

Communication

# Synthesis of 1-[1H,1H,2H,2H-perfluooctyl]-3-[2-(oxiran-2-yl)ethyl]imidazolium 4-[(2-oxiran-2-yl)ethoxy]benzenesulfonate as a New Perfluorinated Ionic Monomer

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**Abstract:** Access to perfluorinated compounds represents a growing challenge in the academic and industrial fields to achieve target compounds with specific physico-chemical properties. Especially, the insertion of a perfluorinated chain within an ionic liquid can provide improvements not just in terms of hydrophobicity and lipophobicity, but also viscosity, density as well as thermal stability. In this research area, we have recently developed new access points to several epoxy imidazolium salts combined with fluorinated anions such as bistriflimide ( $\text{NTf}_2^-$ ), hexafluorophosphate ( $\text{PF}_6^-$ ) or tetrafluoroborate ( $\text{BF}_4^-$ ). Here, we reported the synthesis of a perfluorinated imidazolium cation associated with a sulfonate anion as a new functionalized partner. This sequence required four steps from imidazole (cationic part) and three steps from sodium 4-hydroxybenzenesulfonate (anionic part), respectively. This perfluorinated ionic liquid was fully characterized by nuclear magnetic resonance with  $^1\text{H-NMR}$ ,  $^{19}\text{F-NMR}$ ,  $^{13}\text{C-NMR}$ , DEPT, COSY, HSQC, HMBC and IR spectroscopy. The two parts of the salt were confirmed by high-resolution mass spectrometry (HRMS), and we combined thermogravimetric analysis (TGA) and differential scanning calorimetry (DSC) to determine the thermal properties of this new compound.

**Keywords:** perfluorinated imidazolium; sulfonate; epoxides; ionic liquid



**Citation:** Kui, T.; Livi, S.; Baudoux, J. Synthesis of 1-[1H,1H,2H,2H-perfluooctyl]-3-[2-(Oxiran-2-yl)ethyl]imidazolium 4-[(2-oxiran-2-yl)ethoxy]benzenesulfonate as a New Perfluorinated Ionic Monomer. *Molbank* **2022**, *2022*, M1409. <https://doi.org/10.3390/M1409>

Academic Editor: Fawaz Aldabbagh

Received: 18 June 2022

Accepted: 13 July 2022

Published: 15 July 2022

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## 1. Introduction

In ionic liquids chemistry, the anion strongly influences the physico-chemical properties and the final applications of the salt [1,2]. To this end, fluorinated anions have been particularly studied to achieve tailor-made characteristics by improving thermal and chemical stability. Until now, we have essentially found the imidazolium cation linked to a bis(trifluoromethyl)sulfonimide ( $\text{NTf}_2$ ), hexafluorophosphate ( $\text{PF}_6$ ) and tetrafluoroborate ( $\text{BF}_4$ ), leading to hydrophilic or hydrophobic salts [3]. Among the other fluorinated anions, we report the trifluoromethylsulfonate ion (triflate  $\text{CF}_3\text{SO}_3$ ) and sulfonate analogs which also display interesting properties [4,5]. As an example, the perfluorooctane sulfonate anion (PFOS) is a fluorinated surfactant widely used by industry to provide materials such as textiles or paper with resistance to water, oils and greases [6]. Unfortunately, these PFOS represent a toxic class of persistent organic pollutants that we find in the environment and a large number of wildlife species [7]. In recent decades, some fluorinated ionic liquids have been applied in this sensitive area including the recovery of persistent perfluorocarbon contaminants from industrial effluents [8].

