



Communication

Unveiling Oxygen K-Edge and Cobalt L-Edge Electron Energy Loss Spectra of Cobalt Hydroxide and Their Evolution under Electron Beam Irradiation

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Abstract: Cobalt hydroxides, $\text{Co}(\text{OH})_2$, have attracted considerable attention due to their diverse applications in the fields of energy and the environment. However, probing the electronic structure of $\text{Co}(\text{OH})_2$ is challenging, mainly due to its sensitivity to electron beam irradiation. In this study, we report the unperturbed O K-edge and Co L-edge for $\text{Co}(\text{OH})_2$ by electron beam damage and investigate the electronic structure transformation of $\text{Co}(\text{OH})_2$ under electron beam irradiation, using low current electron energy loss spectroscopy. In particular, the O K-edge pre-peak at 530 eV, which is not found in the undamaged $\text{Co}(\text{OH})_2$, begins to appear with an increasing electron beam current. In addition, the Co L-edge peak shifts to a higher energy, close to Co_3O_4 , indicating that the localized phase transition within $\text{Co}(\text{OH})_2$ leads to the formation of Co_3O_4 .

Keywords: transmission electron microscopy; electron energy loss spectroscopy; electron beam damage; cobalt hydroxide



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1. Introduction

Transition metal hydroxides have attracted interest for various applications, such as batteries, water splitting, and nanomedicine [1–3]. Among transition metal hydroxides, cobalt hydroxide, $\text{Co}(\text{OH})_2$, has received considerable attention as an electrochemical catalyst for oxygen evolution reaction (OER), because of its electrocatalytic activity and efficient ion intercalation [4,5]. $\text{Co}(\text{OH})_2$ exists in two phases depending on its structure: α - $\text{Co}(\text{OH})_2$ and β - $\text{Co}(\text{OH})_2$ phase. The β - $\text{Co}(\text{OH})_2$ phase has a brucite-like arrangement, with hydroxyl ions packed hexagonally and Co(II) ions occupying the octahedral site. The α - $\text{Co}(\text{OH})_2$ phase, on the other hand, resembles a hydrotalcite-like structure, where positively charged $\text{Co}(\text{OH})_{2-x}$ layers are present. These layers are accompanied by charge-stabilizing anions, such as NO_3^{3-} , CO_3^{2-} , Cl^- , etc., in the interlayer space [6]. The α - $\text{Co}(\text{OH})_2$ phase has better electrochemical activity than β - $\text{Co}(\text{OH})_2$ due to the expanded interlayer spacing and intercalated anions [4,5].

Scanning transmission electron microscopy (STEM)–electron energy loss spectroscopy (EELS) measurements have become increasingly popular in atomic structure and chemical analysis as an effective method for characterizing materials in various fields, including for catalysis, batteries, and biotechnology [7–9]. However, one critical drawback of this STEM–EELS technique is electron beam-induced damage to the sample [10–13]. Electron beam-induced damages manifest through a range of effects, including atomic displacement, electron-beam sputtering, radiolysis, heating, and electrostatic charging [14]. These effects hinder the microscope’s ability to fully observe the original structural properties of materials.

$\text{Co}(\text{OH})_2$ is a material that is highly sensitive to electron beam irradiation. Numerous investigations using transmission electron microscopy have been conducted to explore

the structural characteristics of $\text{Co}(\text{OH})_2$; however, the effect of electron beam irradiation has not been seriously considered. Only recently has a study demonstrated the impact of electron beam irradiation in inducing structural modifications from $\text{Co}(\text{OH})_2$ to Co_3O_4 [15]. However, there is still a lack of EELS results for $\text{Co}(\text{OH})_2$, and the reported EELS results vary [16,17], which may be attributed to experimental challenges due to the electron-sensitivity of $\text{Co}(\text{OH})_2$. It is worth noting that electron beam damage could be more severe in EELS experiments than in imaging because spectroscopy typically requires a longer dwell time than imaging. Therefore, great caution must be taken with electron beam-sensitive materials for EELS experiments.

