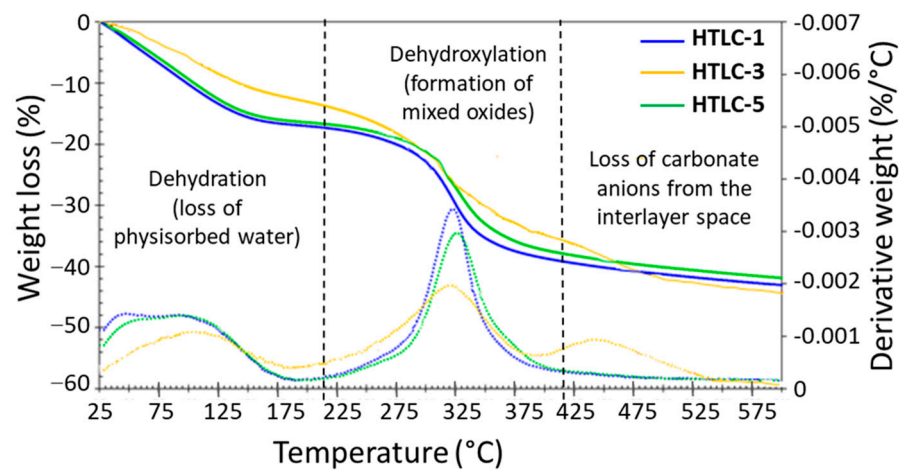
Then, solid-state  $^{27}\text{Al}$  Nuclear Magnetic Resonance (NMR) spectroscopy was performed to probe the local environment and the coordination state of  $^{27}\text{Al}$  into the synthesized LDHs. According to Figure 2, a single resonance can be observed around 9.3 ppm that corresponds to octahedral aluminum (Al VI coordination state). The configuration of aluminum in an octahedral environment is thus confirmed by the presence of this resonance, observed in all the synthesized samples.

Then, Thermogravimetric Analyses (TGA) have been performed by increasing the temperature under argon flow from 25 to 600 °C, as shown in Figure 3, in order to evaluate the thermal stability of the precursors and the evolution of the hydrotalcite into the final mixed oxides. All samples present the typical thermogravimetric (left axes) and derivative thermogravimetric profile (right axis) of hydrotalcites decomposition. The samples exhibit similar profiles with a first weight loss of about 13 to 16.4 wt% between 25 and 200 °C, and related to the dehydration of the LDHs (loss of the surface water). A second weight loss of about 25.5 to 31.3 wt% between 200 and 600 °C was assigned to the dehydroxylation of the lamellar materials to form the corresponding mixed oxides. The third weight loss, observed in the sample HTLC-3 between 390 and 550 °C, was related to the departure of  $\text{CO}_2$  (derived from  $\text{C}_5\text{H}_7\text{O}_2^-$  anions) present in the interlayer space. Compared to similar materials in the literature [50–53], Table 2 shows that the MgAl LDHs sample here prepared are as thermally stable as conventional LDHs that are mainly synthesized by methods with longer synthesis durations.





**Figure 2.** Solid-state  $^{27}\text{Al}$  Nuclear Magnetic Resonance spectra of the non-calcined samples.



**Figure 3.** Thermogravimetric analyses of the hydrotalcite samples, recorded from 25 to 600 °C. The weight losses are shown in solid lines and their derivatives in dotted lines.