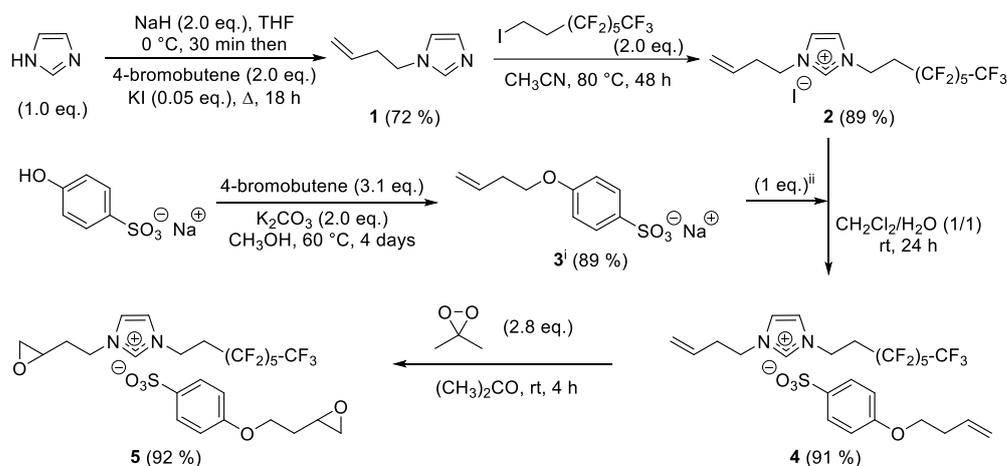
Regarding this environmental and public health aspect, thermosetting resins also represent a sensitive point. These resins are widely used through the bisphenol A diglycidyl ether (DGEBA) prepolymer which results from the reaction of toxic and carcinogenic compounds, i.e., bisphenol A and epichlorohydrin. To avoid the use of these reagents

and promote the emergence of new epoxy resins, we have developed an efficient way to epoxidized imidazoliums as new precursors of poly(ionic liquids) [9]. These ionic liquids exhibit excellent thermal stability when the cation is combined with a triflimide [10]. Nevertheless, this unmodifiable anion requires a more judicious and flexible strategy to functionalize and develop new applications while preserving the properties brought by the perfluorinated chain. In this context, we considered the synthesis of a perfluorinated cation allowing the functionalization of a sulfonate anion by an epoxide.

Based on our previous data [10,11], this work aims to confirm a new rapid and robust way to a diepoxy and perfluorinated ionic liquid. As we noted previously, the sulfonate anion represents a great interest in many applications. Moreover, this salt promotes good cross-linking with the presence of an epoxide on the anion and the cation for the development of ionic perfluorinated polymers. NMR, IR and HRMS analyzes will validate the purity of the salt while TGA and DSC will assess its potential in polymer chemistry (Supplementary Materials).

## 2. Results and Discussion

Although 1-vinylimidazole is a commercially available reagent, we opted for the synthesis of 1-(3-buten-1-yl)imidazole **1**, which makes more sense in this sequence with a more oxidizable alkene. Thus, imidazolium iodide **2** was synthesized by two consecutive nucleophilic substitutions of the imidazole. The first alkylation required sodium hydride to improve the reactivity of the heterocycle and a catalytic amount of iodide potassium to achieve an efficient reaction with 4-bromo-1-butene. After 18 h at reflux, 1-(3-buten-1-yl)imidazole **1** was isolated in a 72% yield—this reaction can be carried out over several grams did not require purification. For the second step, we started from a stoichiometric amount of the previous *N*-alkylated imidazole and 1H,1H,2H,2H-perfluooctyl iodide at reflux of acetonitrile. This quaternization reaction was followed by <sup>1</sup>H-NMR and TLC (see part IVa of the Supplementary Materials) affording 71% of salt after 24 h without significant improvement after 48 h. Thereafter, this alkylation was performed with 2 equivalents of perfluooctyl iodide to reach 89% of imidazolium **2** after 48 h at 80 °C (Scheme 1). After reaction, the unreacted starting material was eliminated by two successive treatments with hydrochloric acid (10%). This fluorinated salt **2** behaved as a surfactant during washing (see part IVb of the Supplementary Materials). Depending on the NMR result, flash chromatography on silica gel can be used as a second complementary method to isolate pure imidazolium **2**.



