

Article

Thermal Properties of Methyl Ester-Containing Poly(2-oxazoline)s

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Received: 13 August 2015 ; Accepted: 30 September 2015 ; Published: 13 October 2015

Academic Editor: Graeme Moad

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Abstract: This paper describes the synthesis and thermal properties in solution and bulk of poly(2-alkyl-oxazoline)s (PAOx) containing a methyl ester side chain. Homopolymers of 2-methoxycarbonylethyl-2-oxazoline (MestOx) and 2-methoxycarbonylpropyl-2-oxazoline (C3MestOx), as well as copolymers with 2-ethyl-2-oxazoline (EtOx) and 2-*n*-propyl-2-oxazoline (nPropOx), with systematic variations in composition were prepared. The investigation of the solution properties of these polymers revealed that the cloud point temperatures (T_{CPS}) could be tuned in between 24 °C and 108 °C by variation of the PAOx composition. To the best of our knowledge, the T_{CPS} of PMestOx and PC3MestOx are reported for the first time and they closely resemble the T_{CPS} of PEtOx and PnPropOx, respectively, indicating similar hydrophilicity of the methyl ester and alkyl side chains. Furthermore, the thermal transitions and thermal stability of these polymers were investigated by DSC and TGA measurements, respectively, revealing amorphous polymers with glass transition temperatures between −1 °C and 54 °C that are thermally stable up to >300 °C.

Keywords: poly(2-oxazoline)s; copolymers; LCST; cloud point temperature; thermal properties; glass transition temperature

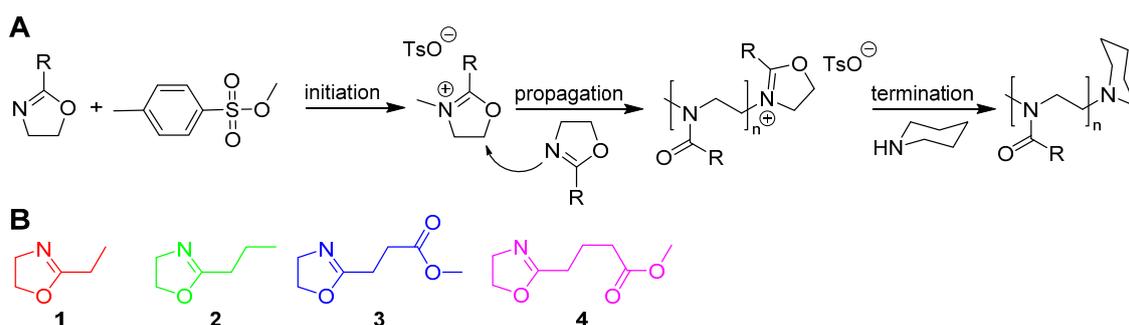
1. Introduction

Smart polymeric materials can change their physical properties in response to external stimuli, such as temperature, ionic strength, pH, chemical and biological stimuli, light or electromagnetic radiation [1–3]. Thermoresponsive polymeric behavior has most extensively been studied and, in particular, polymers that undergo a temperature induced phase transition. One of the best-characterized thermoresponsive polymers is poly(*N*-isopropyl acrylamide) (PNIPAAm), which displays a lower critical solution temperature (LCST) of 32 °C, just below body temperature [2–5]. Other classes of thermoresponsive polymers that have emerged in recent years are for example poly(oligo ethylene glycol acrylate)s [2,6], polyisocyanopeptides grafted with oligo(ethylene glycol) side chains [7–9], poly(2-oxazine)s [10] and poly(2-oxazoline)s [3,5,11–17]. These types of polymers can be applied in temperature sensors, [18] protein chromatography [19] and various biomedical applications such as drug delivery and tissue engineering [1,20]. For these latter applications thermoresponsive polymers are needed that are biocompatible and that can be conveniently functionalized.

A polymer class that fulfills these requirements is the class of poly(2-alkyl/aryl-2-oxazoline)s (PAOx) as they are biocompatible, thermoresponsive, and have tunable properties [15,21,22].

Poly(2-methyl-2-oxazoline) (PMeOx) and poly(2-ethyl-2-oxazoline) (PEtOx) are interesting for biomedical applications because they show stealth behaviour similar to poly(ethylene glycol) (PEG) [15,21–25]. Furthermore, PEtOx and poly(2-propyl-2-oxazoline)s (PPropOx) show thermoresponsive properties. PEtOx has an LCST of ~ 60 °C, poly(2-*n*-propyl-2-oxazoline) (P*n*PropOx) of ~ 25 °C, poly(2-cyclopropyl-2-oxazoline) (PcPropOx) of ~ 30 °C and poly(2-isopropyl-2-oxazoline) (PiPropOx) of ~ 38 °C [3,5,15,26]. Cloud point temperatures (T_{CPS}) can be tuned from 9 °C to 100 °C by varying the molecular weight and composition of the polymer by incorporation of more hydrophobic or hydrophilic monomeric units, either by copolymerization or post-polymerization modification [5,11–13,27,28].

