

Review

Carbon-Based Nanocomposites as Fenton-Like Catalysts in Wastewater Treatment Applications: A Review

Ling Xin ¹, Jiwei Hu ^{2,3,*}, Yiqiu Xiang ¹, Caifang Li ¹, Liya Fu ³, Qiuhua Li ^{1,4} and Xionghui Wei ⁵

- ¹ Guizhou Provincial Key Laboratory for Information Systems of Mountainous Areas and Protection of Ecological Environment, Guizhou Normal University, Guiyang 550001, China; xinling901231@163.com (L.X.); xxiangyiqiu@163.com (Y.X.); lcflcaifang@163.com (C.L.); qiuhua2002@126.com (Q.L.)
 - ² Cultivation Base of Guizhou National Key Laboratory of Mountainous Karst Eco-Environment, Institute of Karst, Guizhou Normal University, Guiyang 550001, China
 - ³ Research Center of Water Pollution Control Technology, Chinese Research Academy of Environmental Sciences, Chaoyang District, Beijing 100012, China; fuliya1986@126.com
 - ⁴ Guizhou International Science and Technology Cooperation Base-International Joint Research Centre for Aquatic Ecology, Guizhou Normal University, Guiyang 550001, China
 - ⁵ Department of Applied Chemistry, College of Chemistry and Molecular Engineering, Peking University, Beijing 100871, China; xhwei@pku.edu.cn
- * Correspondence: jiwei.hu@yahoo.com or jwhu@gznu.edu.cn; Tel.: +86-851-8670-2710

Citation: Xin, L.; Hu, J.; Xiang, Y.; Li, C.; Fu, L.; Li, Q.; Wei, X. Carbon-Based Nanocomposites as Fenton-Like Catalysts in Wastewater Treatment Applications: A Review. *Materials* **2021**, *14*, 2643. <https://doi.org/10.3390/ma14102643>

Academic Editors: Vlassios Likodimos, Xiangyang Ma and Andres Sotelo

Received: 28 April 2021
Accepted: 16 May 2021
Published: 18 May 2021

Publisher's Note: MDPI stays neutral with regard to jurisdictional claims in published maps and institutional affiliations.



Copyright: © 2021 by the authors. Licensee MDPI, Basel, Switzerland. This article is an open access article distributed under the terms and conditions of the Creative Commons Attribution (CC BY) license (<http://creativecommons.org/licenses/by/4.0/>).

Abstract: Advanced oxidation (e.g., fenton-like reagent oxidation and ozone oxidation) is a highly important technology that uses strong oxidizing free radicals to degrade organic pollutants and mineralize them. The fenton-like reactions have the characteristics of low cost, simple operation, thorough reaction and no secondary pollution. Fenton-like reagents refer to a strong oxidation system composed of transition metal ions (e.g., Fe^{3+} , Mn^{2+} and Ag^+) and oxidants (hydrogen peroxide, potassium persulfate, sodium persulfate, etc). Graphene and carbon nanotube possess a distinctive mechanical strength, flexibility, electrical and thermal conductivity and a very large specific surface area, which can work as an excellent carrier to disperse the catalyst and prevent its agglomeration. Fullerene can synergize with iron-based materials to promote the reaction of hydroxyl groups with organic pollutants and enhance the catalytic effect. Fenton-like catalysts influence the catalytic behavior by inducing electron transfer under strong interactions with the support. Due to the short lifespan of free radicals, the treatment effect is usually enhanced with the assistance of external conditions (ultraviolet and electric fields) to expand the application of fenton-like catalysts in water treatment. There are mainly light-fenton, electro-fenton and photoelectric-fenton methods. Fenton-like catalysts can be prepared by hydrothermal method, impregnation and coordination-precipitation approaches. The structures and properties of the catalysts are characterized by a variety of techniques, such as high-resolution transmission electron microscopy, high-angle annular dark-field scanning transmission electron microscopy and X-ray absorption near-edge structure spectroscopy. In this paper, we review the mechanisms, preparation methods, characterizations and applications status of fenton-like reagents in industrial wastewater treatment, and summarize the recycling of these catalysts and describe prospects for their future research directions.

Keywords: hydrogen peroxide; degradation; fenton process; organic pollutants; free radicals

1. Introduction

Water is the source of life, and water quality has a direct impact on our survival and development. Today's economy has been developing rapidly and people's living standards have been improving. However, environmental problems are becoming gradually prominent, and there are increasing types of industrial and agricultural wastewater as well as urban sewage [1,2]. Water pollution not only aggravates the increasingly serious

problem of water shortage, but also poses a great threat to human health [3–5]. The main sources of water pollution are: physical, chemical and biological pollution [6,7]. Among them, chemical pollution is highly prevalent in our society, mainly referring to industrial wastewater containing various toxic contaminants (such as dyes, phenolic compounds and polycyclic aromatic hydrocarbons) and urban sewage containing pharmaceutical and personal care products, and agricultural wastewater contaminated with pesticides, insecticides and herbicides, etc., which are released without treatment [8–10]. Pharmaceutical and personal care products (PPCPs) are emerging pollutants that have a profound and irreversible influence on human health despite their low concentrations in the environment. Polycyclic aromatic hydrocarbons (PAHs) exist frequently in the form of complex mixtures in water and have strong bioaccumulation and persistence, and have been shown in animal studies to cause a carcinogenic effect, thus belonging to the priority pollutants. Dyes are colored organic compounds, which are resistant to photolysis, oxidation and biodegradation, and make them extremely difficult to handle with general water treatment systems. Hence, a deep treatment is necessary before discharging wastewater to make effluent emissions standards more stringent and to achieve the objectives of improving water quality and protecting the ecological environment.

At present, relatively mature wastewater treatment methods are mainly divided into biological technology, physical approaches, chemical processes and advanced oxidation methods [11–13]. Among them, ozone oxidation is an advanced oxidation method where hydroxyl radicals can be generated by catalytic activation of O_3 through catalysts (e.g., Fe^{2+} , Mn^{2+} , Ni^{2+} , Cu^{2+} , Ag^+ and Zn^{2+}) or solid catalysts (metal oxides, porous materials and their composites) [14–16]. Chen et al. [17] examined the operation of O_3/H_2O_2 catalytic system for the treatment of semi-coking wastewater and found its excellent degradation performance, easy magnetic separation and high stability. This proved that this system provided a theoretical basis for the deep treatment of coal chemical wastewater by ozone oxidation and had good practical application prospects and values. Besides, O_3 catalytic oxidation technology is simple to implement and easy to manage, which can achieve the zero emissions [18].

The fenton oxidation reaction is simple to operate and can better oxidize and degrade refractory organics in the wastewater, but there are problems such as insufficient mineralization of the organics, secondary pollution caused by a large amount of iron sludge and low utilization of H_2O_2 [19]. To solve these problems, many fenton-like techniques have been developed based on the conventional fenton method, such as light-fenton, electrical-fenton, and ultrasonic-fenton methods [20]. Fenton-like reagents are composed of oxidants and transition metal ions (Fe^{3+} , Cu^{2+} , Mn^{2+} , Ni^{2+} , etc.), iron-containing minerals, and metal-filled polymers [20,21]. These have certain properties of metals as well as polymers, and can also be made to have special properties such as wear resistance, electrical conductivity, corrosion resistance, interference resistance and radiation protection, which have great potential for further development. Bello et al. [22] summarized the fenton-like reagents, mainly introducing the characteristics of conventional fenton reagents while further illustrating the characteristics and applications of three fenton reagents: heterogeneous fenton, fluidized bed fenton and bioelectro-fenton. Alok et al. [23] exhibited various metal species that decompose hydrogen peroxide in fenton-like reactions, paying particular attention to their redox chemistry, practical advantages, describing in details the complex mechanisms, and highlighting the limitations for their environmental applications. Su et al. [24] used Cu-1 IHSs as fenton-like reagents to adsorb and decompose congo red from aqueous solutions based on the unique porous properties of spheres, which demonstrated their potential applications in water treatment. The above mentioned articles indicate the importance of fenton-like catalysts in wastewater treatment. The present review not only reviews the removal of pollutants from wastewater systems by fenton-like reagents, but also explores the thermodynamic as well as kinetics of fenton-like catalyst removal processes.

Carbon-based nanomaterials, which are made of carbon elements compounded with other nanomaterials, have structural units ranging in size from 1 to 100 nm. Carbon nanomaterials have received a lot of attention for their excellent electrical conductivity, mechanical properties, chemical stability and optical properties [25]. Madima et al. [26] showed the mechanisms and applications of carbon-based nanomaterials in adsorptive and photocatalytic removal of organic and inorganic pollutants from wastewater. The potential applications of carbon-based nanomaterials and their derivatives as adsorbents in wastewater remediation were also highlighted. In their review article, Selvaraj et al. [27] presents an assessment of the available literature on adsorption, catalytic degradation and membrane technologies of water treatment relating to carbon nanotube, GO and rGO. The challenges, limitations and future aspects of their feasibility are also discussed for large-scale applications in the water and wastewater treatment industry. However, earlier papers had focused on the desirability of single carbon nanomaterials for wastewater treatment systems. Therefore, this review concentrates on the use of carbon-based nanomaterials as carriers of fenton-like catalysts to investigate whether there is a synergistic effect when the two are combined to improve the efficiency of wastewater treatment.

As shown above, carbon-based materials and fenton-like catalysts have many unique properties and play an irreplaceable role in wastewater treatment processes; therefore, the combination of carbon-based materials and fenton-like catalysts is the focus of this review. This review aims to (1) introduce carbon-based materials such as fullerene, carbon nanotubes and graphene as carriers combined with fenton-like catalysts; (2) present an analysis of the mechanism, preparation and characterization methods of light-fenton, electro-fenton, ultrasound-fenton and microwave-fenton reactions; (3) summarize the effects of carbon-based nano-fenton catalysts on dye wastewater and pharmaceutical wastewater; (4) exhibit the kinetics (pseudo-first-order kinetics, pseudo-second-order kinetics) and thermodynamics (Langmuir, Freundlich) to study the processes of removing pollutants as well as the properties and reusability of fenton-like reagents in wastewater treatment.

2. Carbon-Based Nanomaterials

2.1. Classification of Carbon Nanomaterials

Nanomaterials refer to materials with a nanoscale greater than 1 nm and less than 100 nm, which have free surfaces and strong or weak interactions between each nano-unit [28]. Due to the particularity of this structure, nanomaterials have some unique effects, including small size effects, quantum effects, surface and interfacial effects [29]. Carbon elements constitute an abundance of carbon materials due to their unique hybridization forms of sp , sp^2 and sp^3 [30]. Carbon nanomaterials can be used as catalyst carriers and they can be classified into three categories based on their structures. Those with grain sizes in both directions in the nanometer range are called one-dimensional nanomaterials, such as nanotubes and nanowires. Two-dimensional materials refer to materials where electrons can move freely (planar motion) in only two dimensions on the nanoscale (1–100 nm), for example, graphene. Those materials that have nanoscale scales in all three dimensions are zero-dimensional nanomaterials (e.g., fullerenes and carbon quantum dots) [31–33].

2.2. Application of Carbon Nanomaterials

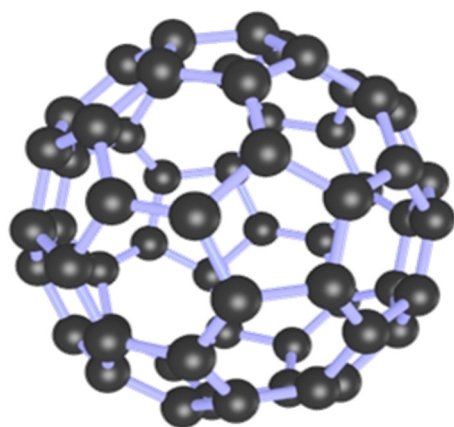
Carbon nanomaterials can inhibit impurities and defects generated by their variable structure and strong interactions between tubes, layers and sheets [34–36]. They exhibit properties in catalytic, mechanical and physical functions (optical, electrical, magnetic and sensitive) that conventional materials do not possess, thus having been applied broadly in the field of nanoscale reinforced fiber composites (Table 1) [37,38].