**Scheme 1.** Synthesis of perfluorinated imidazolium sulfonate salt (**5**). (i) Despite the use of K<sub>2</sub>CO<sub>3</sub> (2 eq.), a sodium sulfonate salt was considered for the next step. (ii) The reaction was carried out with one equivalent of **3** in two stages to allow a more efficient anion exchange.

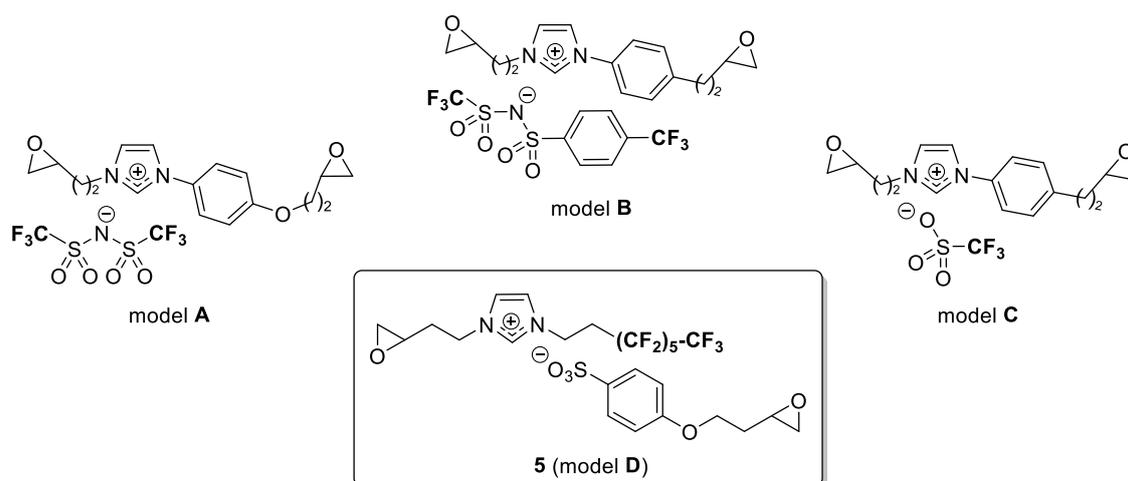
At the same time, sodium 4-hydroxybenzenesulfonate was used as a widely available and inexpensive starting material of functionalized sulfonate anions. Once again, we used 4-bromo-1-butene to generate a particularly interesting terminal alkene in the anionic part of the salt. After optimization, an excess of alkylating reagent was required in the presence of potassium carbonate to give sulfonate **3** in 89% yield after 24 h in methanol. To exchange iodide of imidazolium **2**, an anionic metathesis was carried out in a mixture dichloromethane/water (1/1) using a stoichiometric amount of sulfonate **3** previously synthesized. After 24 h at room temperature, imidazolium sulfonate **4** was mainly observed with a conversion of 59%. This anion exchange was slightly improved by using two equivalents of **3** to reach 70% under similar conditions as the previous attempt. This optimization confirmed the advantage of using an excess of sulfonate **3** but in a two-step addition by stirring a stoichiometric mixture of homemade sulfonate **3** and imidazolium **4** for 24 h before adding one more equivalent of sulfonate to generate a complete exchange of the iodide (see Section 3.5). Under these optimized conditions, several extractions with dichloromethane provided the pure perfluorinated salt with an isolated yield of 91%.

This new ionic liquid **4** was oxidized in the presence of freshly prepared dimethyldioxirane (2.8 eq.) in acetone [12]. The reaction mixture was stirred at room temperature until the reaction was completed ( $^1\text{H-NMR}$  monitoring). After 4 h, perfluorinated imidazolium sulfonate bearing two epoxides **5** was isolated in 92% yield. Additionally, an excess of mCPBA (5 eq.) was also used as another oxidizing reagent to give 90% of **5** after 24 h at 40 °C in acetonitrile.