In this study, we present the undamaged spectra of the O K-edge and Co L-edge for $\text{Co}(\text{OH})_2$, obtained through low-current EELS. In addition, we demonstrate the evolution of the electronic structure of $\text{Co}(\text{OH})_2$ under electron beam irradiation in TEM. The low-current EELS experiment (screen current below 30 pA) showed no pre-peak at 530 eV in O K-edge, an observation not reported prior. Previously reported O K-edge spectra for $\text{Co}(\text{OH})_2$ displayed a pre-peak at 530 eV, which may have been caused by electron beam damage. The electron beam currents exceeding 30 pA induced changes in the O K-edge and the Co L-edge in $\text{Co}(\text{OH})_2$. The Co L-edge underwent an energy shift to the Co_3O_4 position, which was accompanied by the emergence of the O K pre-edge at 530 eV, indicating a phase transition from $\text{Co}(\text{OH})_2$ to Co_3O_4 under electron beam irradiation.

2. Experimental Methods

2.1. Materials Synthesis

The α - $\text{Co}(\text{OH})_2$ and β - $\text{Co}(\text{OH})_2$ particles were prepared under ambient conditions following a previous study [18]. To synthesize α - $\text{Co}(\text{OH})_2$, a mixture of cobalt chloride hexahydrate (10 mM, $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$, Sigma-Aldrich, St. Louis, MO, USA), sodium chloride (50 mM, NaCl, DUKSAN, Ansan, Korea), and hexamethylenetetramine (60 mM, $\text{C}_6\text{H}_{12}\text{N}_4$, 99.0%, Sigma-Aldrich, St. Louis, MO, USA) was dissolved in 150 mL of a 1:9 blend of ethanol and deionized water. This solution was then heated in an oil bath at 90 °C for one hour, resulting in a suspension containing green particles. The suspended product was subsequently centrifuged at 3500 rpm for 15 min, followed by multiple washes using deionized water and anhydrous ethanol. The final green particles were dried completely at room temperature. The synthesis procedure for β - $\text{Co}(\text{OH})_2$ mirrored that of α - $\text{Co}(\text{OH})_2$, with the exception that sodium chloride (NaCl) was omitted from the initial solution. This modification resulted in the formation of pink powder of β - $\text{Co}(\text{OH})_2$. The Co_3O_4 was synthesized through the annealing of α - $\text{Co}(\text{OH})_2$ at 500 °C for a duration of 3 h under atmospheric conditions.

2.2. Electron Microscopy

Scanning transmission electron microscopy (STEM) experiments were conducted using a Thermo Fisher Scientific Spectra 300 operating at an accelerated voltage of 300 kV. EELS experiments were carried out using a Gatan Continuum equipped with a K3 direct electron detector. The electron dose rate for EELS was $4.77 \times 10^7 \text{ e}^- / \text{\AA}^2 \cdot \text{s}$. The energy spread of the zero-loss peak (ZLP) was measured at $\sim 1 \text{ eV}$ full-width at half-maximum (FWHM). The dispersion was 0.18 eV/ch, and a convergence angle of 22.7 mrad and a collection angle of 64.2 mrad were used. Automatic ZLP alignment was enabled to correct the energy shift during the acquisition time in dual EELS mode. To minimize electron beam damage, EELS experiments were performed in the low-current condition; the screen current was in the range of 1–100 pA and the pixel time for acquisition was 0.1 s. The EELS area scans were performed at 50×50 pixels with a step size of 5 nm. The final spectrum was obtained by averaging 2500 spectra. For the sample, the cobalt hydroxides and cobalt oxide powders were effectively dispersed in ethanol and deposited on a lacey carbon grid.

3. Results and Discussion

Figure 1 illustrates the scanning transmission electron microscopy (STEM) images along with the selected area electron diffraction (SAED) patterns for α -Co(OH)₂, β -Co(OH)₂, and spinel Co₃O₄, respectively. In Figure 1a,b, the STEM images show hexagonal microplates of α -Co(OH)₂ and β -Co(OH)₂, possessing an average dimension of approximately 6 μ m. However, the spinel Co₃O₄ hexagonal microplate, synthesized through the annealing of Co(OH)₂, exhibits granular characteristics accompanied by multiple cracks, as evident in Figure 1c. The SAED patterns reveal that both α -Co(OH)₂ and β -Co(OH)₂ consist of a single crystalline phase, while Co₃O₄ displays a polycrystalline phase. The combined results from STEM and SAED confirm the successful high-quality synthesis of α -Co(OH)₂, β -Co(OH)₂, and spinel Co₃O₄, without the presence of any other phases.

