

## Article

# Paper-Based Vapor Detection of Formaldehyde: Colorimetric Sensing with High Sensitivity

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**Abstract:** We report on a novel colorimetric sensor system for highly sensitive detection of formaldehyde (FA) in the gas phase. The sensor is constructed with paper towel as a substrate coated with the sulfuric acid salt of hydroxylamine ((NH<sub>2</sub>OH)<sub>2</sub>·H<sub>2</sub>SO<sub>4</sub>) together with two pH indicators, bromophenol blue and thymol blue. Upon exposure to FA, the hydroxylamine will react with the absorbed FA to form a Schiff base (H<sub>2</sub>C=N-OH), thus releasing a stoichiometric amount of sulfuric acid, which in turn induces a color change of the pH indicator. Such a color change was significantly enriched by incorporating two pH indicators in the system. With the optimized molar ratio of the two pH indicators, the color change (from brown to yellow, and to red) could become so dramatic as to be visible to the eye depending on the concentration of FA. In particular, under 80 ppb of FA (the air quality threshold set by WHO) the color of the sensor substrate changes from brown to yellow, which can even be envisioned clearly by the naked eyes. By using a color reader, the observed color change can be measured quantitatively as a function of the vapor concentration of FA, which produces a linear relationship as fitted with the data points. This helps estimate the limit of detection (LOD), to be 10 ppb under an exposure time of 10 min, which is much lower than the air quality threshold set by WHO. The reported sensor also demonstrates high selectivity towards FA with no color change observed when exposed to other common chemicals, including solvents and volatile organic compounds. With its high sensitivity and selectivity, the proposed paper-based colorimetric sensor thus developed can potentially be employed as a low-cost and disposable detection kit that may find broad application in detecting FA in indoor air and many other environments.

**Keywords:** vapor detection; formaldehyde; colorimetric; paper-based sensor; hydroxylamine



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## 1. Introduction

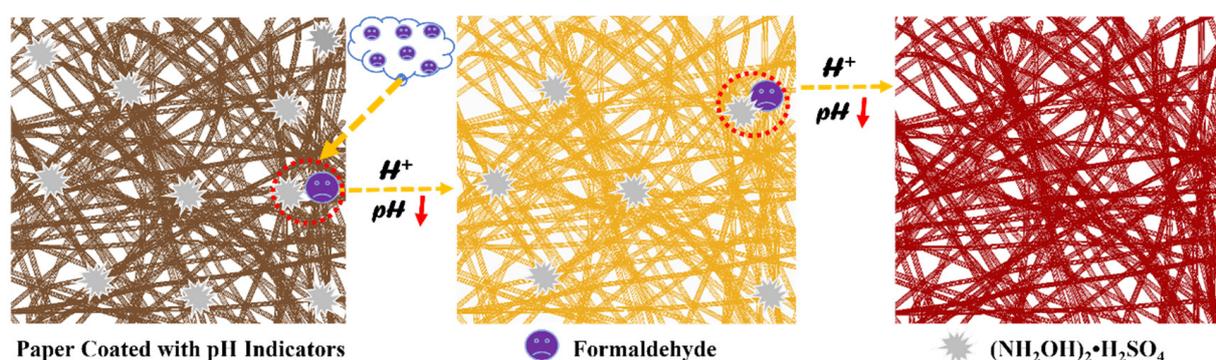
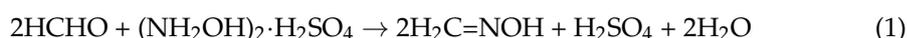
Formaldehyde (FA), a colorless and reactive carbonyl compound that can be produced from various industrial processes via direct release or decomposition of chemical additives. FA has been widely used as a chemical reagent in wood-processing, textiles, construction, medicines, and among others [1]. The wide variety of use provides numerous sources for FA to be released into air environment, resulting in a severe pollution problem. Exposure to FA may cause various adverse health effects, such as skin and eye irritation, central nervous system injury, blindness, and carcinogenicity [2–4]. According to the guidelines set by the World Health Organization (WHO), the safe exposure level of FA in indoor air should not exceed 80 ppb within 30 min [5]. Thus, it is of great significance and urgent to develop gaseous sensors to detect FA at low concentration levels that can be employed for quick, onsite air quality detection, thus helping ensure public health.

