

Review

Gas Sensors Based on Electrospun Nanofibers

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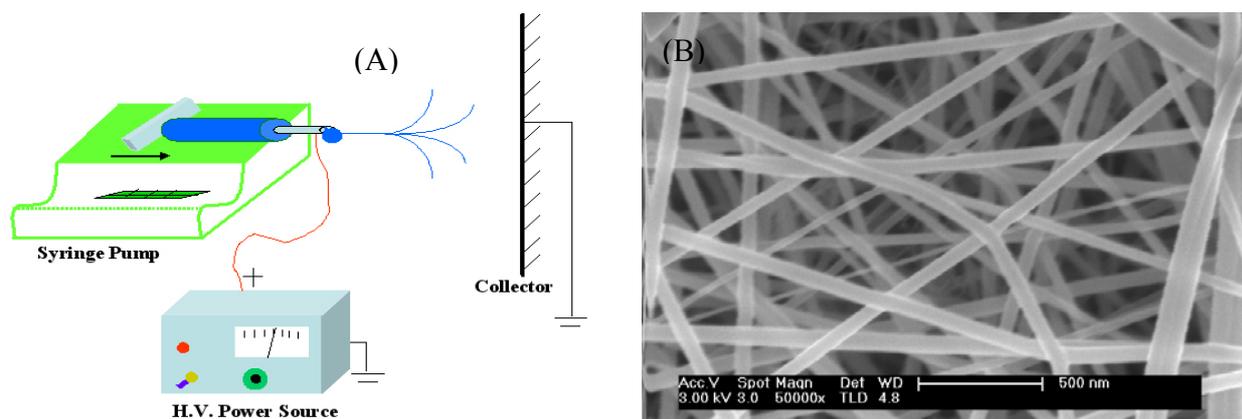
Abstract: Nanofibers fabricated via electrospinning have specific surface approximately one to two orders of the magnitude larger than flat films, making them excellent candidates for potential applications in sensors. This review is an attempt to give an overview on gas sensors using electrospun nanofibers comprising polyelectrolytes, conducting polymer composites, and semiconductors based on various sensing techniques such as acoustic wave, resistive, photoelectric, and optical techniques. The results of sensing experiments indicate that the nanofiber-based sensors showed much higher sensitivity and quicker responses to target gases, compared with sensors based on flat films.

Keywords: Gas sensors; electrospinning; nanofibers; acoustic wave; resistive; photoelectric; optical.

1. Introduction

A gas sensor is a device which detects the presence of different gases in an area, especially those gases which might be harmful to humans or animals. The development of gas sensor technology has received considerable attention in recent years for monitoring environmental pollution. It is well-known that chemical gas sensor performance features such as sensitivity, selectivity, time response, stability, durability, reproducibility, and reversibility are largely influenced by the properties of the sensing materials used [1-3]. Many kinds of materials such as polymers [4,5], semiconductors [6,7], carbon graphites [8,9], and organic/inorganic composites [10,11] have been used as sensing materials to detect the targeted gases based on various sensing techniques and principles. It is worth noting that the sensitivity of chemical gas sensors is strongly affected by the specific surface of sensing materials [12,13]. A higher specific surface of a sensing material leads to a higher sensor sensitivity, therefore many techniques [14-16] have been adopted to increase the specific surface of sensing films with fine structures, especially to form the nanostructures, taking advantage of the large specific surface of nanostructured materials.

Figure 1. (A) A schematic of electrospinning process. (B) A SEM image of typical electrospun fibers.



Electrospinning is an efficient, relatively simple and low cost way to produce polymer and composite fibers with diameters ranging from several nanometers to a few micrometers by applying a high voltage to a polymer solution or melt ejected from a micro-syringe pump [17-19]. The ultrafine fibers produced via electrospinning are assembled as a three dimensional structured fibrous membrane with controllable pore structure and high specific surface area. A schematic of the electrospinning process and a typical SEM image of electrospun fibers are shown in Figure 1. Electrospinning is a special case of the electrospray process which uses an electrostatic field to form and accelerate liquid jets from the tip of a needle [20,21]. The surface of a hemispherical liquid drop suspended in equilibrium will be distorted into a conical shape in the presence of an external electric field. This distortion is caused by a balancing of the repulsive force resulting from the induced charge distribution on surface of drop with the surface tension of liquid [22]. A stable jet of liquid can be ejected and accelerated if the applied voltage exceeds this critical voltage. The jet breaks up into droplets as a

result of the longitudinal Rayleigh instability caused by surface tension in the case of low viscosity liquids. This process is known as electrospray for applications requiring aerosols composed of sub-micron droplets with narrow distribution [23,24]. For high viscosity liquids, polymer solutions or melts, the jet does not break up, but travels as a jet to the grounded target. A transverse instability or splaying of the jet into two or more smaller jets is observed due to the radial charge repulsion [18]. This process is termed as “electrospinning” and it produces the polymer and composite fibers with diameters on the sub-micron scale [17].

Table 1. Types of electrospun nanofibers based gas sensors.

Types	Materials	Array of fibers	Fiber diameter	Gases tested	Operating temperature (°C)	Detection limit	Ref.	
Acoustic wave	PAA-PVA	Nonwoven	100-400 nm	NH ₃	RT	50 ppm	[47]	
	PAA	Nonwoven	1-7 μm	NH ₃	RT	130 ppb	[48]	
	PEI-PVA	Nonwoven	100-600 nm	H ₂ S	RT	500 ppb	[50]	
Resistive	HCSA-PANI/PEO	Single	100-500 nm	NH ₃	RT	500 ppb	[52]	
	PDPA-PMMA	Nonwoven	~400 nm	NH ₃	RT	1 ppm	[53]	
	PANI	Nonwoven	0.3-1.5 μm	Amines	RT	100 ppm	[54]	
	PMMA-PANI	Nonwoven	250-600 nm	(C ₂ H ₅) ₃ N	RT	20 ppm	[55]	
	HCSA-PANI	Single	20-150 nm	Alcohols	RT	No data	[56]	
	HCSA-POT/PS	Nonwoven	0.2-1.9 μm	H ₂ O	RT	No data	[57]	
	MWCNT/nylon	Nonwoven	110-140 nm	VOCs	RT	No data	[58]	
	CB-PECH, PEO, PIB, PVP	Oriented	~3 μm	CH ₃ OH	RT	1000 ppm	[59]	
				C ₅ H ₁₀ Cl ₂		5 ppm		
				C ₆ H ₅ CH ₃		250 ppm		
				C ₂ HCl ₃		500 ppm		
		TiO ₂	Nonwoven	200-500 nm	NO ₂	150-400 °C	500 ppb	[62]
		TiO ₂	Nonwoven	120-850 nm	CO, NO ₂	300-400 °C	50 ppb	[63]
		LiCl-TiO ₂	Nonwoven	150-260 nm	H ₂ O	RT	11%	[64]
		SnO ₂	Single	700 nm	H ₂ O	RT	No data	[65]
	SnO ₂	Nonwoven	~100 nm	C ₂ H ₅ OH	330 °C	10 ppb	[66]	
	MWCNT/SnO ₂	Nonwoven	300-800 nm	CO	RT	47 ppm	[67]	
	WO ₃	Nonwoven	20-140 nm	NH ₃	350 °C	50 ppm	[68]	
	SrTi _{0.8} Fe _{0.2} O _{3-δ}	Nonwoven	~100 nm	CH ₃ OH	400 °C	5 ppm	[69]	
Photoelectric	Co-ZnO	Nonwoven	50-400 nm	O ₂	RT	0.32 Torr	[70]	
Optical	Oxides-PAN	Nonwoven	50-200 nm	CO ₂	RT	700 ppm	[71]	