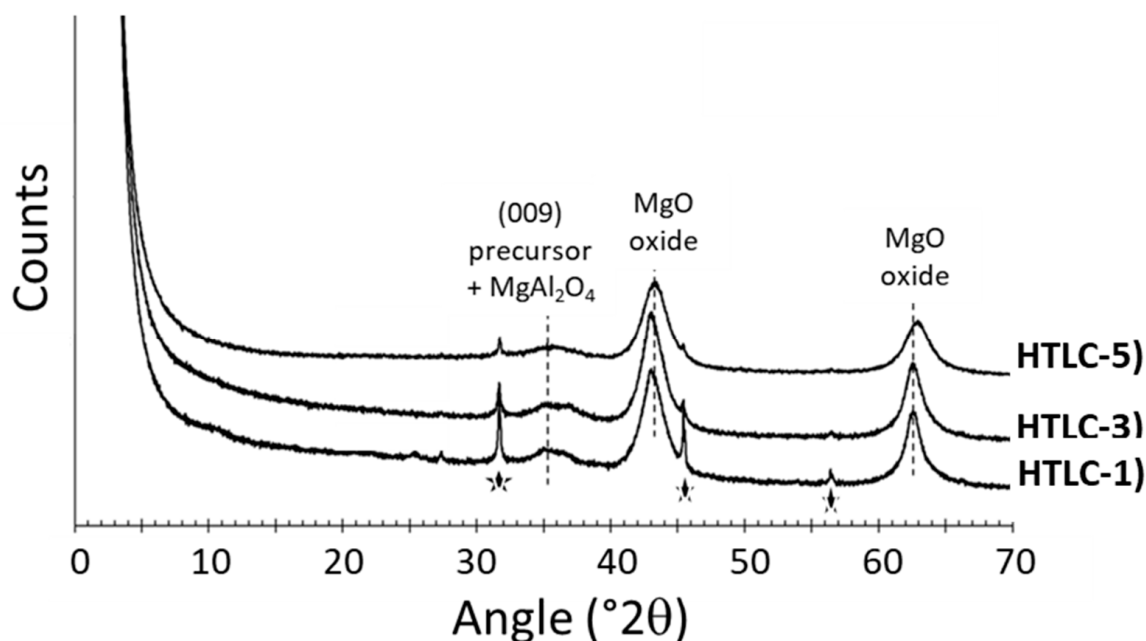
**Table 2.** TGA analyses and the weight losses of hydrotalcite precursors compared to similar materials reported in the literature.

Sample	Weight Loss 25–200 °C (%)	Weight Loss 200–600 °C (%)	Total Weight Loss (%)	Comments	Ref.
HTCL-1	17.02	26.04	43.06	-	This article
HTCL-3	13.03	31.32	44.36	-	This article
HTCL-5	16.44	25.46	41.90	-	This article
LDH1	12.7	29.3	42.00	Commercial MgAl LDH	[50]
LDH2	15.2	26.7	41.90	Commercial MgAl LDH	[50]
LDH_amm	15	38.83	53.83	Coprecipitation with ammonia	[51]
LDH_70	15	29	44.00	Coprecipitation at low supersaturation, then pretreatment at 70 °C	[52]
LDH_140	10	30	40.00	Coprecipitation at low supersaturation, then pretreatment at 140 °C	[52]
Mg/Al_LDH	24	23	47.00	Coprecipitation at low supersaturation	[53]

### 3.2. Characterization of the Hydrotalcite-Derived Mixed Oxides

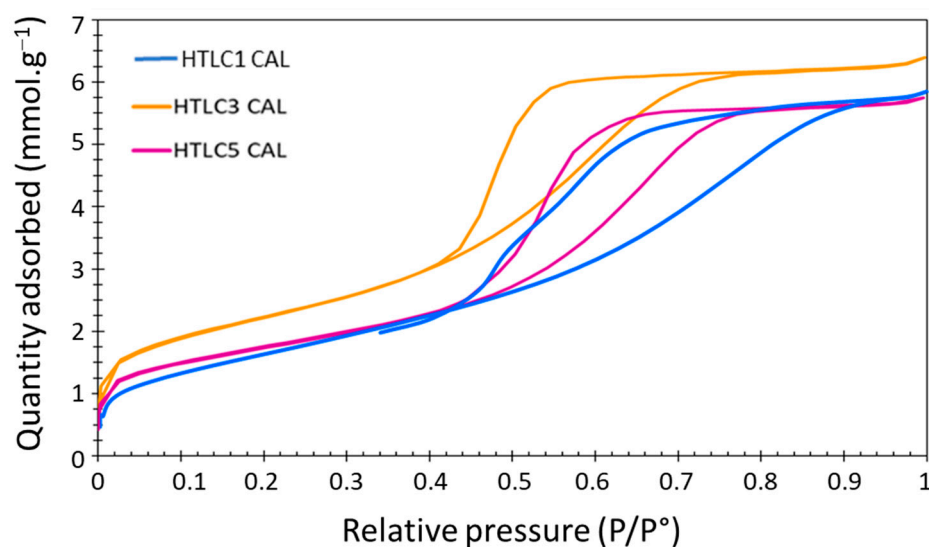
The mixed oxides obtained after calcination of the hydrotalcite precursor at 450 °C for 5 h under air were also characterized.

