

Article

# Ag/SiO<sub>2</sub>- and Ag/Co<sub>3</sub>O<sub>4</sub>-Based Monolithic Flow Microreactors for Hydrogenation of Dyes: Their Activity and Stability

Yasemin Hakat <sup>†</sup>, Trupti V. Kotbagi <sup>†</sup> and Martin G. Bakker <sup>\*</sup>

Department of Chemistry, the University of Alabama, Tuscaloosa, AL 35487-0336, USA;  
E-Mails: yhakat@crimson.ua.edu (Y.H.); trupti.kotbagi@gmail.com (T.V.K.)

<sup>†</sup> These authors contributed equally to this work.

<sup>\*</sup> Author to whom correspondence should be addressed; E-Mail: bakker@ua.edu;  
Tel.: +1-205-348-9116; Fax: +1-205-348-9104.

Academic Editor: Volker Hessel

Received: 31 December 2014 / Accepted: 9 February 2015 / Published: 16 February 2015

---

**Abstract:** Silver nanoparticles supported on hierarchically porous silica and cobalt oxide monoliths have previously been shown to be catalytically active for the hydrogenation of common organic dyes in batch studies. This work presents a detailed investigation of the activity and stability of these monoliths during the hydrogenation of eosin-Y in a continuous flow microreactor. The silver-containing monoliths showed excellent catalytic activity that reached a plateau after a period of approximately 6 h. From SEM particle size distribution studies of the catalysts before and after water and hexane were flowed through them, it was determined that under reaction conditions, silver was removed both by washing off of particles and by dissolution of silver.

**Keywords:** catalyst; flow; hydrogenation; microreactor; monolith; nanoparticle

---

## 1. Introduction

Heterogeneous catalysis is an important component in the development of atom and energy-efficient chemical transformations. The design of processes that are highly selective, energy efficient and environmentally benign requires detailed knowledge of the reaction kinetics, thorough evaluation of the catalytic materials and rapid optimization of the reaction conditions. Microreactors enable precise control of reaction parameters (e.g., temperature, pressure), and the reaction conditions

can be continuously varied and monitored in flow systems during catalytic processes, greatly accelerating reaction kinetic studies and process optimization [1,2]. Additionally, flow microreactors have advantages, such as enhanced atom efficiency, selectivity, yield and purity of products, reproducibility, increased turn-over-number (TON), turn-over-frequency (TOF), process safety, shorter development cycles and rapid scale-up from lab to pilot plant [3]. The use of water as the solvent in such reactors is particularly attractive in cutting operating costs.

The efficiency of heterogeneous catalysts depends on various properties, such as surface area, the pore size of the support and the particle size of the active metals on the support. Hierarchically porous materials have interconnecting pore networks at multiple length scales, making them particularly attractive for reactions that are mass transport limited. When fabricated as monoliths, *i.e.*, as single pieces, they are particularly attractive as continuous flow microreactors, providing superior mass diffusion and increasing the accessibility of the fluids (reactants and products) during organic transformations [4]. A wide variety of hierarchically porous monoliths [5], such as hierarchically porous zeolites ( $\text{SiO}_2/\text{Al}_2\text{O}_3$ , aluminophosphate (AlPO), silico-aluminophosphate SAPO) [6], hierarchically porous metal oxides ( $\text{SiO}_2$  [7–10],  $\text{Co}_3\text{O}_4$  [8–10], NiO [8–10], ZnO [11],  $\text{TiO}_2$  [12]) and hierarchically porous carbon-based catalysts [13,14], have been studied.

With the recent explosion of work in the synthesis and application of metal nanoparticles, there have been numerous reports of the use of metal nanoparticles as catalysts for various organic transformations, including carbonylation [15,16], hydroformylation [17] and hydrogenation [18,19]. However, during solution phase catalytic reactions, nanoparticles tend to aggregate due to high surface energy and van der Waals forces. This is undesirable in terms of catalytic activity and the longevity of the catalyst. Supporting the nanoparticles on a high surface area support is one step towards overcoming the issue of aggregation. Supporting the catalyst also simplifies the separation of the catalyst from the reaction products. Hierarchically porous monolithic materials, like  $\text{SiO}_2$ , can be used as high surface area supports in which the nanoparticles can be formed or entrapped inside the interconnected channels. The formation of metal nanoparticles on supports to give heterogeneous catalysts is somewhat of an art form that has been studied over many years [20]. However, there is considerably less work on hierarchically porous materials [4,5,21,22], particularly monolithic materials, which have been evaluated in only a handful of flow experiments [23–25].

We have recently shown in batch reactions that silver (Ag) nanoparticles supported on hierarchically porous  $\text{SiO}_2$  and  $\text{Co}_3\text{O}_4$  monoliths (Ag/ $\text{SiO}_2$  and Ag/ $\text{Co}_3\text{O}_4$ , respectively) are catalytically active for the hydrogenation of methyl orange (MO) and methylene blue (MB) dyes with rate constants that indicate the potential of studying these reactions in a continuous flow microreactor [26]. Herein, for the first time, we report the hydrogenation of eosin-Y (EO) using Ag/ $\text{SiO}_2$  and Ag/ $\text{Co}_3\text{O}_4$  hierarchically porous monoliths as microreactors in a continuous flow system. The use of a microreactor allows us to assess the stability of these catalysts and explore the origins of the decreases observed in the initial catalytic activities.

## 2. Experimental Section

### 2.1. Materials

All chemicals used were commercially available and used as received. Polyethylene glycol 35000 (PEG 35000), tetraethyl orthosilicate (TEOS, 98%), ammonium hydroxide (NH<sub>4</sub>OH), cobalt nitrate hexahydrate (Co(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O), silver nitrate (AgNO<sub>3</sub>), 30% nitric acid (HNO<sub>3</sub>), potassium hydroxide (KOH), eosin-Y (EO) and sodium borohydride (NaBH<sub>4</sub>) were purchased from Sigma-Aldrich. Octadecyltrimethylammonium bromide (C<sub>18</sub>TAB) was obtained from Genscript. The assay of all of the chemicals was 99.9%, unless otherwise noted.

### 2.2. Synthesis of the SiO<sub>2</sub> Monolith Template

The SiO<sub>2</sub> template monoliths were synthesized according to the procedure described by Småtå *et al.* [8]. In summary: polyethylene glycol (PEG, MW = 35,000 g/mol), 30% v/v HNO<sub>3</sub>, tetraethyl orthosilicate (TEOS) and octadecyltrimethylammonium bromide (C<sub>18</sub>TAB) were mixed with molar ratios of  $5.5 \times 10^{-4}$ : $2.5 \times 10^{-1}$ :14.5:1:0.12, respectively. This solution was transferred to a mold (96-well plate) and allowed to gel for at least 72 h at 45 °C. To improve the stability of these monoliths, they were aged in 1 M NH<sub>4</sub>OH at 90 °C for 8 h, followed by neutralization in 0.1 M HNO<sub>3</sub>, rinsing with deionized water and acetone and drying at 45 °C overnight. Finally, the SiO<sub>2</sub> monoliths were calcined at 550 °C for 5 h with a ramp rate of 1 °C/min.