Abbreviations in Table 1: PAA (polyacrylic acid); PVA (polyvinyl alcohol); PEI (polyethyleneimine); HCSA (10-camphorsulfonic acid); PANI (polyaniline); PEO (polyethylene oxide); PDPA (polydiphenylamine); PMMA (polymethyl methacrylate); POT (polyo-toluidine); PS (polystyrene); MWCNT (multi-walled carbon nano-tube); CB (carbon black); PECH (polyepichlorohydrin); PIB (polyisobutylene); PVP (polyvinylpyrrolidone); PAN (polyacrylonitrile); VOC (volatile organic compound); RT (room temperature).

To date, it is believed that more than one hundred different polymers have been successfully electrospun into nanofibers by this technique. Generally, the electrospun fibers are deposited on a fixed collector in a three-dimensional nonwoven membrane structure with a wide range of fiber diameter distribution. Extensive research on electrospinning show that the fiber collectability,

uniformity of fibers, average fiber diameter, fiber diameter distribution, and fiber porosity are strongly affected by the solution properties [19,25] and processing parameters [26,27]. More recently, aligned electrospun fibers were obtained by using a rotating or pre-patterned collector [28,29]. Besides polymer and composite nanofibers, various inorganic semiconductor nanofibers are fabricated by sintering composite electrospun nanofibers in an elevated temperature [30-32]. The nanofibrous membranes with specific surface approximately 1-2 orders of the magnitude larger than the flat films have been targeted with applications in filtration [33], tissue engineering [34], drug delivery system [35], catalysis [36], dye-sensitized solar cells [37], battery separators [38], etc.

Electrospun fibers with controllable membrane thickness, fine structures, diversity of materials, and large specific surface are expected to be an ideal candidate as the structure of sensing materials. So far, many attempts (listed in Table 1) are carried out to prepare ultrasensitive gas sensors to detect vapors of NH₃, H₂S, CO, NO₂, O₂, CO₂, moisture, and VOCs (CH₃OH, C₂H₅OH, C₅H₁₀Cl₂, C₆H₅CH₃, C₄H₈O, CHCl₃, C₂H₂Cl₂, C₃H₆O, C₃H₇NO, C₂HCl₃, N₂H₄, (C₂H₅)₃N, C₆H₁₄, etc.) with new and improved detection limits using electrospun nanofibrous membranes as sensing structures. The types of prepared gas sensors mainly include acoustic wave, resistive, photoelectric, and optical gas sensors. Electrospun fibers with polyelectrolyte components, conducting polymer composites, and semiconductors are successfully applied as gas sensing interfaces with the fiber arrangement of single fiber, oriented fibers, or nonwoven membranes at room or elevated operating temperature.

The fabrication and applications of nanofibers via electrospinning have been reviewed previously [39-42]. However, to the best of our knowledge, there have been no comprehensive reports on gas sensors based on electrospun nanofibers. In this review, we summarize the applications of electrospun fibers as gas sensing materials in recent years and call the attention of both scientists and engineers to this interesting research area.

2. Sensors

2.1. Acoustic Wave Gas Sensors

Quartz crystal microbalances (QCMs), one of a broad class of acoustic wave techniques, have been applied to detect small mass uptake on the order of nanograms when used for the characterization of thin layers, fluids, and gases due to their high mass sensitivity and durability [43]. According to Sauerbrey Equation (1) [44]:

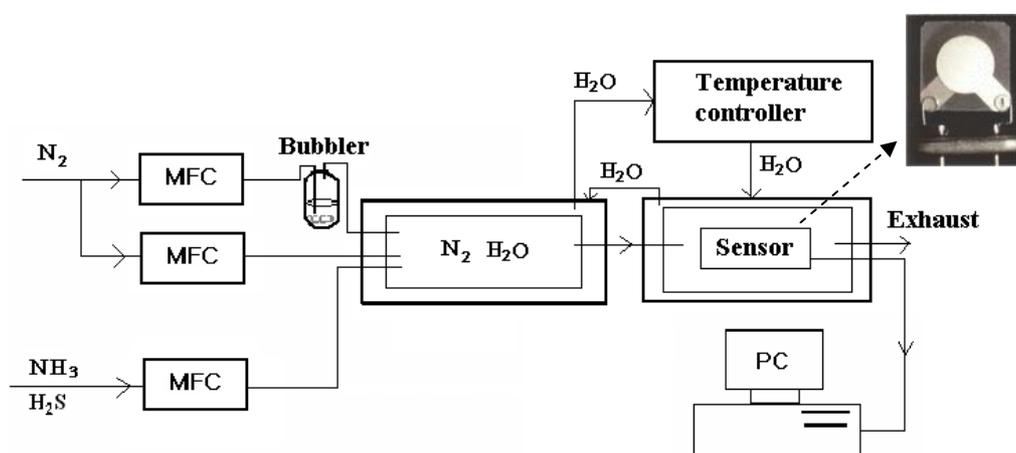
$$\Delta f = - f_0^2 \Delta m / NS\rho \quad (1)$$

where Δf is the frequency change, f_0 is the frequency of the quartz plate, Δm is the weight of applied coating, N is the unique constant of quartz crystal, S is the surface area of coated electrode, and ρ is the density of the quartz plate. QCM technology has been widely applied to the area of environmental monitoring, gas sensors, pressure sensors, and chemical sensors [45]. The sensitivity of QCM gas sensors is highly increased in the presence of a coating that interacts with the target gas [46].

Ding *et al* [47] first applied electrospun nanofibrous membranes on QCM electrodes as highly sensitive gas sensors in 2004. A series of composite fibers of polyacrylic acid (PAA) and poly(vinyl alcohol) (PVA) with diameters of 100-400 nm containing various weight percentages of PAA to PVA

were deposited on QCM via electrospinning and characterized with regards to their morphology and sensitivity to NH_3 . PAA, a weak anionic polyelectrolyte, interacts with NH_3 gas. However, it is difficult to obtain nanofibers by electrospinning a pure PAA solution because of the formation of strong hydrogen bonds between water and the large amount of carboxyl groups in PAA. Therefore, the water-soluble PVA was introduced as a template to form composite fibers. Subsequently, heat treatment was employed to induce the crosslinking between PAA and PVA for reinforcing the fiber anti-water solubility. Sensing experiments were carried out at room temperature (RT) in a flow-type gas testing system (Figure 2) by measuring the resonance frequency shifts of QCM which due to the additional mass loading. The flow rates of dry N_2 , wet N_2 , and ammonia were kept constant by mass flow controllers (MFCs).