The structural modifications due to the calcination process were evidenced by XRD, reported in Figure 4. Despite the presence of NaCl impurities observed in the precursors, the layered structure of the hydrotalcite precursor clearly collapsed after calcination. An amorphous phase characterized by broad peaks was observed. The peaks centered at 35.1, 37.2, 43.4, and 63° 2θ corresponded respectively to the precursor (35.1° 2θ),  $\text{MgAl}_2\text{O}_4$  oxide (37.2° 2θ), and MgO oxide (43.4 and 63° 2θ). These structural changes are associated to the dehydroxylation of the materials observed by TGA (between 200 and 600 °C) and the formation of mixed oxides.



**Figure 4.** X-ray Diffractograms of the hydrotalcite-derived mixed oxides. The star symbols correspond to NaCl impurities.

Then, the nitrogen adsorption/desorption isotherms were collected to determine the specific surface area and the microstructure of the calcined hydrotalcite samples (see Figure 5). Despite the different Al-sources used for the synthesis of the hydrotalcite precursors (respectively, acetylacetonate and chloride for HTLC-3 and HTLC-1), all calcined samples presented isotherms of type IVa and hysteresis loops of type H2b, characteristic of mesoporous adsorbents, according to the IUPAC classification [54]. Such a type of hysteresis is characteristic of a complex and interconnected pore structure. The specific surface areas have been calculated by the BET method. The samples showed high surface areas, respectively, of 138, 180, and 139  $\text{m}^2\cdot\text{g}^{-1}$  for the HTCL-1 CAL, HTCL-3 CAL, and HTCL-5 CAL, with medium pore volumes in the 0.20–0.22  $\text{mmol}\cdot\text{g}^{-1}$  range in all the calcined samples. These results are in agreement with the ones reported in the literature for similar mixed oxides [55–57]. The samples prepared with the precursor containing the acetylacetonate anion lead to a material with a slightly higher surface area. On the other hand, the amount of  $\text{OH}^-$  added had little effect on the surface area of the calcined samples, despite the differences observed for non-calcined samples in terms of crystallinity.