### 2.3. Synthesis of the Co<sub>3</sub>O<sub>4</sub> Replica by Nanocasting

The SiO<sub>2</sub> monoliths were used as a template for the synthesis of Co<sub>3</sub>O<sub>4</sub> replica monoliths by nanocasting. Initially, the SiO<sub>2</sub> monoliths were degassed under vacuum and filled with a 5 M aqueous solution of Co(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O. The infiltrated monoliths were dried at 150 °C overnight under N<sub>2</sub> flow. These monoliths were then heated at 250 °C for 4 h to ensure total Co(NO<sub>3</sub>)<sub>2</sub>·6H<sub>2</sub>O salt decomposition to metal oxide, Co<sub>3</sub>O<sub>4</sub>. The above procedure was repeated three times to completely fill the mesopores and so giving continuous Co<sub>3</sub>O<sub>4</sub> monoliths. The SiO<sub>2</sub> template was then dissolved using 3 M aqueous KOH at room temperature for 24 h. This process was carried out twice to ensure complete SiO<sub>2</sub> removal, and the replicas were then rinsed with water until a neutral pH was achieved.

### 2.4. Synthesis of Ag/SiO<sub>2</sub> and Ag/Co<sub>3</sub>O<sub>4</sub> Composite Monoliths by Solution Infiltration

An aqueous solution of 0.1 M AgNO<sub>3</sub> was prepared and degassed prior to infiltration into SiO<sub>2</sub> and Co<sub>3</sub>O<sub>4</sub>. The SiO<sub>2</sub> monolith appeared greyish in color after the addition of silver (Ag), while the Co<sub>3</sub>O<sub>4</sub> monolith remained black in color after Ag addition. The infiltrated monoliths were dried overnight at 65 °C. The Ag salt was then reduced using ethylene glycol at 165 °C for 30 min, to give the supported catalysts that will be referred to as Ag/SiO<sub>2</sub> for the Ag supported on SiO<sub>2</sub> and Ag/Co<sub>3</sub>O<sub>4</sub> for the Ag supported on Co<sub>3</sub>O<sub>4</sub>.

## 2.5. Characterization

Scanning electron microscopy (SEM) on a JEOL-7000 FE SEM equipped with an Oxford energy dispersive spectroscopy (EDS) was used for characterization and elemental analysis. X-ray diffraction (XRD) was carried out on a Bruker D8 Discover with General Area Detector Diffraction System (GADDS) X-ray (Co K $\alpha$ ,  $\lambda = 1.79$  Å). ImageJ (version: 1.48v; National Institutes of Health: Bethesda, MD, USA, 1997) was used to carry out the determination of the particle sizes from the SEM images.

## 2.6. Catalytic Reduction of EO Using Supported Ag Nanoparticles as Catalyst

The continuous flow monolithic microreactor was constructed as follows: The monolith (5 mm  $\times$  7 mm) was positioned into a heat shrinkable PTFE tube of 30 mm and 60 mm in length, which was heated to shrink the PTFE tubing until a tight fit to the monolith was produced. An aqueous solution of EO (2.9  $\mu$ M) and reducing agent NaBH $_4$  (0.018 M) was pumped with a peristaltic pump through the microreactor at a constant flow rate of 2.1 mL/min. The product solution was passed through a UV cuvette flow cell of 4.5 mL capacity to analyze the reduction of EO using a Cary 50 Scan UV-Vis spectrophotometer in the range of 200–800 nm. Calibration experiments showed that the absorbance recorded at 521 nm was linear for concentrations up to 2.9  $\mu$ M, corresponding to an absorbance of 2.5.

## 2.7. Washing Studies

The washing tests of the Ag/SiO $_2$  and Ag/Co $_3$ O $_4$  monoliths were carried out in a continuous flow microreactor. The fixed bed of the continuous monolithic microreactor was constructed as mentioned in Section 2.6. Neat hexane was pumped through the microreactor with a peristaltic pump at a constant flow rate of 2.1 mL/min. The quantitative analysis of the washing of Ag from the support was determined by measuring the Ag ion concentration in the effluent using a Perkin Elmer AAnalyst 400 atomic absorption spectrophotometer.

# 3. Results and Discussion

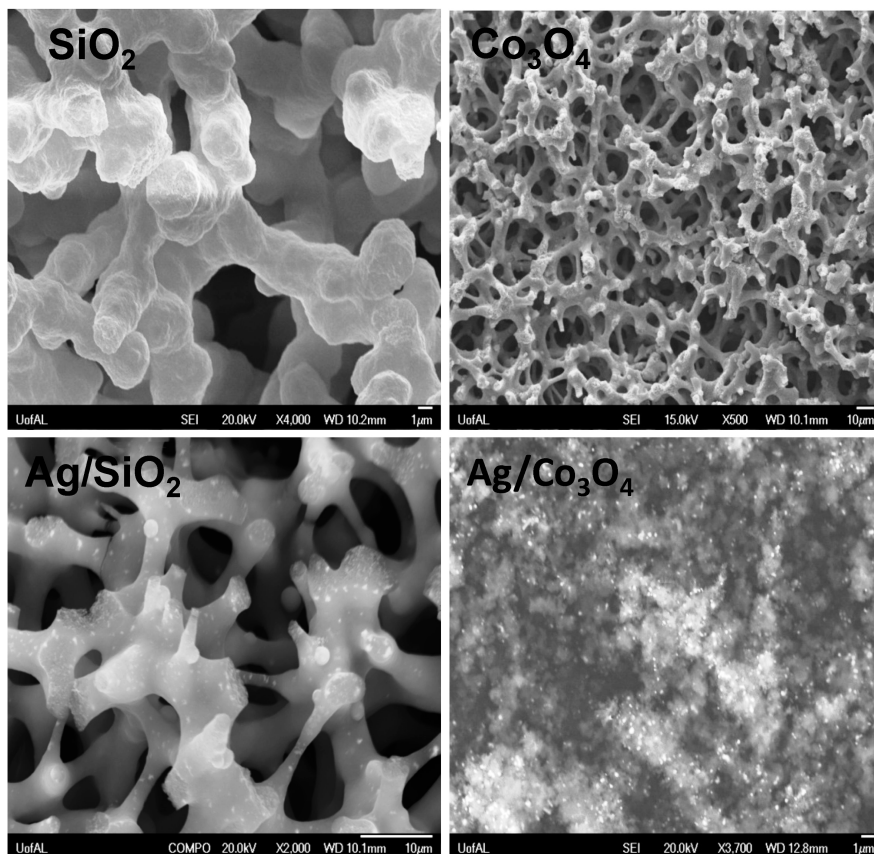
## 3.1. Catalyst Characterization

The detailed characterization of the SiO $_2$ , Ag/SiO $_2$ , Co $_3$ O $_4$  and Ag/Co $_3$ O $_4$  monoliths was performed and reported previously [26]. In brief, SiO $_2$  and Co $_3$ O $_4$  showed high BET surface areas of 741.1 and 71.7 m $^2$ /g, with total pore volumes of 1.26 and 0.34 cm $^3$ /g, respectively. The surface area was found to decrease with the addition of Ag for both monoliths. The structure of the monoliths can be seen in the SEM images (Figure 1). Atomic Absorption Spectroscopic AAS analyses showed the amount Ag on SiO $_2$  and Co $_3$ O $_4$  to be 1.8 and 2.1 wt%, respectively.