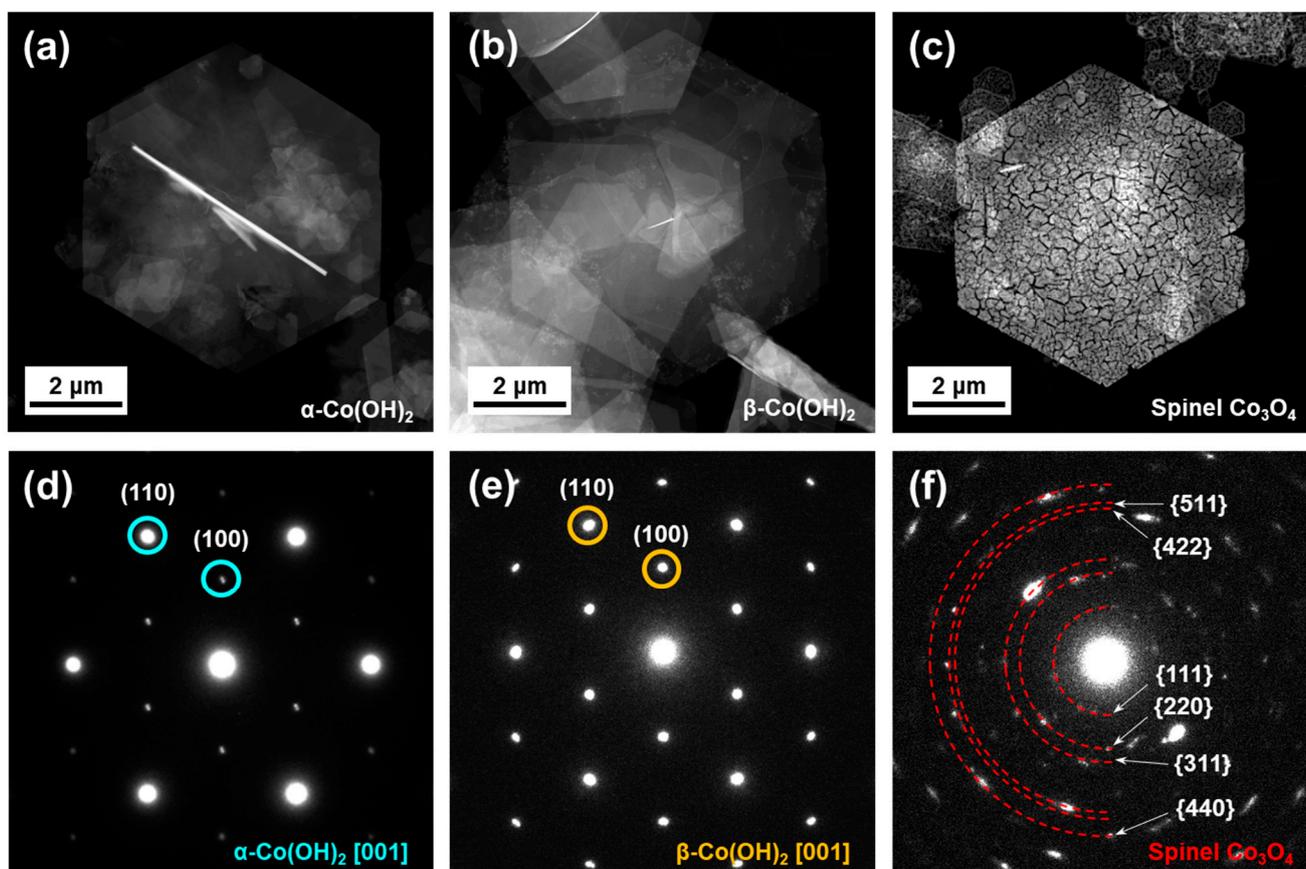


Figure 1. STEM images and typical selected area electron diffraction patterns of representative (a,d) α -Co(OH)₂, (b,e) β -Co(OH)₂, and (c,f) spinel Co₃O₄ microplates, respectively.

