



Article

Ni_{0.5}Cu_{0.5}Co₂O₄ Nanocomposites, Morphology, Controlled Synthesis, and Catalytic Performance in the Hydrolysis of Ammonia Borane for Hydrogen Production

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Supplementary Materials:

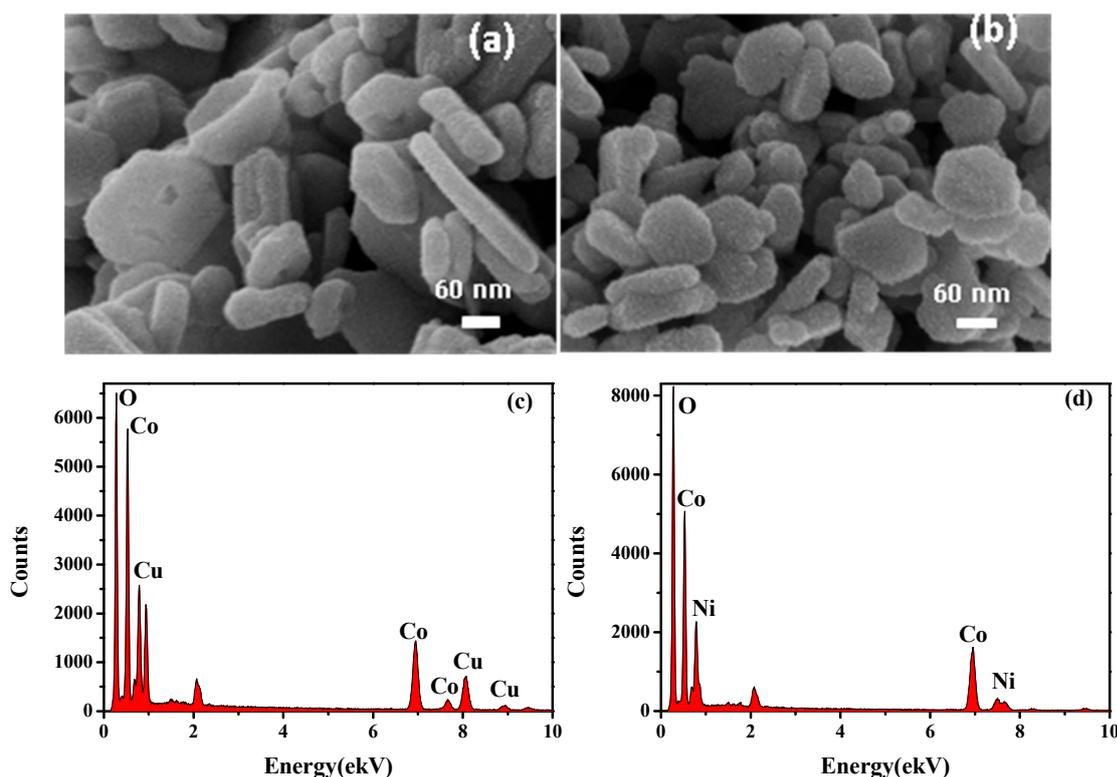


Figure S1 SEM images of the physical mixture of CuCo₂O₄ and NiCo₂O₄ (a,b) and EDS patterns on some selected nanoplatelets.