Table 1. Application fields of carbon-based nanomaterials.

Performance	Applications	References
Chemical catalysis	Catalytic catalyst medium	[39,40]
Mechanical properties	The preferred material for high-strength such as reinforcement and toughening	[41]
Magnetic properties	Magnetic absorption, storage memory element materials, soft ferrite materials, etc.	[42]
Electrical performance	Microelectronic device materials, microelectronic device materials, field emission cathode materials	[43]
Optical performance	Large-capacity supercapacitor materials, superconducting materials, nano-integrated circuit materials	[44]
Mechanism performance	Light absorbing materials, optical communication materials, optical recording, optical display, optoelectronic materials	[45]
Thermal properties	Micro-mechanical component materials such as molecular coils and pistons, damping devices and rotary sealing materials	[46]
Physical properties	Micro weapon materials such as micro engines, micro spy vehicles, micro high-efficiency explosives, and materials for aviation and spacecraft	[47]
Sensitive characteristics	Hydrogen storage materials, metal nanowire template materials	[48]
Other	Sensitive materials (sensors, detectors, sensitive electronic scales)	[49]

2.3. The role of Carbon Nanomaterials in Fenton-Like Reagents

Fullerene (Figure 1), also known as soccerene, is an allotrope of carbon besides diamond and graphite, and it has a unique three-dimensional π -electron delocalization structure with strong electron capture and electron transferability under light radiation [50]. Zou et al. [51] doped iron oxide into fullerene to study the degradation of RhB, MO and phenol by introducing visible light in the presence of oxides and found increased catalytic activity, wider pH range and the stability was better than that of loaded fullerene. Meng et al. [52] prepared Fe-fullerene/TiO₂ with fullerene and titanium (IV) n-butoxide to research its photocatalytic performance for decolorization of methylene blue solution under visible light. The results showed that the fullerene-enhanced TiO₂ doped with Fe improved the decolorization of methylene blue. Xu et al. [53] have fabricated PHF/hydrous ferrite (PHF/Fh) composites and investigated their photo-fenton activity in simulated sunlight irradiation degradation of acidic red 18. The results revealed that the catalytic activity of PHF/Fh was higher than that of Fh, and the appropriate content of PHF/Fh content could not only improve the stability of Fh, but also increase the production of O₂ and \cdot OH.

**Figure 1.** 3D structure diagram of fullerene.

Carbon atoms in carbon nanotubes (Figure 2) are dominated by sp^2 hybridization, while the hexagonal lattice structure has a certain degree of bending, forming a spatial topology. Among carbon atoms, a certain sp^3 hybrid bond can be formed, that is, the chemical bond formed has both sp^2 and sp^3 hybrid states, and these p orbitals overlap each other to form a highly delocalized large π bond outside the carbon nanotube graphene sheet [54]. The large π bond on the outer surface of carbon nanotubes is the chemical basis for the non-covalent binding between carbon nanotubes and some macromolecules with conjugate properties. Yao et al. [25] prepared magnetic metal M (M = Fe, Co, Ni) nanocrystals encapsulated in nitrogen-doped carbon nanotubes using dicyandiamide as a C/N precursor, which was used to degrade dyes. The nanocrystals were found to have high catalytic properties and stability with good reusability. Arshadi et al. [55] degraded methyl orange (MO) by using a non-homogeneous class of fenton reagents immobilized on aluminosilicate nanoparticles and multi-walled carbon nanotubes (Si/Al@Fe/MWCNT) with ferrocene groups. The results demonstrated that the fenton reagents had a good degradation effect on MO and this material had good stability. Samadi et al. [56] prepared magnetic Fe_3O_4 /MWCNTs composites and then investigated their role as non-homogeneous fenton catalysts in the removal of ciprofloxacin from synthetic wastewater.

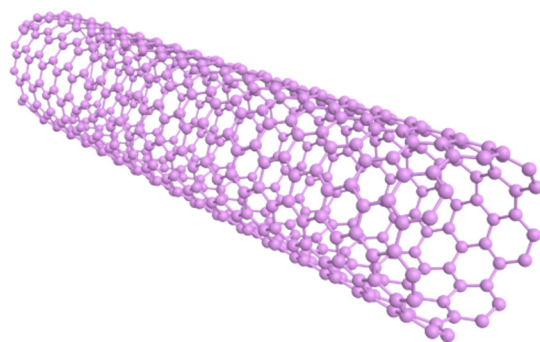


Figure 2. 3D structure diagram of carbon nanotubes.

Graphene (Figure 3) is currently the thinnest but toughest conductive nanomaterial. Graphene also shows a large nonlinear anti-magnetism, even greater than that of graphite [57]. Zubir et al. [58] prepared GO/ Fe_3O_4 fenton-like catalysts for degradation of acid orange 7 (AO7), utilizing GO in transferring electrons to maintain active Fe^{2+} to ensure the decomposition of H_2O_2 into $\cdot OH$ for the degradation of the dye. Zubir et al. [59] also prepared GO and zinc partially substituted magnetite $GO-Fe_{1-x}Zn_xO_y$ ($0 \leq x \leq 0.285$) catalysts for degradation of acid orange 7 (AO7). The study revealed that GO enhanced the ability to combine zinc with metal oxide, while zinc limited the crystal growth and formed smaller microcrystal sizes. Liu et al. [60] added GO to ferrite nickel ($NiFe_2O_4$) to obtain GO-doped $NiFe_2O_4$ ($GO-NiFe_2O_4$) for the catalytic degradation of organic dyes under visible light irradiation. It was concluded that GO played a crucial role in the catalytic light-fenton process.

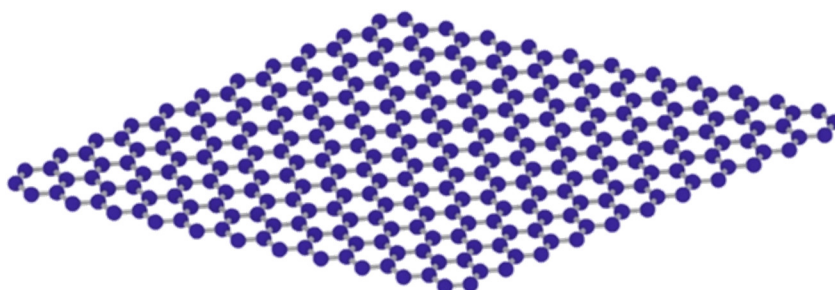


Figure 3. 3D structure diagram of graphene.

3. Classification and Mechanism of Fenton-Like Reagents

3.1. Traditional Fenton Reagents

The conventional fenton reagents are a mixed system consisting of H_2O_2 and Fe^{2+} , which attack organic macromolecules through the $\cdot\text{OH}$ produced by the catalytic disintegration of H_2O_2 [61]. In the catalytic degradation process, free radicals break down large-molecule organic matter into small-molecule organic matter or mineralize them into inorganic substances such as carbon dioxide and water, and the strength of free radical oxidation can be determined by the magnitude of the free radical oxidation-reduction potential [62]. It can be seen from Table 2 that the hydroxyl radical has a high oxidation electrode potential and is rapidly oxidized, which leads to the conclusion that hydroxyl radical reaction is a fast reaction. However, fenton reagents have some problems: the utilization rate of H_2O_2 is low, and the treated water may have color, which means it is relatively difficult to apply the reagents in drinking water treatment [63]. After continuous exploration and research, it is generally agreed that the fenton reaction mechanism is $\cdot\text{OH}$ generated by a fenton reagent through catalytic decomposition to attack organic molecules, as shown by the following chemical equations:

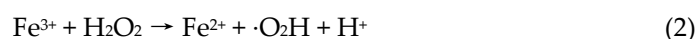


Table 2. Oxidation-reduction potentials of various oxidants.

Oxidants	Equations	Oxidation-Reduction Potential (V)
$\cdot\text{OH}$	$\cdot\text{OH} + \text{H}^+ + \text{e} = \text{H}_2\text{O}$	2.80
O_3	$\text{O}_3 + 2\text{H}^+ + 2\text{e} = \text{H}_2\text{O} + \text{O}_2$	2.07
H_2O_2	$\text{H}_2\text{O}_2 + 2\text{H}^+ + 2\text{e} = 2\text{H}_2\text{O}$	1.77
MnO_4^-	$\text{MnO}_4^- + 8\text{H}^+ + 5\text{e} = \text{Mn}^{2+} + 4\text{H}_2\text{O}$	1.51
ClO_2	$\text{ClO}_2 + \text{e} = \text{Cl}^- + \text{O}_2$	1.50
Cl_2	$\text{Cl}_2 + 2\text{e} = 2\text{Cl}^-$	1.30

3.2. Fenton-Like Reactions

Numerous researchers have continuously improved the traditional fenton oxidation method, and a large number of improved fenton reagents have emerged, such as $\text{H}_2\text{O}_2/\text{Fe}^{3+}$, $\text{H}_2\text{O}_2/\text{O}_3$, light-fenton reagents and electro-fenton reagents [64]. Since the reaction principle of the improved technique is still a hydroxyl radical mechanism (Figure 4), which is similar to the fenton reaction, such techniques are collectively referred to as fenton-like reagents methods [65].

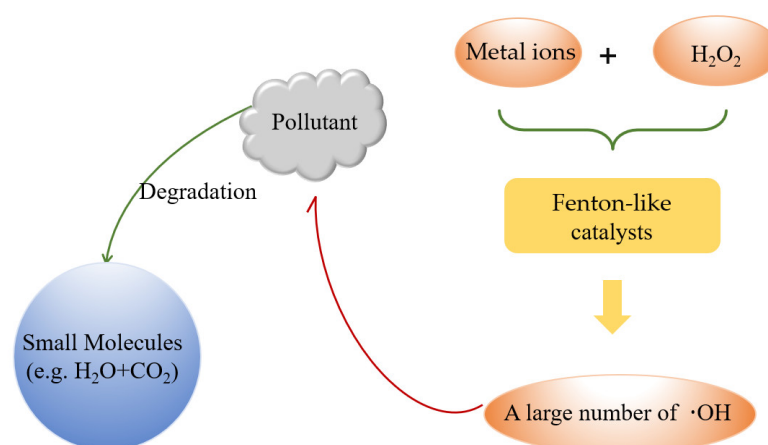


Figure 4. Basic schematic diagram of fenton reactions.