Both SiO $_2$  and Co $_3$ O $_4$  monoliths showed the presence of similar network of 10–15  $\mu$ m macropores, consistent with the replication into Co $_3$ O $_4$  of the macropore structure of the SiO $_2$  template. The formation of Ag nanoparticles on the macropore surfaces of the SiO $_2$  and Co $_3$ O $_4$  replica monoliths can clearly be seen in Figure 1. These images were taken with a back-scatter detector, so that brighter



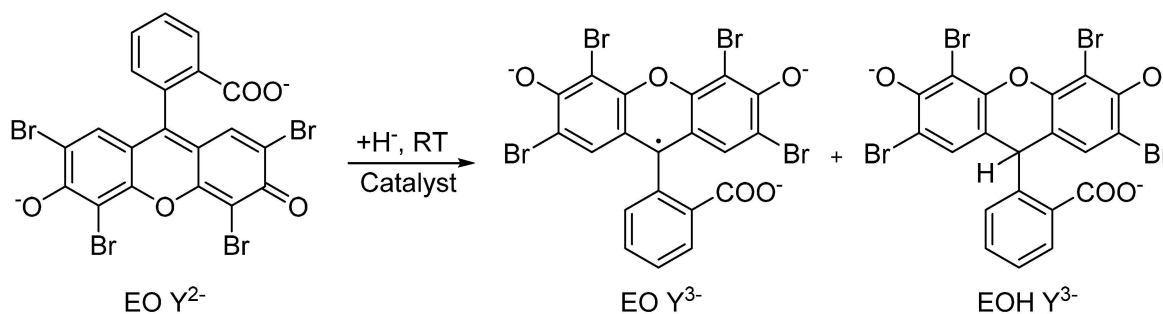
areas correspond to Ag nanoparticles, which have a higher atomic number than the supports. In general, the Ag nanoparticles are <100 nm.



**Figure 1.** SEM images of monolithic catalysts.

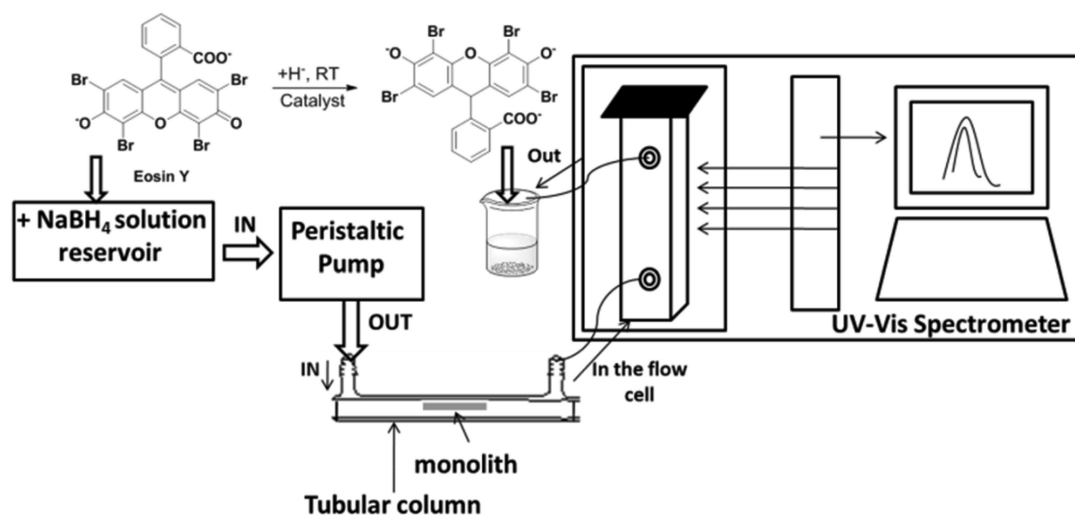
### 3.2. Hydrogenation of EO in Continuous Monolithic Microreactors

Hydrogenation of EO dye has been studied to evaluate the activity and stability of these materials in a continuous reactor (Scheme 1).



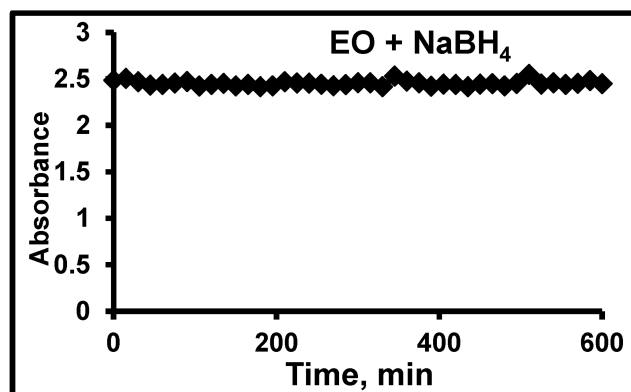
**Scheme 1.** Hydrogenation of eosin (EO)-Y<sup>2-</sup> to EO-Y<sup>3-</sup> and EOH-Y<sup>3-</sup>.

The continuous monolithic microreactor set-up used is shown in Figure 2 and incorporates a UV-visible spectrophotometer as the detector.



**Figure 2.** The typical continuous monolithic microreactor used for the hydrogenation of EO.

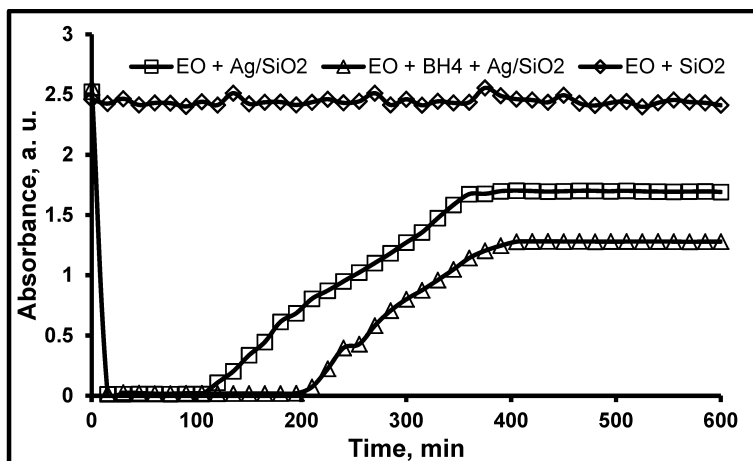
Careful control experiments were carried out to confirm the stability of the experimental setup. The EO hydrogenation was performed in the absence of the monolithic catalyst using  $\text{NaBH}_4$  as the reducing agent under flow conditions. The absorbance was recorded at 525 nm, the absorption maximum of EO. The absorbance ( $A$ ) vs. time ( $t$ ) plot is shown in Figure 3. No change in the absorbance of the initial EO solution was observed over 600 min. This clearly indicated that the rate of the reaction of EO is negligible in the absence of catalyst.



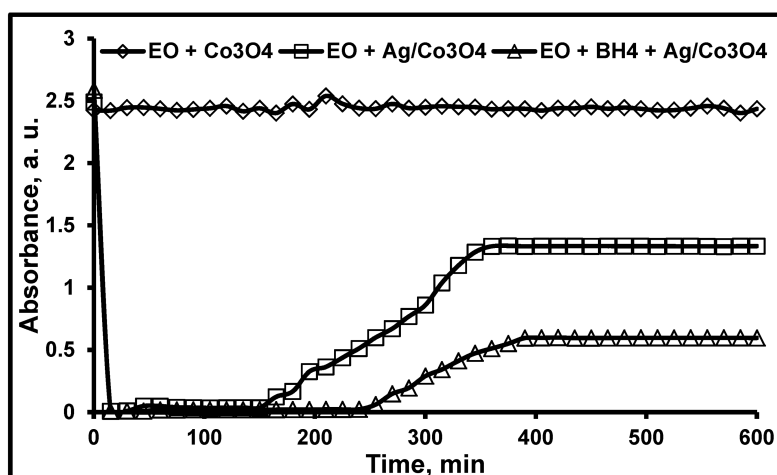
**Figure 3.** Reaction of EO and  $\text{NaBH}_4$  in the absence of catalyst (absorbance ( $A$ ) vs. time ( $t$ ) plot).