Figure 2 shows the O K-edge and Co L-edge spectra of α -Co(OH)₂, β -Co(OH)₂, and Co₃O₄, respectively. The spectra of α -Co(OH)₂ and β -Co(OH)₂ show distinct features at the O K- and Co L-edges, in contrast to the spectrum of Co₃O₄. However, the difference between the α - and β -Co(OH)₂ spectra was minimal. For Co₃O₄, the pre-peak of the O K-edge is located at \sim 530 eV, while the pre-peak of the O K-edge for Co(OH)₂ cannot be found, but only a small bump appears at \sim 534 eV. In the transition metal oxide, the O K-edge pre-peak contributes to the hybridization of O-2p with transition metal 3d states [19]. Previously, several EELS results have been reported for Co(OH)₂ in which the O K-edge possesses a pre-peak close to 530 eV, as in the case of Co₃O₄, indicating the fact that the phase change occurred due to electron beam irradiation [15]. The main peak for the O K-edge is located at \sim 542 eV in Co₃O₄, while the main peak is shifted, towards lower energy, to 539 eV in Co(OH)₂ marked by the arrows in Figure 2a. The Co L-edge consists of L₃ and L₂ white lines, which originate from the transitions of 2p_{3/2} to 3d_{3/2}3d_{5/2} and

$2p_{1/2}$ to $3d_{3/2}$, respectively. These L_3 and L_2 white lines are affected by the valence states of the transition metal, due to their relationship to unoccupied states in the transition metal 3d band. Similar to the O K-edge, α -Co(OH)₂ and β -Co(OH)₂ show almost identical Co L-edges, but clearly different peak positions and shapes from Co₃O₄, as shown in Figure 2b,c. The peak of the Co L_3 -edge of Co(OH)₂ is at ~779 eV (peak a), and the shoulder peak b is found at ~780 eV. Note that the shoulder peak b in Co(OH)₂ is almost coincident with the Co L_3 -edge peak for Co₃O₄. The white line ratios (L_3/L_2) are calculated to be 4.5 for Co(OH)₂ and 2.6 for Co₃O₄, in agreement with previously reported values [20].