In this sequence, the perfluorinated imidazolium sulfonate **5** was synthesized in 54% overall yield from imidazole, 4-bromo-1-butene, 1H,1H,2H,2H-perfluorooctyl iodide, and sodium 4-hydroxybenzenesulfonate. This procedure does not require any purification by silica gel column chromatography during the entire sequence. This efficient pathway is reproducible and can be performed on a large scale in the laboratory. The structure and the purity of this new ionic liquid was confirmed by nuclear magnetic resonance using  $^1\text{H-NMR}$ ,  $^{19}\text{F-NMR}$ ,  $^{13}\text{C-NMR}$ , DEPT, COSY, HSQC, HMBC and Infrared spectroscopy (see part I of the Supplementary Materials). High-resolution mass spectrometry confirmed the two subparts of this salt with a cation  $[\text{M}]^+$  at 485.0910 with the formula  $\text{C}_{15}\text{H}_{14}\text{N}_2\text{OF}_{13}$  and an anion  $[\text{M}]^-$  at 243.0324 with the formula  $\text{C}_{10}\text{H}_{11}\text{O}_5\text{S}$ .

The last part of this work was devoted to the thermal properties of this perfluorinated imidazolium salt. The thermal stability was carried out by thermogravimetric analysis (TGA) using a Perkin Elmer Pyris 1 TGA working under nitrogen atmosphere with flow rate  $20\text{ mL}\cdot\text{min}^{-1}$  and heating  $20\text{ K}\cdot\text{min}^{-1}$ . The weight loss as a function of temperature was analyzed to determine the degradation temperature and the percentage of degradation of this salt. In previous studies, we reported the excellent thermal behavior for several diepoxy aryl-imidazolium triflimide salts (Figure 1, model A) with a maximal degradation temperature above 390 °C. Homemade perfluorinated sulfonimides (Figure 1, model B) afforded similar thermal stability. Very recently, new promising diepoxidized salts were described with sulfonate as a versatile counterion. In this unprecedented series, good thermal properties were observed around 375 °C to reach 411 °C with trifluoromethylsulfonate (Figure 1, model C) as counter-anion [13]. Here, by transferring the fluorinated chain to the imidazolium, the highly functionalized ionic liquid (Figure 1, model D) showed a mass loss of only 10% at 245 °C to reach a  $T_{\text{max}}$  of 346 °C.

Finally, the thermal behavior of this salt was investigated by differential scanning calorimetry to determine the glass transition temperature, i.e.,  $T_g$ . A glass transition temperature of  $-32\text{ }^\circ\text{C}$  was obtained corresponding to the  $T_g$ s obtained in our previous studies [9–11] opening some perspectives in the field of solid electrolytes where a lower  $T_g$  is required to reach the optimal ionic conductivity.



**Figure 1.** Example of diepoxy perfluorinated salts in relation to sulfonate 5. Model A: Unmodified commercial triflimide; Model B: Homemade perfluorinated sulfonimide; Model C: Unmodified commercial trifluoromethylsulfonate; Model D: Homemade functionalized sulfonate with perfluorinated imidazolium.