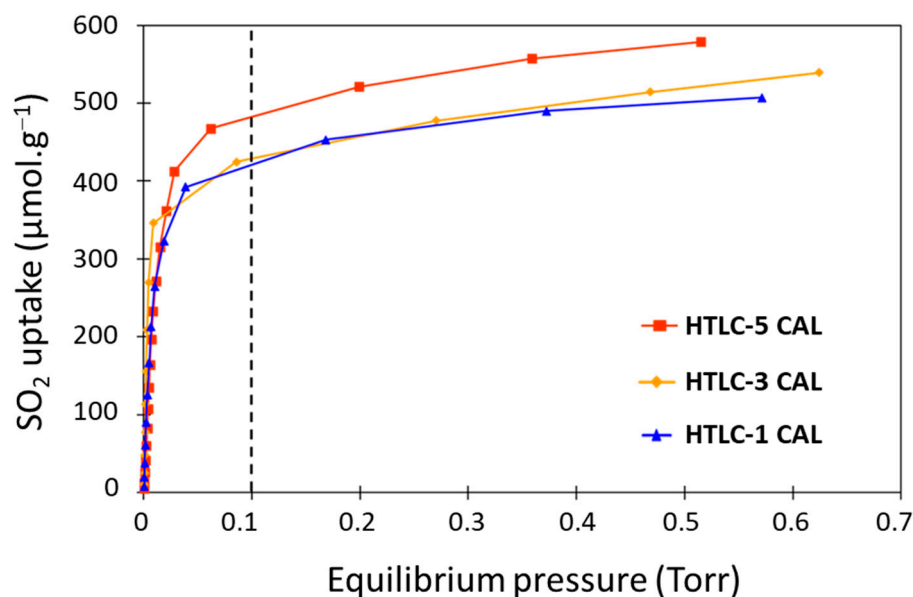


**Figure 5.** N<sub>2</sub>-adsorption isotherms for the three investigated samples.

### 3.2.1. SO<sub>2</sub> Adsorption Calorimetry

Adsorption calorimetry is the most adapted technique for determining the concentration (from the adsorption volumetric isotherms), strength, and strength distribution (using the differential heat of adsorption as a function of coverage) of the basic sites [58] involved in CO<sub>2</sub> adsorption. Then, the hydrotalcite-derived mixed oxides surface basicity has been probed by adsorption calorimetry of SO<sub>2</sub>.

Figure 6 exhibits the SO<sub>2</sub> adsorption isotherms of the calcined hydrotalcite samples HTLC-1 CAL, HTLC-3 CAL, and HTLC-5 CAL. Some differences can be observed between the samples: HTLC-5 CAL has a higher SO<sub>2</sub> uptake (up to 578  $\mu\text{mol}\cdot\text{g}^{-1}$  at 0.5 Torr), which is associated to a higher basicity than the other samples. After a fast uptake at low equilibrium pressure ( $p < 0.1$  Torr) due to the adsorption on the strongest sites, the SO<sub>2</sub> uptake reaches a plateau around 550  $\mu\text{mol}\cdot\text{g}^{-1}$ , even increasing the equilibrium pressure. Despite the difference in terms of specific surface area of the samples HTLC-1 CAL and HTLC-3 CAL (138.3 and 174.8  $\text{m}^2\cdot\text{g}^{-1}$ , respectively), their SO<sub>2</sub> adsorption isotherms are quite similar.



**Figure 6.** SO<sub>2</sub> adsorption isotherms of the calcined hydrotalcite samples.



In Figure 7, the SO<sub>2</sub> differential heat curve shows a pseudo plateau placed at around 140 kJ·mol<sup>−1</sup>, indicating the homogeneity in the strength of the basic sites present in the mixed oxide. This is not a common result for such kinds of samples, which generally present a more heterogeneous strength site distribution [59,60]. For more insights: at low coverage, the curves present a few points characterized by a relatively high adsorption heat (between 150 and 180 kJ·mol<sup>−1</sup>, which can be connected to the presence of Lewis adsorption sites (very strong sites)). At a coverage higher than 50 μmol·g<sup>−1</sup>, a plateau (that ends up, respectively, around 350, 400, and 450 μmol·g<sup>−1</sup> for HTLC-3 CAL, HTLC-1, and HTLC-5) can be ascribed to SO<sub>2</sub> adsorption on relatively strong sites and characterized by an almost constant heat. The sites represented by this plateau correspond to homogeneous Brønsted sites (most probably −OH groups). After the plateau, a region characterized by steep decreasing in the differential heat curve can be observed and assigned to the presence of a small number of heterogeneous sites (probably of Lewis type) [60]. The last part of the curve corresponds to the reversible adsorption domain (physisorption of the SO<sub>2</sub> probe) or to very weak Lewis acid sites.

