

A Simple Iron-catalyst for Alkenylation of Ketones Using Primary Alcohols

Motahar Sk, Ashish Kumar, Jagadish Das, and Debasis Banerjee*

Department of Chemistry

Laboratory of Catalysis and Organic Synthesis

Indian Institute of Technology Roorkee, Roorkee-247667, India

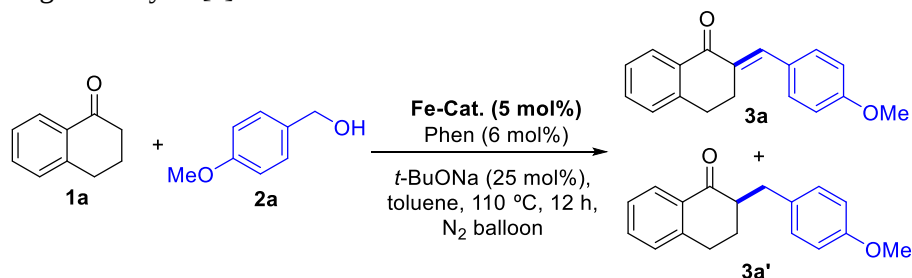
E-mail: debasis.banerjee@cy.iitr.ac.in

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[1.1] Alkenylation of ketones with alcohols:

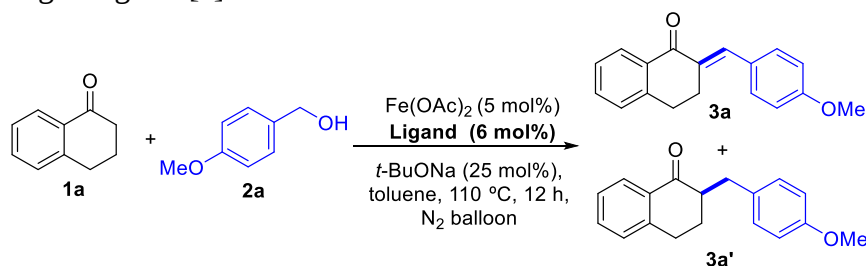
Table S1: Screening of catalysts [a]



Entry	Fe-Catalyst	GC-MS Conversion 3a (%)	GC-MS Conversion 3a' (%)
1^b	Fe(OAc)₂	77(75)	3
2	Fe(acac) ₃	68	4
3	Fe ₂ (CO) ₉	64	4
4	Fe(OAc) ₂ (2.5 mol%), Phen (3 mol%)	53	2
5	No Catalyst, No Ligand	20	-

Reaction conditions: [a] *p*-methoxy benzyl alcohol (0.25 mmol), α -tetralone (0.375 mmol), **Fe-catalyst (5 mol%)**, Phen (6 mol%), *t*-BuONa (0.0625 mmol), toluene (1.0 mL), Schlenk tube under nitrogen atmosphere, 110 °C oil bath, 12 h reaction time. [b] Isolated yield.

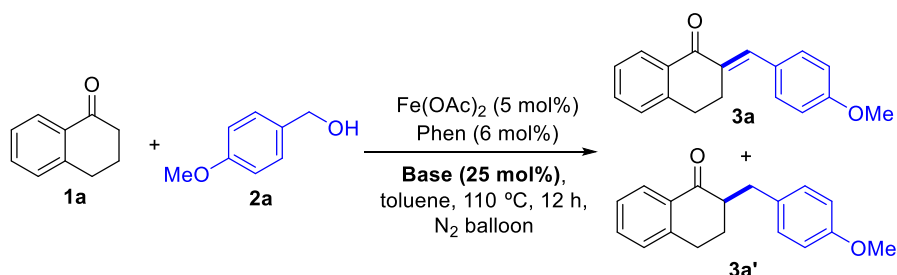
Table S2: Screening of ligands[a]



Entry	Ligand	GC-MS Conversion 3a (%)	GC-MS Conversion 3a' (%)
1^b	L1	77(75)	3
2	L2	68	4
3	L3	54	5
4	L4	67	6
5	PPh₃ (L5)	70	17
6	L6	60	30
7	No Ligand	43	2

Reaction conditions: [a] *p*-methoxy benzyl alcohol (0.25 mmol), α -tetralone (0.375 mmol), Fe(OAc)₂ (5 mol%), **Ligand (6 mol%)**, *t*-BuONa (0.0625 mmol), toluene (1.0 mL), Schlenk tube under nitrogen atmosphere, 110 °C oil bath, 12 h reaction time. PPh₃ (10 mol%) was used. [b] Isolated yield.