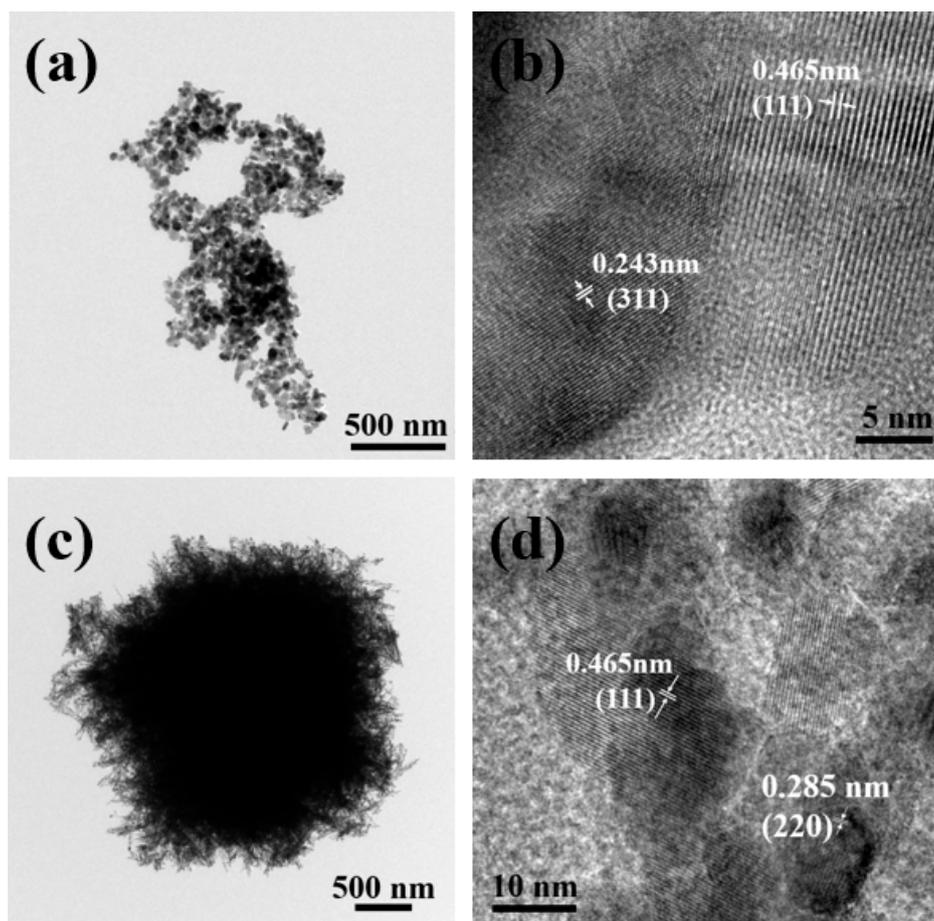


Figure S2. TEM image of nanoparticles (a), HRTEM image of nanoparticles (b), TEM image of urchin-like microspheres (c), HRTEM image of urchin-like microspheres (d).