### 3. Materials and Methods

#### 3.1. Chemistry

All reagents were purchased from Sigma Aldrich, Alfa Aesar, Acros Organic or TCI and were used as received: mCPBA ( $\leq 77\%$  from Sigma Aldrich), 4-bromo-1-butene (97% from Alfa Aesar), anisole ( $>99\%$  from Sigma Aldrich), potassium carbonate (99% from Acros Organic), imidazole (from TCI), perfluorooctyle iodide (96% from Sigma Aldrich), sodium hydride (60% from Sigma Aldrich), Sodium 4-Hydroxybenzenesulfonate ( $<98\%$  from TCI). DMDO was prepared according to the procedure described by D. F. Taber [12]. Solvents were used in RPE grade without further purification. Anhydrous solvents were obtained from a PURESOLV SPS400 apparatus developed by Innovative Technology Inc.  $^1\text{H}$ -,  $^{19}\text{F}$ - and  $^{13}\text{C}$ -NMR spectra were recorded on a Bruker Avance III 500 MHz or Avance NEO 600 MHz spectrometer. Samples were dissolved in an appropriate deuterated solvent ( $\text{CDCl}_3$ ,  $\text{CD}_3\text{CN}$ ,  $\text{D}_2\text{O}$  and acetone- $d_6$ ). The chemical shifts ( $\delta$ ) are expressed in ppm relative to internal tetramethylsilane for  $^1\text{H}$  and  $^{13}\text{C}$  nuclei and coupling constants are indicated in Hz. Abbreviations for signal coupling are as follows: s = singlet; d = doublet; dd = doublet of doublets; t = triplet; q = quartet; m = multiplet. To assign the signals to the different proton and carbon atoms, additional 2D NMR experiments (COSY, HSQC, HMBC and DEPT) were performed. High-resolution mass spectra (HRMS) were performed on Acquity UPLC H-Class Xevo G2-XS QToF (WATERS) by electrospray ionization (ESI). Infrared (IR) spectra were recorded with a Perkin Elmer 16 PC FTIR ATR spectrometer, using the pure product (oil or solid). Thermographic analyses were recorded with Perkin Elmer Pyris 1 TGA, working under azote atmosphere with flow rate 20 mL/min and heating 20 °C/min. Thin layer chromatography (TLC) was run on pre-coated aluminum plates of silica gel 60 F-254 (Merck). Flash chromatography was performed on silica gel column (Merck silica gel, 40–63 mm).

#### 3.2. Synthesis of 1-(3-buten-1-yl)imidazole 1

To a round-bottom flask purged with Ar, imidazole (5.00 g, 73.5 mmol, 1.0 eq) was added to a suspension of NaH (5.88 g, 147 mmol, 2.0 eq) in dry THF (100 mL) at 0 °C. The reaction was stirred for 30 min and we completed the mixture with 4-bromobutene (1.06 mL, 10.54 mmol, 2.0 eq) and potassium iodide (0.61 g, 3.60 mmol, 0.05 eq). After 18 h at 60 °C, the solvent was removed under reduced pressure. The residue was partitioned between  $\text{CH}_2\text{Cl}_2$  and water and the organic layer was washed with  $\text{NH}_4\text{Cl}$  saturated (2  $\times$  10 mL). The combined organic

extracts were dried over  $\text{MgSO}_4$ , and concentrated under reduced pressure to afford product **1** as a clear yellow oil (6.45 g, 72%).

$^1\text{H-NMR}$  (600 MHz,  $\text{CDCl}_3$ )  $\delta$  7.44 (s, 1H), 7.02 (s, 1H), 6.91 (s, 1H), 5.75–5.68 (m, 1H), 5.07–5.04 (m, 2H), 3.98 (t,  $J = 6.9$  Hz, 2H), 2.50 (q,  $J = 6.9$  Hz, 2H).

$^{13}\text{C-NMR}$  (151 MHz,  $\text{CDCl}_3$ )  $\delta$  136.9, 133.5, 129.1, 118.7, 118.0, 46.3, 35.2.

IR (neat)  $\text{cm}^{-1}$  3379, 3108, 2979, 1641, 1507, 1438, 1359, 1283, 1228, 1108, 1077, 1036, 907, 814.

HRMS  $m/z$  (ESI): calcd. for  $\text{C}_7\text{H}_{11}\text{N}_2$   $[\text{MH}]^+$ : 123.0922, found: 123.0923.

