

Supporting Information

for

Role of Cobalt(III) Cationic Complexes in the Self-assembling Process of a Water Soluble Porphyrin

¹ Dipartimento di Scienze Chimiche, Biologiche, Farmaceutiche ed Ambientali, University of Messina and C.I.R.C.M.S.B V.le F. Stagno D'Alcontres, 31 - 98166 Messina, Italy E.mail nadiamanganaro87@gmail.com (N.M.); anromeo@unime.it (A.R.); lmonsu@unime.it (L.M.S.)

² CNR - ISMN Istituto per lo Studio dei Materiali Nanostrutturati c/o Dipartimento di Scienze Chimiche, Biologiche, Farmaceutiche ed Ambientali, University of Messina, V.le F. Stagno D'Alcontres, 31 - 98166 Messina, Italy roberto.zagami@ismn.cnr.it (R.Z.); mariachiara.trapani@cnr.it (M.T.); maria.castriciano@cnr.it (M.C.)

* Correspondence: lmonsu@unime.it; Tel.: +39 090 6765711

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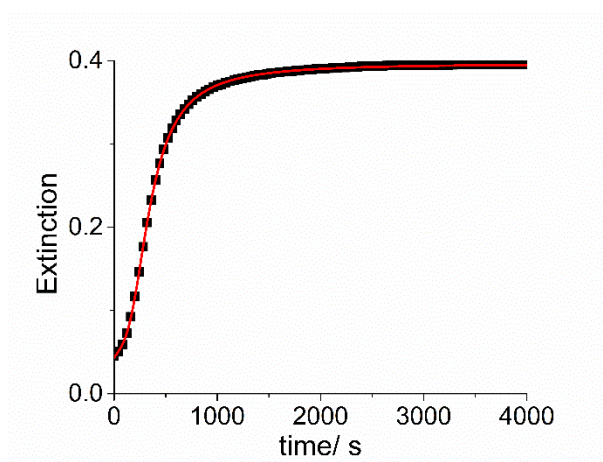


Figure S1. Typical extinction time trace for the formation of J-aggregates of H_2TPPS_4 at $\text{pH} = 2$ upon addition of $[\text{Co}(\text{phen})_3]^{3+}$ (Experimental conditions: $[\text{H}_2\text{TPPS}_4] = 3 \mu\text{M}$; $[\text{HCl}] = 0.01 \text{ M}$, $[\text{Co}(\text{phen})_3]^{3+} = 200 \mu\text{M}$, $T = 298 \text{ K}$). The solid line represents the best-fitted curve to the experimental data according eq. 1.