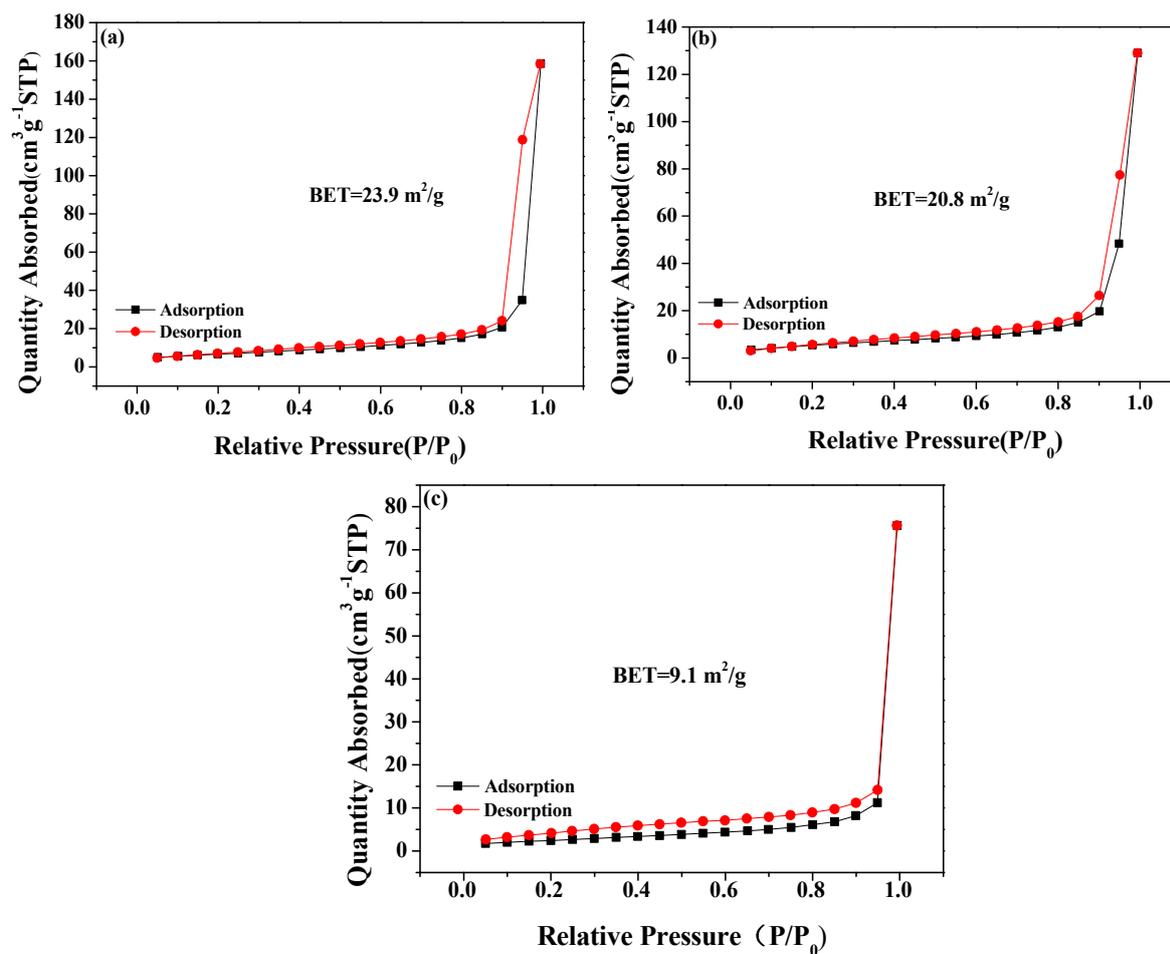


Figure S3. N₂ adsorption-desorption isotherms curves of nanoparticles (a), urchin-like microspheres (b) and nanoplatelets (c).