Many instrumental analysis methods such as spectrophotometry, gas chromatography, and others derived from them have been developed and applied in FA vapor detection. However, most of these instruments are expensive, complicated to operate, and have time consuming sample treatment and detection processes, and are thus not well suited for the quick, portable detection that is crucial for real-time detection of air quality, particularly in remote areas [6–8]. In contrast, chemosensors based on electrochemistry, surface acoustic waves, chemiresistors and other methods provide great potential for being developed into a device that is small in size, and easier to use [9–11], though further development of these sensors is hampered by the typical poor selectivity and low sensitivity (with the limit of detection, LOD usually >300 ppb for detection of FA). Among all the chemosensors developed thus far, optical chemosensors based on colorimetric or fluorescence modulation usually provide relatively more options for improving the detection sensitivity and selectivity by unique molecular design and structural engineering of the sensors [12–15], which can in turn enhance the chemical binding with gas analytes like FA. Optical chemosensors are also simple in sensing response, ideally based on direct, one-step interaction with the analyte, and can be fabricated into a small, easy-to-use system in low cost, which is adaptable for portable detection or integration into a current air inspection system for quick, quantitative detection of a specific pollutant like FA.

In comparison to fluorescence-based sensors, colorimetric sensors can become more competitive with simplicity in device construction or system miniaturization by removing the need of photoexcitation and detection parts as required for a fluorescence sensor. Indeed, the signal change of colorimetric chemosensors can often be visually recognized by naked eyes without involving complicated optics, thus making the whole detection system simple and quick in operation, as well as low cost. At present, colorimetric sensors for the detection of FA are mostly based on some color change induced directly or indirectly by the chemical reactions between FA and the sensor components, such as pH indicators [16–20], chromotropic acid [21,22], Fluoral-P [23] or Tollens' reactions [24–26]. Among these sensors, aldimine condensation, i.e., reaction between aldehyde and primary amine producing a Schiff base ( $R-C=N-R'$ ), represents a simple, and highly selective sensing approach to detecting aldehydes. In particular, the acidic salts of hydroxylamine have attracted increasing interest in recent years in gas phase detection of FA as the sensing reaction (response) can occur quickly, and the acid thus produced can be measured simultaneously using pH indicators or chemiresistive sensors based on carbon nanotubes or other materials that are sensitive to interactions with acids (protons) [13,27]. Although the chemiresistive sensors with carbon nanotubes gave low LOD for detection of FA (16 ppb), the colorimetric sensors based on pH indicators showed much less sensitivity, with LOD barely around the WHO safety threshold (80 ppb) set for FA [20]. The low detection sensitivity is mainly due to that only one pH indicator (methyl yellow) was employed in the sensor, and the homochromatic response generated from the single-component was not sufficient for detecting small changes in pH [28]. Therefore, it remains imperative to explore new colorimetric compounds that may provide a heterochromatic response, thus enhancing the measurement of color gradient and further improving the detection sensitivity for FA.

Herein, we report on a novel colorimetric sensor system that can detect FA down to a LOD of 10 ppb with only 10 min of exposure. As illustrated in Figure 1, the sensor is composed of a paper towel coated with the sulfate salt of hydroxylamine  $(NH_2OH)_2 \cdot H_2SO_4$  mixed with two pH indicators, thymol blue (TB) and bromophenol blue (BPB). The fibril framework of paper towel not only allows for homogeneous coating of the sensing reagents, but also provides open, interconnected porosity for expedient diffusion of gas analytes, which is crucial for vapor detection of chemicals like FA. The colorimetric sensing relies on the specific aldimine condensation reaction between  $(NH_2OH)_2 \cdot H_2SO_4$  and FA to form the corresponding Schiff base ( $H_2C=N-OH$ ), thus a releasing stoichiometric amount of  $H_2SO_4$  (Equation (1)), which in turn induces a color change of the two pH indicators. With appropriate optimization of the molar ratio of the two pH indicators, the color change upon exposure to FA can be maximized in terms of both color gradient and the number

of discernible colors involved. This results in highly efficient detection of FA with LOD down to 10 ppb, which is among the lowest for all the colorimetric sensors reported so far for FA [13]. The sensor also demonstrates high selectivity towards FA in the presence of other common organic solvents and volatile chemicals. Moreover, the enriched color changes associated with the multiple pH indicators used will also help improve the visual resolution for direct detection by naked eyes. Combination of the high sensitivity, selectivity, simplicity and low cost makes the paper-based colorimetric sensor reported herein to have a great potential to be broadly applied in air quality inspection for FA. As the aldimine condensation reaction between  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  and FA is irreversible, the paper-based sensor thus developed can be employed as a disposable detection kit, taking advantage of the low-cost.



**Figure 1.** Schematic illustration of the colorimetric sensor for FA based on paper towel coated with  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  and two pH indicators, TB and BPB.