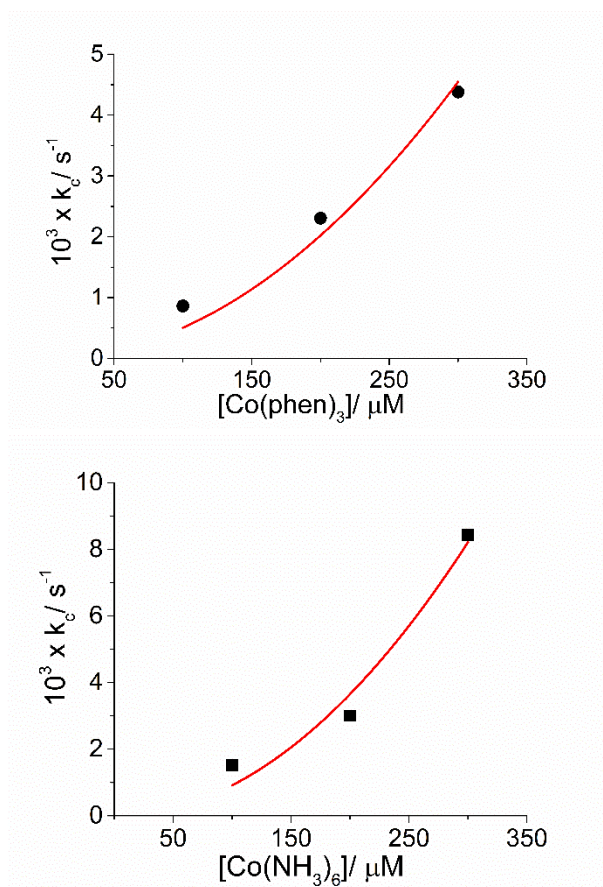


Figure S2. Plot of the rate constants k_c (s^{-1}) for the catalyzed growth of J-aggregates of H_2TPPS_4 at $\text{pH} = 2$ upon addition of $[\text{Co(phen)}_3]^{3+}$ (circles) and $[\text{Co(NH}_3)_6]^{3+}$ (solid squares). (Experimental conditions: $[\text{H}_2\text{TPPS}_4] = 3 \mu\text{M}$; $[\text{HCl}] = 0.01 \text{ M}$, $T = 298 \text{ K}$). The solid lines represent the non-linear best-fits to the law $k_c = k_c' \times [\text{CoL}_n^{3+}]^2$ ($[\text{Co(NH}_3)_6]^{3+}$: $k_c'' = (9.13 \pm 0.66) \times 10^{-8} \text{ s}^{-1}\mu\text{M}^{-2}$, $R^2 = 0.9681$; $[\text{Co(phen)}_3]^{3+}$: $k_c'' = (5.05 \pm 0.35) \times 10^{-8} \text{ s}^{-1}\mu\text{M}^{-2}$, $R^2 = 0.9622$).

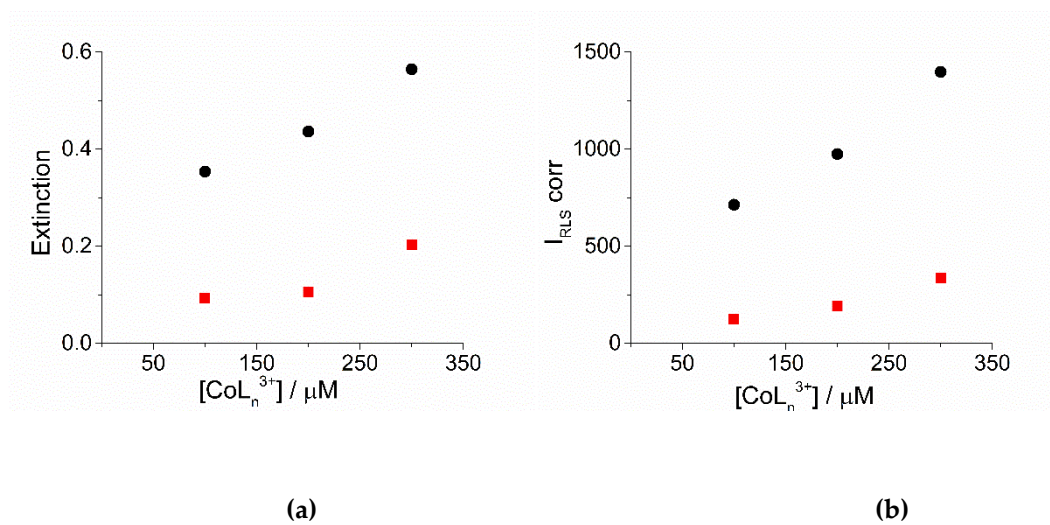


Figure S3. Plot of (a) the extinction values and (b) the RLS intensity at maxima corrected for the extinction of the samples as function of $[\text{Co}(\text{phen})_3]^{3+}$ (circles) and $[\text{Co}(\text{NH}_3)_6]^{3+}$ (solid squares) (Experimental conditions: $[\text{H}_2\text{TPPS}_4] = 3 \mu\text{M}$; $[\text{HCl}] = 0.01 \text{ M}$, $T = 298 \text{ K}$).