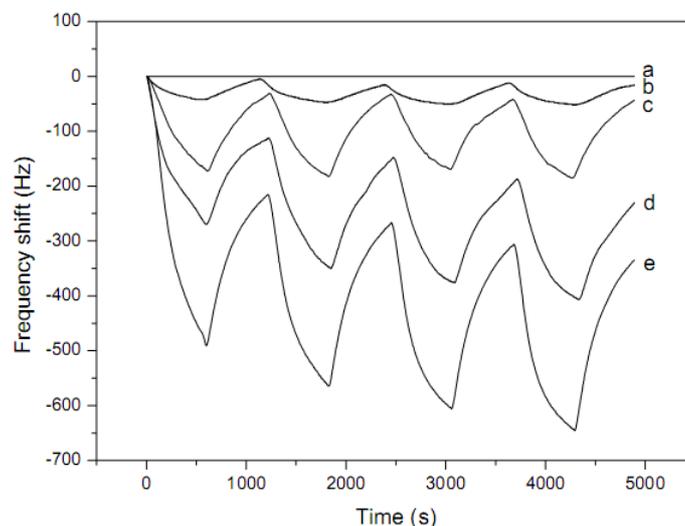
Figure 2. Schematic of a flow-type gas testing system for polyelectrolyte nanofiber gas sensors. (reproduced with permission from ref. [47]).



The fiber sensors showed fast responses to NH_3 after adding the PAA component in fibers (Figure 3). The average resonance frequency shifts of fibers coated QCM sensors with 11, 18, 25, and 33 wt% of PAA to PVA were 40, 150, 240, and 380 Hz, respectively. The sensing results indicated the content of PAA component, concentration of NH_3 , and relative humidity in gas testing chamber were all important parameters influencing the sensitivity of QCM sensors for NH_3 . Compared with the casted film-coated QCM sensors, the nanofiber-coated QCM sensors showed much higher sensitivity.

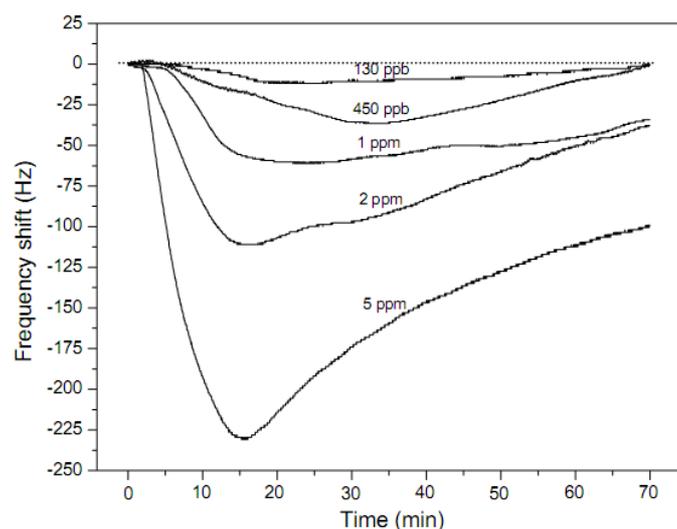
In order to further increase the fiber sensor sensitivity towards ammonia, microsized pure PAA fibers were deposited on the QCM electrodes with various solvent compositions of H_2O and ethanol [48]. The connectivity between the electrode and fibrous membranes was designed to be enhanced by the incomplete evaporation of solvent in fibers with a short tip-to-electrode distance of 5 cm.

Figure 3. Response of nanofibers coated QCM sensors with various weight percentage of PAA to PVA exposed to 50 ppm of NH_3 at the relative humidity of 55%. (a) 0 wt%; 11 wt%; 18 wt%; (d) 25 wt%; (e) 33 wt%. (reproduced with permission from ref. [47]).



The results of sensing experiments indicated that the sensitivity of the fibrous PAA membrane-coated QCM sensors was four times higher than that of the casted PAA film-coated QCM sensors. Moreover, the fiber-coated QCM sensors with small average fiber diameters had higher sensitivity towards low concentrations of ammonia, as low as 130 ppb at a relative humidity of 40 % (Figure 4).

Figure 4. Response of the fibers coated QCM sensors with 10,000 Hz coating load exposed to various concentration of ammonia. The PAA fibrous membrane was obtained from the solvent of H_2O and the relative humidity was 40% (reproduced with permission from ref. [48]).



However, the fiber QCM sensors could not detect low concentrations of ammonia gas at very low relative humidity. Therefore, the pre-sorbed moisture in fibrous membranes was proven to be the key

factor affecting the sensitivity of QCM sensors for ammonia. Furthermore, the fiber-coated QCM sensors were found to have cross-sensitivity to ammonia and H₂O, and the sensitivity of these sensors to moisture was much higher than to ammonia. Hence, they have a strong potential application in humidity sensors. Nanoporous films consisting of weak polyelectrolytes and TiO₂ nanoparticles fabricated by a layer-by-layer (LBL) self-assembly method were applied on the electrode of QCM as ammonia gas sensors were prepared by Kim *et al* [49]. Compared to the sensing performance of nanoporous sensing films (60 Hz response to 10 ppm ammonia), the fiber-based gas sensors showed much higher sensitivity (230 Hz response to 5 ppm ammonia). However, the polyelectrolyte fibers are not as stable as the nanoporous hybrid LBL films in high humid atmospheres. Further design to enhance the insolubility of the fiber in water is needed.

Additionally, a novel H₂S gas sensor [50] with a detection limit of 500 ppb was successfully fabricated by electrospinning deposition on QCM of nanofibrous polyethyleneimine (PEI)/PVA membranes with diameters of 100-600 nm as sensing materials. The morphology of sensing electrospun nanofibrous membranes was strongly influenced by the molar composition ratios of PEI to PVA. The sensor sensitivity was demonstrated to be affected by the morphology of sensing materials – it increases as the specific surface and porosity increase. Moreover, the ambient temperature and relative humidity were found to play an important role in the sensor sensitivity. The sensor sensitivity decreased with the conditions of high temperature and low relative humidity. The selectivity of fiber-based sensors was examined with H₂S, CH₃SH, NH₃ and benzene. The sensors were insensitive to NH₃ and benzene, but sensitive to H₂S and CH₃SH. The frequency shift of sensors response to 1 ppm of H₂S, CH₃SH, NH₃ and benzene, is 40, 27, 1 and 0.3 Hz, respectively, thus relatively good selectivity of the fiber sensors towards the sulfide gases was found.