## 2. Experimental Methods

### 2.1. Materials

Formalin solution (37 wt.% aqueous solution of FA) and  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  were purchased from J&K Scientific Ltd. (Beijing, China). Glycerin, methanol, sodium bicarbonate ( $\text{NaHCO}_3$ ), and other chemicals were all in AR grade and purchased from Aladdin Ltd., (Shanghai, China). The gas sample, such as  $\text{CO}_2$ ,  $\text{SO}_2$ , etc. were purchased from Dalian Special Gases Ltd. (Liaoning, China). The paper towels (white, K-MB05) were purchased from Hangzhou Yuanmian Technology Ltd. (Zhejiang, China). Ultrapure water was generated by an ultrapure water treatment system (EPED-S2-90DF, Nanjing EPED Technology Ltd., Jiangsu, China). The optical images of sensor strips (cut from paper towels) were recorded with a digital camera (Canon 750D; Canon, Tokyo, Japan).

### 2.2. Fabrication of Colorimetric Sensing Strips

The sensor was fabricated as a strip ( $1.2 \text{ cm} \times 1.2 \text{ cm}$ ) cut from a paper towel sheet and coated with  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  mixed with two pH indicators, TB and BPB. The paper towel substrate possesses a three-dimensional (3D) interconnected network formed by the intertwined cellulose microfibers. Such a microstructure provides not only large surface area, but also continuous open porosity that is conducive to the diffusion and absorption of gas analytes. Such 3D bulk diffusion allows us to use thick paper towel in order to increase the color density, which helps improve the measurement accuracy, and thus detection sensitivity. Moreover, the paper towel and the sensor reagents involved are all in extremely low cost, especially considering the small size and amount used. This will facilitate the development of the sensor composite into a disposable detection kit, making the onsite measurement to be even easier. In this study, three brands of paper towels were chosen for initial sensor testing, and all the samples showed similar color responses, as shown in Figure S1 in the Supporting Information. One of the three paper towels (No. 1)

possesses a relatively tighter and more uniform fibril network structure, which combined helps produce stable, homogeneous coating of  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  and the pH indicators as revealed by the uniform color distribution throughout the whole area both before and after exposure to FA (Figure S1).

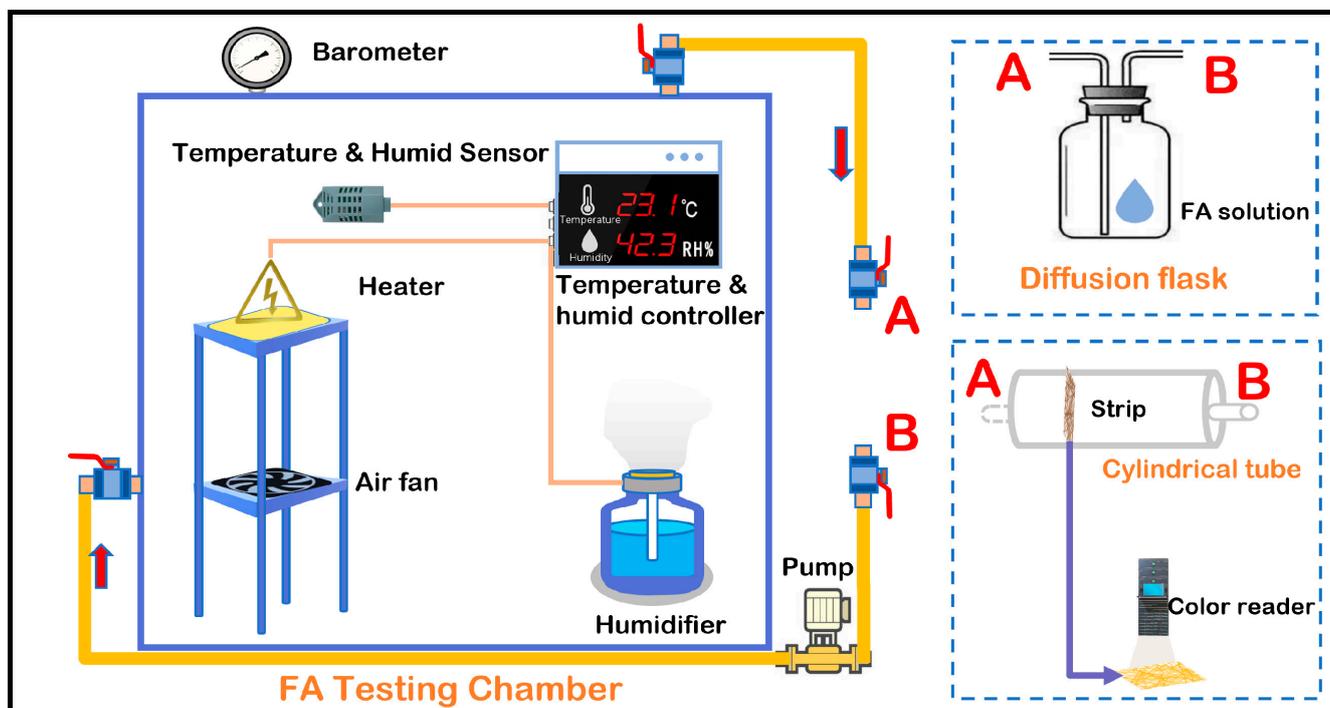
$(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  was chosen as the hydroxylamine acidic salt for development as the detection kit mainly because of its relatively large range of pH change upon reaction with FA, as well as its high solubility and stability in aqueous solution for repeatable sample preparation. Four other hydroxylamine acidic salts (Table S1) were also selected for screening characterization in aqueous solution under the same conditions. Upon reaction with the same stoichiometric amount of FA, the solution of  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  demonstrated the largest scale of pH change, from pH 4.0 to 1.3, whereas all the other four salts showed only one unit of pH change. The large scale of pH change allows for more options of pH indicators to be selected. In this study, TB and BPB were selected as the pH indicators to provide colorimetric change corresponding to the large pH change associated with  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$ . The color of TB is pink at pH below 1.2, changes to orange at pH 1.2–2.8, and finally becomes yellow at pH above 2.8. The color of BPB is yellow at pH below 3.0, changes to turquoise at pH 3.0–4.6, and finally becomes blue at pH above 4.6. Using the two pH indicators in the system would provide multiple colors and combinations along the pH change from 4.0–5.0 to ~1.0, which covers the range of sensor response between  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  and FA as fabricated in the paper towel kit (vide infra). To facilitate the absorption of FA, a hydrophilic reagent such as glycerin was also incorporated into the sensor composition. The high hygroscopicity of glycerin also helps maintain water in the film, which is conducive to the molecular diffusion to expedite the aldimine condensation reaction (Equation (1)), and diffusion of protons to trigger the color change of pH indicators.