3.2.1. Light-Fenton Method

Light-fenton method introduces light sources (e.g., UV, visible light) into the fenton reaction to form a UV-fenton system [66]. UV and fenton catalysts can have a synergistic effect on the catalytic decomposition of H_2O_2 , which has a great improvement on the oxidation of fenton reagents. However, the proportion of UV light in natural light is relatively small (about 4%), and visible light accounts for approximately 43%, thus the introduction of visible light into the fenton system is of great significance. The fenton-like catalyst absorbs photons and oxidizes H_2O molecules coordinated to Fe(III) to produce hydroxyl radicals [67]. The electrons of iron atoms undergo charge transfer with oxygen atoms, and then the Fe(II) is oxidized to Fe(III) by the dissolved molecular oxygen (O_2) in solution, or the $\text{HOO}\cdot$ and $\cdot\text{OH}$ generated during the reaction [68]. The mechanism diagram is shown in Figure 5, which is represented by the following chemical equations:

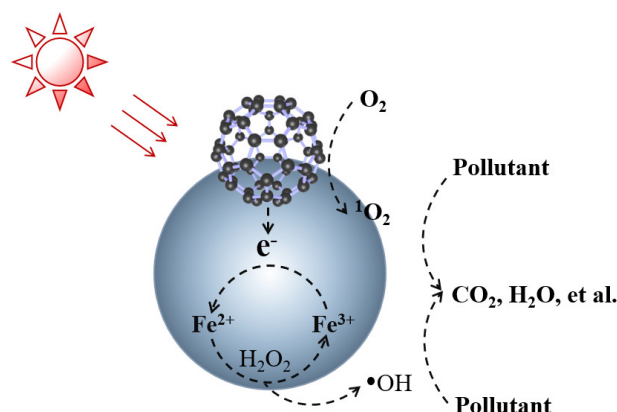
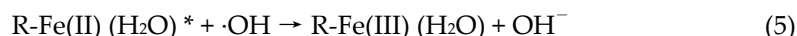
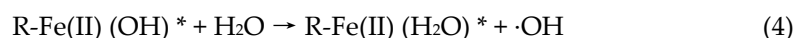
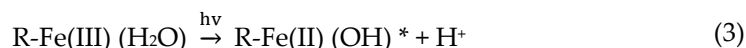


Figure 5. Light-fenton reactions mechanism diagram.

3.2.2. Electro-Fenton Method

The electro-fenton reaction generates H_2O_2 or Fe^{2+} through electrolysis in an electrolytic cell, thereby forming fenton reagents [69]. After wastewater flowing into the electrolyzer, the reaction mechanism can be improved under the electrochemical effect. The continuous regeneration of Fe^{2+} at the cathode increases the decomposition rate of organic pollutants thus improving the treatment efficiency of the reagents. The electro-fenton method is to spray oxygen to the cathode of the electrolytic cell to produce H_2O_2 , which rapidly reacts with Fe^{2+} produced by the oxidation of the Fe anode applied to the solution to generate $\cdot\text{OH}$ and Fe^{3+} . This makes use of the strong oxidizing power of $\cdot\text{OH}$ to catalyze the degradation of organic matter [70]. Besides, Fe^{3+} on the cathode can be reduced to Fe^{2+} , so that the oxidation reaction cyclically proceeds [71]. The mechanism diagram is shown in Figure 6, and the reaction process is as follows:



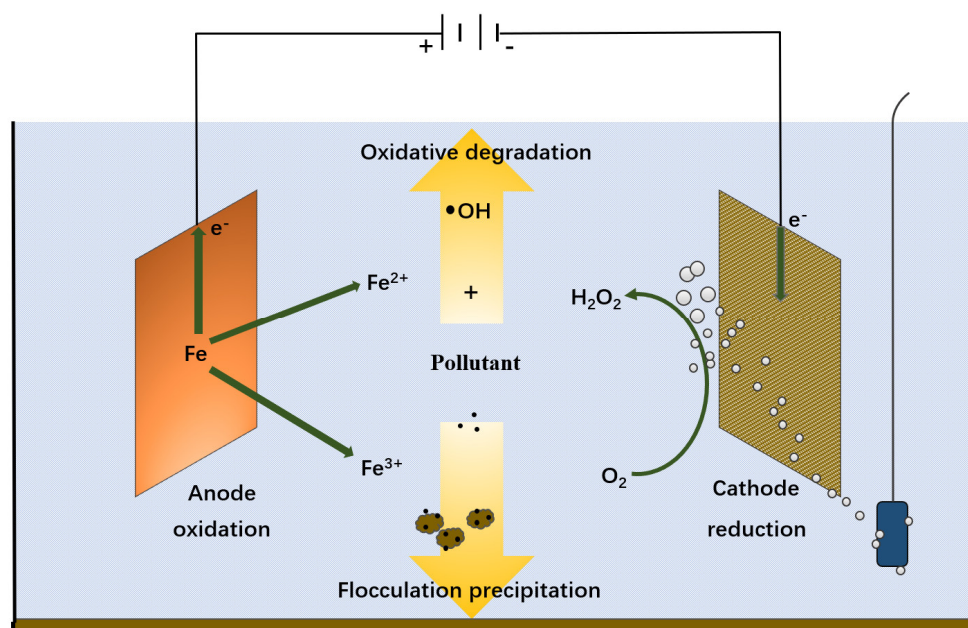


Figure 6. Electrical-fenton reactions mechanism diagram.

3.2.3. Ultrasound-Fenton Method

The ultrasonic and vibration of ultrasound can play the role of agitator and mass transfer, and the ultrasound/fenton synergistic method can utilize the cavitation effect of ultrasound radiation to improve degradation rate of pollutants [72]. The mechanism of the ultrasound-fenton method consists of the following two parts: pyrolysis of pollutants due to local high temperature and high pressure that generated by the cavitation of ultrasonic. The strong oxidation potential of hydroxyl radicals generated in the high temperature and high pressure environment has an oxidizing effect on pollutants [73].

3.2.4. Microwave-Fenton Method

The mechanism of action of the microwave-fenton method is similar to that of the ultrasound-fenton method [74]. The microwave-fenton method is the utilization of the free radical oxidation and microwaves (thermal and electromagnetic field effects) to decompose pollutants, to promote the decomposition of H₂O₂ and the yield of •OH, and accelerate the polarization of pollutant molecules to achieve synergy with the oxidant, catalyzing destruction of organic matters in wastewater [75].

4. Preparation of Carbon-Based Nano Fenton-Like Reagents

4.1. Preparation Methods of Carbon Materials

Carbon atoms form many carbon isomers and their bonding is richly varied under high pressure. Carbon nanomaterials have received much attention from researchers due to their excellent properties, such as a wide source of raw materials, controllable structure, good chemical stability and wonderful electrical conductivity, and have shown a great potential in the fields of electronic devices, biology and catalysis. The following are the methods of preparing carbon nanomaterials, and the advantages and limitations of these methods are shown in Table 3.

- (1) Method of laser evaporation of graphite, which means that with the help of metal catalysts, pulsed laser bombard on graphite surface to manufacture carbon nanomaterials [76]. For example, Plotnikov et al. [77] fabricated diamond amorphous carbon films with this method.

- (2) Plasma spray deposition technique means that the plasma-sprayed tungsten electrode (cathode) and copper electrode (anode) are cooled by water. When Ar/He gas carries benzene vapor through the plasma torch, it deposits nano-carbon soot materials on the anode.
- (3) Condensed electrolytic generation method, with argon playing the role of protective gas, is where the graphite electrode (electrolytic cell as the anode) is employed to electrolyze the molten halide alkali salt and generate various forms of carbon nano-materials at a temperature of about 600 °C, as well as a certain voltage and current.
- (4) Graphite arc method is to apply graphite electrodes to discharge in an atmosphere to gather carbon nano-molecular materials from cathode deposits. Xingke et al. [78] employed this method to prepare carbon nanotubes.
- (5) Chemical vapor deposition is a widely adopted technique for the preparation of carbon materials and can be classified into catalytic and non-catalytic chemical vapor deposition methods. Catalytic chemical vapor deposition is the catalytic decomposition of a gas (or vapor) containing a carbon source as it flows over the surface of the catalyst. Non-catalytic vapor phase deposition requires no catalyst and directly decomposes the gas phase carbon containing organics thermally under a protective atmosphere.

Table 3. Comparison of methods for the preparation of carbon nanomaterials.

Method	Advantages	Limitations
Method of laser evaporation of graphite	High purity of the products	Low output and valuable equipment
Plasma spray deposition technique means	Long electrode life; stable combustion; independent airflow and pressure control; higher efficiency	Expensive carrier gas; small spraying rate; high quality requirements for spraying materials
Graphite arc method	No harmful product formation	Obtain high purity products and consumption of too large amount of energy
Chemical vapor deposition	Simple process, low cost, high yield, suitable for industrial production	Due to the low reaction temperature, the prepared material is defective and requires some post-treatment

4.2. Preparation of Carbon-Based Nano Fenton-Like Catalysts

The fabrication methods of fenton-like catalysts include impregnation, electrostatic spinning, solvothermal synthesis, hydrothermal method, co-precipitation, electrolysis, ion exchange, etc. The impregnation method refers to immerse solid powder or solid of a certain shape and size (carrier or catalyst containing the main body) in a solution of soluble compounds containing active components (main and auxiliary catalytic components) and separating the residual liquid after contacting for a certain period time [79]. Zou et al. [51] immersed Fe₂O₃ on fullerenes to form a C₆₀/Fe₂O₃ combined with hydrogen peroxide (H₂O₂) and visible light to degrade RhB, MO and phenol.

Electrostatic spinning is a special fiber manufacturing process in which a polymer solution or melt is jet spun in a strong electric field. Under the action of an electric field, the droplet at the tip of the needle changes from a sphere to a cone and extends from the tip of the cone to obtain fibrous filaments [80]. In this way, polymer filaments of nanometer diameter can be produced. Akhi et al. [81] demonstrated that the carbon/MWCNT/Fe₃O₄ composite nanomaterials were prepared. Multi-walled carbon nanotubes were synthesized by catalytic chemical vapor deposition in advance, and nitric acid and sulfuric acid were added to make MWCNT undergoes carboxylation reaction, and then FeCl₃·6H₂O and FeCl₂·4H₂O were respectively added to prepare this composite nanomaterial. The solvothermal method in the original mixture is reacted at a certain temperature and autogenous pressure of the solution, with organic substances or non-water

as the solvent [82]. Deng et al. [83] reported the preparation of Fe_3O_4 -multiwalled carbon nanotube hybrids applying a solvothermal method with acid treatment of MWCNTs and iron acetylacetonate in a mixed solution of ethylene glycol and ultrapure water.

The hydrothermal method refers to a preparation way in which water as a solvent and the powder is dissolved and recrystallized in a sealed pressure vessel [84]. Xu et al. [85] revealed that the $\text{ZnO}/\text{GO}/\text{SiO}_2$ composite catalyst was synthesized by hydrothermal loading of ZnO nanorods on the GO/SiO_2 surface, and this composite catalyst with H_2O_2 constituted a photocatalytic fenton-like system for the degradation of methylene blue. The co-precipitation is the presence of two or more cations in solution, which are present in a homogeneous phase in the solution [86]. Homogeneous precipitation of various components can be obtained by adding a precipitating agent, and it is an excellent approach to prepare ultrafine powders containing two or more metal elements. Song et al. [87] synthesized ultrasmall graphene oxide (usGO) by chemical oxidation of graphite powder and then synthesized $\text{Fe}_3\text{O}_4/\text{usGO}$ on the surface of usGO by co-precipitation method.

5. Characterization of Carbon-Based Fenton-Like Nanocomposites

In order to obtain the mechanism as well as the physicochemical properties (e.g., structure, shape, particle size, size distribution, valence and specific surface area) of carbon-based nanocomposites fenton-like reagents, they can be analyzed by some characterization techniques [88]. The main methods described in this article are as follows.

Electron spin resonance (ESR) is the most direct and effective way to detect radicals, and the ESR spectra can give the hyperfine splittings to provide structural information [89]. To verify whether hydroxyl radicals in fenton-like reagents play a role in the degradation of substrates. Peng et al. [90] employed ESR assays to analyze the mechanism because testing 2-propanol was an effective solvent to remove $\cdot\text{OH}$. Then, by increasing the amount of 2-propanol, the destruction rate of TCS became slower and the removal of TCS was almost completely inhibited, indicating that $\cdot\text{OH}$ is the main radical for TCS degradation in the $\text{Fe}/\text{H}_2\text{O}_2/\text{GO}$ system. The phase analysis of the crystal material based on the diffraction peak intensity, angular position, relative intensity sequence and diffraction peak shape was carried out by X-ray diffraction (XRD) [88]. The Bragg equation reflects the relationship between the direction of diffraction lines and the crystal structure. For a particular crystal, only the angle of the incident ray satisfying the Bragg equation can produce interference and exhibit diffraction fringes. Zhu et al. [91] performed the XRD analysis on 3D- MoS_2 nanospheres and GO-loaded 3D- MoS_2 , respectively, and deduced that the preparation of GO-loaded MoS_2 nanospheres was successful. This material had more surface defects, which was expected to expose more metal active centers, and exhibited an excellent degradation efficiency in AOPs.