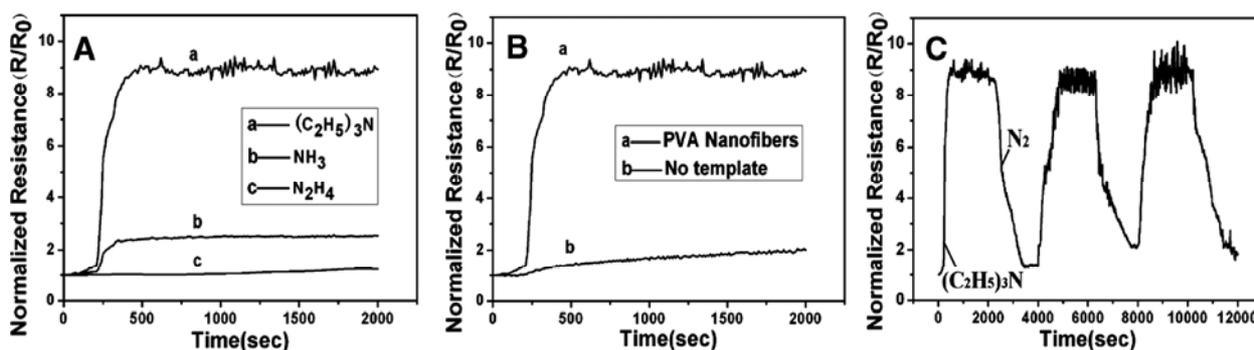
2.2. Resistive Gas Sensors

Various kinds of conducting polymers such as polyaniline (PANI), poly(diphenylamine) (PDPA), polypyrrole, polythiophene, etc. are currently used in various applications including metallic interconnects in circuits, chemical sensors, and electromagnetic radiation shielding [51]. The conductivity of such polymers results from the existence of charge carriers and from the ability of those charge carriers to move along the bonds of the polymer chains. These conducting polymers show chemical selectivity, which allows them to act as excellent materials for the immobilization of gas molecules, and exhibit highly reversible redox behavior with a distinguishable chemical memory.

In 2004, Craighead *et al.* [52] reported the fabrication of ammonia gas sensors via a scanned-tip electrospinning method using a single 10-camphorsulfonic acid (HCSA) doped PANI/poly(ethylene oxide) (PEO) nanofiber with diameter of 100-500 nm on gold electrodes. The well-defined single fiber geometry allowed for the characterization of the fiber material and the sensor response. The sensors showed a rapid and reversible resistance change upon exposure to NH₃ gas at concentrations as low as 500 ppb through the protonation and deprotonation of PANI. The response time of nanofiber sensors was understood by considering the diffusion of ammonia into the fiber and the reaction of ammonia with doped PANI. The response times of sensors with various diameters corresponded to radius-dependent differences in the diffusion time of ammonia gas into the fibers.

Another ammonia gas sensor with a detection limit of 1 ppm was prepared by Manesh *et al.* [53] using electrospun PDPA/poly(methyl methacrylate) (PMMA) nanofibers as sensing materials. The changes in resistance of the nonwoven membrane showed linearity with the concentration of ammonia in the range of 10–300 ppm. Furthermore, the detection target was expanded from ammonia to other amines (Figure 5) by Gong *et al.* [54] using PANI nanotubes prepared using electrospun PVA fiber mats membrane as the template. The high surface areas, small diameter, and porous nature of the PANI nanotubes gave significantly better performance with regards to both sensitivity and time response. It was observed that the responses follow the orders: $(C_2H_5)_3N > NH_3 > N_2H_4$. Compared with PANI prepared without a template, the PANI nanotubes showed higher sensitivity and quicker response to $(C_2H_5)_3N$. Additionally, the reversible circulation response change of PANI nanotubes has a reasonable reproducibility.

Figure 5. (A) Resistance change of PANI nanotubes exposed to 100 ppm of different gases (NH_3 , N_2H_4 , and $(C_2H_5)_3N$). (B) Response of PANI prepared by using PVA fiber mats as the template and without a template upon exposure to 100 ppm of $(C_2H_5)_3N$. (C) The reversible circulation response change of PANI nanotubes upon exposure to 100 ppm of $(C_2H_5)_3N$. (reproduced with permission from ref. [54]).



Li *et al.* [55] obtained coaxial PANI/PMMA composite nanofibers using the electrospinning technique and an *in situ* polymerization method. The electrical responses of the gas sensors based on these composite nanofibers towards triethylamine (TEA) vapor were investigated at RT. It was revealed that the sensors showed a sensing magnitude as high as 77 towards TEA vapor of 500 ppm. In addition, the responses were linear, reversible and reproducible towards TEA vapor concentrations ranging from 20 to 500 ppm. Furthermore, it was found that the concentration of doping acids only led to changes in resistance of the sensor, but did not affect its sensing characteristics. The gas sensor with toluene sulfonic acid as the doping acid exhibited the highest sensing magnitude, which is explained by taking into account of the sensing mechanism and the interactions of doping acids with TEA vapor.

Other conducting composite electrospun fibers such as HCSA doped PANI [56], HCSA doped poly(*o*-toluidine) (POT)/polystyrene (PS) [57], and multi-walled carbon nanotubes (MWCNT) doped nylon [58] were fabricated as sensing materials to detect alcohols, moisture, hexane, acetone, ethyl acetate, dichloromethane, trichloromethane, tetrahydrofuran, and toluene vapors by measuring the resistance changes of fibers. However, there are no detailed concentrations of target vapors described in such reports. Therefore, the effects of temperature and vapor concentration should be investigated in

the future work to fulfill the study of new sensing fibers. In addition, Tepper *et al.* [59] prepared electrospun fibers of carbon black (CB) doped polymers of poly(epichlorohydrin) (PECH), PEO, poly(isobutylene) (PIB), and poly(vinylpyrrolidone) (PVP) as sensing membranes to detect toluene, trichloroethylene, methanol, and dichloropentane vapors by measuring the electrical resistance of each element of the four-component sensor array. The sensor elements exhibited a linear response as a function of vapor concentration and response time ranged from several seconds to over 1 min.

Besides the conducting polymer composites, semiconducting metal oxides such as TiO₂, SnO₂, ZnO, WO₃, and other wide band gap metal oxides are known for their ability to detect trace concentrations of various gases in air via charge transfer interactions between the sensor and chemisorbed species that modify the sensor's resistance [60]. Kolmakov *et al.* [61] produced highly uniform SnO₂ nanowires with bulk electronic properties directed by their surface chemistry through isolating and oxidizing tin nanowires selected from a template-synthesized array. The nanowires act as sensitive, fast, stable and reproducible gas sensors that can be easily integrated into a multi-component array. Recent efforts have been focused on the development of other nanostructured sensors to achieve increased surface-to-volume ratios and reduced cross sections, offering more effective gas modulation of device resistance.