In a typical process of fabricating the sensor strip, a piece of paper towel (cut in 1.2 cm × 1.2 cm square shape) was immersed for 20 min in a solution that contains hydroxylamine sulfate, pH indicators and glycerin. The strip was then taken out of the solution and dried under vacuum at 30 °C for 30 min. The immersion solution was prepared by dissolving TB (32.1 μmol; 0.015 g) and BPB (32.1 μmol; 0.021 g) in 45 mL methanol, followed by adding 15 mL of water, hydroxylamine sulfate (3.6 mmol; 0.6 g) and 10 mL of glycerin were added. The solution was then ultrasonicated for 5 min before immersing in the paper towel strip. The pH of the solution can be adjusted as needed by adding appropriate amount of  $\text{NaHCO}_3$ . Although the above preparation process used equal molar amount of the two pH indicators (1:1), the molar ratio of TB and BPB was adjustable. Indeed, in this study we have also fabricated three more sensor strips with TB/BPB ratios of 1:1.5, 1.5:1 and 2:1, and tested their colorimetric sensing along with the one of 1:1 ratio, aiming to identify the composition that gives the best sensor performance.

### 2.3. Construction of an Environment Chamber for FA Vapor Generation and Sensor Testing

An environmental chamber (Figure 2) was designed and constructed to produce FA vapor at varying concentrations under highly controlled temperature and humidity conditions. The chamber consisted of a transparent polymethyl methacrylate (PMMA) box with a volume of 252 L, which houses a heater, a humidifier, temperature and humidity sensors both connected to a controller for adjusting and controlling the inside temperature and humidity. A micro-fan is also included for facilitating the diffusion of gas species, producing homogeneous concentration distribution. The attached pumping system allows for easy, quick generation of FA vapor by connecting to a diffusion flask containing a solution of FA. Varying vapor concentrations of FA in the chamber can be generated by purging the head-space FA from the diffusion flask containing different volumes of formalin solution with a constant flow of air for a controlled length of time. The vapor concentration of FA generated can be calibrated by a PPM-400ST FA Metre (PPM Technology Ltd., Caernarfon, UK). Replacing the diffusion flask and reconnecting with a glass tube containing the paper-based sensor strip allows for testing of the colorimetric response towards FA. The sensor

strip is fixed at a right position in the glass tube allowing for maximal surface contact to the air flow containing FA. The closed flow circulation helps maintain concentration equilibrium between the chamber and the glass tube, as well as expedite the adsorption of FA. After certain amount of time (say 10 min), the sensor strip will be taken out and measured by a color reader. Under such a flow circulation system, the sensor exposure time can be easily adjusted and controlled.



**Figure 2.** Schematic illustration of the environmental chamber for generation of FA vapor and testing of the colorimetric sensor.