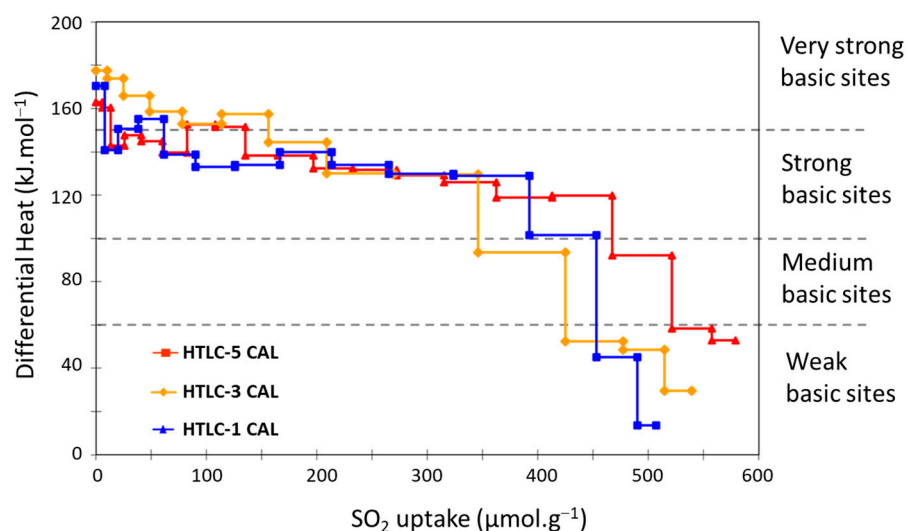


Figure 7. Differential heats of SO<sub>2</sub> adsorption of the calcined hydrotalcite samples.

More in detail, the basic sites can be divided into four groups according to their strength: very strong (above 150 kJ·mol<sup>−1</sup>), strong (in the range 100–150 kJ·mol<sup>−1</sup>), medium (in the range 60–100 kJ·mol<sup>−1</sup>), and weak (below 60 kJ·mol<sup>−1</sup>). The distribution is reported in Figure 8.