An ultrasensitive chemiresistor based on electrospun TiO₂ nanofibers [62] was fabricated by directly depositing TiO₂/poly(vinyl acetate) (PVAc) composite nanofiber mats onto interdigitated Pt electrode arrays with subsequent hot pressing at 120 °C and calcining at 450 °C (Figure 6). The hot pressed TiO₂ fibers had much higher specific surface area (138 m²/g) than that of unpressed fibers (31 m²/g) and led to an exceptionally high gas sensitivity. Exemplary results of the resistance response of one of the sensors upon cyclic exposure to NO₂ at different concentrations and operating temperatures are shown in Figure 7. One can see that the resistance increased, for the most part, monotonically with NO₂ concentrations of 500 ppb to 12.5 ppm and recovered reasonably fast, returning to the baseline resistance in dry air following each pulse of NO₂-enriched air. The sensitivity versus temperature histogram went through a peak at 300 °C. The histogram's shape resulted from the competition between slow kinetics at low temperatures and enhanced desorption at high temperatures. Unusual response patterns were observed at high NO₂ concentrations (>12.5 ppm), consistent with n to p inversion of the surface-trap limited conduction facilitated by the high surface-to-volume ratio of this material.

A further study [63] on TiO₂ nanofibers as sensing membrane to detect traces of CO and NO₂ in air was reported to demonstrate the processing-microstructure-properties correlation of ultrasensitive gas sensors. The sensor was more sensitive at 300 °C than at 400 °C, similar to the previous report in [62], demonstrating exceptional sensitivity toward NO₂ displaying a change of 405% in the sensor's resistance on exposure to 50 ppb of NO₂. By extrapolating the results to lower gas concentrations they estimated the detection limit to be 5 ppb for NO₂ and 20 ppb for CO with operating temperature at 300 °C. Additionally, a highly sensitive and stable humidity nanosensors based on LiCl doped TiO₂ electrospun nanofibers was reported by Wang *et al.* [64]. The as-prepared humidity sensor exhibited excellent sensing characteristics, including ultrafast response time (≤ 3 s) and recovery time (≤ 7 s) for measuring relative humidity in a wide range of 11-95% in air at RT with the impedance changing from 107 to 104 Ω . Moreover, the nanosensor has good reproducibility, linearity, and stability.

Figure 6. (a) Schematic diagram of the processing steps used to fabricate TiO₂ nanofiber mats on Al₂O₃ substrates with interdigitated Pt electrode arrays. (b) Scanning confocal laser micrograph of a calcined TiO₂ nanofiber mat on top of the Al₂O₃ substrate (dark region) and Pt electrode (bright region). (c) Optical micrographs of gas sensor test devices (10 mm × 15 mm) with TiO₂ nanofiber mats after different processing steps. (reproduced with permission from ref. [62]).

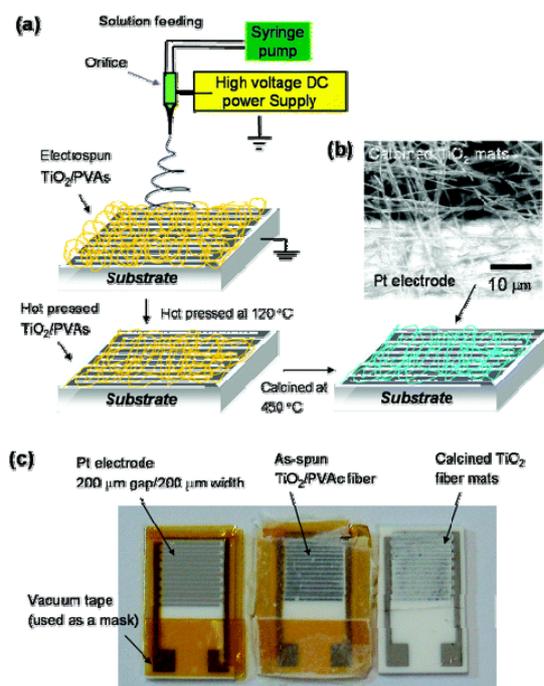
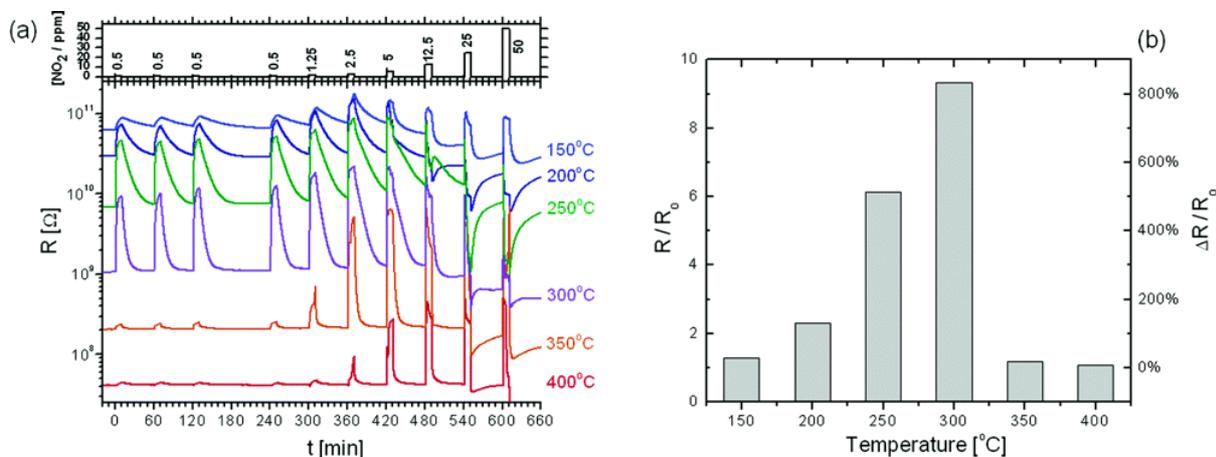


Figure 7. (a) The resistance response during cyclic exposure to 10 min pulses with increasing concentrations of NO₂ mixed in dry air at various operating temperatures. (b) Sensitivity versus temperature histogram upon exposure to 500 ppb NO₂ in dry air. R_0 designates the baseline resistance in dry air. (reproduced with permission from ref. [62]).



Wang *et al* [65] reported the fabrication of a gas sensor to detect moisture and methanol gas based on a single SnO₂ nanofiber made from C₂₂H₄₄O₄Sn/PEO using electrospinning and metallorganic

deposition techniques. The sensors showed high sensitivity to both gases and response times of the complete testing system were in the range of 108-150 s for moisture, and 10-38 s for methanol gas, respectively. Nonwoven type fibrous SnO₂ gas sensors [66] for detection of ethanol were prepared on a micro hot plate by “near-field” electrospinning PVA/SnCl₄ solution with a subsequent annealing at 300, 500, and 700 °C. The fiber resistance response increased with a raise of operating temperature and attained a maximum at 330 °C, followed by a decrease with further increases of operating temperature. The SnO₂ nanofibers with an average diameter of similar to 100 nm exhibited large responses, low detection limits, fast response/recovery, and good reproducibility. The detection limit was <10 ppb and the response/recovery time towards 10 ppm ethanol was <14 s. Tao *et al.* [67] demonstrated a sensitive CO gas sensor for use at room temperature using polycrystalline SnO₂ and MWCNT doped SnO₂ nanofibers via electrospinning. The *n*-type MWCNT/SnO₂ nanofibers are able to detect CO in the concentration range of 47-547 ppm.