### 3.3. Synthesis of 1-[1H,1H,2H,2H-perfluooctyl]-3-(3-buten-1-yl)imidazolium Iodide **2**

To a solution of 1-(3-buten-1-yl)imidazole (0.24 g, 2 mmol, 1.0 eq) in  $\text{CH}_3\text{CN}$  (10 mL) was added 1H,1H,2H,2H-perfluooctyl iodide (0.9 mL, 4 mmol, 2.0 eq). The mixture was refluxed at 80 °C and regularly monitored by  $^1\text{H-NMR}$  or TLC with a solution of dichloromethane/methanol (9/1). After cooling to room temperature, the solvent was removed under reduced pressure. The residue was partitioned between  $\text{CH}_2\text{Cl}_2$  and water and the organic layer was washed with HCl (1M) ( $2 \times 10$  mL). The combined organic extracts were dried over  $\text{MgSO}_4$ , and concentrated under reduced pressure to afford product **2** as a clear yellow oil (1.04 g, 89%). If necessary, the product can be purified by flash chromatography on silica gel with  $\text{CH}_2\text{Cl}_2/\text{MeOH}$  (9/1).

$^1\text{H-NMR}$  (600 MHz,  $\text{CDCl}_3$ )  $\delta$  10.24 (s, 1H), 7.56 (s, 1H), 7.36 (s, 1H), 5.83–5.76 (m, 1H), 5.14–5.08 (m, 2H), 4.87 (t,  $J = 6.5$  Hz, 2H), 4.30 (t,  $J = 6.7$  Hz, 2H), 2.98–2.90 (m, 2H), 2.72–2.69 (m, 2H).

$^{13}\text{C-NMR}$  (151 MHz,  $\text{CDCl}_3$ )  $\delta$  137.4, 131.9, 122.7, 122.1, 120.1, 118.1 ( $\text{CF}_3$ ), 116.1 (5  $\text{CF}_2$ ), 49.7, 42.6, 34.2, 31.9.

$^{19}\text{F-NMR}$  (376 MHz,  $\text{CDCl}_3$ )  $\delta$  –80.8, –113.4, –121.8, –122.8, –123.2, –126.2.

IR (neat)  $\text{cm}^{-1}$  3054, 2882, 1772, 1563, 1422, 1265, 1240, 1145, 896, 735.

HRMS  $m/z$  (ESI): calcd. for  $\text{C}_{15}\text{H}_{14}\text{F}_{13}\text{N}_2$   $[\text{M}]^+$ : 469.0949, found: 469.0953.

### 3.4. Synthesis of Sodium 4-(3-buten-1-yloxy)benzenesulfonate **3**

To a solution of sodium 4-hydroxybenzenesulfonate (3.0 g, 12.9 mmol, 1 eq) and potassium carbonate (4.10 g, 117 mmol, 2.0 eq) in methanol (30 mL) was added 4-bromobut-1-ene (4.06 mL, 40.1 mmol, 3.1 eq). The reaction mixture was stirred at 65 °C for 4 days. The solid residue was filtered and the solvent was concentrated under pressure. The product **3** was obtained as a white solid (2.92 g, 89%).

$^1\text{H-NMR}$  (500 MHz,  $\text{D}_2\text{O}$ )  $\delta$  7.66 (d,  $J = 8.5$  Hz, 2H), 6.99 (d,  $J = 8.5$  Hz, 2H), 5.90–5.82 (m, 2H), 5.14–5.04 (m, 1H), 4.11 (t,  $J = 6.4$  Hz, 2H), 2.49–2.45 (m, 2H).

$^{13}\text{C-NMR}$  (126 MHz,  $\text{D}_2\text{O}$ )  $\delta$  160.5, 135.2, 127.5 ( $\times 2$ ), 117.2, 115.6, 115.0 ( $\times 2$ ), 67.9, 32.8.

IR (neat)  $\text{cm}^{-1}$  3432, 3059, 2970, 1738, 1643, 1638, 1499, 1473, 1231, 1180, 1139, 1065, 1031, 1007, 900.

Mp: >300 °C

HRMS  $m/z$  (ESI): calcd. for  $\text{C}_{10}\text{H}_{11}\text{O}_4\text{S}$   $[\text{M}]^-$ : 227.0378, found: 227.0370.