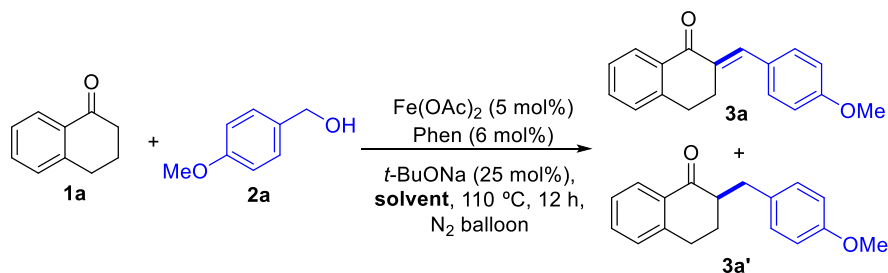
Table S3: Screening of base [a]



Entry	Base	GC-MS Conversion 3a (%)	GC-MS Conversion 3a' (%)
1^b	<i>t</i>-BuONa	77(75)	3
2	<i>t</i> -BuOK	58	30
3	Na ₂ CO ₃	25	<1
4	K ₂ CO ₃	30	<1
5	Cs ₂ CO ₃	12	<1
6	K ₃ PO ₄	7	0
7	No Base	0	0

Reaction conditions: [a] *p*-methoxy benzyl alcohol (0.25 mmol), α -tetralone (0.375 mmol), Fe(OAc)₂ (5 mol%), Phen (6 mol%), **base** (0.0625 mmol), toluene (1.0 mL), Schlenk tube under nitrogen atmosphere, 110 °C oil bath, 12 h reaction time. [b] Isolated yield.

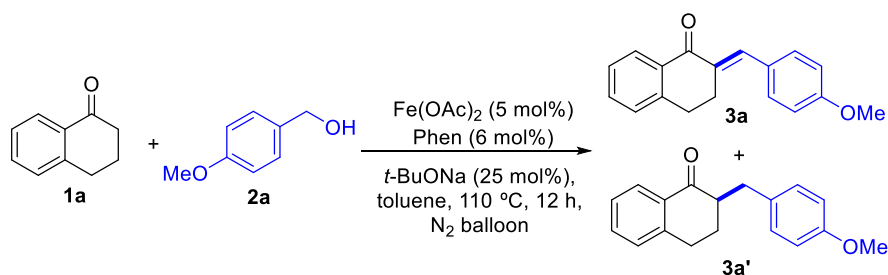
Table S4: Screening of solvents [a]



Entry	Solvent	GC-MS Conversion 3a (%)	GC-MS Conversion 3a' (%)
1[b]	toluene	77(75)	3
2	<i>p</i> -xylene	45	5
3	1,4-dioxane	33	2
4	DMA	8	0
5	<i>t</i> -amylalcohol	25	16

Reaction condition: [a] *p*-methoxy benzyl alcohol (0.25 mmol), α -tetralone (0.375 mmol), Fe(OAc)₂ (5 mol%), Phen (6 mol%), *t*-BuONa (0.0625 mmol), **solvent** (1.0 mL), Schlenk tube under nitrogen atmosphere, 110 °C oil bath, 12 h reaction time. [b] Isolated yield.

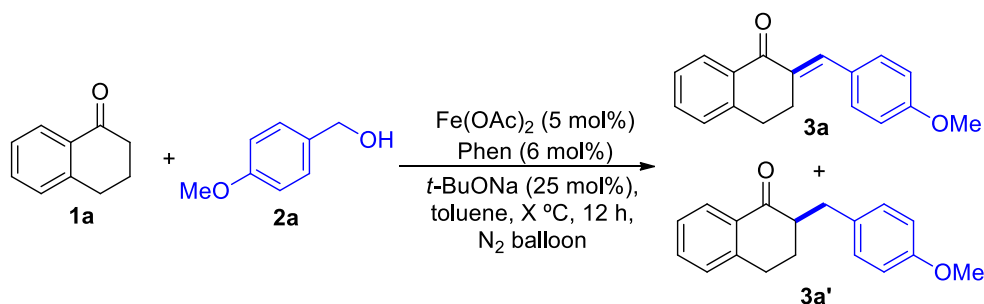
Table S5: Screening of ketone equivalents [a]



Entry	α -tetralone (X equiv.)	GC-MS Conversion 3a (%)	GC-MS Conversion 3a' (%)
1	1.0	54	4
2 ^b	1.5	77(75)	3

Reaction condition: [a] *p*-methoxy benzyl alcohol (0.25 mmol), α -tetralone (**0.25 mmol, 0.375 mmol**), Fe(OAc)₂ (5 mol%), Phen (6 mol%), *t*-BuONa (0.0625 mmol), toluene (1.0 mL), Schlenk tube under nitrogen atmosphere, 110 °C oil bath, 12 h reaction time. [b] Isolated yield (average of two run).