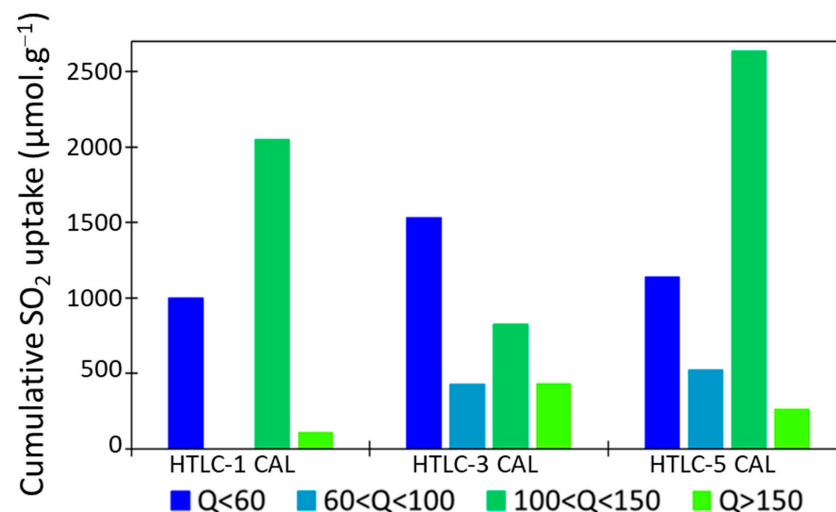


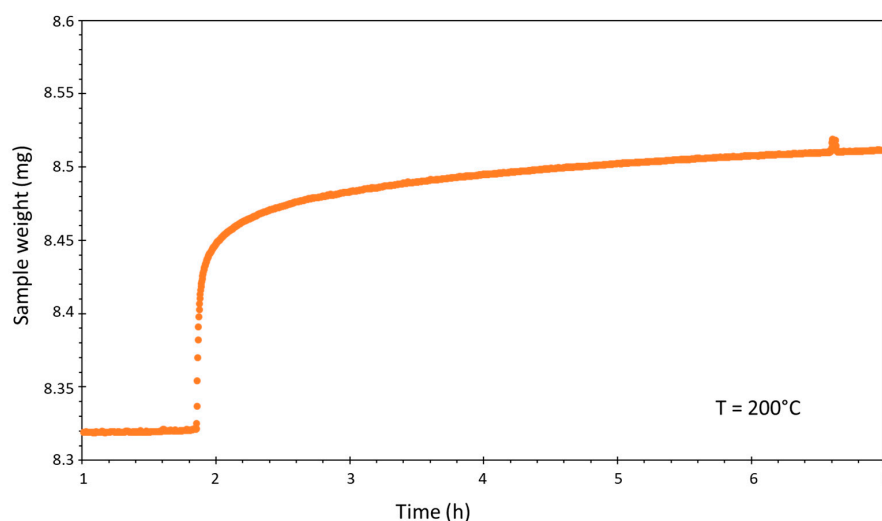
Figure 8. Probe uptake of the calcined samples depending on the range of energy.

Some differences can be observed among the samples. The sample HTLC-5 CAL, for example, presents the highest SO<sub>2</sub> adsorption and the highest ratio of strong basic sites. As a general observation, the three samples present a low proportion of very strong sites ( $Q > 150 \text{ kJ}\cdot\text{mol}^{-1}$ ). This is a positive feature in view of the application of these materials as CO<sub>2</sub> adsorbents, because the strong site can lead to the formation of surface carbonates that react in an irreversible way with CO<sub>2</sub>.

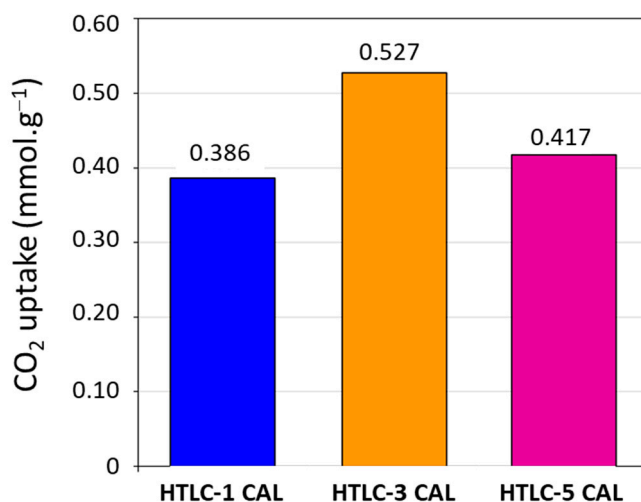
The amount of SO<sub>2</sub> adsorbed is not directly related to the specific surface area of the materials: the sample HTLC-3 CAL that exhibits the highest surface area ( $179.9 \text{ m}^2\cdot\text{g}^{-1}$ ) presents an intermediate total SO<sub>2</sub> adsorption value of  $539.2 \text{ }\mu\text{mol}\cdot\text{g}^{-1}$  at 0.5 torr. The highest SO<sub>2</sub> sorption value is observed in the sample HTLC-5 CAL that exhibits a lower surface area ( $139.4 \text{ m}^2\cdot\text{g}^{-1}$ ).

### 3.2.2. CO<sub>2</sub> Adsorption Tests

Finally, a screening of the adsorption capacity of CO<sub>2</sub> has been performed on the calcined samples. As an example, the CO<sub>2</sub> uptake curve (performed at 200 °C) and the pretreatment and adsorption temperature program are shown in Figure 9 for the sample HTLC-3 CAL. The results are shown in Figure 10, which shows that the studied samples exhibit CO<sub>2</sub> adsorption capacities between 0.38 and 0.53 mmol·g<sup>−1</sup>. The reproducibility of the experiments was verified and the error on the measurements estimated at  $\pm 0.002 \text{ mmol}\cdot\text{g}^{-1}$ .