#### 2.4. Measurement of Color Change

The color change of sensor strips was measured by a color reader (3nh, Model NR60CP, Shenzhen 3nh Technology Ltd., Guangdong, China). A D65 light source normalized by the International Commission on Illumination (CIE) was used in the color reader, which closely matches the daylight, helping minimizing the measurement errors due to lighting fluctuation [29,30].  $\Delta E^*_{ab}$  value as an index to describe the total color difference (TCD) [31], is calculated following (Equation (2)):

$$\Delta E^*_{ab} = [(\Delta L)^2 + (\Delta a)^2 + (\Delta b)^2]^{1/2} \quad (2)$$

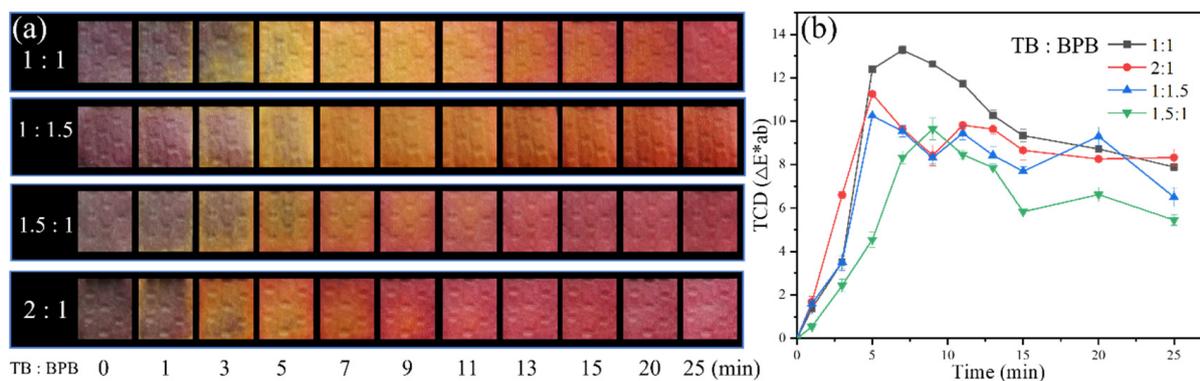
where  $\Delta L$  is the brightness,  $\Delta a$  the red-green degree, and  $\Delta b$  the yellow-blue degree.

### 3. Results and Discussions

#### 3.1. Effect of TB/BPB Ratio on Colorimetric Sensing

The color change of the sensor strips was found to be highly affected by the molar ratio of the two pH indicators used, TB/BPB. As shown in Figure 3a, for all the four strips with different molar ratios of TB/BPB (1:1, 1:1.5, 1.5: 1 and 2:1) a distinct color change could be easily envisioned, even by naked eyes, when exposed to 10 ppm FA for varying times. Remarkably, all the strips at different exposure stages showed uniform color density throughout the whole area, indicating homogeneous distribution of TB/BPB and  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  within the paper matrix. By measuring the total color difference using a color reader, the value of  $\Delta E^*_{ab}$  can be plotted as a function of exposure time as shown in Figure 3b, from which one can clearly identify the best TB/BPB ratio of 1:1 that

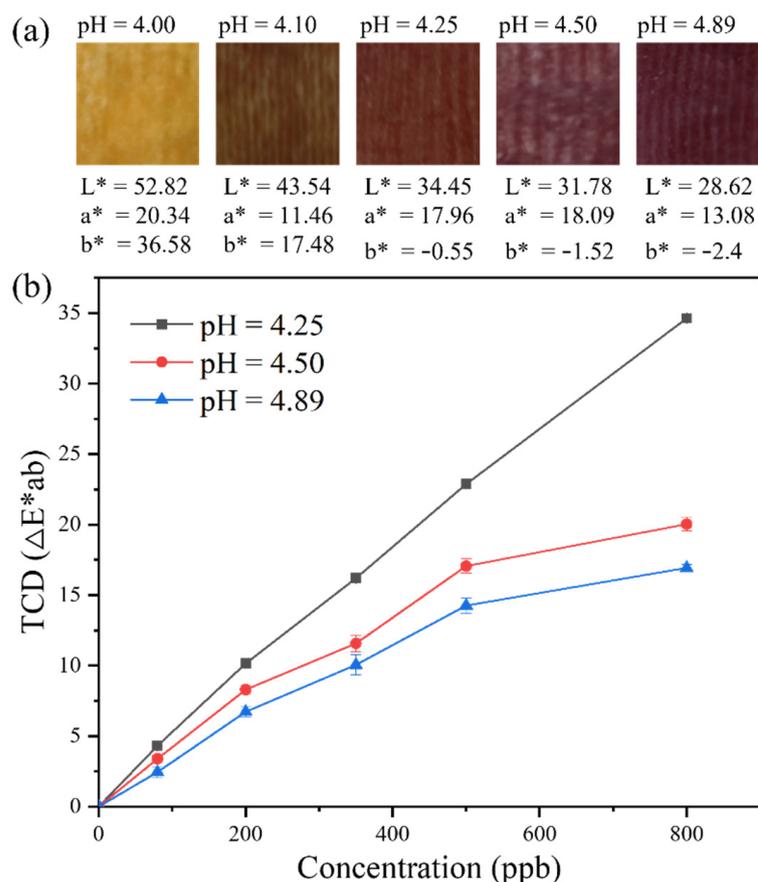
demonstrates the strongest color discrimination as indicated by the highest value of  $\Delta E^*ab$ . For all the strips under test, the maximal value of  $\Delta E^*ab$  was reached after 5–10 min of exposure. Further exposure to FA produces more acids, or lower pH, wherein the combined color of TB and BPB becomes less distinct from the initial color before exposure. In the rest of experiments of this study, we kept the molar ratio of TB/BPB as 1:1 in fabricating the sensor strips.