X-ray absorption near edge structure (XANES) is typically a spectrum between the absorption edge at the absorption threshold and about 50 eV above the absorption edge and can be used to provide information about the oxidation state and local atomic environment of the system [92]. Extended x-ray absorption fine structure (EXAFS) is the absorption spectrum beyond the XANES region, ranging from 50 eV above the threshold up to 1000 eV. EXAFS is an element-specific structure characterization technology that can accurately characterize the local environment around the absorbing atom (the type of adjacent atoms, bond length, coordination number, and the degree of disorder in the structure of the specific atomic shell) [93].

Transmission electron microscopy (TEM) is capable of observing various defects and atomic structures inside materials [94]. High-resolution transmission electron microscopy (HRTEM) can be extensively implemented for microstructure (including general morphology and lattice images) assessment of materials (especially nanomaterials) and qualitative and quantitative analysis of micro-zone components. [95]. It also can obtain lattice fringe image (reflecting the information of interplanar spacing), structure image and single atom image (representing the atoms or atomic groups in the crystal structure configu-

ration situation) and other higher resolution image information. Yao et al. [96] further revealed the tightly hybridized structure of the catalyst by transmission electron microscopy images, with ZnFe_2O_4 NPs deposited on the surface of large rGO sheets. Both the edges of rGO and the nanostructures of ZnFe_2O_4 NPs can be inspected in the TEM images with high resolution. Zubir et al. [97] investigated the structure of $\text{GO}/\text{Fe}_3\text{O}_4$ adopting HRTEM and revealed that the lattice striations of this material were clearly visible, demonstrating the crystallinity nature of the nanomaterials, while it was observed that the relative pressure of the nanocomposites was closely related to the variation of Fe_3O_4 nanoparticles.

High-angle annular dark field scanning transmission electron microscopy (HAADF-STEM) can not only distinguish the spatial atomic distribution of the crystal structure, but also directly display the distribution of atomic species and defect structure within the crystal [98]. For strongly correlated materials, the charge density wave, vacancy sequence, ion sequence modulation can be recognized, which are of great significance for understanding microstructure defects and local symmetry breaking. Li et al. [99] obtained HAADF-STEM images and EDS of the same regions in order to study the loading conditions of Bi-FePc on the rGO surface. The EDS element mapping image of Bi-FePc/rGO nanocomposite material demonstrates that the distribution of carbon and iron are consistent, which reported that the Bi-FePc supported on the rGO surface is well dispersed and does not form aggregates. X-ray photoelectron spectroscopy (XPS) allows not only the study of elements and their different valence states on the surface of materials, but also the analysis of information on chemical states, molecular structures, and chemical bonds. In order to better understand the chemical composition and electronic structure of the Fe/C/BN catalyst synthesized before and after the degradation experiment, Yao et al. [100] conducted XPS research, the analysis revealed that the co-doping of B, N and Fe in the carbon skeleton by Fe/C/BN is the main active catalytic center for the effective oxygen reduction in the fenton-like process.

Nitrogen adsorption is the most common and reliable technique for specific surface and pore size distribution to determine the porosity, surface properties and other parameters of the material [101,102]. Wang et al. [103] investigated the N_2 adsorption-desorption isotherms of Fe_3O_4 nanoparticles and $\text{Fe}_3\text{O}_4/\text{GO}$ nanocomposites. The results revealed that $\text{Fe}_3\text{O}_4/\text{GO}$ has similar isotherms to Fe_3O_4 and type III isotherm according to the classification of IUPAC, which is a typical nonporous material. Fourier transform infrared spectroscopy (FTIR) observes the bond structure of the sample and the interaction in the mixed film, and can obtain the tensile and flexural vibration absorption spectra of the chemical bonds in the sample, which is used for the analysis of the group structure, the qualitative and quantitative analysis of the material [104]. Zhou et al. [105] compared fullerenes and Fe(III)-fullerenes by FT-IR spectroscopy. Original fullerenes are rich in hydroxyl groups and fullerene after loading show lower C–O stretching response and C–O–C and C=O, which revealed that the oxygen-containing functional groups of fullerenes have changed from $\cdot\text{OH}$ to C–O–C and C=O.

6. Application of Carbon-Based Fenton-Like Nanocomposites

The carbon-based nanocomposite fenton reagents have a stable and effective removal function for refractory biological wastewater, toxic wastewater and biological inhibition wastewater. Many factors affect their reaction rate: temperature, acidity and alkalinity, catalyst type, hydrogen peroxide concentration, carrier type, etc. A summary of the applications for fenton-like reagents is shown in Table 4.

6.1. Application of Light-Fenton Reagents

The light-fenton method can directly degrade part of organic matter under the action of light, improving the utilization rate of H_2O_2 , reducing the amount of fenton-like catalyst and further enhance the degree of mineralization of organic matter. Zhou et al. [105] investigated a method to degrade dimethyl phthalate (DMP) based on the high photosensitivity and water solubility of fullerene-enhanced Fe(III)/ H_2O_2 fenton system under visible

light conditions. In addition, the effects of experimental parameters, such as initial pH, iron dosage, fullerene dosage, H_2O_2 dosage and light intensity, on DMP degradation were evaluated, and the DMP degradation pathways were analyzed by monitoring the concentration changes of major intermediates. Hybrid nanomaterials formed by nanostructured $\beta\text{-FeOOH}$ were prepared by Fang et al. [106]. The photocatalytic activity of carbon nanotubes/ $\beta\text{-FeOOH}$ was examined for the decomposition of MO in the presence of hydrogen peroxide and visible light irradiation, and the photocatalytic oxidation mechanism of the fenton-like system was discussed. The findings illustrated that CNTs/ $\beta\text{-FeOOH}$ composites have significant visible light photocatalytic activity than bare $\beta\text{-FeOOH}$ for the oxidation of MO in the aqueous phase. In addition, the growth of visible light absorption intensity and the decrease of $\beta\text{-FeOOH}$ particle size favored the photocatalytic and photo-fenton-like reaction for the destruction of MO. Yu et al. [107] demonstrated $\text{Fe}_3\text{O}_4/\text{GO}$ as an efficient photo-fenton catalyst to degrade phenol under UV light irradiation. In this photo-fenton process, the effects of GO content, catalyst dosage, hydrogen peroxide concentration and initial pH of the solution on the degradation of phenol were analyzed. Additionally, $\text{Fe}_3\text{O}_4/\text{GO}$ still displayed high catalytic activity as well as good stability after five cycles.

6.2. Application of Electric-Fenton Reagents

The electro-fenton method is highly regarded owing to its in situ hydrogen peroxide production and its ability to effectively control degradation costs [108]. In addition to the oxidation of $\cdot\text{OH}$ in the system, there are also electrochemical anodic oxidation and electrosorption to improve the removal rate. Ai et al. [109] combined copper oxide nanotubes and multi-walled carbon nanotubes with polytetrafluoroethylene to fabricate a composite copper oxide/carbon nanotube/polytetrafluoroethylene cathode and developed a novel electro-fenton system with an oxygen-containing gas diffusion cathode. The electro-fenton system generates copper ions in situ from Cu_2O nanocubes while electrochemically reducing oxygen to hydrogen peroxide. The fenton reagents further react together to generate hydroxyl groups, which effectively degrade rhodamine B (RhB) at neutral pH, and it would not lose activity after long-term operation. Effect of H_2O_2 and Fe^{3+} oxidation on the removal efficiency of anionic (AR14) and cationic (MB) dyes in graphite electrodes modified by electrochemical deposition of GO and rGO by Akerdi et al. [110]. The effects of key parameters—wastewater pH, dye concentration, magnetic nanoparticle loading (Fe_3O_4) and current—on dye decontamination were investigated. Furthermore, the chemical behavior of GE-GO and GE-rGO in the oxidation-reduction of MB and AR14 was discussed by cyclic voltammetry. The conclusion was that the removal of AR14 was greater than MB, the removal of rGO-GE system was better than GO-GE, and the magnetic nanoparticles were also effective in generating radical hydroxyl groups in alkaline pH.

6.3. Application of Ultrasonic-Fenton Reagents

The strong oxidation potential hydroxyl radicals generated by the ultrasound-fenton method under high temperature and pressure can degrade organic pollutants in wastewater more efficiently with simple operation and low cost. In the case of ammonium chloride as the nitrogen source, Yang et al. [111] simply carbonized the mixture of glucose and iron salt precursor $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ to synthesize Fe/N-C-2 hybrids and then removed tetracycline. The good magnetic properties facilitate the recovery from aqueous media. After six consecutive runs, Fe/N-C-2 maintained more than 88% of its catalytic capacity, which illustrated the high stability of the Fe/N-C-2 mixture in an aqueous solution. Zhang et al. [112] integrated ultrasound and FeGAC to degrade crystal violet. The consequence revealed that the decolorization efficiency decreased with increasing pH, but increased with increasing power density and catalytic dosage. In optimal conditions, the decolorization efficiency reached 88%, and the toxicity decreased sharply at ultrasonic time greater than 50 min. After three consecutive cycles, the decolorization efficiency did not

change much and the leaching rate of iron was low, indicating the stability of the catalyst in the ultrasonic-fenton system.

6.4. Application of Microwave-Fenton Reagents

The microwave-fenton method combines microwave radiation and fenton reagents to reduce the activation energy of chemical reactions, so that the structure of organic molecules can be more easily destroyed during the reaction [113]. Yang et al. [114] combined microwave technology with fenton reaction for the removal of high-concentration pharmaceutical wastewater. They found that the microwave power was 300 W, the amount of H_2O_2 was 1300 mg/L, the reaction time was 6 min, the dosage of $\text{Fe}_2(\text{SO}_4)_3$ was 4900 mg/L, the removal rate of COD can reach 57.53%, which demonstrated that microwave can shorten the reaction time and improve its oxidation efficiency. Liu et al. [115] used microwave-enhanced fenton-like treatment of methylene blue in aqueous solution. Operating parameters such as initial pH, MB concentration, and the amount of hydrogen peroxide were investigated to obtain optimal conditions for improving the degradation rate of MB.

Table 4. Summary of some applications of fenton-like degradation of pollutants.