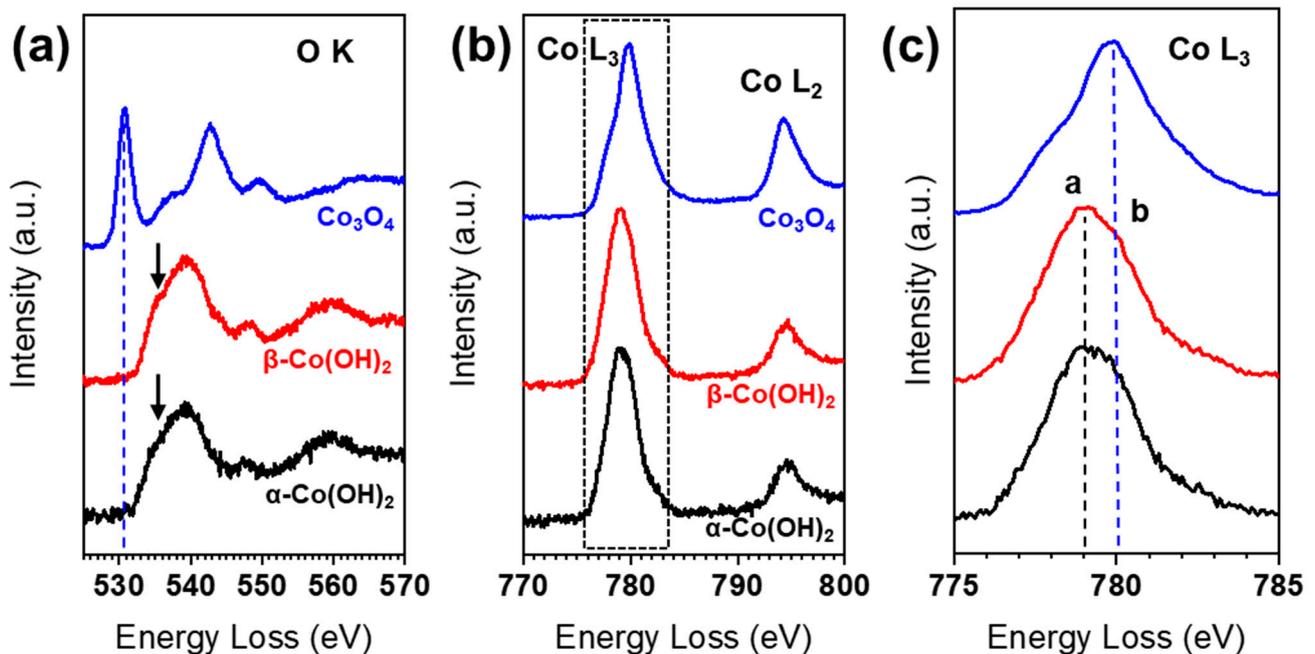


Figure 2. (a) EELS O K-edge and (b) Co $L_{3,2}$ -edge of α -Co(OH)₂, β -Co(OH)₂, and Co₃O₄, respectively. (c) Enlarged spectra for the Co L_3 -edge from the dotted box in (b). Each spectrum is shifted vertically for better comparison.

To investigate changes in electronic structure upon electron beam irradiation, we systematically measured the O K- and Co L-edges by varying the beam current. Figure 3 shows the evolution of the O K- and Co L-edges upon electron beam irradiation. At the electron beam current of 1 pA, the pre-peak of the O K-edge is not observed, and this absence persists up to 30 pA. However, at a current exceeding 35 pA, the pre-peak begins to appear at 530 eV. Furthermore, the main peak of the O K-edge at ~538 eV begins to show some spectral weight at 543 eV, which is consistent with the main peak of Co₃O₄. The Co L-edge shows a similar trend. The position of the L_3 peak remains the same up to 30 pA. At currents above 35 pA, the L_3 peak shifts towards higher energy, indicating a partial conversion of Co(OH)₂ to Co₃O₄ under high-current electron beam irradiation. When the L_3 peak shifts, the L_3/L_2 ratio also changes. The L_3/L_2 ratio value decreases from 4.6 at 1 pA to 3.3 at 100 pA, indicating that Co could be oxidized during the EELS experiment.

Finally, we examined the changes in the O K-edge during the phase transition in real time through multiple acquisitions. Figure 4 displays the alterations of the pre-peak in the O K-edge over 20 acquisitions. The scans were conducted within a fixed area of 50×50 pixels with a step size of 5 nm. The screen current was 20 pA and each spectrum was scanned for a total of 8 min. In the initial EELS acquisitions, Co(OH)₂ did not exhibit a pre-peak in the O K-edge, as previously described. However, after the third acquisition, the O K-edge pre-peak, indicative of spinel Co₃O₄, began to appear at 530 eV, and its intensity increased with the number of acquisitions. This change suggests that the low-current electron beam also causes phase transition under persistent irradiation.

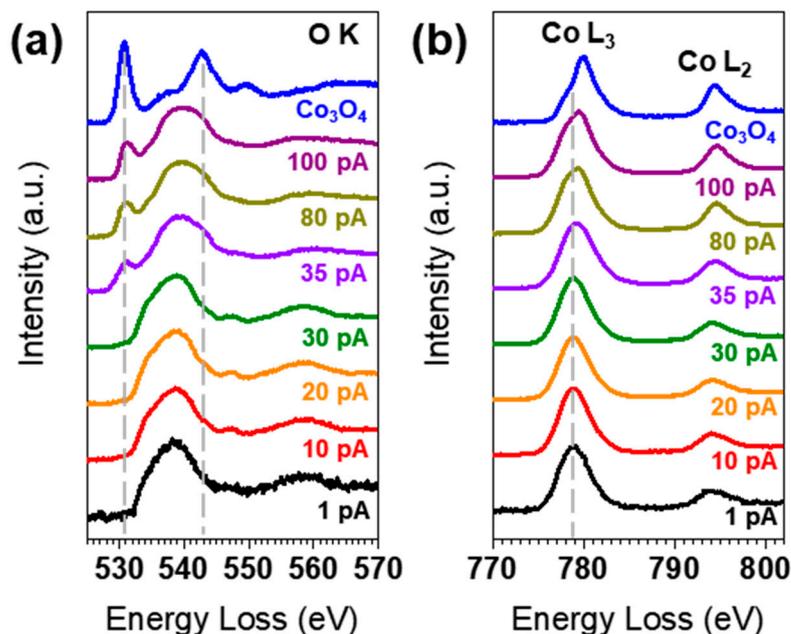


Figure 3. Changes in (a) the O K-edge and (b) the Co L-edge with increasing electron beam current.