**Figure 3.** (a) Photos of the four sensor strips with different TB/BPB molar ratios upon exposure to the vapor of FA (10 ppm) for varying times, showing rich color changes. (b) Values of  $\Delta E^*ab$  measured for the samples in (a) plotted as a function of exposure time. The initial pH value of the immersion solution used for preparing the sensor strip was  $4.7 \pm 0.1$ , and all measurement were performed at 25 °C and 60% RH.

### 3.2. Effect of Initial pH on Colorimetric Sensing

As the colorimetric sensing (value of  $\Delta E^*ab$ ) is determined by the color difference of a sensor strip before and after exposure to FA, it is critical to define the initial color of the strip, which in turn can be optimized by adjusting the pH. The color of the pH indicator TB remains as yellow around pH 4.6, while the color of BPB changes from turquoise to blue from below to above pH 4.6, so the color of the strip (defined as the combined color of TB and BPB) may be dramatically adjusted around this pH. As shown in Figure 4a, in this study five sensor strips were fabricated by immersing the pieces of paper towel into the solutions adjusted at five different pH values, 4.00, 4.10, 4.25, 4.50 and 4.89, all containing the same amount of TB/BPB,  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  and glycerol. For the two strips fabricated at pH = 4.00 and 4.10, the pristine color was light yellow and dark yellow, respectively. When exposed to FA, the color would eventually change to yellow as observed in Figure 3, so the color difference attained for these two strips would not be significant, thus not conducive to enhancing the sensing efficiency. As the pH increases, the three strips fabricated at pH 4.25, 4.50 and 4.89 developed a dark brown color with increased optical density. When tested for colorimetric sensing upon exposure to the vapor of FA at varying concentrations (80, 200, 350, 500, 800 ppb) for 10 min, the three strips demonstrated significant color change with the pH 4.25 one showing the highest value of  $\Delta E^*ab$  (Figure 4b). The decreased color change observed for the strips of pH 4.50 and 4.89 is likely due to the increased color density at pristine state, which otherwise diminishes the color difference in comparison to the dark red color obtained after extensive exposure to FA (as shown Figure 3). With the highest colorimetric sensitivity and richness in color changes, the strip fabricated at pH 4.25 has been used for all the subsequent sensor tests.



**Figure 4.** (a) Photos of the five sensor strips fabricated by immersing in the solutions of different pH values. Values marked underneath the photo are for the brightness ( $L^*$ ), position between red and green ( $a^*$ ), and position between yellow and blue ( $b^*$ ). (b) The values of  $\Delta E^*_{ab}$  measured for the three strips after 10 min of exposure to FA vapor at different concentrations (80, 200, 350, 500, 800 ppb). The measurements were performed at 25 °C and 45% RH.

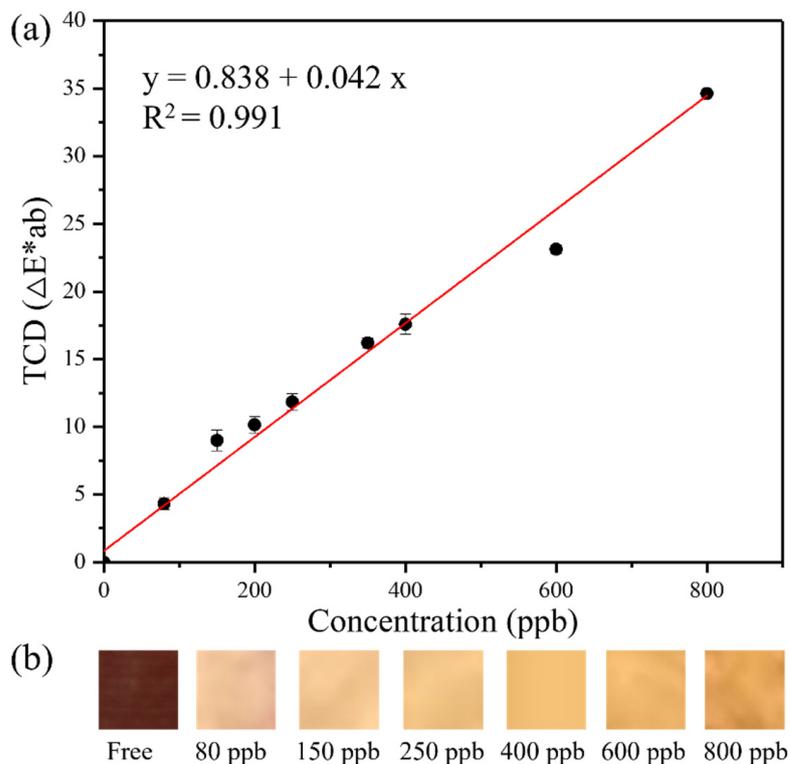
### 3.3. Quantitative Assessment of the Sensitivity of Colorimetric Sensing