Pollutant	Materials	Results	References
Dimethyl phthalate	F/fenton; fullerene-Fe(III)/ H_2O_2 and Fe(III)/ H_2O_2 fenton systems;	Under visible light conditions, the F/fenton system can almost completely degrade DMP within 50 min, and its large capacity eventually reaches 0.0771 min^{-1} , which is 18.5 and 45.4 times higher than that of fullerene-Fe(III)/ H_2O_2 and Fe(III)/ H_2O_2 , respectively	[105]
MO	CNT/ β -FeOOH	The results revealed that the increase of visible light absorption intensity and the decrease of β -FeOOH particle size were favorable to the photocatalytic and photo-fenton reaction degradation.	[106]
Phenol	Fe_3O_4 -GO	Under optimal conditions (pH 5.0, hydrogen peroxide concentration 10.0 mmol/L, catalyst dose 0.25 g/L), 98.8% of phenol in phenol solution can be removed after 120 min	[107]
Rhodamine B	Cu_2O /CNTs/PTFE	The degradation of RhB in this E-fenton system reached 80.2% and 89.3% in 120 min at neutral pH and pH of 3, respectively	[109]
AR14; MB	Fe_3O_4 /GO; Fe_3O_4 /rGO	In the experiments, the removal of AR14 exceeded that of MB and the rGO-GE system exceeded that of GO-GE, demonstrating that magnetic nanoparticles are also effective in generating free radical hydroxyl groups in alkaline pH.	[110]
Tetracycline	Fe/N-C-2/ H_2O_2 /US system	The maximum removal of TC in this type of fenton system was 92.77%, and the catalytic capacity of Fe/N-C-2 remained above 88% after six consecutive runs, which indicates the high stability of Fe/N-C-2 composites in aqueous solutions.	[111]
Crystal Violet	FeGAC/ H_2O_2	The optimal conditions for the removal of crystal violet by this class of fenton reagents were an initial pH of 3, a hydrogen peroxide concentration of 1.8 mmol/L, a catalyst loading of 2.5 g/L and a power density of 141 W/L, and a maximum removal rate of 88%.	[112]

7. Studies on Pollutants Removal Processes

7.1. Studies on Thermodynamics

The Langmuir model is based on the assumption of adsorption between gas and solid phases and explains how adsorption occurs on uniform surfaces and it is valid for single-layer adsorption on a limited number of interfaces at the same location. The Langmuir isotherm can also visually describe the adsorption mechanism during the derivation, which lays the foundation for other adsorption models [116]. Freundlich isotherm model is an empirical law obtained when studying the adsorption of gases by porous materials. It applies to liquids and gases. It is a semi-empirical equation, which expresses the non-uniform multilayer adsorption behavior of adsorbates on the surface of adsorbents [117]. Temkin equation considers that when the adsorbent adsorbs the adsorbate in the solution, it interacts with the adsorbed solute and affects the adsorption behavior and process. The specific energy relationship is that the heat of adsorption decreases linearly with the amount of adsorption [118]. The Dubinin-Radushkevich (D-R) isotherm was used to evaluate whether the adsorption process had a physical or chemical mechanism [119]. The equations and descriptions of these adsorption isotherm models are in Table 5.

Table 5. Commonly used adsorption isotherm models.

Adsorption Type	Isotherm Models	Description	References
Langmuir	$\frac{C_e}{q_e} = \frac{1}{k_L q_m} + \frac{C_e}{q_m}$	where C_e is the final concentration; q_e is defined as the mole of the adsorbate sorbed per mole of adsorbent at equilibrium; q_m (mg/g) is the maximum adsorption capacity; K_L refers to the constant of the proposed Langmuir model.	[120]
Freundlich	$q_e = k_F (C_e)^{1/n}$	where C_e defines the equilibrium concentration (mg/L); q_e is the amount of adsorbance (mg/g); K_F (mg/g) and $1/n$ are empirical constants.	[121]
Temkin	$q_e = \frac{B \ln(AC_e)}{B = RT/b}$	In this equation, q_e is the amount of adsorbance (mg/g); A (L/g) represents the binding constant (maximum binding energy); B (J/mol) stands for the Temkin constant; R is 8.314 J/mol/K and T (K) refers to the Kelvin temperature; the constant b represents the heat of adsorption.	[122]
D-R	$\ln q_e = \ln q_m$ $\varepsilon = RT \ln(1 + \frac{1}{C_e})$ $E = \frac{1}{\sqrt{-2\beta}}$	where q_e denotes the amount of adsorbate adsorbed by the adsorbent at equilibrium (mg/g); q_m represents the isothermal saturation capacity (mg/g) and β is known as the D-R constant (mol ² /kJ ²); ε is Polanyi's potential; the value of E gives information about chemical and physical sorption.	[123,124]

7.2. Studies on Kinetics

The adsorption and desorption rates are mainly determined by the interaction between adsorbent and adsorbate, temperature and pressure. The studies on adsorption kinetics are helpful to explore the mechanism of chemical adsorption and heterogeneous catalysts. The removal kinetic model helps to compare the predicted adsorption parameters with various adsorption behaviors under different experimental conditions. Wang et al. [125] used fenton-like reagents for the oxidation process of Acid Black 1 following the pseudo-first-order kinetics [126].

The form of the pseudo-first-order kinetic rate equation is as follows:

$$\frac{dC_e}{dC_t} = k_{p1}(C_e - C_t) \quad (10)$$

where C_e (mg/g) and C_t (mg/g) refer to the amount of adsorption at equilibrium and at time t , respectively, K_{p1} is the equilibrium rate constant for pseudo-first-order adsorption.

Equation (1) can be transformed into Equation (2) or Equation (3).

$$\log(C_e - C_t) = \log C_e - \frac{k_p t}{2.303} \quad (11)$$

$$\ln(C_e - C_t) = \ln C_e - k_1 t \quad (12)$$

In order to verify whether the experimental data conforms to the pseudo-first-order rate equation, it is necessary to know the calculated equilibrium adsorption capacity. Even if the adsorption capacity changes slowly, its value is usually less than the equilibrium concentration capacity. Therefore, in many cases, the equation does not match the experimental data in the entire range.

Ifelebuegu et al. [127] investigated the degradation of two EDCs (17 β -estradiol and 17 α -ethylestradiol) by fenton-like reagents, which conformed to pseudo-second-order kinetics. The pseudo-second-order model is represented by the following equation [128]:

$$\frac{dC_t}{dt} = -k_2 C_t^2 \quad (13)$$

$$\frac{1}{C_t} = \frac{1}{C_0} - k_2 t \quad (14)$$

where C_0 and C_t (mg/L) are the concentrations of solutes at equilibrium and at time t (min), respectively; k_2 (L (mg min)⁻¹) is the rate constant.

The typical intra-particle diffusion model is the uniform solid diffusion model, which can describe the mass transfer in amorphous and uniform spheres, and it has been widely used in the analysis of adsorption kinetics. The equation is expressed as [129]:

$$q_t = k_{id} t^{1/2} + C \quad (15)$$

where q_t is the amount of adsorption at time t ; $t^{1/2}$ is the square root of time; k_{id} is the intra-particle diffusion rate constant; C is a constant.

8. Summary and Future Perspectives

This paper reviews the preparation and characterization methods of carbon-based nanocomposite fenton-like reagents along with their mechanisms and applications in light-fenton, electro-fenton, ultrasound-fenton and microwave-fenton reactions. With the continuous advancement of human industrialization, a large number of chemical products are manufactured and employed. These chemicals enter the environment through industrial wastewater and urban sewage, which contains many organic pollutants that are difficult to degrade. If they are not effectively treated, the harm to human society is enormous. Fenton-like reactions are currently recognized as an effective technology for the treatment of refractory organic pollutants, but a further exploration about the mechanism, efficient application and the integration with artificial intelligence still needs to be carried out; therefore, the investigation of fenton-like catalysts is of utmost importance for environmental protection.

From this review, it is evident that the previous studies have demonstrated a great breadth and depth, but the following aspects still need to be addressed for fenton-like catalysts. Firstly, the mechanism of action of the fenton-like reagent oxidation reactions system should be more clearly understood, it is necessary to strengthen the study on the micro-kinetics, interfacial effects and synergistic effects of external conditions on the catalyst surface. Secondly, research and evaluation of catalyst reaction intermediates are conducted to develop new efficient and inexpensive catalysts, especially those that do not cause secondary pollution and environmentally friendly, which can further improve the utilization rate of $\cdot\text{OH}$ and the biochemical properties of wastewater. Thirdly, the effective applications of fenton-like catalysts are improved based on the full consideration of practicality. From the perspective of practical applications, the main purpose of broadening the use range of fenton systems is to achieve efficient performance of catalysts in neutral conditions. The effect of the new porous catalyst carrier on the reaction active material and its stability ought to be investigated, and they need to be used selectively to improve

the activity and enhance the force between the active component and the carrier. Lastly, the development and research of relevant treatment equipment and facilities should be strengthened, and the combination of fenton-like and related technologies should be applied to the treatment of solid and gas wastes, for instance, real-time supervision of water quality and quantity, and link it with the dosage controlling system to achieve the computerized and modernized management, and making the operation simple, saving human resources and improving efficiency. Carbon-based fenton-like nanocomposite catalysts have the characteristics of fast reaction speed, mild reaction conditions and reusability, which will be more widely used in the field of water treatment. It is hoped that this review will provide a valuable reference for future research to gain further understanding and accelerate the development of this field.

Author Contributions: Writing—original draft preparation, L.X.; writing—review, J.H., X.W. and Q.L.; searching—literature, Y.X., C.L. and L.F. All authors have read and agreed to the published version of the manuscript.

Funding: This work was supported by the Government of Guizhou Province (Project No. [2017] 5726-42), the National Natural Science Foundation of China under Grant No. 21667012, the National 111 Project of China under Grant No. D17016, and the Science and Technology Foundation of Guizhou Province (Project No. [2020] 6009).

Institutional Review Board Statement: Not applicable.

Informed Consent Statement: Not applicable.

Data Availability Statement: In this paper, the data from Google Scholar, Web of Science, China National Knowledge Infrastructure, China Wan Fang Literature Database, and China WeiPu Literature Database.

Conflicts of Interest: The authors declare no conflict of interest.

Abbreviations

PPCPs	Pharmaceutical and Personal Care Products
PAHs	Polycyclic Aromatic Hydrocarbons
COD	Chemical Oxygen Demand
CWAO	Catalytic Wet Oxidation
CNT	Carbon Nanotubes
MWCNT	Multi-Walled Carbon Nanotubes
GO	Graphene Oxide
usGO	Ultrasmall Graphene Oxide
MO	Methyl Orange
AO7	Acid Orange 7
ESR	Electron Spin Resonance
XRD	X-ray Diffraction
XANES	X-ray Absorption Near Edge Structure
EXAFS	Extended X-Ray Absorption Fine Structure
TEM	Transmission Electron Microscopy
HRTEM	High-Resolution Transmission Electron Microscopy
HAADF-STEM	High-Angle Annular Dark Field Scanning Transmission Electron Microscopy
XPS	X-ray Photoelectron Spectroscopy
FTIR	Fourier Transform Infrared Spectroscopy
DMP	Dimethyl Phthalate
D-R	Dubinini-Radushkevich

References

1. Wang, Z.; Jiang, Q.; Fu, Q.; Jiang, X.; Mo, K. Eco-environmental effects of water resources development and utilization in the Sanjiang Plain, Northeast China. *Water. Sci. Technol. Water Supply* **2018**, *18*, 1051–1061.
2. Osman, A.; Kloas, W. Water Quality and Heavy Metal Monitoring in Water, Sediments, and Tissues of the African Catfish *Clarias gariepinus* (Burchell, 1822) from the River Nile, Egypt. *J. Environ. Prot. Ecol.* **2010**, *1*, 389–400.