The fabrication and characterization of polycrystalline WO₃ nanofibers and their application for ammonia sensing were reported by Wang *et al.* [68]. Pure tungsten oxide nanofibers with controllable diameters of around 100 nm were obtained by electrospinning PVAc/tungsten isopropoxide followed by a calcinations step. The relationship between solution concentration and ceramic nanofiber morphology has been studied. It has been shown that the as-prepared tungsten oxide ceramic nanofibers have a quick response to ammonia with various concentrations at 350 °C, suggesting potential applications of the electrospun tungsten oxide nanofibers as a sensor material for gas detection. Sahner *et al.* [69] fabricated a perovskite *p*-type conducting electrospun SrTi_{0.8}Fe_{0.2}O_{3-δ} nanofibrous membrane as a model sensing material for application in methanol sensing at the operating temperature of 400 °C. When exposed to reducing gases, the sensor resistance increased, as expected for a *p*-type sensor membrane. Very good sensor characteristics such as short response and recovery times, and a stable baseline can be observed. At 10 ppm methanol, sensor response R/R_0 was observed to be 5.5 (electrospun) and 2.9 (electrosprayed). In the case of the electrospun membrane, the signal for methanol with concentrations ranging from 5 to 50 ppm in dry air was found to be reproducible.

2.3. Photoelectric Gas Sensors

Most gas sensors based on one dimensional nanosized ZnO have to be operated at an elevated temperature in order to exhibit their sensing behavior, which is significantly inconvenient for practical applications. Xie *et al.* [70] thought to use light excitation instead of increasing temperature to change conductance owing to the distinguished optical and electric performance of ZnO. To extend the photoelectric response of wide-band gap semiconductors into the visible spectral range and utilize solar energy effectively, nonwoven type Co doped ZnO nanofibers with diameter of 50-400 nm were deposited on ITO comb-like electrodes by electrospinning PVP/zinc acetate/cobalt acetate solution with a subsequent annealing at 500 °C for 3 h. The surface photocurrent (I_{SPC}) of electrospun Co-ZnO fibers was measured with the comb-like electrodes using the light source-monochromator-lock-in detection technique. To exploit the application of Co doped ZnO fibers for a photoelectric sensor, four values of oxygen pressure: 9.28, 4.25, 4.06, and 1.16 Torr were used. The sensors showed a very short response and recovery time to oxygen, as well as the good reproducibility under the illumination of a

Xe lamp. The response time, corresponding to 70% of the equilibrium value, was about 1 min. Meanwhile, the change of I_{SPC} was linear to the oxygen concentrations with a detection limit of 0.322 Torr.

2.4. Optical Gas Sensors

Optical fiber sensors have many advantages compared to traditional types of sensors such as the absence of the electromagnetic interference in sensing and electric contacts in the probe, multiplexity on a single net work, etc. Fourier transform infrared (FTIR) spectroscopy is one of the widely used optical methods to study the interaction of electromagnetic radiation in the infrared region with chemical compounds. The large band width (400–4500 cm^{-1}) and the distinct absorbance bands render the FTIR technology suitable for sensor applications. Hahn *et al.* [71] first utilized electrospun nanocomposite fiber mats as optical sensors in conjunction with FTIR spectroscopy to detect CO_2 gas. The nanocomposite fiber mats were prepared by electrospinning polyacrylonitrile (PAN) solutions containing nanoparticles including iron oxide, antimony tin oxide and zinc oxide with diameters ranging from 10 to 70 nm. The porosity increased from 70% for PAN fiber mat to an average of 86% for nanocomposite fiber mats. The absorbance spectra showed a higher sensitivity with a fiber mat, regardless of its type, than without, indicating gas adsorption on the fiber mat. The highest sensitivity was obtained with PAN/ Fe_2O_3 fiber mat. As the underlying mechanism is physisorption rather than chemisorption, the response time was short and the sensor could be used repeatedly.

3. Conclusions

The utilization of electrospun nanofibers with large specific surface as gas sensing materials has received great attention since 2004. Various gas sensors comprising electrospun nanofibers of polyelectrolytes, conducting polymer composites, and semiconducting metal oxides were successfully fabricated with ultrahigh sensitivity, very short response and recovery time, and good reversibility, reproducibility, and stability based on various sensing techniques and principles. The fine structure of electrospun fibers was proven to be an excellent candidate instead of the current widely used solid flat films to further increase the sensor sensitivity. It opens a new way to fabricate the ultrasensitive sensors. It can be foreseen that in the near future more and more ultrasensitive and practical electrospun nanofiber sensors will be fabricated based on the above reviewed publications. Furthermore, we suggest studying the specific surface area of fibrous membranes and the connectivity between the electrodes and nanofibrous membranes since they would have the large effects on sensing performance.

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References and Notes

1. Kong, J.; Franklin, N.; Zhou, C.; Chapline, M.; Peng, S.; Cho, K.; Dai, H. Nanotube molecular wires as chemical sensors. *Science* **2000**, *287*, 622-625.
2. Gopel, W.; Schierbaum, K. SnO₂ sensors current status and future prospects. *Sensor Actuator B-Chem* **1995**, *26*, 1-12.
3. Yamada, Y.; Seno, Y.; Masuoka, Y.; Yamashita, K. Nitrogen oxides sensing characteristics of Zn₂SnO₄ thin film. *Sensor Actuator B-Chem* **1998**, *49*, 248-252.
4. Slater, J.; Watt, E.; Freeman, N.; May, I.; Weir, D. Gas and vapor detection with poly(pyrrole) gas sensors. *Analyst* **1992**, *117*, 1265-1270.
5. Slater, J.; Paynter, J.; Watt, E. Multilayer conducting polymer gas sensor arrays for olfactory sensing. *Analyst* **1993**, *118*, 379-384.
6. Schierbaum, K.; Weimar, U.; Gopel, W. Comparison of ceramic, thick-film and thin-film chemical sensors based upon SnO₂. *Sens. Actuat. B-Chem* **1992**, *7*, 709-716.
7. Savage, N.; Chwierogh, B.; Ginwalla, A.; Patton, B.; Akbar, S.; Dutta, P. Composite n-p semiconducting titanium oxides as gas sensors. *Sens. Actuat. B-Chem* **2001**, *79*, 17-27.
8. Varghese, O.; Kichambre, P.; Gong, D.; Ong, K.; Dickey, E.; Grimes, C. Gas sensing characteristics of multi-wall carbon nanotubes. *Sens. Actuat. B-Chem* **2001**, *81*, 32-41.
9. Lukaszewicz, J. Carbon materials for chemical sensors: a review. *Sens. Lett.* **2006**, *4*, 53-98.
10. Lonergan, M.; Severin, E.; Doleman, B.; Beaber, S.; Grubb, R.; Lewis, N. Array-based vapor sensing using chemically sensitive, carbon black-polymer resistor. *Chem. Mater.* **1996**, *8*, 2298-2312.
11. Unde, S.; Ganu, J.; Radhakrishnan, S. Conducting polymer-based chemical sensor: characteristics and evaluation of polyaniline composite films. *Adv. Funct. Mater.* **1996**, *6*, 151-157.
12. Gopel, W. Supramolecular and polymeric structures for gas sensors. *Sens. Actuat. B-Chem* **1995**, *24*, 17-32.
13. Lee, D.; Han, S.; Huh, J.; Lee, D. Nitrogen oxides-sensing characteristics of WO₃-based nanocrystalline thick film gas sensor. *Sens. Actuat. B-Chem* **1999**, *60*, 57-63.
14. Jin, Z.; Zhou, H.; Jin, Z.; Savinell, R.; Liu, C. Application of nano-crystalline porous tin oxide thin film for CO sensing. *Sens. Actuat. B-Chem* **1998**, *52*, 188-194.
15. Dong, L.; Cui, Z.; Zhang, Z. Gas sensing properties of nano-ZnO prepared by arc plasma method. *Nanostruct. Mater.* **1997**, *8*, 815-823.
16. Matsumiya, M.; Shin, W.; Izu, N.; Murayama, N. Nano-structured thin-film Pt catalyst for thermoelectric hydrogen gas sensor. *Sens. Actuat. B-Chem* **2003**, *93*, 309-315.
17. Reneker, D.; Chun, I. Nanometer diameter fibers of polymer, produced by electrospinning. *Nanotechnology* **1996**, *7*, 216-223.
18. Doshi, J.; Reneker, D. Electrospinning process and applications of electrospun fibers. *J. Electrostat.* **1995**, *35*, 151-160.
19. Fong, H.; Chun, I.; Reneker, D. Beaded nanofibers formed during electrospinning. *Polymer* **1999**, *40*, 4585-4592.
20. Reneker, D.; Yarin, A.; Fong, H.; Koombhongse, S. Bending instability of electrically charged liquid jets of polymer solutions in electrospinning. *J. Appl. Phys.* **2000**, *87*, 4531-4547.