In our earlier work on the hydrogenation of dyes under batch conditions, it was observed that  $\text{SiO}_2$  and  $\text{Co}_3\text{O}_4$  were not catalytic for the hydrogenation of either MO or MB and so would not be expected to catalyze the hydrogenation of EO. However, it was observed that both MB and MO adsorb on  $\text{SiO}_2$  and  $\text{Co}_3\text{O}_4$  supports, which significantly complicated the data analysis. Based on the negative surface charges of both  $\text{SiO}_2$  and  $\text{Co}_3\text{O}_4$  at neutral pH, anionic EO was predicted not to significantly adsorb on either support. The adsorption behavior of EO on the  $\text{SiO}_2$  and  $\text{Co}_3\text{O}_4$  monolithic supports was studied under flow conditions in the absence of  $\text{NaBH}_4$ . The results can be seen in Figures 4 and 5, which show no decrease in the absorbance value from the initial EO solution. Previously strong adsorption of MO and MB on Ag nanoparticles supported on  $\text{SiO}_2$  and  $\text{Co}_3\text{O}_4$  was observed, in agreement with other

reports. Accordingly, the adsorption of EO on Ag nanoparticles supported on  $\text{SiO}_2$  and  $\text{Co}_3\text{O}_4$  was studied as a function of time in the absence of  $\text{NaBH}_4$ . As can be seen in Figures 4 and 5, initially, 100% of the EO is adsorbed on both  $\text{Ag}/\text{SiO}_2$  and  $\text{Ag}/\text{Co}_3\text{O}_4$  monoliths. The EO was completely absorbed until 105 min for  $\text{Ag}/\text{SiO}_2$  and 150 min in the case of  $\text{Ag}/\text{Co}_3\text{O}_4$ . Subsequently, the absorbance gradually increased and attained a steady-state value less than that of the initial EO solution, indicating continuing adsorption of EO on the Ag nanoparticles. The steady state for  $\text{Ag}/\text{SiO}_2$  was achieved at 345 min with approximately 33% of the EO being adsorbed. The steady state for  $\text{Ag}/\text{Co}_3\text{O}_4$  was achieved at 345 min with 46% of the EO being adsorbed.



**Figure 4.** Absorbance (A) vs. time (t) plots of EO reaction studies using  $\text{SiO}_2$ -based monolithic catalysts.



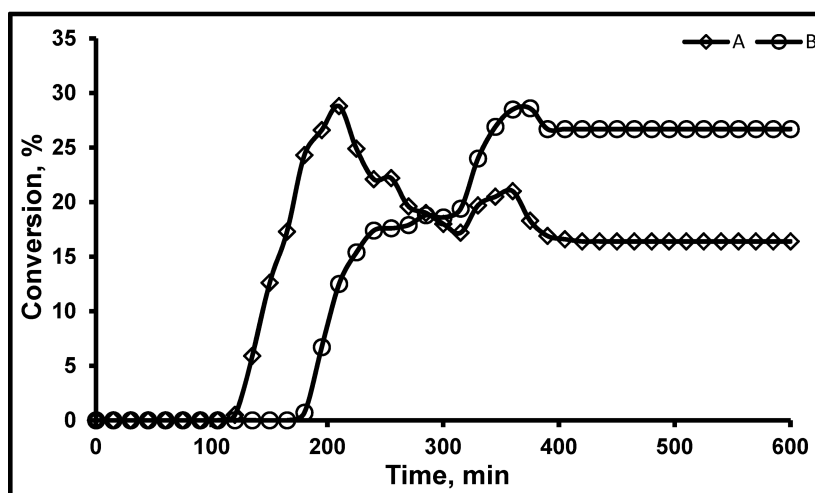
**Figure 5.** Absorbance (A) vs. time (t) plots of EO reaction studies using  $\text{Co}_3\text{O}_4$ -based monolithic catalysts.

The absence of adsorption on the bare supports indicates clearly that the adsorption of EO is taking place on the Ag nanoparticles.

Figures 4 and 5 also show the absorbance vs. time plots when  $\text{NaBH}_4$  is added to the EO solution. For  $\text{Ag}/\text{SiO}_2$ , complete removal of EO from the solution increases from 100 min in the absence of  $\text{NaBH}_4$  to 195 min in the presence of  $\text{NaBH}_4$ . For  $\text{Ag}/\text{Co}_3\text{O}_4$ , complete removal of EO increases from

150 min to 205 min with the addition of  $\text{NaBH}_4$ . Our earlier work with MO and MB confirmed that reaction with  $\text{NaBH}_4$  directly converted these dyes to their hydrogenation products and did not indicate any changes in the extent of dye adsorption. Similarly, batch studies of EO confirmed the stoichiometric conversion to the products, with no indication of changes in the extent of EO adsorption on Ag. Accordingly, the difference between the amount of EO leaving the microreactor in the absence and presence of  $\text{NaBH}_4$  can be ascribed to the catalytic reaction and conversion of EO to products.

Figure 6 shows the extent of catalytic conversion of EO to products. For  $\text{Ag}/\text{SiO}_2$ , the catalytic conversion reaches a peak of 29% at 210 min and then decreases to a constant efficiency of 16.4%. For  $\text{Ag}/\text{Co}_3\text{O}_4$ , the catalytic conversion reaches an efficiency of 28% at 360 min and then decreases slightly to a constant efficiency of 26.7%. The initial increases in conversion can be readily ascribed to the decrease in the extent of EO adsorption, but in the absence of other factors, the conversion efficiency would be expected to monotonically increase to a steady-state value. That it does not suggests that something is occurring beyond the decrease in the extent of absorption seen in the absence of  $\text{NaBH}_4$ . There are a number of possible causes for this change in catalytic activity, including alterations in the nature of the monolith in terms of the structure and the porosity of the support and changes in the number and size distribution of the Ag nanoparticles.