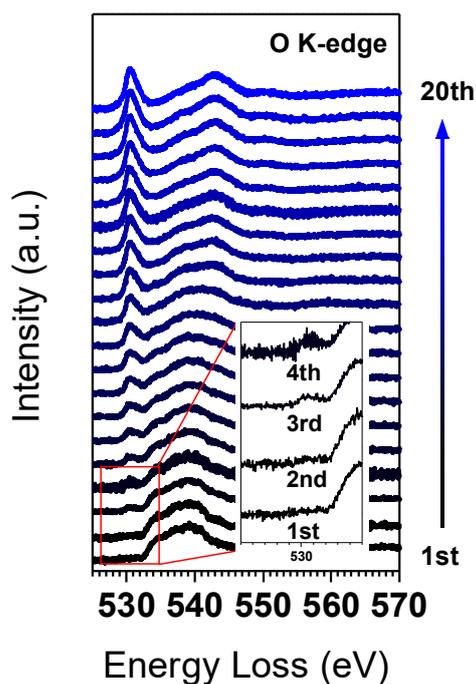


Figure 4. Changes in O K-edge of Co(OH)_2 over a number of multiple EELS acquisitions.

The phase transition of Co(OH)_2 to Co_3O_4 is observed in the thermal decomposition process, following two steps [21]: (1) $\text{Co(OH)}_2 + 0.25\text{O}_2 \rightarrow 2\text{CoOOH} + 0.5\text{H}_2\text{O}$; (2) $3\text{CoOOH} \rightarrow \text{Co}_3\text{O}_4 + 1.5\text{H}_2\text{O} + 0.25\text{O}_2$. However, the phase transition induced by electron beam irradiation could follow a different reaction path, due to the different environment in the TEM chamber. The electron beam could easily kick off the hydrogen atoms in Co(OH)_2 , resulting in the formation of CoOOH phases, which had been found in an in situ electron diffraction experiment under electron beam irradiation [15]. However, in this study, the existence of CoOOH was not clearly shown in the O K-edge and Co L-edge EELS results, which could be attributed to the relatively fast phase transition from Co(OH)_2 to CoOOH , compared to the subsequent phase transition to Co_3O_4 . It is worth noting that the multiple

EELS acquisitions in Figure 4 show continuous change of the O K-edge and the Co L-edge over the irradiation time, and no discernible change that could indicate a CoOOH phase is found during these multiple acquisitions. The CoOOH to Co₃O₄ phase transition is anticipated to follow a reaction similar to the thermal decomposition process, and the water and oxygen molecules could be extracted in the gas phases in TEM.

4. Conclusions

In this study, we reported the original electronic structure of Co(OH)₂, and systematically investigated the changes in the electronic structure of Co(OH)₂ under an electron beam by changing the electron beam current for the EELS experiment. Employing low-current EELS conditions (with an electron current below 30 pA) revealed the unperturbed O-K edge and Co-L edge spectra for Co(OH)₂, thereby contrasting with previous EELS data. However, the electron beam currents exceeding 30 pA induced alterations in the O-K edge and Co-L edge for both α -Co(OH)₂ and β -Co(OH)₂. The O-K pre-edge at 530 eV, indicative of Co₃O₄, began to appear with an increasing electron beam current, accompanying the energy shift of the Co-L edge to the position for Co₃O₄. These observed EELS spectral changes indicate the facile transformation of Co(OH)₂ into Co₃O₄ under electron beam irradiation. Our results emphasize the need for a cautious interpretation of results from EELS experiments and highlight the potential impact of electron beam influence on the electronic structure of materials.

Author Contributions: J.-H.K. conceived the project and designed the experiments; TEM and EELS experiments have been carried out by J.H.S.; data analysis has been done by J.-H.K. and J.H.S.; the manuscript has been written by J.-H.K. and J.H.S. All authors have read and agreed to the published version of the manuscript.

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Data Availability Statement: The datasets used and/or analyzed during the current study are available from the corresponding author on reasonable request.

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Conflicts of Interest: The authors declare no conflict of interest.

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