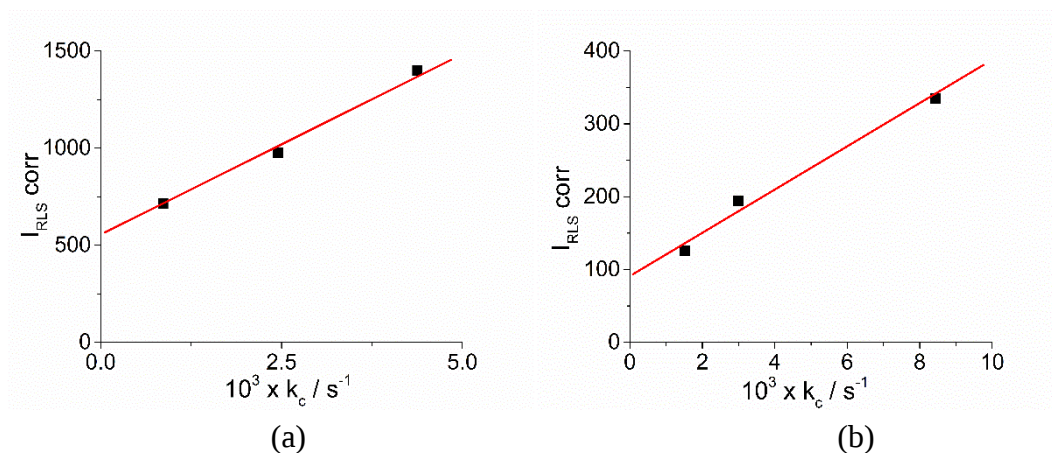


Figure S4. Plot of the RLS intensity at maxima corrected for the extinction of the samples as function of the corresponding rate constants k_c for the catalyzed growth of J-aggregates of H_2TPPS_4 at $\text{pH} = 2$ upon addition of a) $[\text{Co}(\text{phen})_3]^{3+}$ and b) $[\text{Co}(\text{NH}_3)_6]^{3+}$. (Experimental conditions: $[\text{H}_2\text{TPPS}_4] = 3 \mu\text{M}$; $[\text{HCl}] = 0.01 \text{ M}$, $[\text{CoL}_n^{3+}] = 100 \mu\text{M}$, $200 \mu\text{M}$ and $300 \mu\text{M}$, $T = 298 \text{ K}$). The solid lines represent the linear best-fits to the equations: a) $I_{\text{RLS}}^{\text{corr}} = (527 \pm 47) + (1.95 \pm 0.16) \times 10^5 \times [\text{Co}(\text{phen})_3]$ ($R^2 = 0.9934$); b) $I_{\text{RLS}}^{\text{corr}} = (93 \pm 18) + (2.90 \pm 0.35) \times 10^4 \times [\text{Co}(\text{NH}_3)_6]$.

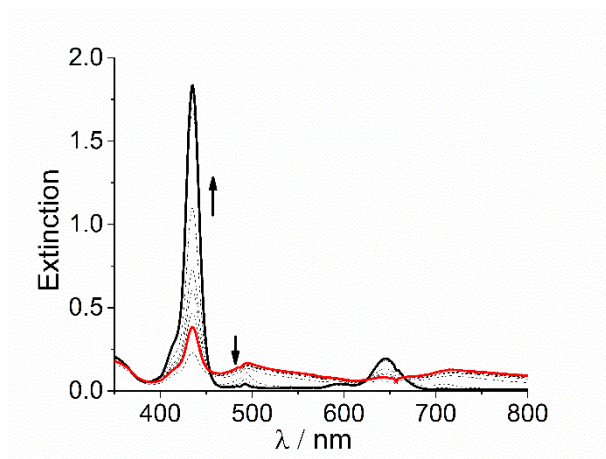


Figure S5. UV/Vis extinction spectral changes for the disassembling of J-aggregates of H_2TPPS_4 stabilized with $[\text{Co}(\text{NH}_3)_6]^{3+}$ at pH = 2 upon addition of PVS (Experimental conditions: $[\text{H}_2\text{TPPS}_4] = 3 \mu\text{M}$; $[\text{HCl}] = 0.01 \text{ M}$, $[\text{Co}(\text{NH}_3)_6] = 300 \mu\text{M}$, $[\text{PVS}] = 1 \text{ mM}$, $T = 298 \text{ K}$, total scanning time 1800 s).