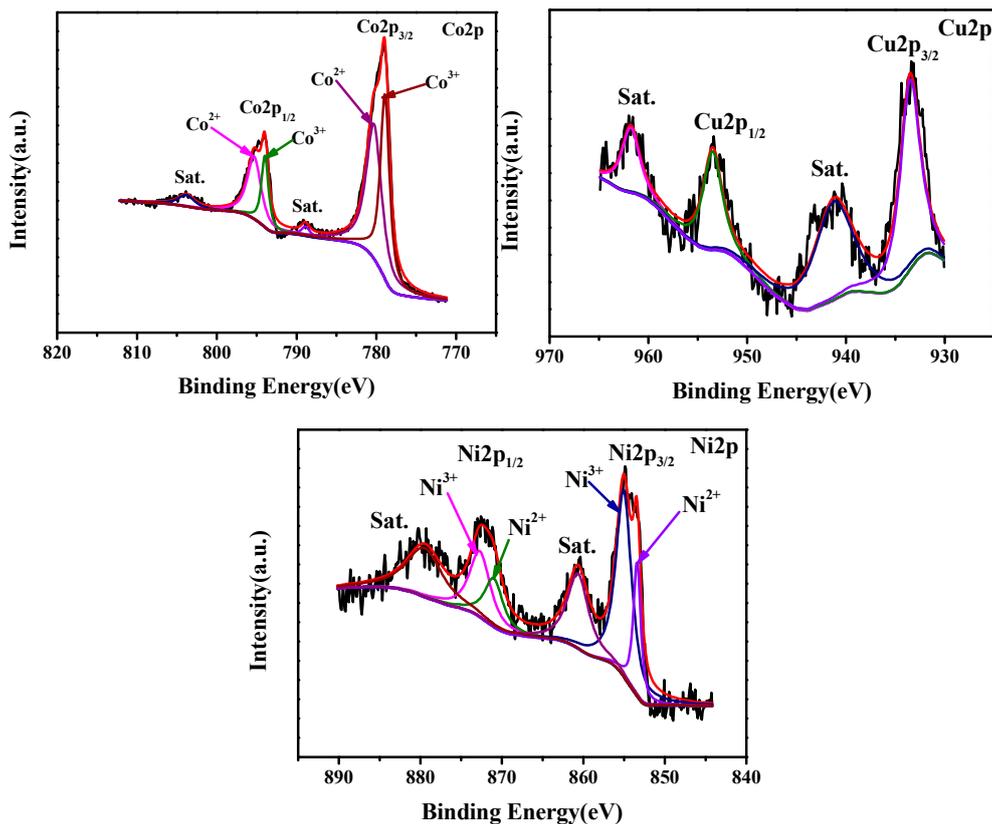


Figure S4 XPS spectra of the CuCo₂O₄ and NiCo₂O₄ mixture

Table S1 Comparison of the relative contents of Ni²⁺ and Co²⁺ on the surface of the composition and mixture

Catalyst	Ni ²⁺ /Ni(%)	Ni ³⁺ /Ni(%)	Co ²⁺ /Co(%)	Co ³⁺ /Co(%)
Composition	35.5	64.5	27.7	72.3
Mixture	23.8	76.2	42.1	57.9

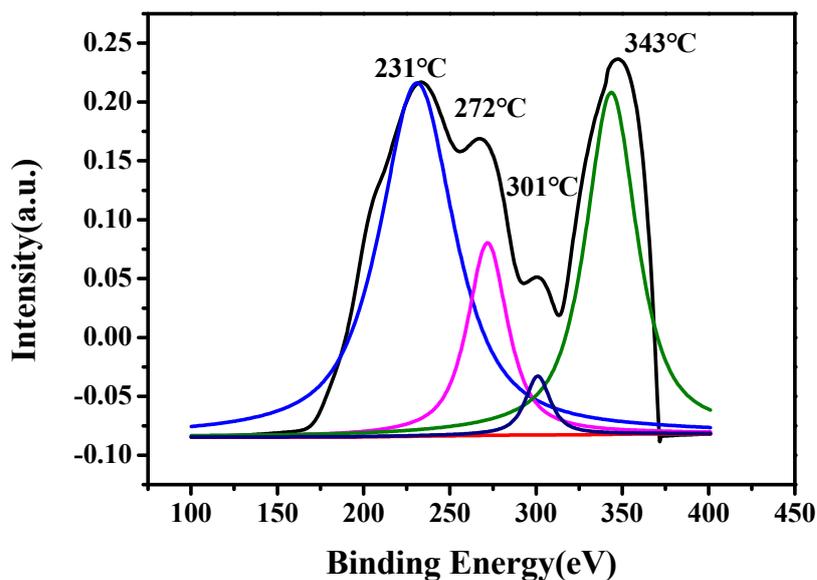


Figure S5 H₂-TPR curve of the mixture of CuCo₂O₄ and NiCo₂O₄.