**Figure 9.** CO<sub>2</sub> uptake curve for the sample HTLC-3 CAL, reported as an example.



**Figure 10.** Quantity of CO<sub>2</sub> adsorbed by the calcined hydrotalcite samples (HTCL-1 CAL in blue, HTCL-3 CAL in orange, and HTCL-5 in fuchsia).

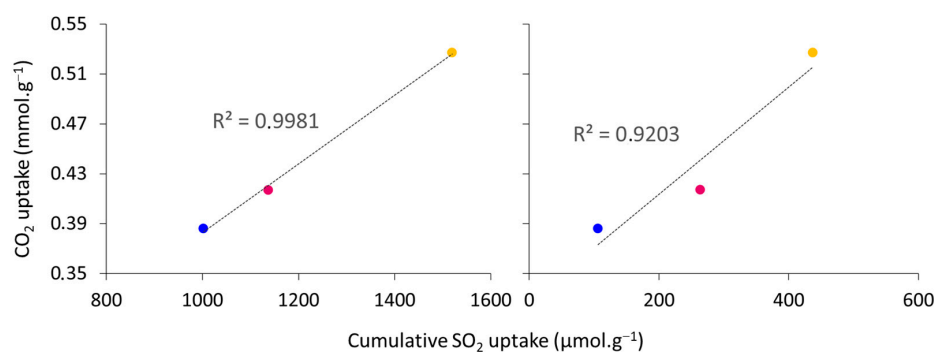
Table 3 compares these adsorption values with similar materials found in the literature, and proves that the samples here synthesized in mild conditions are as efficient (under similar analyses conditions) as other hydrotalcites synthesized by other methods for the adsorption of CO<sub>2</sub>. They present adsorption capacities around 0.5 mmol·g<sup>−1</sup>, without the need of additional treatments or structural modifications.

**Table 3.** Comparisons of the BET surface areas and CO<sub>2</sub> adsorption capacities between the calcined hydrotalcite samples and similar materials reported in the literature.

Sample	Mg/Al Molar Ratio	T <sub>pretreatment</sub> (°C)	T <sub>adsorption</sub> (°C)	SSA (m <sup>2</sup> ·g <sup>−1</sup> )	CO <sub>2</sub> Adsorption Capacity (mmol·g <sup>−1</sup> )	Ref.
HTLC-1 CAL	3	300	200	138	0.386	This work
HTLC-3 CAL	3	300	200	180	0.527	This work
HTLC-5 CAL	3	300	200	139	0.417	This work
HTLC_Low	2	500	450	63	0.270	[61]
HTLC_High	2	500	450	154	0.260	[61]
LDO200	3	400	200	167	0.486	[62]
Htlc-200	3	200	200	66	0.604	[63]
Htlc-400	3	400	200	184	0.896	[63]
Htlc	3	-	300	74	0.5	[64]

Sample HTLC-3 CAL, which exhibits a large amount of weak basic sites (mainly below 60 kJ·mol<sup>−1</sup>) and the highest specific surface area (179.9 m<sup>2</sup>·g<sup>−1</sup>), shows the highest CO<sub>2</sub> adsorption capacity (0.527 mmol·g<sup>−1</sup>).

The interaction of CO<sub>2</sub> with the basic sites of the material depends on their strength. Typically, low-energy basic sites ( $Q < 60$  kJ·mol<sup>−1</sup>), measured by SO<sub>2</sub> adsorption calorimetry, correspond to weak physisorption sites and are related to reversible sorption. On the other hand, high-energy basic sites ( $Q > 150$  kJ·mol<sup>−1</sup>) are related to strong chemisorption. Thus, low-energy sites remain accessible after thermal regeneration, which is important for the reusability of the materials. High-energy sites cannot be regenerated and are irreversibly poisoned by acid/amphoteric probes, such as CO<sub>2</sub>. In the present sample series, a very good linear correlation between the concentration of low- ( $Q < 60$  kJ·mol<sup>−1</sup>) and very high- ( $Q > 150$  kJ·mol<sup>−1</sup>) energy sites and the CO<sub>2</sub> adsorption capacity can be noticed. This observation is illustrated in Figure 11: the higher the concentration of these sites, the higher the sorption capacity of the material. This demonstrates that the sorption capacity of these hydrotalcites is strongly influenced by the type of sites. Even if the correlation with the amount of weak and very strong adsorption sites is linear, very strong sites can irreversibly adsorb CO<sub>2</sub>. Then, the reusability of the sorbents might be maximized when the sample presents a high number of weak adsorption sites and a limited amount of strong energy sites. This point will need further investigation prior to the application of such materials in real applications.