Functional groups, often protected, can be introduced into PAOx by making use of a functional monomer or initiator during the living cationic ring opening polymerisation (CROP, Scheme 1) of 2-oxazoline monomers or by using a functional terminating agent, yielding well-defined PAOx with control over number and type of functionalities [15,29–38]. Methyl esters are especially interesting, because they can undergo a direct amidation with a variety of amines to easily introduce other functional groups such as alcohols, hydrazide and amines [39–42]. Moreover, the ester can be hydrolyzed to the corresponding carboxylic acid providing a versatile handle for conjugation as well as a means to introduce pH responsiveness. Although the potential of side-chain methyl ester-containing PAOx copolymers is well established, there is remarkably little known about their thermal features and thermoresponsive behavior [33,43–47].



Scheme 1. (A) Cationic ring-opening polymerisation of 2-oxazolines with methyl *p*-toluenesulfonate as initiator and piperidine as terminating agent and (B) Monomer structures of EtOx (1), *n*PropOx (2), MestOx (3) and C3MestOx (4).

In this paper we describe the synthesis and thermal properties of methyl ester-containing PAOx. Homopolymers and statistical copolymers, with near ideal random monomer distributions [38], of 2-ethyl-2-oxazoline (EtOx, 1) and 2-*n*-propyl-2-oxazoline (*n*PropOx, 2) with 2-methoxycarbonylethyl-2-oxazoline (MestOx, 3) and 2-methoxycarbonylpropyl-2-oxazoline (C3MestOx, 4) were prepared with 10, 20, 30, 50, and 70 mol % methyl ester content. The thermal solution and bulk properties of these polymers were determined by turbidimetry, differential scanning calorimetry (DSC), and thermal gravimetric analysis (TGA), respectively.

2. Experimental Section

2.1. Materials

2-Chloroethylamine hydrochloride, methyl *p*-toluenesulfonate (MeOTs), and sodium carbonate were purchased from Acros Organics. EtOx was kindly donated by Polymer Chemistry Innovations. All other reagents were purchased from Sigma Aldrich and used as received. EtOx and methyl *p*-toluenesulfonate (MeOTs) were purified by distillation over barium oxide and stored under argon. Dry solvents were obtained from a solvent purification system from J.C. Meyer, with aluminum oxide

drying system and a nitrogen flow. MestOx [37], C3MestOx [38], and *n*PropOx [48] were prepared according to literature procedures.

2.2. Instrumentation

Nuclear magnetic resonance (NMR) spectra were recorded on a Bruker DMC300 (300 MHz for ^1H , 75 MHz for ^{13}C) (Bruker: Billerica, MA, USA).

Polymerization reaction mixtures were prepared in a VIGOR Sci-Lab SG 1200/750 Glovebox system (Vigor: Houston, TX, USA), with purity levels of less than 1 ppm for O_2 and H_2O .

Polymerizations were carried out in a Biotage Initiator Microwave System with Robot Sixty utilizing capped reaction vials. These vials were heated to 120 °C overnight, allowed to cool to room temperature and filled with nitrogen prior to use. All microwave polymerizations were performed with temperature control (IR sensor).

Size exclusion chromatography (SEC) was performed on an Agilent 1260 (Agilent Technologies: Santa Clara, CA, USA): series HPLC system equipped with a 1260 online degasser, a 1260 ISO-pump, a 1260 automatic liquid sampler, a temperature controlled column compartment, a 1260 diode array detector (DAD) and a 1260 refractive index detector (RID). Analyses were performed on a PSS Gram30 column (PSS Polymer Standards Service GmbH: Mainz, Germany) in series with a PSS Gram1000 column at 50 °C. *N,N*-dimethylacetamide (DMA, Sigma-Aldrich: Zwijndrecht, The Netherlands), containing 50 mM of LiCl, was used as an eluent, at a flow rate of 0.593 mL·min⁻¹. The SEC traces were analyzed using the Agilent Chemstation software with the GPC add on. Number average molecular weights (M_n) and dispersity (D) values were calculated against poly(methyl methacrylate) PMMA standards.

Turbidity measurements were performed on an Avantium Crystal 16 platform (Avantium: Geleen, The Netherlands). Solutions of the polymers were prepared in milliQ (WaterPro PS polisher from Labconco: Kansas City, MO, USA) water with a conductivity of 18.2 MΩ/cm at 5 mg·mL⁻¹ and were stirred at room temperature until all polymer was dissolved. Three heating cycles were applied with a heating/cooling rate of 1 °C·min⁻¹, temperature ranges of 5–45, 20–60, 20–75, 60–105 or 70–110 °C were used depending on the polymer, and withhold steps of 5 min at the extreme low temperatures, no hold time was applied when the samples were heated above 100 °C; all measurements were performed with a stirring rate of 700 rpm. The cloud point temperatures are given as the 50% transmittance point during heating. GC vials with open caps and septa were used for these measurements allowing heating until 110 °C with minor pressure build up as noticed by slight expansion of the septum.