Further assessment of the sensing sensitivity was performed on the sensor strip fabricated from the solution of pH 4.25 containing 1:1 TB/BPB as optimized from the above. The color change of the sensor strip was examined by exposure to various concentrations of FA vapor (0–800 ppb) for 10 min. As shown in Figure 5a, the value of  $\Delta E^*_{ab}$  measured by the sensor strip increases linearly with the vapor concentration of FA within the testing range. Linear fitting of the data points gives  $\Delta E^*_{ab} = 0.838 + 0.042C$ , with a correlation coefficient of 0.991. If defining three times the standard deviation as the detectable value of  $\Delta E^*_{ab}$ , the LOD can be projected to be 10 ppb, which is eight times lower than the air quality threshold (80 ppb) set for FA by the WHO. Figure 5b shows the photos of the sensor strips after 10 min of exposure to different concentrations of FA vapor. Even under the lowest concentration of 80 ppb a distinct color change can be easily envisioned by naked eyes, implying the practical feasibility of applying the sensor kit in air detection of FA.

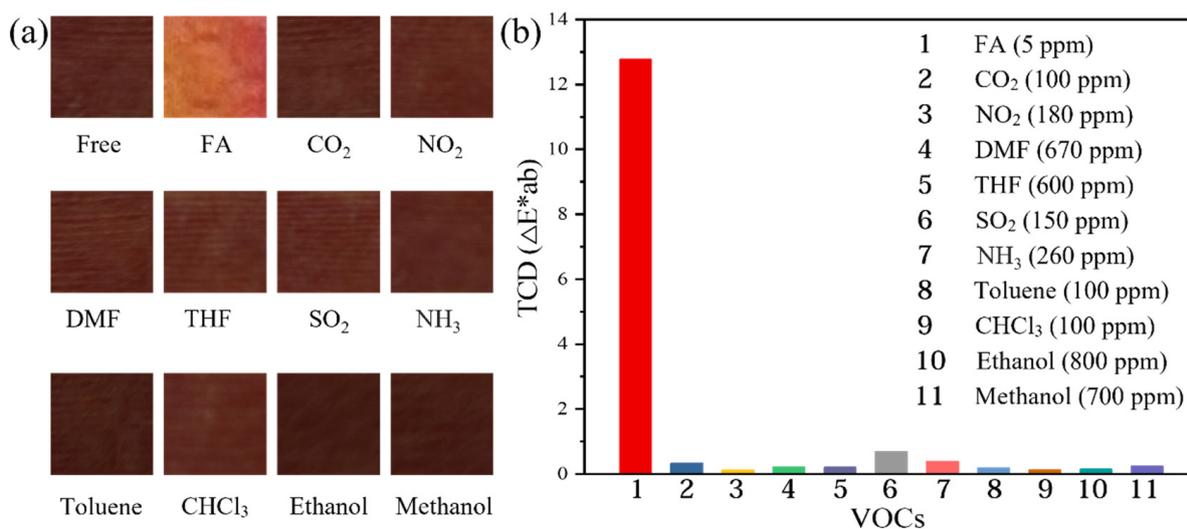
### 3.4. Assessment of the Selectivity of Colorimetric Sensing

To investigate the selectivity of the colorimetric sensor, the same testing was performed by exposing the sensor strip towards many other gases and volatile organic compounds (VOCs), such as  $\text{CO}_2$ ,  $\text{SO}_2$ ,  $\text{NO}_2$ ,  $\text{NH}_3$ , DMF, THF,  $\text{CHCl}_3$ , toluene, ethanol and methanol under the same experimental conditions. As shown in Figure 6a, despite the much higher vapor concentrations employed, none of the interferent gases or VOCs demonstrated a visible colorimetric response when tested under the same experimental conditions as done for the FA vapor, for which a distinct color change was otherwise observed, easily

visualized by naked eyes. Comparison of the  $\Delta E^*ab$  values thus measured for FA and all the interferents further proves the strong selectivity of the sensor towards FA, as shown in Figure 6b.