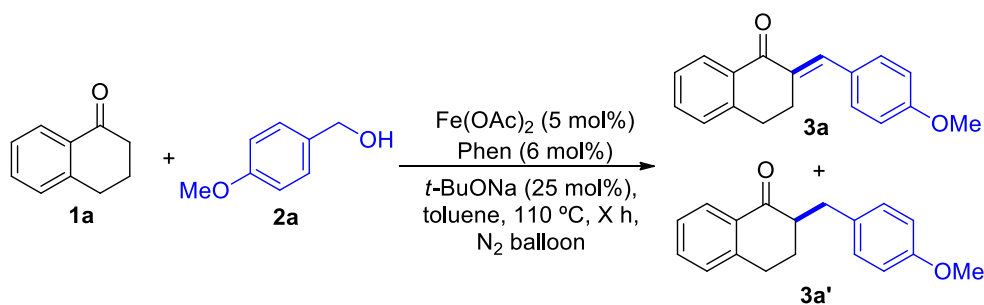
Table S6: Screening of temperature [a]



Entry	Temperature (X °C)	GC-MS Conversion 3a (%)	GC-MS Conversion 3a' (%)
1	110	77(75) ^[b]	3
2	80	58	-
3	50	9	-

Reaction conditions: [a] *p*-methoxy benzyl alcohol (0.25 mmol), α -tetralone (0.375 mmol), Fe(OAc)₂ (5 mol%), Phen (6 mol%), NaO^tBu (0.0625 mmol), toluene (1.0 mL), Schlenk tube under nitrogen atmosphere, X °C oil bath, 12 h reaction time. [b] Isolated yield.

Table S7: Screening of time [a]

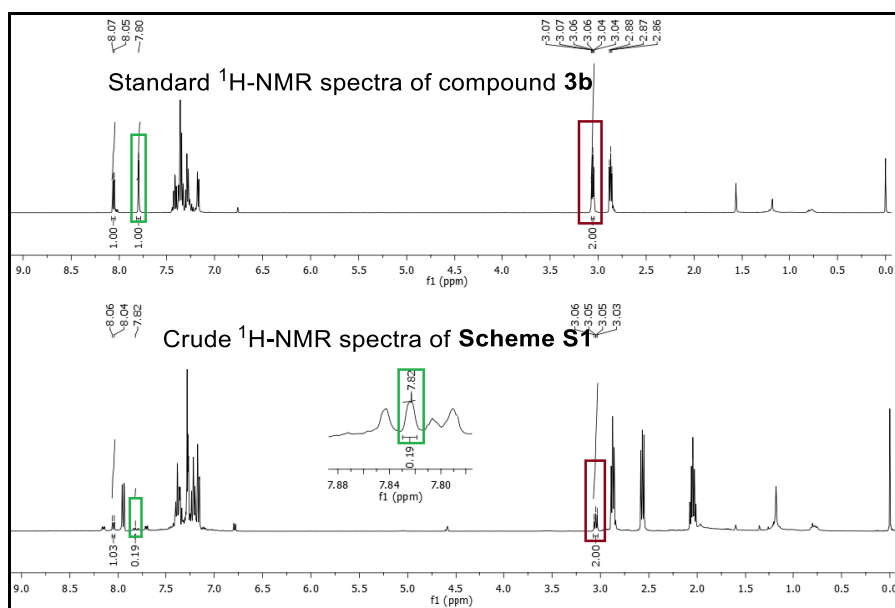
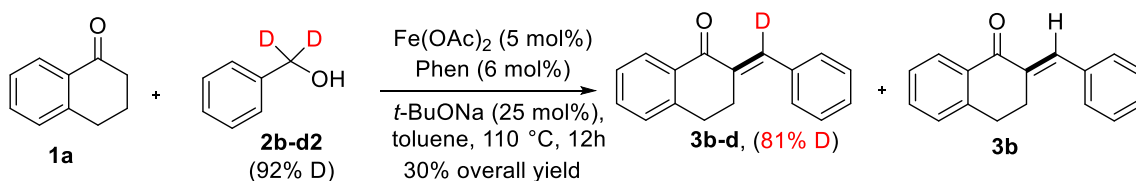


Entry	Time (X h)	GC-MS Conversion 3a (%)	GC-MS Conversion 3a' (%)
1	12	77(75)^[b]	3
2	9	62	2
3	6	45	4

Reaction conditions: [a] *p*-methoxy benzyl alcohol (0.25 mmol), α -tetralone (0.375 mmol), $\text{Fe}(\text{OAc})_2$ (5 mol%), Phen (6 mol%), NaO^tBu (0.0625 mmol), toluene (1.0 mL), Schlenk tube under nitrogen atmosphere, 110 °C oil bath, X h reaction time. [b] Isolated yield.