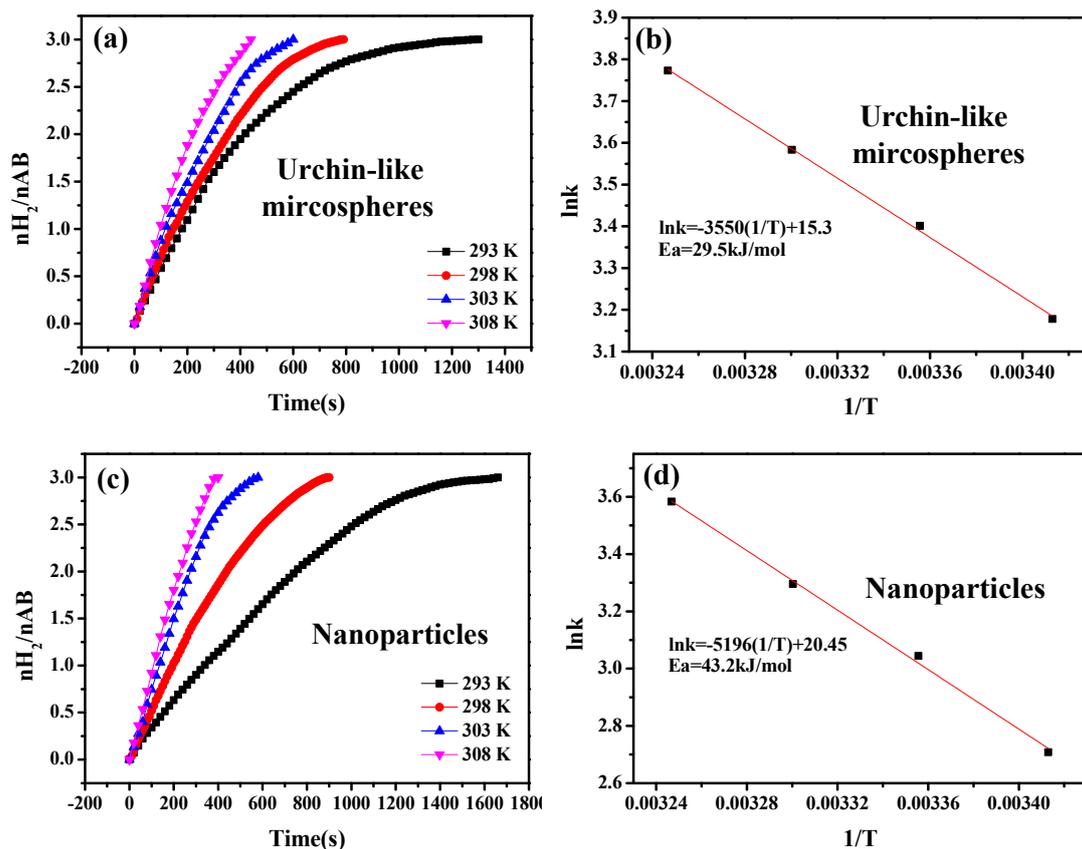


Figure S6. Hydrogen evolution at different temperature (a,c) and the calculation of the activation energy for different catalysts (b,d).

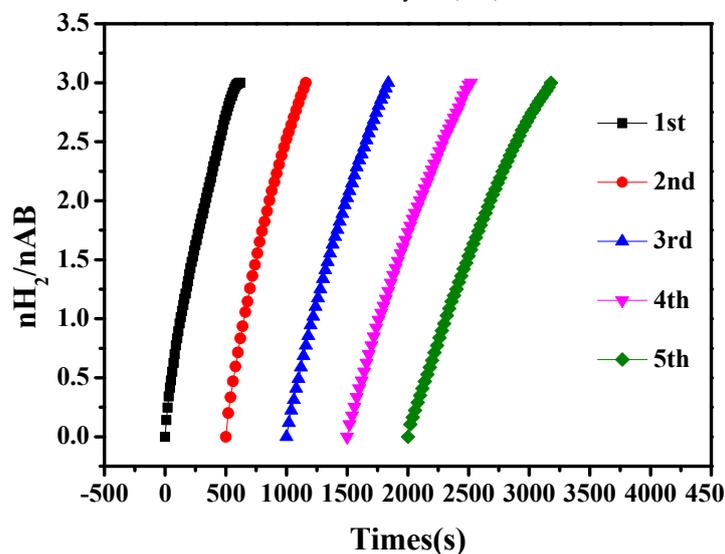


Figure S7. Hydrogen evolution at different recycle number when the CuCo₂O₄/NiCo₂O₄ nanoplatelets act as catalysts.

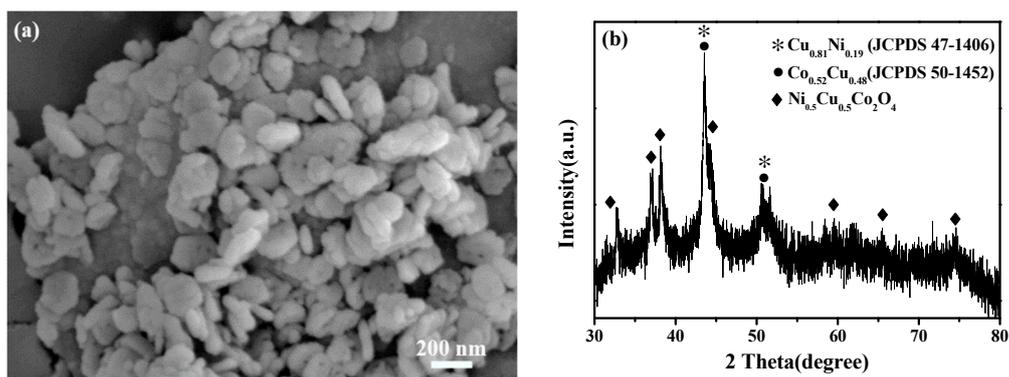


Figure S8. SEM image (a) and the XRD pattern (b) of the used $\text{Ni}_{0.5}\text{Cu}_{0.5}\text{Co}_2\text{O}_4$ nanoplatelets after catalytic reaction.