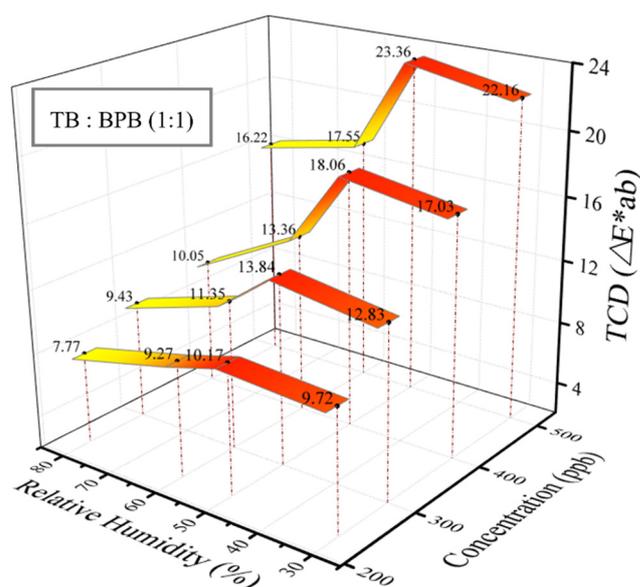
**Figure 5.** (a) The colorimetric response ( $\Delta E^*ab$ ) versus vapor concentrations of FA. The error bars represent the standard deviations of measurement. (b) Photos of the sensor strips after 10 min of exposure to different concentrations of FA vapor. The measurements were performed at 25 °C and 45% RH.



**Figure 6.** (a) Photos and (b) colorimetric response ( $\Delta E^*ab$  values) of the sensor strips after 10 min of exposure to FA vapor (5 ppm), in comparison with other gases and volatile organic compounds under much higher concentration (marked in the parenthesis). The measurements were performed at 25 °C and 45% RH.

### 3.5. Effect of Humidity on Colorimetric Sensing

Humidity effects are common for all the chemosensors considering the relatively high concentration of moisture under ambient conditions. To explore the humidity effect on the colorimetric sensing described above, the same sensor tests were repeated but under different relative humidity (RH) levels. As shown in Figure 7, within the range of 30–50% RH the color response of the sensor strips was almost unaffected, as tested for all four vapor concentrations of FA. However, when the humidity increases above 50% RH,  $\Delta E^*ab$  values start dropping sharply, and this trend seems to become more obvious when tested for higher concentrations of FA. The decreased sensor response under increased humidity is likely due to the competitive reaction (hydrolysis) between FA and water, which converts FA to methylene glycol, thus losing the reactivity with  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$ , and preventing the generation of free acids [32]. Regarding the humidity effect, the colorimetric sensor reported herein should be calibrated under different humidity levels before it can be deployed in practical applications. There remains a possibility that the color of pH indicator may be affected by the humidity change. To explore this potential effect, the same sensor strip was tested under the wide range of humidity (up to 95%). As shown in Figure S2 in the Supporting Information, the color value of the sensor strip remains little changed under the varying humidities, excluding any significant effect of humidity on the initial color of sensor strip.



**Figure 7.** Colorimetric response ( $\Delta E^*ab$ ) of the sensor strips upon exposure to four different vapor concentrations of FA (230, 300, 400, 500 ppb), each tested under varying humidity levels. All experiments were performed at 28 °C.

## 4. Conclusions

We have developed an efficient colorimetric sensing system for the detection of FA in the gas phase. The sensory material is primarily based on cheap paper towels, the cellulose fibril network of which provides a large, open interface for modification with  $(\text{NH}_2\text{OH})_2 \cdot \text{H}_2\text{SO}_4$  and pH indicators that combined enable selective sensor reaction with FA. This disposable paper-based sensor kit provides a simple and quick detection of FA with both high sensitivity and selectivity. A LOD of 10 ppb is projected for the sensor kit when using a color reader, which is much lower than the air quality threshold (80 ppb) set by the WHO. Actually, in the presence of 80 ppb of FA a distinct, color change obtained for the sensor strip can even be visualized by naked eyes. With appropriate calibration of the humidity effect, the reported sensor kit will find broad applications in air detection of FA as a close-to-zero cost approach in comparison to many of the current detection technologies.

**Supplementary Materials:** The following are available online at <https://www.mdpi.com/article/10.3390/chemosensors9120335/s1>, Figure S1: Comparison of colorimetric response of sensor strips fabricated from different paper towels, Table S1: Selection of hydroxylamine acid salts, Figure S2: The stability of the initial color of the sensor strip under different humidity.

**Author Contributions:** Conceptualization, C.L., S.C., L.Z., C.W.; investigation and original draft preparation, C.L.; revising and test assisting, M.Z., N.G., Q.T., J.S.; review, editing, and project administration, L.Z. and S.C. All authors have read and agreed to the published version of the manuscript.

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**Conflicts of Interest:** The authors declare no conflict of interest.

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