### 3.5. Synthesis of 1-[1H,1H,2H,2H-perfluooctyl]-3-(3-buten-1-yl)imidazolium 4-(3-buten-1-yloxy)benzenesulfonate **4**

To a stirred solution of sodium sulfonate **3** (40 mg, 0.16 mmol, 1.0 eq) in biphasic solution  $\text{CH}_2\text{Cl}_2/\text{H}_2\text{O}$  (1:1,  $v/v$ ) (8 mL) was added imidazolium iodide **2** (100 mg, 0.16 mmol, 1.0 eq). The mixture was stirred at room temperature for 24 h. The organic compounds were extracted with dichloromethane, dried with  $\text{MgSO}_4$  and concentrated under reduced pressure. After NMR analysis, an excess of sodium sulfonate (1.0 eq) was used to complete the reaction. This mixture was stirred at room temperature for another 24 h. The product **4** was extracted with dichloromethane, dried with  $\text{MgSO}_4$  and then concentrated under reduced pressure to afford a yellow clear oil (101 mg, 91%).

$^1\text{H-NMR}$  (600 MHz,  $\text{CDCl}_3$ )  $\delta$  9.82 (s, 1H), 7.77 (d,  $J = 8.8$  Hz, 2H), 7.58 (s, 1H), 7.37 (s, 1H), 6.83 (d,  $J = 8.8$  Hz, 2H), 5.90–5.84 (m, 1H), 5.70–5.63 (m, 1H), 5.16–5.08 (m, 2H),

5.02–4.95 (m, 2H), 4.70 (t,  $J = 6.5$  Hz, 2H), 4.28 (t,  $J = 6.8$  Hz, 2H), 3.98 (t,  $J = 6.7$  Hz, 2H), 2.81–2.73 (m, 2H), 2.56–2.50 (m, 4H).

$^{13}\text{C}$ -NMR (151 MHz,  $\text{CDCl}_3$ )  $\delta$  159.9, 138.7, 137.8 ( $\times 2$ ), 134.2, 132.2 ( $\times 2$ ), 127.4, 122.8, 122.2, 119.5, 118.1 ( $\text{CF}_3$ ), 117.1, 110.8 (5  $\text{CF}_2$ ), 113.9, 67.3, 49.2, 42.2, 34.2, 33.5, 26.9.

$^{19}\text{F}$ -NMR (376 MHz,  $\text{CDCl}_3$ )  $\delta$  –85.8, –118.46, –126.8, –127.9, –128.2, –131.2.

IR (neat)  $\text{cm}^{-1}$  3050, 2772, 1780, 1610, 1423, 1265, 1240, 1188, 1121, 1078, 1004, 835, 735.

HRMS  $m/z$  (ESI): calcd. for  $\text{C}_{15}\text{H}_{14}\text{F}_{13}\text{N}_2$   $[\text{M}]^+$ : 469.0949, found: 469.0958, calcd. for  $\text{C}_{10}\text{H}_{11}\text{O}_4\text{S}$   $[\text{M}]^-$ : 227.0378, found: 227.0378.

### 3.6. Synthesis of 1-[1H,1H,2H,2H-perfluorooctyl]-3-[2-(Oxiran-2-yl)ethyl]imidazolium 4-[(2-oxiran-2-yl)ethoxy]benzenesulfonate 5

To a solution of compound 4 (50 mg, 0.071 mmol, 1.0 eq) in acetone (0.50 mL) freshly prepared DMDO (0.044 mol/L) (4.57 mL, 0.20 mmol, 2.8 eq) was added. The reaction mixture was stirred at room temperature for 4 h. Two drops of DMS were added to neutralize the excess of DMDO. The solvent was evaporated under reduced pressure and the product 5 was obtained as a red clear oil (48 mg, 92%).