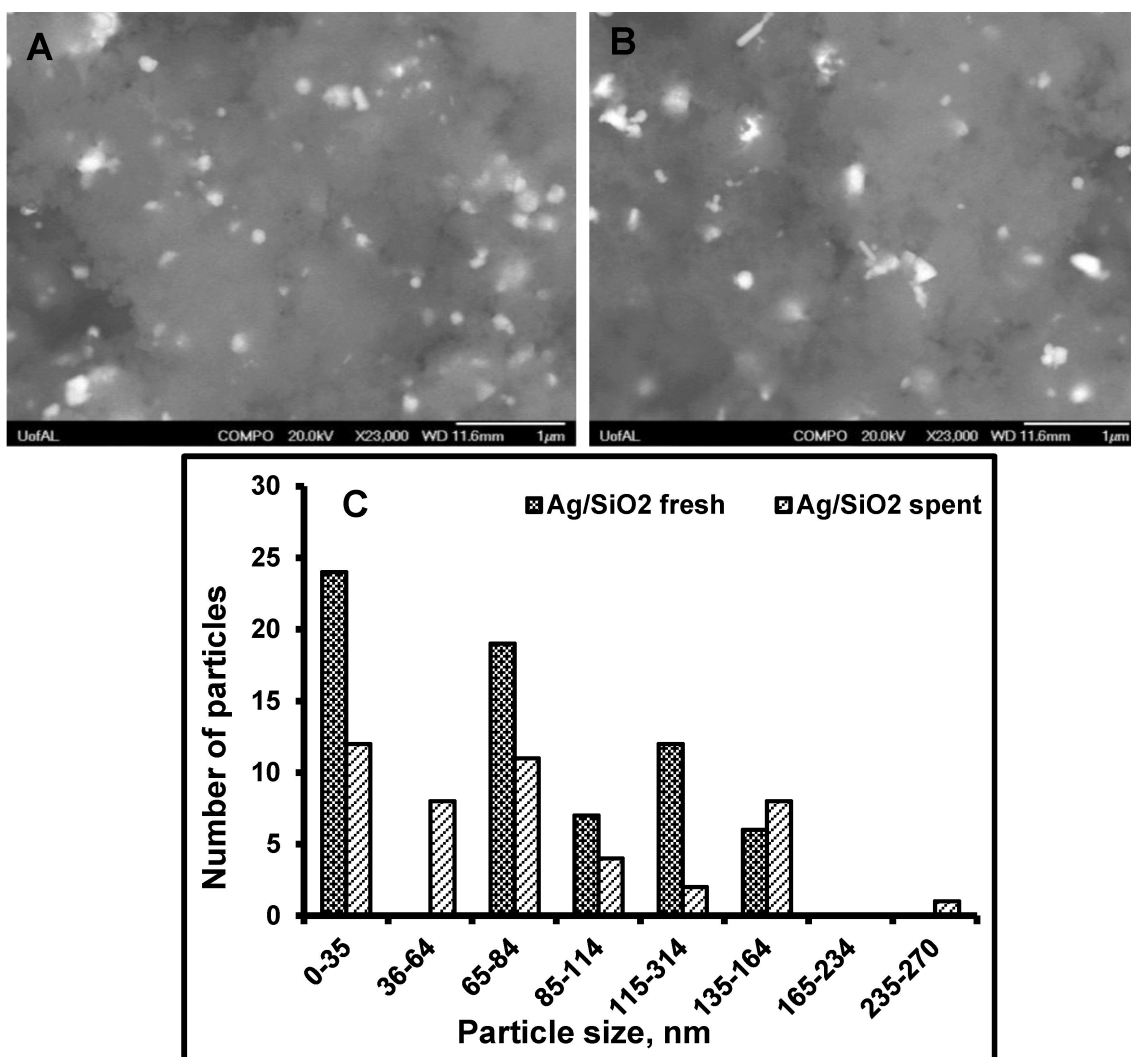


**Figure 6.** Conversion of EO in the presence of  $\text{NaBH}_4$  using (A)  $\text{Ag}/\text{SiO}_2$  and (B)  $\text{Ag}/\text{Co}_3\text{O}_4$  as the catalysts.

### 3.3. Particle Size Distribution and Washing/Leaching Studies

Detailed SEM analysis was carried out on the spent catalysts and compared to that of fresh catalyst. No changes in the structure of the supports were seen. Figure 7 shows SEM images of the fresh and spent  $\text{Ag}/\text{SiO}_2$  monolith at a magnification at which the Ag nanoparticles can be resolved. Particle size distributions for the Ag nanoparticles are also shown in Figure 7. The size distribution of Ag nanoparticles in fresh  $\text{Ag}/\text{SiO}_2$  monoliths is observed to be somewhat bimodal, with a significant number of Ag nanoparticles with a size  $<35$  nm. No Ag nanoparticles in the 36–64 nm range were observed, with the remainder of the Ag nanoparticles falling in the 65–164 nm size range. The spent  $\text{Ag}/\text{SiO}_2$  catalyst showed a decrease in the number of the smallest nanoparticles (from 24 to 12) and an approximately 50% decrease of Ag nanoparticles  $<84$  nm. Agglomerates of Ag nanoparticles in

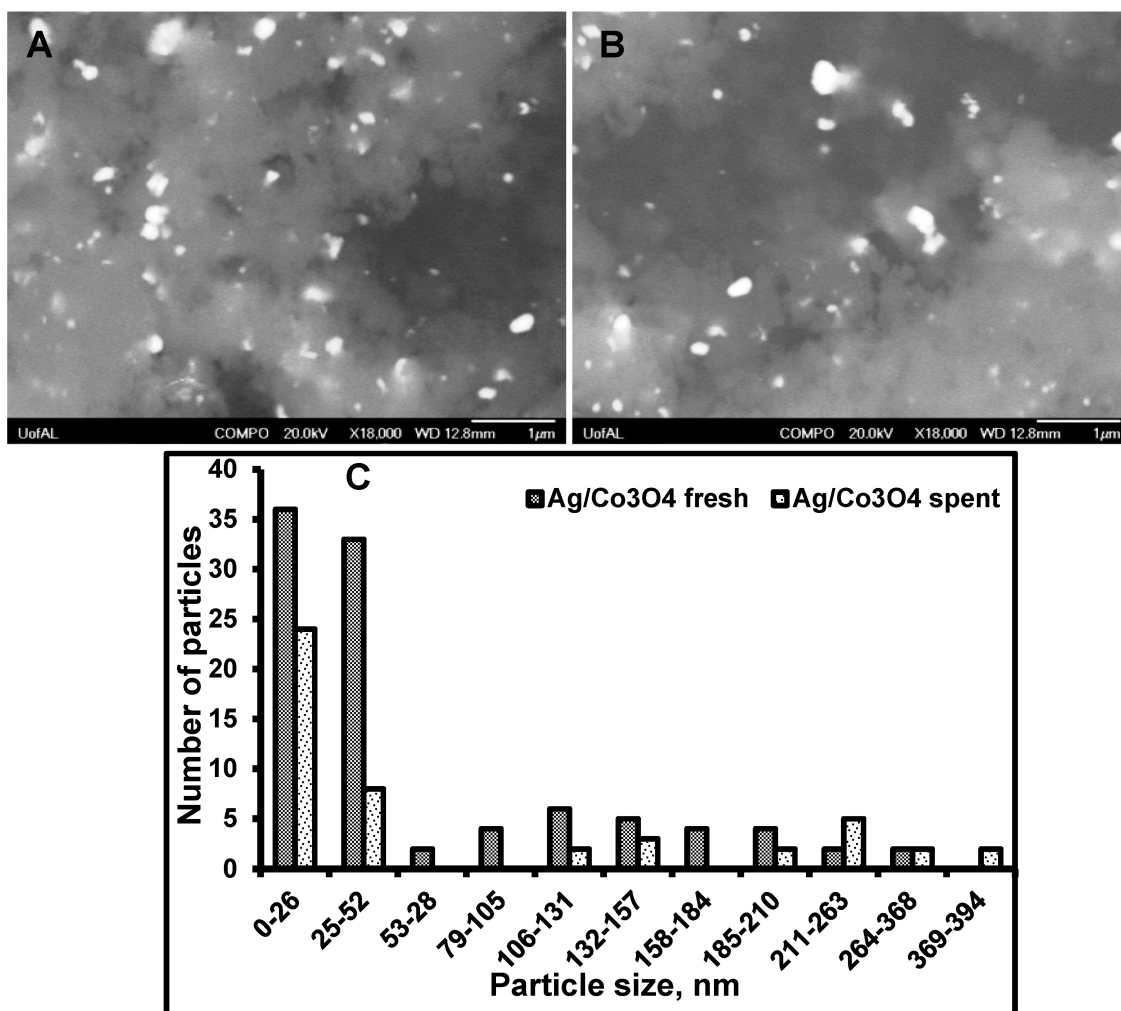
the range of 235–270 nm were seen in the spent catalyst. Overall, there is a significant decrease in the number of Ag nanoparticles, accompanied by an increase in average nanoparticle size, as well as the observation of agglomerates for the spent catalyst.



**Figure 7.** SEM particle size distributions of (A) fresh; (B) spent Ag/SiO<sub>2</sub> monoliths; (C) particle size distribution of fresh and spent Ag/SiO<sub>2</sub>.