3. Salmivaara, A.; Porkka, M.; Kumm, M.; Keskinen, M.; Guillaume, J.H.A.; Varis, O. Guillaume, Exploring the Modifiable Areal Unit Problem in Spatial Water Assessments: A Case of Water Shortage in Monsoon Asia. *Water* **2015**, *7*, 898–917.
4. Chen, C.; Cheng, T.; Shi, Y. Adsorption of Cu(II) from Aqueous Solution on Fly Ash Based Linde F (K) Zeolite. *Iran. J. Chem. Chem. Eng.* **2014**, *33*, 29–35.
5. Nan, Y.A.; Chen, C.B.; Djl, A. Cobalt nanoparticles embedded over periodic mesoporous organosilica functionalized with benzotriazolium ionic liquid for efficient and heterogeneous catalytic transformation of carbon dioxide to cyclic carbonates. *J. Environ. Chem. Eng.* **2020**, *8*, 103953.
6. Lian, D.; Keith, A.; Maruya, B.; Eddy, Y.; Zeng, A. China's water pollution by persistent organic pollutants-ScienceDirect. *Environ. Pollut.* **2012**, *163*, 100–108.
7. Zhang, X.; Chen, C.; Lin, P.F.; Hou, A.X.; Niu, Z.B.; Wang, J. Emergency Drinking Water Treatment during Source Water Pollution Accidents in China: Origin Analysis, Framework and Technologies. *Environ. Sci. Technol.* **2011**, *45*, 161–167.
8. Guittony-Philippe, A.; Masotti, V.; Höhener, P.; Boudenne, J.-L.; Viglione, J.; Laffont-Schwob, I. Constructed wetlands to reduce metal pollution from industrial catchments in aquatic Mediterranean ecosystems: A review to overcome obstacles and suggest potential solutions. *Environ. Int.* **2014**, *64*, 1–16.
9. Lettinga, G.; Roersma, R.; Grin, P. Anaerobic treatment of raw domestic sewage at ambient temperatures using a granular bed UASB reactor. *Biotechnol. Bioeng.* **1983**, *25*, 1701–1723.
10. Zhao, Y.Q.; Zhao, X.H.; Babatunde, A.O. Use of dewatered alum sludge as main substrate in treatment reed bed receiving agricultural wastewater: Long-term trial. *Bioresour. Technol.* **2009**, *99*, 644–648.
11. Hong, M.L.; Rong, L.F.; Li, W.J. Biological treatment of high-pH and high-concentration black liquor of cotton pulp by an immediate aerobic-anaerobic-aerobic process. *Water Sci. Technol.* **2009**, *60*, 3275–3284.
12. Santos, A.; Sagüillo, E.G.; Martínez-Villacorta, Á.M.; Rosas, J.M.; Romero, A. Diuron abatement in contaminated soil using Fenton-like process. *Chem. Eng. J.* **2012**, *183*, 357–364.
13. Fan, M.; Hu, J.; Cao, R.; Xiong, K.; Wei, X. Modeling and prediction of copper removal from aqueous solutions by nZVI/rGO magnetic nanocomposites using ANN-GA and ANN-PSO. *Sci. Rep.* **2017**, *7*, 1–14.
14. Rezaei, E.; Soltan, J.; Chen, N.; Lin, J. Effect of noble metals on activity of MnOx/ γ -alumina catalyst in catalytic ozonation of toluene. *Chem. Eng. J.* **2019**, *214*, 219–228.
15. Einaga, H.; Ogata, A. Benzene oxidation with ozone over supported manganese oxide catalysts: Effect of catalyst support and reaction conditions. *J. Hazard. Mater.* **2009**, *164*, 1236–1241.
16. Keen, O.S.; Ferrer, I.; Thurman, E.M.; Linden, K.G. Degradation pathways of lamotrigine under advanced treatment by direct UV photolysis, hydroxyl radicals, and ozone. *Chemosphere* **2014**, *117*, 316–323.
17. Liu, Z.Y.; Yue, T.; Xu, Y.H.; Zheng, Y.Q. Ozone Catalytic Oxidation of Biologically Pretreated Semi-Coking Wastewater (BPSCW) by Spinel-type MnFe₂O₄ Magnetic Nanoparticles-ScienceDirect. *Sep. Purif. Technol.* **2020**, in press.
18. Chen, L.; Xu, Y.; Sun, Y. Combination of Coagulation and Ozone Catalytic Oxidation for Pretreating Coking Wastewater. *Int. J. Environ. Res. Public Health* **2019**, *16*, 1705.
19. Rui, C.M.; Rossi, A.F.; Quinta-Ferreira, R.M. Fenton's oxidation process for phenolic wastewater remediation and biodegradability enhancement. *J. Hazard. Mater.* **2010**, *180*, 716–721.
20. Garrido-Ramirez, E.G.; Theng, B.K.G.; Mora, M.L. Clays and oxide minerals as catalysts and nanocatalysts in Fenton-like reactions—A review. *Appl. Clay. Sci.* **2010**, *47*, 182–192.
21. Wei, J.; Zhu, W.; Li, H.; Chao, Y.; Zhao, Z. Mechanism and optimization for oxidative desulfurization of fuels catalyzed by Fenton-like catalysts in hydrophobic ionic liquid. *J. Mol. Catal. A Chem.* **2014**, *382*, 8–14.
22. Bello, M.M.; Raman, A.A.; Asghar, A. A review on approaches for addressing the limitations of Fenton oxidation for recalcitrant wastewater treatment. *Process Saf. Environ. Prot.* **2019**, *126*, 119–140.
23. Bokare, A.D.; Choi, W. Review of iron-free Fenton-like systems for activating H₂O₂ in advanced oxidation processes. *J. Hazard. Mater.* **2014**, *275*, 121–135.
24. Su, Y.C.; Ma, S.H.; Zhao, X.L.; Dong, M.D.; Han, X.J. Template-free synthesis of inorganic hollow spheres at water/"water-brother" inter-faces as Fenton-like reagents for water treatment. *J. Environ. Sci.* **2017**, *55*, 331–338.
25. Mesarič, T.; Baweja, L.; Drašler, B.; Drobne, D.; Makovec, D.; Dušak, P.; Dhawan, A.; Sepčić, K. Effects of surface curvature and surface characteristics of carbon-based nanomaterials on the adsorption and activity of acetylcholinesterase. *Carbon* **2013**, *62*, 222–232.
26. Madima, N.; Mishra, S.B.; Inamuddin, I. Carbon-based nanomaterials for remediation of organic and inorganic pollutants from wastewater. A review. *Environ. Chem. Lett.* **2020**, *18*, 1169–1191.
27. Ms, A.; Ah, A.; Fb, A. Application and prospects of carbon nanostructured materials in water treatment: A review-ScienceDirect. *J. Water Process Eng.* **2020**, *33*, 100996.
28. Zhang, L.B.; Wang, E.K. Metal nanoclusters: New fluorescent probes for sensors and bioimaging-ScienceDirect. *Nano Today* **2014**, *9*, 132–157.
29. Yang, C.C.; Mai, Y.-W. Thermodynamics at the nanoscale: A new approach to the investigation of unique physicochemical properties of nanomaterials. *Mater. Sci. Eng. R Rep.* **2014**, *79*, 1–40.
30. You, J.Y.; Ma, X.Y.; Zhang, Z.; Hao, K.R.; Yan, Q.B.; Sheng, X.L.; Su, G. Carboneyane: A nodal line topological carbon with sp²-sp³ chemical bonds. *Carbon* **2019**, *152*, 909–914.

31. Hall, B.; Zhuo, C.; Levendis, Y.A.; Richter, H. Influence of the fuel structure on the flame synthesis of carbon nanomaterials. *Carbon* **2011**, *49*, 3412–3423.
32. Zhao, Y.S.; Fu, H.; Hu, F.; Peng, A.; Yang, W.; Yao, J. Tunable Emission from Binary Organic One-Dimensional Nanomaterials: An Alternative Approach to White-Light Emission. *Adv. Mater.* **2010**, *20*, 79–83.
33. Shemella, P.T.; Zhang, Y.; Mailman, M.; Ajayan, P.M.; Nayak, S.K. Energy gaps in zero-dimensional graphene nanoribbons. *Appl. Phys. Lett.* **2007**, *91*, 042101.
34. Salisu, N.; Mohd, H.; Zulkarnain, Z.; Nor, Y. Carbon-Based Nanomaterials/Allotropes: A Glimpse of Their Synthesis, Properties and Some Applications. *Materials* **2018**, *11*, 295.
35. Kosa, S.A.; Al-Zhrani, G.; Salam, M.A. Removal of heavy metals from aqueous solutions by multi-walled carbon nanotubes modified with 8-hydroxyquinoline. *Chem. Eng. J.* **2012**, *181*–182, 159–168.
36. Shahsavari, R.; Sakhavand, N. Junction configuration-induced mechanisms govern elastic and inelastic deformations in hybrid carbon nanomaterials. *Carbon* **2015**, *95*, 699–709.
37. Yuan, Q.; Xu, Z.; Jakobson, B.I.; Ding, F. Efficient Defect Healing in Catalytic Carbon Nanotube Growth. *Phys. Rev. Lett.* **2012**, *108*, 2157–2161.
38. Yu, H.; Liu, J.; Wang, Z.; Jiang, Z.W.; Tang, T. Combination of Carbon Nanotubes with Ni₂O₃ for Simultaneously Improving the Flame Retardancy and Mechanical Properties of Polyethylene. *J. Phys. Chem. C* **2009**, *113*, 13092–13097.
39. Shen, K.; Chen, X.; Chen, J.; Li, Y. Development of MOF-Derived Carbon-Based Nanomaterials for Efficient Catalysis. *ACS Catal.* **2016**, *6*, 5887–5903.
40. Xu, D.Y.; Lu, P.; Dai, P.; Wang, H.Z.; Ji, S.F. In Situ Synthesis of Multiwalled Carbon Nanotubes over LaNiO₃ as Support of Cobalt Nanoclusters Catalyst for Catalytic Applications. *J. Phys. Chem. C* **2012**, *116*, 3405–3413.
41. Karandikar, P.G. Nanotube-Containing Composite Bodies, and Methods for Making Same. U.S. Patent No 10/832,823, 18 August 2005.
42. Wang, K.F.; Chen, Y.J.; Tian, R.; Zhou, H.; Duan, Y. Porous Co-C Core-Shell Nanocomposites Derived from Co-MOF-74 with Enhanced Electromagnetic Wave Absorption Performance. *ACS Appl. Mater. Interfaces* **2018**, *10*, 11333–11342.
43. Kang, W.; Kim, N.H.; Lee, D.Y.; Chang, S.T.; Cho, J.H. Micropatterned Single-Walled Carbon Nanotube Electrodes for Use in High-Performance Transistors and Inverters. *ACS Appl. Mater. Interfaces* **2014**, *6*, 9664–9670.
44. Zhang, X.; Selkirk, A.; Zhang, S.; Huang, J.; Li, Y.; Xie, Y.; Dong, N.; Cui, Y.; Zhang, L.; Blau, W.J. Inside Cover: MoS₂/Carbon Nanotube Core-Shell Nanocomposites for Enhanced Nonlinear Optical Performance. *Chem. Eur. J.* **2017**, *23*, 3223–3223.
45. Zhang, J.; Wang, R.; Liu, E.; Gao, X.; Sun, Z.; Xiao, F.-S.; Girgsdies, F.; Su, D.S. Spherical Structures Composed of Multiwalled Carbon Nanotubes: Formation Mechanism and Catalytic Performance. *Angew. Chem. Int. Ed.* **2012**, *51*, 7581–7585.
46. Balandin, A.A. Thermal properties of graphene and nanostructured carbon materials. *Nat. Mater.* **2011**, *10*, 569–581.
47. Fan, J.; Wan, M.; Zhu, D.; Chang, B.; Pan, Z.; Xie, S. Synthesis, characterizations, and physical properties of carbon nanotubes coated by conducting polypyrrole. *J. Appl. Polym. Sci.* **1999**, *74*, 2605–2610.
48. Naik, N.; Ramasimha, R.; Arya, H.; Prabhu, S.; ShamaRao, N. Impact response and damage tolerance characteristics of glass-carbon/epoxy hybrid composite plates. *Compos. Part B Eng.* **2001**, *32*, 565–574.
49. Sang, L.; Liao, M.; Sumiya, M. A Comprehensive Review of Semiconductor Ultraviolet Photodetectors: From Thin Film to One-Dimensional Nanostructures. *Sensors* **2013**, *13*, 10482–10518.
50. Han, G.; Guo, Y.; Song, X.; Wang, Y.; Yi, Y. Terminal π - π stacking determines three-dimensional molecular packing and isotropic charge transport in an A- π -A electron acceptor for non-fullerene organic solar cells. *J. Phys. Chem. C* **2017**, *5*, 4852–4857.
51. Zou, C.-Y.; Meng, Z.-D.; Ji, W.-C.; Liu, S.-Q.; Shen, Z.; Zhang, Y.; Jiang, N.-S. Preparation of a fullerene[60]-iron oxide complex for the photo-fenton degradation of organic contaminants under visible-light irradiation. *Chin. J. Catal.* **2018**, *39*, 1051–1059.
52. Meng, Z.D.; Cho, K.Y.; Oh, W.C. Photocatalytic Degradation of Methylene Blue on Fe-Fullerene/TiO₂ Under Visible-Light Irradiation. *J. Law Med. Ethics* **2011**, *34*, 793–808.
53. Xu, T.; Zhu, R.; Liu, J.; Zhou, Q.; Zhu, J.; Liang, X.; Xi, Y.; He, H. Fullerol modification ferrihydrite for the degradation of acid red 18 under simulated sunlight irradiation. *J. Mol. Catal. A Chem.* **2016**, *424*, 393–401.
54. Li, Z.; Zheng, L.; Yan, W.; Pan, Z.; Wei, S. Spectroscopic Characteristics of Differently Produced Single-Walled Carbon Nanotubes. *ChemPhysChem* **2009**, *10*, 2296–2304.
55. Arshadi, M.; Abdolmaleki, K.; Mousavinia, M.; Khalafi-Nezhad, F.; Firouzabadi, A. Degradation of methyl orange by heterogeneous Fenton-like oxidation on a nano-organometallic compound in the presence of multi-walled carbon nanotubes. *Chem. Eng. Res. Des.* **2016**, *112*, 113–121.
56. Samadi, M.T.; Shokoohi, R.; Poormohammadi, A.; Salmani, B.; Harati, R. Ciprofloxacin oxidation by magnetic Fe₃O₄/multi walled carbon nano tubes composite as an effective heterogeneous fenton catalysts. *J. Solid State Electr.* **2016**, *7*, 253–259.
57. Li, Z.; Chen, L.; Meng, S.; Guo, L.; Huang, J.; Liu, Y.; Wang, W.; Chen, X. Field and temperature dependence of intrinsic diamagnetism in graphene: Theory and experiment. *Phys. Rev. B* **2015**, *91*, 094429.
58. Zubir, N.A.; Yacou, C.; Motuzas, J.; Zhang, X.; Zhao, X.S.; Da Costa, J.C.D. The sacrificial role of graphene oxide in stabilising a Fenton-like catalyst GO-Fe₃O₄. *Chem. Commun.* **2015**, *51*, 9291–9293.
59. Zubir, N.A.; Motuzas, J.; Yacou, C.; Zhang, X.; Costa, J. Graphene oxide with zinc partially substituted magnetite (GO-Fe_{1-x}Zn_xO_y) for the UV-assisted heterogeneous Fenton-like reaction. *RSC Adv.* **2016**, *6*, 44749–44757.