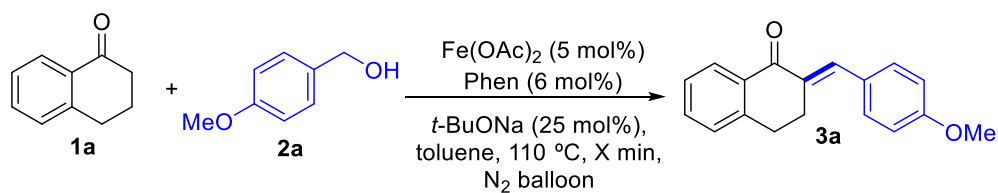
Deuterium Incorporation Experiments:

Scheme S1:



Scheme S2: Determination of rate and order of reaction

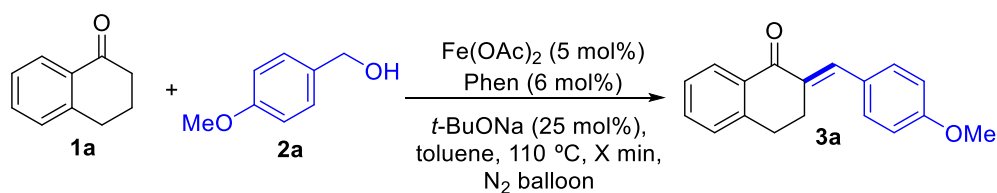
Run 1: Reaction was carried out in 1 mL of toluene and yield was calculated by GC



No.	1a (mmol)	2a (mmol)	$\text{Fe}(\text{OAc})_2$ (mmol)	Phen (mmol)	$t\text{-BuONa}$ (mmol)	Toluene (mL)
Run 1	0.3	0.2	0.01	0.012	0.05	1.0

Sl. No.	Time (min)	Concentration of 2a (mM)
1	4	190
2	8	179
3	12	172
4	16	165
5	20	160

Run 2: Reaction was carried out in 1 mL of toluene and yield was calculated by GC



No.	1a (mmol)	2a (mmol)	Fe(OAc) ₂ (mmol)	Phen (mmol)	t-BuONa (mmol)	Toluene (mL)
Run 2	0.375	0.25	0.0125	0.015	0.0625	1.0

Sl. No.	Time (min)	Concentration of 2a (mM)
1	4	229
2	8	216
3	12	203
4	16	196
5	20	191

Considering steady state approximation for benzyl alcohol

$$\text{From Run 1: Slope} = k [2a]^x$$

$$-1.85 = k [0.20]^x$$

$$\text{From Run 2: Slope} = k [2a]^x$$

$$-2.4 = k [0.25]^x$$

$$-2.4/-1.85 = [0.25]^x / [0.2]^x$$

$$1.297 = [1.25]^x$$

$$\text{Log}(1.297) = x \cdot \text{Log}(1.25)$$

$$x = 0.113 / 0.0969$$

$$= 1.16 \approx 1$$

$$\text{Rate} = k [2a]^1$$

Scheme S3: Quantitative determination of hydrogen gas produced in the reaction

In a 10 mL oven dried Schlenk tube, 4-methoxy benzyl alcohol (0.5 mmol), Fe(OAc)₂ (5 mol%), Phen (6 mol%), α -tetralone (0.75 mmol) and *t*-BuONa (0.125 mmol), were added followed by toluene 2.0 mL and connected to the gas burette as shown in figure 3. Then the reaction mixture was heated at 110 °C until the production of hydrogen gas ceased. The procedure was repeated three times to get concordant reading.



Experimental setup for H₂ determination

Total volume of water displaced, V = 0.007 L

Vapor pressure of water at 298K, P_{H₂O} = 23.7695 Torr

Atmospheric pressure at 298K, P_{atm} = 758.3124 Torr

Pressure of H₂ gas, P_{H₂} = P_{atm} - P_{H₂O} = (758.3124 - 23.7695) Torr
= 734.5429 Torr

$$P_{H_2} * V = n_{H_2} * R * T$$

$$n_{H_2} = P_{H_2} * V / R * T$$

$$= 734.5429 \text{ Torr} * 0.007 \text{ L} / 62.3635 \text{ L Torr K}^{-1} \text{ mol}^{-1} * 298 \text{ K}$$

$$= 0.000277 \text{ mol}$$

$$= 0.277 \text{ mmol}$$

[1.2] Copies of ^1H NMR, ^{13}C NMR and HRMS Spectra for selected compounds