A similar analysis for Ag/Co<sub>3</sub>O<sub>4</sub> is shown in Figure 8. Again, the distribution of Ag nanoparticles is bimodal, even more so than for Ag/SiO<sub>2</sub>, with a majority of the Ag nanoparticles having diameters <52 nm. Along with the fine Ag nanoparticles, there is also a broad distribution of Ag nanoparticles centered at 106–131 nm, extending out to 368 nm. The number of small (<52 nm) Ag nanoparticles in the spent Ag/Co<sub>3</sub>O<sub>4</sub> was found to decrease to 50% of that in the fresh Ag/Co<sub>3</sub>O<sub>4</sub>. Similarly, the average size distribution for the larger Ag nanoparticles was also observed to shift to larger average particle sizes, with some agglomerates in the 369–394 nm range being observed in the spent Ag/Co<sub>3</sub>O<sub>4</sub>. The overall number of Ag nanoparticles was also observed to decrease. Ag/Co<sub>3</sub>O<sub>4</sub> shows the same trends as found for Ag/SiO<sub>2</sub>: fewer, larger, Ag nanoparticles in the spent catalyst compared to fresh catalyst. Ag/Co<sub>3</sub>O<sub>4</sub> shows a larger overall decrease in the number of Ag nanoparticles, which may reflect

differences in metal-support interactions or could simply be a result of how the initially different particle size distribution changes during the course of the reaction.



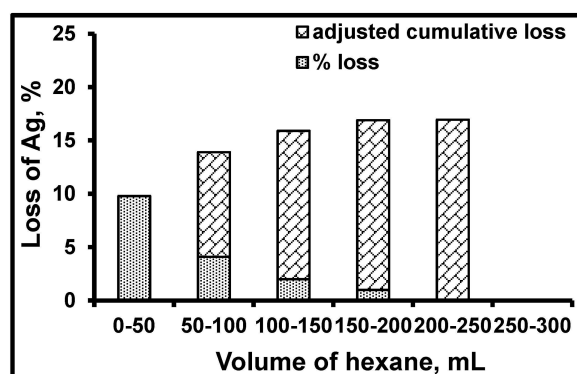
**Figure 8.** SEM particle size distributions of fresh (A) and spent (B) Ag/Co<sub>3</sub>O<sub>4</sub> monoliths; (C) particle size distribution of fresh and spent Ag/Co<sub>3</sub>O<sub>4</sub>.

The extent of the changes in the number and size of the Ag nanoparticles is large enough that this appears likely to explain the changes in the catalytic activity and extent of dye adsorption observed.

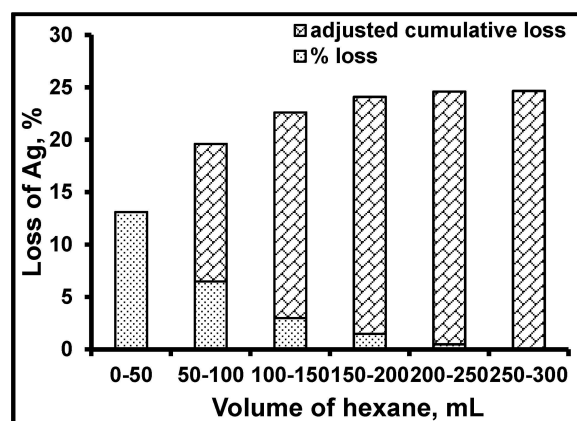
The changes in the number and size of the Ag nanoparticles could result from a number of physical mechanisms. The observation of large nanoparticles that are clearly agglomerates suggests that the collision and merging of two or more nanoparticles into an agglomerate (with a smaller surface area) is likely to be occurring. This could be either by the movement of nanoparticles across the surface of the support or by the nanoparticle being dislodged from the support and suspended in the solvent. If the latter is occurring, then some loss of Ag from the monolith would be expected. We will refer to this mechanism as “washing” of the nanoparticles. A second possible mechanism would involve Ostwald ripening. Dissolution/ionization of Ag would occur preferentially from the highest curvature surfaces, and precipitation/reduction would occur preferentially on the Ag surfaces with the lowest surface energy/surface curvature. The formation of Ag ions is believed to play an important role in the bactericidal properties of Ag nanoparticles [27], although the presence of NaBH<sub>4</sub> would be expected to

suppress this ionization by reducing the Ag ions back to Ag. We will refer to this mechanism as “leaching”, and it would also be predicted to result in a net loss of Ag from the monolith.

In order to evaluate the relative importance of these possible mechanisms, the amount of Ag lost when different solvents were passed through the column was determined. The use of a non-polar organic solvent, such as hexane, would be expected to substantially reduce or possibly completely suppress any Ag ionization. For such solvents, the “leaching contribution” should be significantly reduced, if not completely eliminated. “Washing” of Ag nanoparticles would be expected to be relatively unaffected by the nature of the solvent. Neat hexane was flowed through fresh Ag/SiO<sub>2</sub> and Ag/Co<sub>3</sub>O<sub>4</sub> monoliths at the same constant rate as used in the flow experiments, and the effluent was collected and analyzed using AAS to evaluate the concentration of washed Ag particles. The percent loss and cumulative loss of Ag nanoparticles during the washing with hexane are shown in Figures 9 and 10.



**Figure 9.** Percent loss of Ag from Ag/SiO<sub>2</sub> under flow conditions using hexane as the solvent.



**Figure 10.** Percent loss of Ag from Ag/Co<sub>3</sub>O<sub>4</sub> under flow conditions using hexane as the solvent.

From Figures 9 and 10, it is clear that Ag nanoparticles are being suspended in the hexane and that some fraction is then being washed completely out of the monolith.

The particle size distribution studies of fresh and spent Ag/SiO<sub>2</sub> and Ag/Co<sub>3</sub>O<sub>4</sub> for the hexane washing experiments are shown in Figures 11 and 12, respectively. For both Ag/SiO<sub>2</sub> and Ag/Co<sub>3</sub>O<sub>4</sub>, the particle size distribution shifts markedly towards larger particle sizes. Larger Ag particles in the 300–500-nm range were observed in the spent catalysts, which were not seen in fresh catalysts. In total, 1/6th of the Ag particles were washed from the fresh Ag/SiO<sub>2</sub> and 1/4th of the initial Ag particles were washed from the fresh Ag/Co<sub>3</sub>O<sub>4</sub> when hexane was used. This is significantly less than