60. Liu, S.-Q.; Xiao, B.; Feng, L.-R.; Zhou, S.-S.; Chen, Z.-G.; Liu, C.-B.; Chen, F.; Wu, Z.-Y.; Xu, N.; Oh, W.-C.; et al. Graphene oxide enhances the Fenton-like photocatalytic activity of nickel ferrite for degradation of dyes under visible light irradiation. *Carbon* **2013**, *64*, 197–206.
61. Teixeira, A.P.C.; Tristão, J.C.; Araujo, M.H.; Oliveira, L.C.A.; Moura, F.C.C.; Ardisson, J.D.; Amorim, C.C.; Lago, R.M. Iron: A versatile element to produce materials for environmental applications. *J. Braz. Chem. Soc.* **2012**, *23*, 1579–1593.
62. Ria, B.; Coca, O.M.; García-González, M. Evaluation of Fenton method and ozone-based processes for colour and organic matter removal from biologically pre-treated swine manure. *Chemosphere* **2014**, *117*, 193–199.
63. Yang, X.; Lee, J.; Yuan, L.; Chae, S.; Peterson, V.K.; Minett, A.I.; Yin, Y.; Harris, A.T. Removal of natural organic matter in water using functionalised carbon nanotube buckypaper. *Carbon* **2013**, *59*, 160–166.
64. Vaishnav, P.; Kumar, A.; Ameta, R.; Punjabi, P.; Ameta, S.C. Photo oxidative degradation of azure-B by sono-photo-Fenton and photo-Fenton reagents. *Arab. J. Chem.* **2014**, *7*, 981–985.
65. Thomas, C.; Mackey, M.M.; Diaz, A.A.; Cox, D.P. Hydroxyl radical is produced via the Fenton reaction in submitochondrial particles under oxidative stress: Implications for diseases associated with iron accumulation. *Redox Rep.* **2009**, *14*, 102–108.
66. Freitas, D.; Adriane, M.; Peralta, A. Microcystin-LR degradation by solar photo-Fenton, UV-A/photo-Fenton and UV-C/H₂O₂: A comparative study. *Photochem. Photobiol. Sci.* **2013**, *12*, 696–702.
67. Zhou, X.; Zhang, Y.; Wang, C.; Wu, X.; Yang, Y.; Zheng, B.; Wu, H.; Guo, S.; Zhang, J. Photo-Fenton Reaction of Graphene Oxide: A New Strategy to Prepare Graphene Quantum Dots for DNA Cleavage. *ACS Nano* **2012**, *6*, 6592–6599.
68. Calza, P.; Di Sarro, J.; Magnacca, G.; Prevot, A.B.; Laurenti, E. Use of Low-Cost Magnetic Materials Containing Waste Derivatives for the (Photo)-Fenton Removal of Organic Pollutants. *Materials* **2019**, *12*, 3942.
69. El-Ghenymy, A.; Rodríguez, R.; Brillas, E.; Oturan, N.; Oturan, M.A. Electro-Fenton degradation of the antibiotic sulfanilamide with Pt/carbon-felt and BDD/carbon-felt cells. Kinetics, reaction intermediates, and toxicity assessment. *Environ. Sci. Pollut. R* **2014**, *21*, 8368–8378.
70. Scialdone, O.; Galia, A.; Gattuso, C.; Sabatino, S.; Schiavo, B. Effect of air pressure on the electro-generation of H₂O₂ and the abatement of organic pollutants in water by electro-Fenton process. *Electrochim. Acta* **2015**, *182*, 775–780.
71. Brillas, E.; Sirés, I.; Oturan, M.A. Electro-Fenton Process and Related Electrochemical Technologies Based on Fenton's Reaction Chemistry. *Chem. Rev.* **2009**, *109*, 6570–6631.
72. Liu, Y.; Pan, J.; Zhang, J.; Tang, A.; Liu, Y. Study on Mass Transfer-Reaction Kinetics of NO Removal from Flue Gas by Using a UV/Fenton-like Reaction. *Ind. Eng. Chem. Res.* **2012**, *51*, 12065–12072.
73. Liu, T.; He, F.W.; Zhang, Y.Q. Synergistic Degradation of Acid Scarlet Dyeing Wastewater by the Ultrasound/Fenton Method. *Appl. Mech. Mater.* **2013**, *448–453*, 34–37.
74. Liu, X.; Liang, J.; Wang, X. Kinetics and reaction pathways of formaldehyde degradation using the UV-fenton method. *Water Environ. Res.* **2011**, *83*, 418–426.
75. Xi, B.; Shi, Q. Removal of thiocyanate from industrial wastewater by microwave-Fenton oxidation method. *J. Environ. Sci.* **2013**, *25*, S201–S204.
76. Cebriano, T.; Méndez, B.; Piqueras, J. Study of luminescence and optical resonances in Sb₂O₃ micro- and nanotriangles. *J. Nano-particle Res.* **2012**, *14*, 1–8.
77. Plotnikov, V.; Dem'Yanov, B.; Yeliseyev, A.; Makarov, S.; Zyryanova, A. Structural state of diamond-like amorphous carbon films, obtained by laser evaporation of carbon target. *Diam. Relat. Mater.* **2019**, *91*, 225–229.
78. Cai, X.; Cong, H.; Liu, C. Synthesis of vertically-aligned carbon nanotubes without a catalyst by hydrogen arc discharge. *Carbon* **2012**, *50*, 2726–2730.
79. Shi, S.J.; Tu, J.P.; Tang, Y.Y.; Liu, X.Y.; Gu, C.D. Enhanced cycling stability of Li[Li_{0.2}Mn_{0.54}Ni_{0.13}Co_{0.13}]O₂ by surface modification of MgO with melting impregnation method. *Electrochim. Acta* **2013**, *88*, 671–679.
80. Bai, Y.; Liu, Y.; Li, Y.; Ling, L.; Wu, F.; Wu, C. Mille-feuille shaped hard carbons derived from polyvinylpyrrolidone via environmentally friendly electrostatic spinning for sodium ion battery anodes. *RSC Adv.* **2017**, *7*, 5519–5527.
81. Akhi, Y.; Irani, M.; Olya, M.E. Simultaneous degradation of phenol and paracetamol using carbon/MWCNT/Fe₃O₄ composite nanofibers during photo-like-Fenton process. *J. Taiwan Institute Chem. E* **2016**, *63*, 327–335.
82. Chen, D.; Qiao, X.; Qiu, X.; Chen, J.; Jiang, R. Convenient synthesis of silver nanowires with adjustable diameters via a solvothermal method. *J. Colloid Interface Sci.* **2010**, *344*, 286–291.
83. Deng, J.; Wen, X.; Wang, Q. Solvothermal in situ synthesis of Fe₃O₄-multi-walled carbon nanotubes with enhanced heterogeneous Fenton-like activity. *Mater. Res. Bull.* **2012**, *47*, 3369–3376.
84. Thomas, R.; Davidson, P.; Badanina, E. Water-and Boron-Rich Melt Inclusions in Quartz from the Malkhan Pegmatite, Transbaikalia, Russia. *Minerals* **2012**, *2*, 435–458.
85. Xu, B.; Maimaiti, H.; Wang, S.; Awati, A.; Wang, Y.; Zhang, J.; Chen, T. Preparation of coal-based graphene oxide/SiO₂ nanosheet and loading ZnO nanorod for photocatalytic Fenton-like reaction. *Appl. Surf. Sci.* **2019**, *498*, 143835.
86. Zhang, M.; Sheng, G.; Fu, J. Novel preparation of nanosized ZnO-SnO₂ with high photocatalytic activity by homogeneous coprecipitation method. *Mater. Lett.* **2005**, *59*, 3641–3644.
87. Song, S.; Wang, Y.; Shen, H.; Zhang, J.; Mo, H.; Xie, J.; Zhou, N.-L.; Shen, J. Ultrasmall Graphene Oxide Modified with Fe₃O₄ Nanoparticles as a Fenton-Like Agent for Methylene Blue Degradation. *ACS Appl. Nano Mater.* **2019**, *2*, 7074–7084.
88. Xiang, Y.; Xin, L.; Hu, J.; Li, C.; Wei, X. Advances in the Applications of Graphene-Based Nanocomposites in Clean Energy Materials. *Crystals* **2021**, *11*, 47.