21. Shin, Y.; Hohman, M.; Brenner, M.; Rutledge, G. Experimental characterization of electrospinning: the electrically forced jet and instabilities. *Polymer* **2001**, *42*, 9955-9967.
22. Deitzel, J.M.; Kleinmeyer, J.; Harris, D.; Beck Tan, N.C. The effect of processing variables on the morphology of electrospun nanofibers and textiles. *Polymer* **2001**, *42*, 261-272.
23. Dejuan, L.; Delamora, J. Charge and size distributions of electrospray drops. *J. Colloid Interface Sci.* **1997**, *186*, 280-293.
24. Chen, D.; Pui, D. Experimental investigation of scaling laws for electrospaying: dielectric constant effect. *Aerosol Sci. Tech.* **1997**, *27*, 367-380.
25. Ding, B.; Kim, H.; Lee, S.; Lee, D.; Choi, K. Preparation and characterization of nanoscaled poly(vinyl alcohol) fibers via electrospinning. *Fiber. Polym.* **2002**, *3*, 73-79.
26. McCann, J.; Marquez, M.; Xia, Y. Highly porous fibers by electrospinning into a cryogenic liquid. *J. Am. Chem. Soc.* **2006**, *128*, 1436-1437.
27. Ding, B.; Li, C.; Miyauchi, Y.; Kuwaki, O.; Shiratori, S. Formation of novel 2D polymer nanowebs via electrospinning. *Nanotechnology* **2006**, *17*, 3685-3691.
28. Theron, A.; Zussman, E.; Yarin, A.L. Electrostatic field-assisted alignment of electrospun nanofibers. *Nanotechnology* **2001**, *12*, 384-390.
29. Kameoka, J.; Orth, R.; Yang, Y.; Czaplewski, D.; Mathers, R.; Coates, G.; Craighead, H. A scanning tip electrospinning source for deposition of oriented nanofibers. *Nanotechnology* **2003**, *14*, 1124-1129.
30. Ding, B.; Kim, H.; Kim, C.; Khil, M.; Park, S. Morphology and crystalline phase study of electrospun TiO₂-SiO₂ nanofibers. *Nanotechnology* **2003**, *14*, 532-537.
31. Shao, C.; Kim, H.; Gong, J.; Lee, D. A novel method for making silica nanofibers by using electrospun fibers of polyvinyl alcohol/silica composite as precursor. *Nanotechnology* **2002**, *13*, 635-637.
32. Ding, B.; Kim, C.; Kim, H.; Seo, M.; Park, S. Titanium dioxide nanofibers prepared by using electrospinning method. *Fiber. Polym.* **2004**, *5*, 105-109.
33. Hajra, M.; Mehta, K.; Chase, G. Effects of humidity, temperature, and nanofibers on drop coalescence in glass fiber media. *Sep. Purif. Technol.* **2003**, *30*, 79-88.
34. Wnek, G.; Carr, M.; Simpson, D.; Bowlin, G. Electrospinning of nanofiber fibrinogen structures. *Nano Lett.* **2003**, *3*, 213-216.
35. Kenawy, E.; Bowlin, G.; Mansfield, K.; Layman, J.; Simpson, D.; Sanders, E.; Wnek, G. Release of tetracycline hydrochloride from electrospun poly(ethylene-co-vinylacetate), poly(lactic acid), and a blend. *J. Control. Release* **2002**, *81*, 57-64.
36. Dersch, R.; Steinhart, M.; Boudriot, U.; Greiner, A.; Wendorff, J.H. Nanoprocessing of polymers: applications in medicine, sensors, catalysis, photonics. *Polym. Advan. Technol.* **2005**, *16*, 276-282.
37. Onozuka, K.; Ding, B.; Tsuge, Y.; Naka, T.; Yamazaki, M.; Sugi, S.; Ohno, S.; Yoshikawa, M.; Shiratori, S. Electrospinning processed nanofibrous TiO₂ membranes for photovoltaic applications. *Nanotechnology* **2006**, *17*, 1026-1031.
38. Choi, S.; Lee, Y.; Joo, C.; Lee, S.; Park, J.; Han, K. Electrospun PVDF nanofiber web as polymer electrolyte or separator. *Electrochim. Acta* **2004**, *50*, 339-343.