Differential scanning calorimetry (DSC) traces were recorded under nitrogen with a Mettler-Toledo DSC1 module. Glass transition temperatures were obtained from the second heating run with heating/cooling rates of 10 °C·min⁻¹. Indium was used as a standard for temperature and enthalpy calibrations.

Thermogravimetric analysis (TGA) was performed with a Mettler Toledo TGA/SDTA851e instrument (Mettler Toledo: Columbus, OH, USA) under nitrogen atmosphere at a heating rate of 10 °C·min⁻¹ between 25 °C and 800 °C.

2.3. Polymer Synthesis

Polymerization mixtures with a total monomer concentration of 4 M in acetonitrile, an overall [M]/[I] ratio of 100 and MeOTs as initiator, were prepared in the glovebox under argon and the reactor vials were crimped air-tight inside the glovebox. Polymerizations were carried out under microwave irradiation at 140 °C for 20 min aiming for full conversion [37,38]. After cooling to room temperature, the polymerizations were quenched by the addition of piperidine. After removal of the solvents under reduced pressure, the polymers were re-dissolved in dichloromethane and precipitated in cold diethyl ether twice. The polymers were dried in the vacuum oven at 50 °C before further analysis.

3. Results and Discussion

Series of homopolymers and copolymers of EtOx and nPropOx with MestOx and C3MestOx were prepared [37,38] with varying feed ratios of 10, 20, 30, 50, and 70 mol % MestOx or C3MestOx and a constant monomer to initiator ratio of $[M]/[I] = 100$, yielding four sets of copolymers. All polymers were terminated with piperidine instead of the widely used methanolic sodium hydroxide to avoid saponification of the methyl ester. All polymers were characterized by ^1H NMR spectroscopy and SEC and were obtained with the desired composition and a narrow dispersity ($\bar{D} < 1.21$, Table 1).

Cloud point temperatures (T_{CPs}) of the homo- and copolymers were determined by turbidimetry of $5 \text{ mg} \cdot \text{mL}^{-1}$ aqueous solutions, under stirring as this gives more reliable dissolution data when cooling down the samples. All polymers, except EtOx₄₉–MestOx₅₁, showed reversible LCST behavior covering a broad temperature range (24–108 °C, Table 1) depending on the polymer composition with nearly no hysteresis between heating and cooling cycles (See Supplementary Information Figures S1–S21). In some cases the T_{CP} in the cooling cycles was slightly higher than in the heating cycles; this can be possibly explained by the fact that re-dissolution of these polymers was a rather slow process, leading to temporary lower concentrations in solution during cooling and thus an increase of T_{CP} . However, the precipitates were fully dissolved when the solution cooled down completely.

Interestingly, the homopolymers of MestOx (PMestOx) and C3MestOx (PC3MestOx) also showed thermo-responsive properties. PMestOx displayed a shallow transition with a T_{CP} of 102 °C and 80% transmission at 92 °C similar to PEtOx with a degree of polymerization of 100 (Figure 1A), whereas PC3MestOx exhibits a sharp transition at 26 °C (Figure 1B) comparable to PnPropOx. When the polymers were heated above 100 °C during the turbidimetry measurements, the T_{CP} slightly increased during every run (Figure 1 and Figures S2–S11 in Supplementary Information), which is ascribed to partial hydrolysis of the methyl ester groups resulting in the corresponding carboxylic acid, as was confirmed by ^1H NMR spectroscopy (Supplementary Information Figures S22, S23). Partial precipitation or crystallization of the sample could be excluded after visual inspection of the samples. The determined average T_{CP} of all heating and cooling cycles are reported in Table 1, except for the polymers that were heated above 100 °C for which the T_{CP} in the first heating and cooling runs are listed.