**Figure 11.** Linear correlations between the concentrations of weak (below 60 kJ·mol<sup>−1</sup>) and very strong (above 150 kJ·mol<sup>−1</sup>) energy sites and the CO<sub>2</sub> adsorption capacities of the calcined samples (HTLC-1 CAL in blue, HTLC-3 CAL in orange, and HTLC-5 in fuchsia).

#### 4. Conclusions

Layered Double Hydroxides, especially hydrotalcites, are very interesting adsorbents for carbon dioxide thanks to their lamellar structure and strong basicity. In this work, hydrotalcites with a Mg/Al molar ratio of 3 have been successfully synthesized by coprecipitation in mild conditions and characterized by complementary physico-chemical techniques. The influence of the pH and of the Al source has been investigated. Their thermal stability has been evaluated by DTA/TGA, showing a major structural change around 325 °C, due to the formation of mixed oxides. These oxides show a high specific surface area (up to 180 m<sup>2</sup>·g<sup>−1</sup> for the sample prepared starting from MgCl<sub>2</sub>·6H<sub>2</sub>O and Al(C<sub>5</sub>H<sub>7</sub>O<sub>2</sub>)<sub>3</sub>), as well as a surface basicity (high quantity of medium and weak basic sites) adapted to the reversible adsorption of CO<sub>2</sub>. The calcined hydrotalcites exhibit CO<sub>2</sub> adsorption capacities of around 0.5 mmol·g<sup>−1</sup>, a value similar to those reported in the literature for samples generally synthesized by conventional coprecipitation methods. The use of the Al(C<sub>5</sub>H<sub>7</sub>O<sub>2</sub>)<sub>3</sub> gave rise to the sample with the highest specific surface area and containing the optimum quantity of the basic site of adapted energy (measured by adsorption microcalorimetry) for CO<sub>2</sub> adsorption.

On the other hand, the modification of the pH of condensation during synthesis did not impact the main microstructural properties of the samples that presented the same specific surface area and pore volume for the synthesis performed at pH = 10 and 11. Only the surface basicity was modified; the calcined sample derived from the hydrotalcite prepared at pH 10 presented a slightly higher surface basicity and, in particular, a prominence of the basic site in the 100–150 kJ·mol<sup>−1</sup> range, and associated with a higher CO<sub>2</sub> adsorption capacity, when compared to the sample prepared at pH = 11. This result can be explained by the fact that at pH = 11, a small portion of the aluminum precursor could form Al(OH)<sub>4</sub><sup>−</sup> and not be integrated in the condensation reaction, this resulting in the enrichment of the sample in Mg, known for its basicity. This hypothesis could not be proved by the detection of Al<sub>2</sub>O<sub>3</sub> in the calcined sample (HTCL-5 CAL), but probably the very low quantity of alumina and its presence in an amorphous or highly dispersed form can prevent its detection by XRD analysis.

This work demonstrates that the coprecipitation method in mild conditions leads to the formation of hydrotalcites with similar (sometimes higher) performances than conventional LDHs prepared by time- and energy-consuming methods. The linear correlations between the surface basicity (concentration of low- and high-energy sites) of these lamellar materials with the adsorption capacity shows their possibility to be reused several times in the same application. The possibility to tune the Mg/Al molar ratio is a promising feature to obtain a wide range of LDHs (with tuned surface basicity) that might reveal to be more efficient towards the adsorption of CO<sub>2</sub> in the future.

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