$^1\text{H}$ -NMR (500 MHz,  $\text{CD}_3\text{CN}$ )  $\delta$  8.73 (s, 1H), 7.56 (d,  $J = 9.2$  Hz, 2H), 7.41 (s, 1H), 7.37 (s, 1H), 6.77 (d,  $J = 9.1$  Hz, 2H), 4.45 (t,  $J = 7.2$  Hz, 2H), 4.25–4.22 (m, 2H), 4.05–4.03 (m, 2H), 2.99–2.96 (m, 1H), 2.88–2.85 (m, 1H), 2.81–2.72 (m, 2H), 2.66–2.65 (dd,  $J = 4.7$  and 4.0 Hz, 1H), 2.6–2.61 (dd,  $J = 4.7$  and 4.0 Hz, 1H), 2.44–2.43 (dd,  $J = 5.0$  and 2.7 Hz, 1H), 2.36–2.35 (dd,  $J = 5.0$  and 2.7 Hz, 1H), 2.14–2.09 (m, 1H), 1.94–1.89 (m, 1H), 1.84–1.79 (m, 2H).

$^{13}\text{C}$ -NMR (126 MHz,  $\text{CD}_3\text{CN}$ )  $\delta$  159.7, 141.4, 137.4, 127.8 ( $\times 2$ ), 123.4, 123.2, 123.2 ( $\text{CF}_3$ ), 114.1 ( $\times 2$ ), 111.2 (5  $\text{CF}_2$ ), 65.5, 49.8, 49.3, 47.6, 46.9, 46.4, 42.2, 32.9 ( $\times 2$ ), 32.7.

$^{19}\text{F}$ -NMR (376 MHz,  $\text{CD}_3\text{CN}$ )  $\delta$  –81.5, –114.4, –122.3, –123.3, –124.1, –126.6.

IR (neat)  $\text{cm}^{-1}$  3100, 2769, 1598; 1188, 1144, 1121, 1078, 1004, 961, 835, 809, 746, 699.

HRMS  $m/z$  (ESI): calcd. for  $\text{C}_{15}\text{H}_{14}\text{N}_2\text{OF}_{13}$   $[\text{M}]^+$ : 485.0899, found: 485.0910, calcd. for  $\text{C}_{10}\text{H}_{11}\text{O}_5\text{S}$   $[\text{M}]^-$ : 243.0327, found: 243.0324.

## 4. Conclusions

In conclusion, we have synthesized a perfluorinated imidazolium cation combined with a sulfonate as a new functionalized counter-anion. This ionic liquid incorporating an aromatic ring and terminal epoxides on the two parts of the salt is a promising candidate for perfluorinated poly(ionic liquid)s. This diepoxide salt was obtained in four steps from imidazole and sulfonate 3. After optimization, the final compound was characterized by  $^1\text{H}$ -,  $^{13}\text{C}$ -, and  $^{19}\text{F}$ -NMR but also COSY, HMBC, HSQC and IR spectroscopy. We analyzed this ionic liquid by HRMS spectrometry, thermogravimetric analysis (TGA) and differential scanning calorimetry (DSC) to confirm its high purity and thermal behavior.

**Supplementary Materials:** The following supporting information are available online: copies of the NMR spectra for all the compounds, IR spectra and TGA curve.

**Author Contributions:** Conceptualization, methodology and validation, J.B. and S.L.; formal analysis, investigation and data curation, J.B. and T.K.; writing—original draft preparation, writing—review and editing, visualization, J.B.; supervision, project administration, funding acquisition, J.B. and S.L. All authors have read and agreed to the published version of the manuscript.

**Funding:** This research was funded by the French Agence Nationale de la Recherche through the project 3DSMART (fellowships to T.K.), the Ministry of Higher Education and Research, the “Region Basse Normandie”, the LABEX SynOrg (ANR11-LABX-0029), the European FEDER and the Groupement de Recherches Liquides Ioniques et PolymèreS (GDR LIPSCNRS #5223).

**Institutional Review Board Statement:** Not applicable.

**Informed Consent Statement:** Not applicable.

**Data Availability Statement:** Not applicable.

**Conflicts of Interest:** The authors declare no conflict of interest.

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