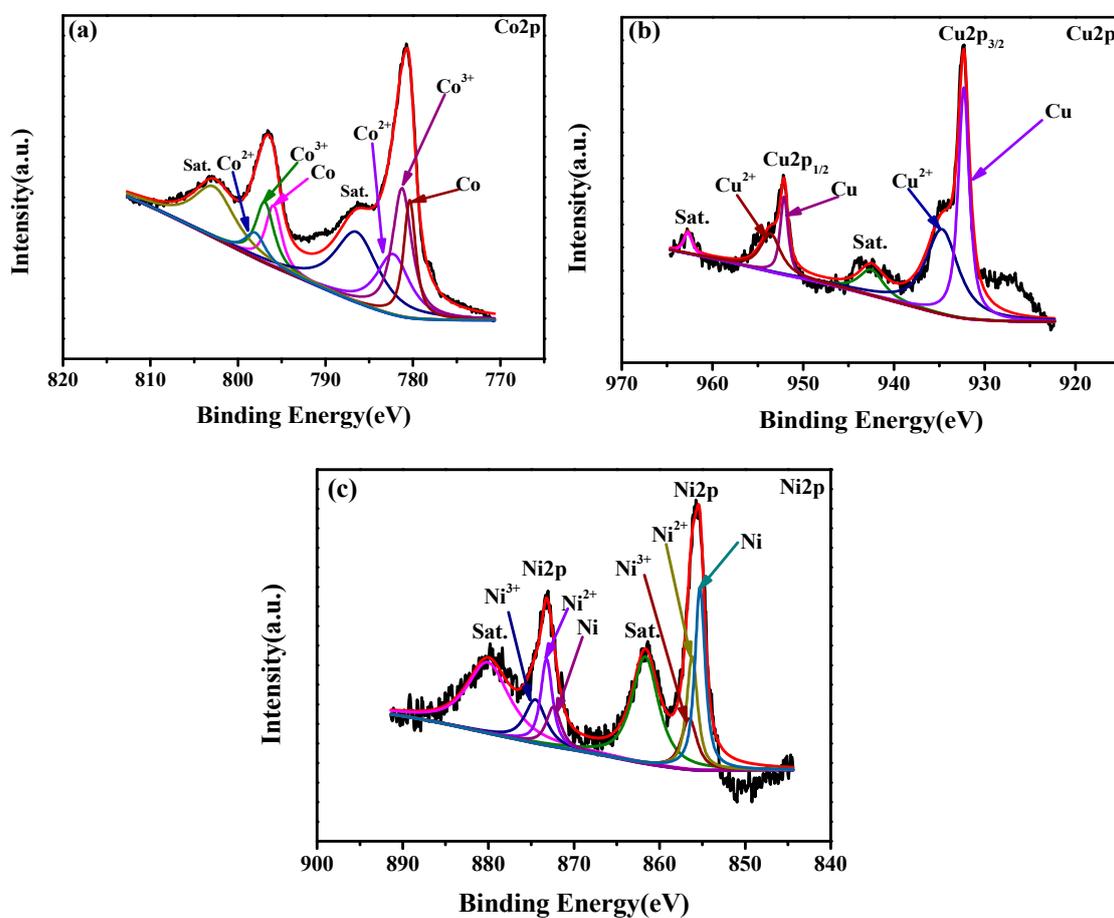


Figure S9 XPS spectra of $\text{CuCo}_2\text{O}_4/\text{NiCo}_2\text{O}_4$ nanoplatelets after catalytic reaction.