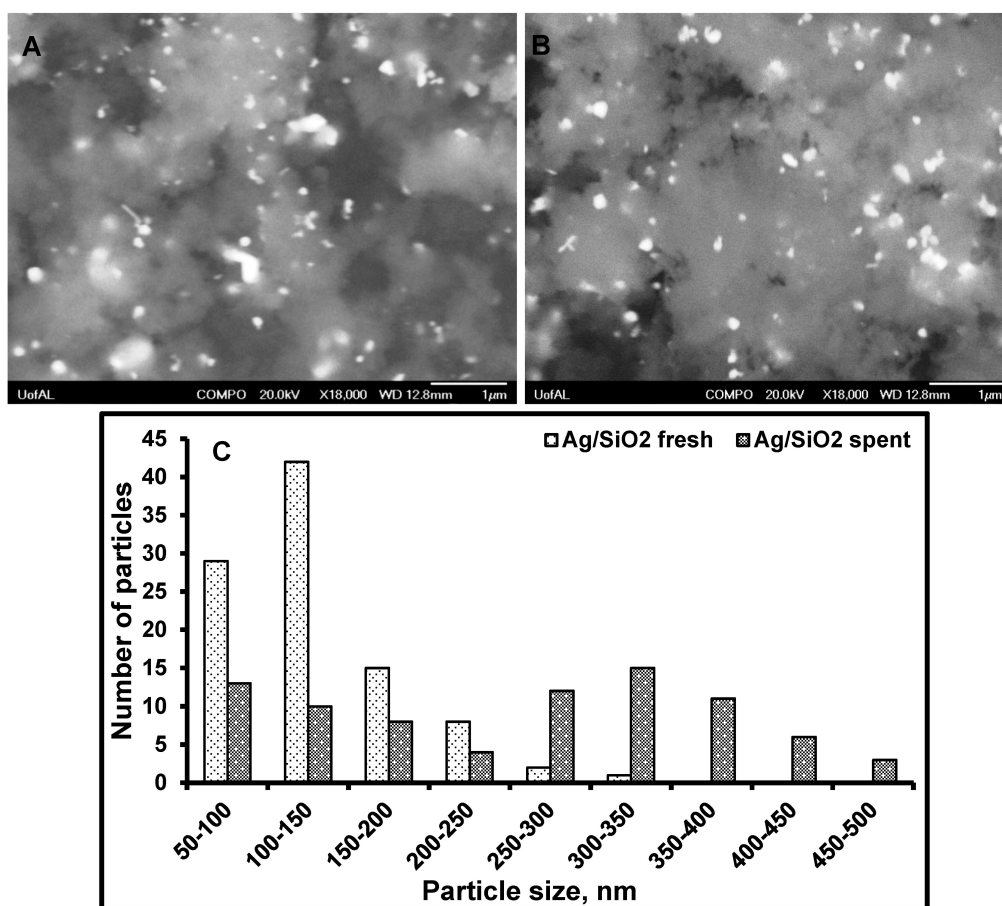
seen for aqueous media. This indicates that there is significant leaching/dissolution occurring in water, such that some of the small Ag nanoparticles are not washed off, but rather dissolved completely. The growth of larger agglomerates for the hexane-washed samples suggests that many nanoparticles that might otherwise dissolve are instead washed into contact with other Ag nanoparticles, forming the larger agglomerates observed.

From these observations, it is evident that in aqueous media, the loss of Ag particles is due to the washing of the loosely-bound Ag nanoparticles along with the leaching of Ag ions due to their dissolution in aqueous medium. The contribution of the leaching to Ag loss can be estimated as the difference between the loss of Ag from the aqueous samples and that washed off in hexane. These values are given in Table 1.

Comparison of the data for Ag/SiO<sub>2</sub> and Ag/Co<sub>3</sub>O<sub>4</sub> shows that a significantly larger portion of the Ag is leached from the Co<sub>3</sub>O<sub>4</sub> than from Ag/SiO<sub>2</sub>, suggesting that there is some metal-support interaction.

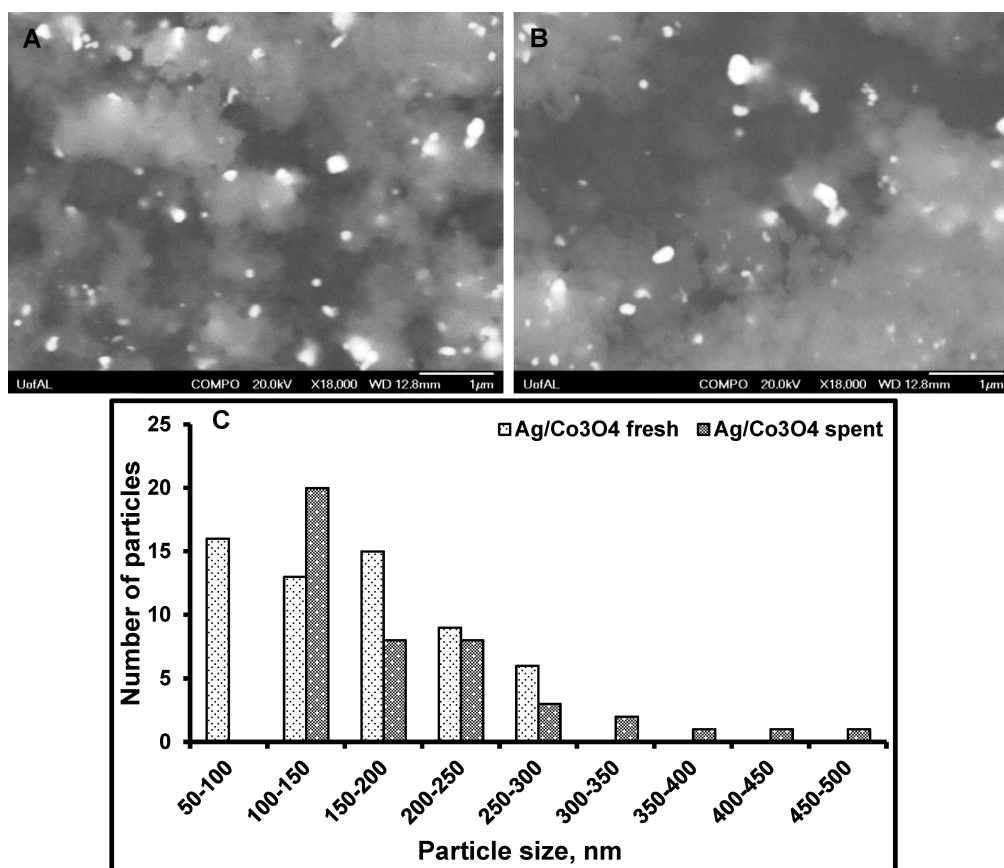
**Table 1.** Percentage loss of washed and leached Ag from SiO<sub>2</sub> and Co<sub>3</sub>O<sub>4</sub> monolithic supports.

Monolithic Catalyst	Number of Ag Nanoparticles Lost from Support, %		Ag Leached from Support in Aqueous Reaction Medium, %
	Polar Solvent (Water)	Non-Polar Solvent (Hexane)	
Ag/SiO <sub>2</sub>	27	17	10
Ag/Co <sub>3</sub> O <sub>4</sub>	44	25	19



**Figure 11.** Washing studies of Ag/SiO<sub>2</sub> in hexane: (A) SEM of fresh Ag/SiO<sub>2</sub>; (B) SEM of spent Ag/SiO<sub>2</sub>; (C) particle size distribution of fresh and spent Ag/SiO<sub>2</sub>.





**Figure 12.** Washing studies of Ag/Co<sub>3</sub>O<sub>4</sub> in hexane: (A) SEM of fresh Ag/Co<sub>3</sub>O<sub>4</sub>; (B) SEM of spent Ag/Co<sub>3</sub>O<sub>4</sub>; (C) particle size distribution of fresh and spent Ag/Co<sub>3</sub>O<sub>4</sub>.

From the data, it is clear that for the hierarchically porous catalysts studied, the loss of catalytic activity is caused by both washing and leaching of the Ag nanoparticles. For non-polar solvents where leaching is expected to be much less prevalent, agglomeration may not produce such a large change in metal surface area as it would for aqueous media, since there would be less Ostwald ripening [28]. This would predict that in non-aqueous media, the decrease in catalytic activity would be closer to proportionate to the amount of metal lost. It is also likely that the flow rate could impact the fraction of particles that are washed off, as higher flow rates would be expected to dislodge more strongly-bound particles.

#### 4. Conclusions

Silica and cobalt oxide monoliths supporting silver nanoparticles were successfully used as continuous flow monolithic microreactors for the hydrogenation of eosin-Y by sodium borohydride. The dye was found to adsorb on silver nanoparticles in the absence of reducing agent. Catalysis and absorption occurred simultaneously on the silver-based monoliths in the presence of reducing agent. Ag/Co<sub>3</sub>O<sub>4</sub> was found to be more active than Ag/SiO<sub>2</sub> for the hydrogenation of eosin-Y. A loss of catalytic activity was observed for silver nanoparticles on both silica and cobalt oxide supports in the aqueous reaction media. From the comparison of the particle size distributions of the spent and fresh catalysts and the impact of the solvent, it was concluded that the loss of activity is a result of the loss of silver from the support and of particle size growth due to washing and leaching.