39. Huang, Z.; Zhang, Y.; Kotaki, M.; Ramakrishna, S. A review on polymer nanofibers by electrospinning and their applications in nanocomposites. *Composites Sci. Technol.* **2000**, *287*, 622-625.
40. Li, D.; Xia, Y. Electrospinning of nanofibers: reinventing the wheel? *Advan. Mater.* **2004**, *16*, 1151-1170.
41. Greiner, A.; Wendorff, J. Electrospinning: a fascinating method for the preparation of ultrathin fibers. *Angew. Chem. Int. Ed.* **2007**, *46*, 5670-5703.
42. Lu, P.; Ding, B. Applications of electropun fibers. *Recent Pat. Nanotechnol.* **2008**, *2*, 169-182.
43. Syritski, V.; Reut, J.; Opik, A.; Idla, K. Environmental QCM sensors coated with polypyrrole. *Synthetic Meth.* **1999**, *102*, 1326-1327.
44. Muratsugu, M; Ohta, F.; Miya, Y.; Hosokawa, T.; Kurosawa, S.; Kamo, N.; Ikeda, H. Quartz crystal microbalance for the detection of microgram quantities of human serum albumin: relationship between the frequency change and the mass of protein adsorbed. *Anal. Chem.* **1993**, *65*, 2933-2937.
45. Kang, K.H.; Kim, J.M.; Kim, D.K.; Jung, S.B.; Chang, J.S.; Kwon, Y.S. Effect of pH on the properties of palmitic acid LB films for gas sensors. *Sens. Actuat. B-Chem* **2001**, *77*, 293-296.
46. Gomes, M.; Nogueira, P.; Oliveira, J. Quantification of CO₂, SO₂, NH₃, and H₂S with a single coated piezoelectric quartz crystal. *Sens. Actuat. B-Chem* **2000**, *68*, 218-222.
47. Ding, B.; Kim, J.; Miyazaki, Y.; Shiratori, S. Electrospun nanofibrous membranes coated quartz crystal microbalance as gas sensor for NH₃ detection. *Sens. Actuat. B-Chem* **2004**, *101*, 373-380.
48. Ding, B.; Yamazaki, M.; Shiratori, S. Electrospun fibrous polyacrylic acid membrane-based gas sensors. *Sens. Actuat. B-Chem* **2005**, *106*, 477-483. 47.
49. Kim, J.H., Kim, S.H., Shiratori, S. Fabrication of nanoporous and hetero structure thin film via a layer-by-layer self assembly method for a gas sensor. *Sens. Actuat. B-Chem* **2004**, *102*, 241-247.
50. Ding, B.; Kikuchi, M.; Li, C.; Shiratori, S. Electrospun nanofibrous polyelectrolytes membranes as high sensitive coatings for QCM-based gas sensors. In *Nanotechnology at the Leading Edge*; Dirote, E.V.; Nova Science Publishers: New York, USA, 2006; pp. 1-28.
51. Stutzmann, N.; Friend, R.; Siringhaus, H. Self-aligned, vertical-channel, polymer field-effect transistors. *Science* **2003**, *299*, 1881-1884.
52. Liu, H.; Kameoka, J.; Czaplewski, D.; Craighead, H. Polymeric nanowire chemical sensor. *Nano Lett.* **2004**, *4*, 671-675.
53. Manesh, K.; Gopalan, A.; Lee, K.; Santhosh, P.; Song, K.; Lee, D. Fabrication of functional nanofibrous ammonia sensor. *IEEE Trans. Nanotechnol.* **2007**, *6*, 513-518.
54. Gao, Y.; Gong, J.; Fan, B.; Su, Z.; Qu, L. Polyaniline nanotubes prepared using fiber mats membrane as the template and their gas-response behavior. *J. Phys. Chem. C* **2008**, *112*, 8215-8222.
55. Ji, S.; Li, Y.; Yang, M. Gas sensing properties of a composite composed of electrospun poly(methyl methacrylate) nanofibers and in situ polymerized polyaniline. *Sens. Actuat. B-Chem* **2008**, *133*, 644-649.
56. Pinto, N.; Ramos, I.; Rojas, R.; Wang, P.; Johnson, A.; Jr. Electric response of isolated electrospun polyaniline nanofibers to vapors of aliphatic alcohols. *Sens. Actuat. B-Chem* **2008**, *129*, 621-627.

57. Aussawasathien, D.; Sahasithiwat, S.; Menbangpung, L. Electrospun camphorsulfonic acid doped poly(*o*-toluidine)-polystyrene composite fibers: chemical vapor sensing. *Synth. Met.* **2008**, *158*, 259-263.
58. Lala, N.; Thavasi, V.; Ramakrishna, S. Preparation of surface adsorbed and impregnated multi-walled carbon nanotube/nylon-6 nanofiber composites and investigation of their gas sensing ability. *Sensors* **2009**, *9*, 86-101.
59. Kessick, R.; Tepper, G. Electrospun polymer composite fiber arrays for the detection and identification of volatile organic compounds. *Sens. Actuat. B-Chem* **2006**, *117*, 205-210.
60. Yamazoe, N. Toward innovations of gas sensor technology. *Sens. Actuat. B-Chem* **2005**, *108*, 2-14.
61. Kolmakov, A.; Zhang, Y.; Cheng, G.; Moskovits, M. Detection of CO and O₂ using tin oxide nanowire sensors. *Advan. Mater.* **2003**, *15*, 997-1000.
62. Kim, I.; Rothschild, A.; Lee, B.; Kim, D.; Jo, S.; Tuller, H. Ultrasensitive chemiresistors based on electrospun TiO₂ nanofibers. *Nano Lett.* **2006**, *6*, 2009-2013.
63. Landau, O.; Rothschild, A.; Zussman, E. Processing-microstructure-properties correlation of ultrasensitive gas sensors produced by electrospinning. *Chem. Mater.* **2009**, *21*, 9-11.
64. Li, Z.; Zhang, H.; Zheng, W.; Wang, W.; Huang, H.; Wang, C.; MacDiarmid, A.; Wei, Y. Highly sensitive and stable humidity nanosensors based on LiCl doped TiO₂ electrospun nanofibers. *J. Am. Chem. Soc.* **2008**, *130*, 5036-5037.
65. Wang, Y.; Ramos, I.; Santiago-Aviles, J. Detection of moisture and methanol gas using a single electrospun tin oxide nanofiber. *IEEE Sens. J.* **2007**, *7*, 1347-1348.
66. Zhang, Y.; He, X.; Li, J.; Miao, Z.; Huang, F. Fabrication and ethanol-sensing properties of micro gas sensor based on electrospun SnO₂ nanofibers. *Sens. Actuat. B-Chem* **2008**, *132*, 67-73. 65.
67. Yang, A.; Tao, X.; Wang, R. Room temperature gas sensing properties of SnO₂/multiwall-carbon-nanotube composite nanofibers. *Appl. Phys. Lett.* **2007**, *91*, 133110.
68. Wang, G.; Ji, Y.; Huang, X.; Yang, X.; Gouma, P.; Dudley, M. Fabrication and characterization of polycrystalline WO₃ nanofibers and their application for ammonia sensing. *J. Phys. Chem. B* **2006**, *110*, 23777-23782.
69. Sahner, K.; Gouma, P.; Moos, R. Electrodeposited and sol-gel precipitated p-type SrTi_{1-x}Fe_xO_{3-δ} semiconductors for gas sensing. *Sensors* **2007**, *7*, 1871-1886.
70. Yang, M.; Xie, T.; Peng, L.; Zhao, Y.; Wang, D. Fabrication and photoelectric oxygen sensing characteristics of electrospun Co doped ZnO nanofibers. *Appl. Phys. A-Mat. Sci. Process.* **2007**, *89*, 427-430.
71. Luoh, R.; Hahn, H.T. Electrospun nanocomposite fiber mats as gas sensors. *Composites. Sci. Technol.* **2006**, *66*, 2436-2441.