89. Gardner, P.T.; McPhail, D.B.; Duthie, G.G. Electron spin resonance spectroscopic assessment of the antioxidant potential of teas in aqueous and organic media. *J. Sci. Food Agric.* **1998**, *76*, 257–262.
90. Peng, J.; Zhang, Y.; Zhang, C.; Miao, D.; Li, J.; Liu, H.; Wang, L.; Gao, S. Removal of triclosan in a Fenton-like system mediated by graphene oxide: Reaction kinetics and ecotoxicity evaluation. *Sci. Total. Environ.* **2019**, *673*, 726–733.
91. Zhu, L.; Ji, J.; Liu, J.; Mine, S.; Matsuoka, M.; Zhang, J.; Xing, M. Designing 3D-MoS₂ Sponge as Excellent Cocatalysts in Advanced Oxidation Processes for Pollutant Control. *Angew. Chem.* **2020**, *5933*, 13968–13976.
92. Campi, G.; Innocenti, D.; Bianconi, A. CDW and Similarity of the Mott Insulator-to-Metal Transition in Cuprates with the Gas-to-Liquid-Liquid Transition in Supercooled Water. *J. Supercond. Novel Magn.* **2015**, *28*, 1355–1363.
93. Sciortino, L.; Alessi, A.; Messina, F.; Buscarino, G.; Gelardi, F.M. Structure of the FeBTC Metal–Organic Framework: A Model Based on the Local Environment Study. *J. Phys. Chem. C* **2015**, *119*, 7826–7830.
94. Zhou, W.; Wachs, I.E.; Kiely, C.J. Nanostructural and chemical characterization of supported metal oxide catalysts by aberration corrected analytical electron microscopy. *Curr. Opin. Solid State Mater. Sci.* **2012**, *16*, 10–22.
95. Kim, C.S.; Lee, J.Y.; Yong, T.K. Atomic crystal structure of ordered In₃Sb₁Te₂ ternary alloy studied by high-resolution transmission electron microscopy. *Appl. Phys. Lett.* **2012**, *100*, 151903.
96. Yao, Y.; Qin, J.; Cai, Y.; Wei, F.; Lu, F.; Wang, S. Facile synthesis of magnetic ZnFe₂O₄–reduced graphene oxide hybrid and its photo-Fenton-like behavior under visible irradiation. *Environ. Sci. Pollut. Res.* **2014**, *21*, 7296–7306.
97. Zubir, N.A.; Yacou, C.; Motuzas, J.; Zhang, X.; Diniz, D. Structural and functional investigation of graphene oxide-Fe₃O₄ nanocomposites for the heterogeneous Fenton-like reaction. *Sci. Rep.* **2018**, *4*, 1–8.
98. Sourty, E.; Bavel, S.V.; Lu, K.; Guerra, R.; Bar, G.; Loos, J. High-Angle Annular Dark Field Scanning Transmission Electron Microscopy on Carbon-Based Functional Polymer Systems. *Mater. Trans.* **2009**, *15*, 251–258.
99. Li, R.; Zhang, D.; Zhou, Y.; Wang, X.; Guo, G. Synthesis and characterization of a novel binuclear iron phthalocyanine/reduced graphene oxide nanocomposite for non-precious electrocatalyst for oxygen reduction. *Sci. China Ser. B Chem.* **2016**, *59*, 746–751.
100. Qin, J.C.; Lian, C.; Yao, Y.J.; Chen, H.; Wang, S.B. Iron encapsulated in boron and nitrogen codoped carbon nanotubes as synergistic catalysts for Fenton-like reaction. *Water Res.* **2016**, *101*, 281–291.
101. Yu, G.; Matthias, G.; Cychosz, T.K. Quenched solid density functional theory method for characterization of mesoporous carbons by nitrogen adsorption. *Carbon* **2012**, *50*, 1583–1590.
102. Wu, X.; Cao, R.; Hu, J.; Zhou, C.; Fu, L.; Wei, X. Graphene-supported iron-based composites: A review of applications to wastewater treatment. *Desalin. Water Treat.* **2020**, *181*, 300–320.
103. Wang, W.; Qi, H.; Xiao, K.; Liang, Z. Electrostatic self-assembly of Fe₃O₄/GO nanocomposites and their application as an efficient Fenton-like catalyst for degradation of rhodamine B. *Mater. Res. Express* **2018**, *5*, 035510.
104. Garriques, L.N.; Frokjaer, S.; Carpenter, J.F.; Brange, J. The Effect of Mutations on the Structure of Insulin Fibrils Studied by Fourier Transform Infrared (FTIR) Spectroscopy and Electron Microscopy. *J. Pharm. Sci.* **2002**, *91*, 2473–2480.
105. Zhou, P.; Zhang, J.; Xiong, Z.; Liu, Y.; Huo, X.; Cheng, X.; Li, W.; Cheng, F.; Zhang, Y. C60 Fullerol promoted Fe(III)/H₂O₂ Fenton oxidation: Role of photosensitive Fe(III)-Fullerol complex. *Appl. Catal. B Environ.* **2020**, *265*, 118264.
106. Fang, D.; Yu, Y.Q.; Xu, Z.H.; Liang, J.G.; Zhou, L.X. Enhanced catalytic performance of β-FeOOH by coupling with single-walled carbon nanotubes in a visible-light-Fenton-like process. *Sci. Eng. Compos. Mater* **2018**, *25*, 9–15.
107. Yu, L.; Chen, J.; Liang, Z.; Xu, W.; Chen, L.; Ye, D. Degradation of phenol using Fe₃O₄-GO nanocomposite as a heterogeneous photo-Fenton catalyst. *Sep. Purif. Technol.* **2016**, *171*, 80–87.
108. García-Rodríguez, O.; BaUelos, J.A.; Rico-Zavala, A.; Godínez, L.; Rodríguez-Valadez, F. Electrocatalytic Activity of Three Carbon Materials for the In-situ Production of Hydrogen Peroxide and Its Application to the Electro-Fenton Heterogeneous Process. *Int. J. Chem. React. Eng.* **2016**, *14*, 843–850.
109. Ai, Z.; Xiao, H.; Mei, T.; Liu, J.; Zhang, L.; Deng, K.; Qiu, J. Electro-Fenton Degradation of Rhodamine B Based on a Composite Cathode of Cu₂O Nanocubes and Carbon Nanotubes. *J. Phys. Chem. C* **2008**, *112*, 11929–11935.
110. Akerdi, A.G.; Es'Haghzade, Z.; Bahrami, S.; Arami, M. Comparative study of GO and reduced GO coated graphite electrodes for decolorization of acidic and basic dyes from aqueous solutions through heterogeneous electro-Fenton process. *J. Environ. Chem. Eng.* **2017**, *5*, 2313–2324.
111. Yang, Y.; Zhang, X.; Chen, Q.; Li, S.; Chai, H.; Huang, Y. Ultrasound-Assisted Removal of Tetracycline by a Fe/N-C Hybrids/H₂O₂ Fenton-like System. *ACS Omega* **2018**, *3*, 15870–15878.
112. Zhang, H.; Gao, H.; Cai, C.; Zhang, C.; Chen, L. Decolorization of Crystal Violet by ultrasound/heterogeneous Fenton process. *Water Sci. Technol.* **2013**, *68*, 2515–2520.
113. Wang, N.; Zheng, T.; Jiang, J.P.; Lung, W.S.; Miao, X.L.; Wang, P. Pilot-scale treatment of p-Nitrophenol wastewater by microwave-enhanced Fenton oxidation process: Effects of system parameters and kinetics study-ScienceDirect. *Chem. Eng. J.* **2014**, *239*, 351–359.
114. Yu, Y.; Wang, P.; Shi, S.; Yuan, L. Microwave enhanced Fenton-like process for the treatment of high concentration pharmaceutical wastewater. *J. Hazard. Mater.* **2009**, *168*, 238–245.
115. Liu, S.-T.; Huang, J.; Ye, Y.; Zhang, A.-B.; Pan, L.; Chen, X.-G. Microwave enhanced Fenton process for the removal of methylene blue from aqueous solution. *Chem. Eng. J.* **2013**, *215–216*, 586–590.
116. Berhe, D.M. Batch Sorption Experiments: Langmuir and Freundlich Isotherm Studies for the Adsorption of Textile Metal Ions onto Teff Straw (*Eragrostis tef*) Agricultural Waste. *J. Thermodyn.* **2013**, *2013*, 9–17.

-
117. Sadaf, S.; Bhatti, H.N. Batch and fixed bed column studies for the removal of Indosol Yellow BG dye by peanut husk. *J. Taiwan Inst. Chem. Eng.* **2014**, *45*, 541–553.
 118. Akl, M.; Awwad, N.; Salem, M. Kinetics and thermodynamics of Cd(II) biosorption onto loquat (*Eriobotrya japonica*) leaves. *J. Saudi Chem. Soc.* **2014**, *18*, 486–493.
 119. Qi, J.; Hou, Y.; Hu, J.; Ruan, W.; Xiang, Y.; Wei, X. Decontamination of methylene Blue from simulated wastewater by the mesoporous rGO/Fe/Co nanohybrids: Artificial intelligence modeling and optimization. *Mater. Today Commun.* **2020**, *24*, 100709.
 120. Backov, R.; Lee, C.M.; Khan, S.R.; Mingotaud, C.; Fanucci, A.G.E.; Talham, D.R. Calcium Oxalate Monohydrate Precipitation at Phosphatidylglycerol Langmuir Monolayers. *Langmuir* **2000**, *16*, 6013–6019.
 121. Ngwabebhoh, F.A.; Erdem, A.; Yildiz, U. Synergistic removal of Cu(II) and nitrazine yellow dye using an eco-friendly chi-tosan-montmorillonite hydrogel: Optimization by response surface methodology. *J. Appl. Polym. Sci.* **2016**, *29*, 133.
 122. Wang, W.; Tian, G.; Zong, L.; Zhou, Y.; Kang, Y.; Wang, Q.; Wang, A. From illite/smectite clay to mesoporous silicate adsorbent for efficient removal of chlortetracycline from water. *J. Environ. Sci.* **2017**, *51*, 31–43.
 123. Bhaumik, R.; Mondal, N.K.; Das, B.K.; Roy, P.; Pal, K.C.; Das, C.K.; Baneerjee, A.; Datta, J.K. Eggshell Powder as an Adsorbent for Removal of Fluoride from Aqueous Solution: Equilibrium, Kinetic and Thermodynamic Studies. *J. Chem.* **2015**, *9*, 1457–1480.
 124. Fan, M.; Li, T.; Hu, J.; Cao, R.; Wei, X.; Shi, X.; Wenqian, R. Artificial Neural Network Modeling and Genetic Algorithm Optimization for Cadmium Removal from Aqueous Solutions by Reduced Graphene Oxide-Supported Nanoscale Zero-Valent Iron (nZVI/rGO) Composites. *Materials* **2017**, *10*, 544.
 125. Wang, S. A Comparative study of Fenton and Fenton-like reaction kinetics in decolourisation of wastewater. *Dye. Pigment.* **2008**, *76*, 714–720.
 126. Cao, R.; Fan, M.; Hu, J.; Ruan, W.; Xiong, K.; Wei, X. Optimizing Low-Concentration Mercury Removal from Aqueous Solutions by Reduced Graphene Oxide-Supported Fe₃O₄ Composites with the Aid of an Artificial Neural Network and Genetic Algorithm. *Materials* **2017**, *10*, 1279.
 127. Ifelebuegu, A.O.; Ezenwa, C.P. Removal of Endocrine Disrupting Chemicals in Wastewater Treatment by Fenton-Like Oxidation. *Water Air Soil Pollut.* **2010**, *217*, 213–220.
 128. Ruan, W.; Shi, X.; Hu, J.; Hou, Y.; Fan, M.; Cao, R.; Wei, X. Modeling of Malachite Green Removal from Aqueous Solutions by Nanoscale Zerovalent Zinc Using Artificial Neural Network. *Appl. Sci.* **2017**, *8*, 3.
 129. Cheng, T.; Chen, C.; Rong, T. Competitive Adsorption of Cu, Ni, Pb, and Cd from Aqueous Solution Onto Fly Ash-Based Linde F(K) Zeolite. *Iran. J. Chem. Chem. Eng.* **2018**, *37*, 61–72.