## Acknowledgments

Support by the Research Stimulation Program at The University of Alabama (T.V.K.) is gratefully acknowledged. A seed grant from the Central Analytical Facility at The University of Alabama is acknowledged, the required cost match for which was provided by the Bakker Research Gift fund.

## Author Contributions

T.V.K. and M.G.B. designed the study and constructed the flow reactor. Y.H. carried out the catalyst preparation and the flow studies. Y.H., T.V.K. and M.G.B. carried out the data analysis and preparation of the manuscript.

## Conflicts of Interest

The authors declare no conflict of interest.

## References

1. Wirth, T. *Microreactors in Organic Chemistry and Catalysis*, 2nd ed.; Wiley: New York, NY, USA, 2013; p. 478.
2. Liu, X.; Unal, B.; Jensen, K.F. Heterogeneous catalysis with continuous flow microreactors. *Catal. Sci. Technol.* **2012**, *2*, 2134–2138.
3. Escriba, M.; Ibadon, A.O. Recent developments in catalytic micro process engineering for fine chemicals synthesis. *Recent Patents Catal.* **2013**, *2*, 101–115.
4. Parlett, C.M.A.; Wilson, K.; Lee, A.F. Hierarchical porous materials: Catalytic applications. *Chem. Soc. Rev.* **2013**, *42*, 3876–3893.
5. Li, X.; Sun, M.; Rooke, J.C.; Chen, L.; Su, B.L. Synthesis and applications of hierarchically porous catalysts. *Chin. J. Catal.* **2013**, *34*, 22–47.
6. Chen, L.-H.; Li, X.-Y.; Rooke, J.C.; Zhang, Y.-H.; Yang, X.-Y.; Tang, Y.; Xiao, F.-S.; Su, B.-L. Hierarchically structured zeolites: Synthesis, mass transport properties and applications. *J. Mater. Chem.* **2012**, *22*, 17381–17403.
7. Nakanishi, K. Pore structure control of silica gels based on phase separation. *J. Porous Mat.* **1997**, *4*, 67–112.
8. Smått, J.-H.; Sayler, F.M.; Grano, A.; Bakker, M.G. Formation of hierarchically porous metal oxide and metal monoliths by nanocasting into silica monoliths. *Adv. Eng. Mat.* **2012**, *14*, 1059–1073.
9. Grano, A.J.; Sayler, F.M.; Smått, J.-H.; Bakker, M.G. Hierarchically porous monoliths of carbon and metal oxides with ordered mesopores. *J. Porous Mat.* **2014**, *21*, 1113–1122.
10. Sayler, F.M.; Grano, A.J.; Smått, J.-H.; Lindén, M.; Bakker, M.G. Nanocasting of hierarchically porous NiO, Co<sub>3</sub>O<sub>4</sub>, Ni, Co and Ag monoliths: Impact of processing conditions on fidelity of replication. *Micropor. Mesopor. Mater.* **2014**, *184*, 141–150.
11. Grano, A.J.; Sayler, F.M.; Smått, J.-H.; Bakker, M.G. Alternative etching methods to expand nanocasting, and use in the synthesis of hierarchically porous nickel oxide, zinc oxide and copper monoliths. *J. Mater. Res.* **2013**, *28*, 2483–2489.

12. Konishi, J.; Fujita, K.; Nakanishi, K.; Hirao, K. Monolithic  $\text{TiO}_2$  with controlled multiscale porosity via a template-free sol-gel process accompanied by phase separation. *Chem. Mater.* **2006**, *18*, 6069–6074.
13. Chuenchom, L.; Kraehnert, R.; Smarsly, B.M. Recent progress in soft-templating of porous carbon materials. *Soft Matter* **2012**, *8*, 10801–10812.
14. Fang, B.; Kim, J.H.; Kim, M.-S.; Yu, J.-S. Hierarchical nanostructured carbons with meso-macroporosity: Design, characterization, and applications. *Acc. Chem. Res.* **2013**, *46*, 1397–1406.
15. Halttunen, M.E.; Niemelä, M.K.; Krause, A.O.I.; Vaara, T.; Vuori, A.I. Rh/C catalysts for methanol hydrocarbonylation i. Catalyst characterisation. *Appl. Catal. A* **2001**, *205*, 37–49.
16. Magano, J.; Dunetz, J.R. Large-scale applications of transition metal-catalyzed couplings for the synthesis of pharmaceuticals. *Chem. Rev.* **2011**, *111*, 2177–2250.
17. Neves, A.N.C.B.; Calvete, M.R.J.F.; Pinho e Melo, T.M.V.D.; Pereira, M.M. Immobilized catalysts for hydroformylation reactions: A versatile tool for aldehyde synthesis. *Eur. J. Org. Chem.* **2012**, *2012*, 6309–6320.
18. Thomas, J.M.; Johnson, B.F.G.; Raja, R.; Sankar, G.; Midgley, P.A. High-performance nanocatalysts for single-step hydrogenations. *Acc. Chem. Res.* **2003**, *36*, 20–30.
19. Alonso, F.; Riente, P.; Yus, M. Nickel nanoparticles in hydrogen transfer reactions. *Acc. Chem. Res.* **2010**, *44*, 379–391.
20. De Jong, K.P. *Synthesis of Solid Catalysts*; Wiley-VCH: Weinheim, Germany, 2009.
21. Takahashi, R.; Sato, S.; Tomiyama, S.; Ohashi, T.; Nakamura, N. Pore structure control in Ni/SiO<sub>2</sub> catalysts with both macropores and mesopores. *Micropor. Mesopor. Mater.* **2006**, *98*, 107–114.
22. Zhang, S.; Chen, L.; Zhou, S.; Zhao, D.; Wu, L. Facile synthesis of hierarchically ordered porous carbon via *in situ* self-assembly of colloidal polymer and silica spheres and its use as a catalyst support. *Chem. Mater.* **2010**, *22*, 3433–3440.
23. El Kadib, A.; Chimenton, R.; Sachse, A.; Fajula, F.; Galarneau, A.; Coq, B. Functionalized inorganic monolithic microreactors for high productivity in fine chemicals catalytic synthesis. *Angew. Chem. Int. Ed.* **2009**, *121*, 5069–5072.
24. Numata, M.; Takahashi, R.; Yamada, I.; Nakanishi, K.; Sato, S. Sol-gel preparation of Ni/TiO<sub>2</sub> catalysts with bimodal pore structures. *Appl. Catal. A* **2010**, *383*, 66–72.
25. Linares, N.; Hartmann, S.; Galarneau, A.; Barbaro, P. Continuous partial hydrogenation reactions by Pd@unconventional bimodal porous titania monolith catalysts. *ACS Catal.* **2012**, *2*, 2194–2198.
26. Hakat, Y.; Kotbagi, T.V.; Bakker, M.G. Catalytic activity of Ag/SiO<sub>2</sub> and Ag/Co<sub>3</sub>O<sub>4</sub> hierarchically porous monoliths for hydrogenation of dyes. *Curr. Catal.* **2014**, *3*, 286–295.
27. Morones, J.R.; Elechiguerra, J.L.; Camacho, A.; Holt, K.; Kouri, J.B.; Ramírez, J.T.; Yacaman, M.J. The bactericidal effect of silver nanoparticles. *Nanotechnology* **2005**, *16*, 2346–2353.
28. Evans, J.W.; Thiel, P.A. A little chemistry helps the big get bigger. *Science* **2010**, *330*, 599–600.