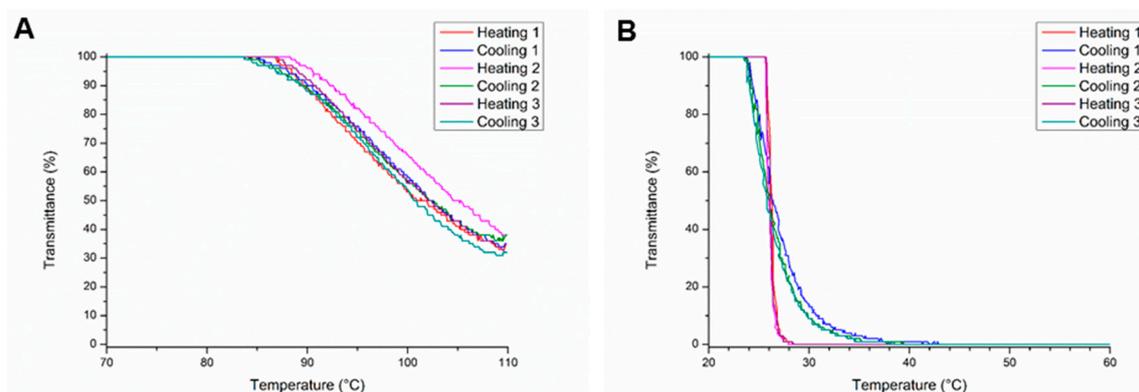


Figure 1. Turbidimetry measurements for (A) PMestOx and (B) PC3MestOx.

The T_{CPs} of the different copolymers (Table 1) are plotted as a function of the copolymer composition in Figure 2. The relationship between the T_{CPs} and EtOx content of the EtOx–C3MestOx copolymers is linear (Figure 2A), and, thus with this monomer combination the T_{CP} can straightforwardly be tuned between the T_{CPs} of the corresponding homopolymers by varying the polymer composition. A linear relationship between copolymer composition and T_{CP} was previously also described for P(iPropOx–nPropOx), P(iPropOx–EtOx) and P(cPropOx–EtOx) copolymers [13,14,27].

Table 1. Overview of synthesis and characterisation data of studied (co)polymers.

Monomer 1	Monomer 2	Feed ratio (M1:M2)	Composition ^a (M1:M2)	SEC ^b		T_{CP} (°C) ^c		T_g ^d (°C)	TGA (°C) ^e	
				M_n (10 ³ g/mol)	\bar{D}	Heating	Cooling		5%	50%
MestOx	-	100:0	100:0	19.4	1.10	101	102	39	258	353
C3MestOx	-	100:0	100:0	24.1	1.12	26	26	-1	325	360
EtOx *	-	100:0	100:0	14.0*	1.19 *	91 *	91 *	54 *	-	-
<i>n</i> PropOx	-	100:0	100:0	16.6	1.10	27	23	35	312	406
EtOx	MestOx	90:10	91:9	18.9	1.14	93	96	53	322	403
EtOx	MestOx	80:20	82:18	16.9	1.14	99	102	45	293	389
EtOx	MestOx	70:30	70:30	17.4	1.13	103	106	44	310	376
EtOx	MestOx	50:50	49:51	19.3	1.14	#	#	39	306	365
EtOx	MestOx	30:70	29:71	22.8	1.16	102	104	38	296	359
<i>n</i> PropOx	MestOx	90:10	90:10	21.0	1.12	25	24	30	327	393
<i>n</i> PropOx	MestOx	80:20	79:21	15.5	1.12	27	30	30	306	385
<i>n</i> PropOx	MestOx	70:30	68:32	16.3	1.13	32	36	30	303	377
<i>n</i> PropOx	MestOx	50:50	48:52	20.6	1.15	40	40	32	310	367
<i>n</i> PropOx	MestOx	30:70	27:73	29.3	1.19	55	54	35.0	309	367
EtOx	C3MestOx	90:10	90:10	21.1	1.13	89	89	43	334	399
EtOx	C3MestOx	80:20	82:18	21.8	1.14	78	79	39	313	382
EtOx	C3MestOx	70:30	70:30	21.6	1.14	71	78	27	316	377
EtOx	C3MestOx	50:50	50:50	22.7	1.14	56	56	17	322	369
EtOx	C3MestOx	30:70	30:70	24.1	1.14	42	43	-2	320	363
<i>n</i> PropOx	C3MestOx	90:10	90:10	19.2	1.21	24	25	28	321	399
<i>n</i> PropOx	C3MestOx	80:20	79:21	20.7	1.15	25	32	18	328	385
<i>n</i> PropOx	C3MestOx	70:30	68:32	21.0	1.15	25	31	15	332	379
<i>n</i> PropOx	C3MestOx	50:50	48:52	16.7	1.21	26	27	12	311	368
<i>n</i> PropOx	C3MestOx	30:70	28:72	18.8	1.16	25	30	5	304	361

* Literature values [27]; # no cloud point temperature observed up to 110 °C; ^a determined by ¹H NMR spectroscopy; ^b determined by SEC against PMMA standards; ^c determined by turbidimetry measurements; ^d determined by DSC measurements; 5%: loss of 5 wt % after loss of solvents; for some polymers it was difficult to determine the point where all solvents were removed and for these the start of the flatter area was used for calculation of the 5% weight loss temperature. 50%: loss of 50 wt % from the total mass.

All T_{CP} s of the n PropOx–C3MestOx copolymers were in between 24 and 26.5 °C (Figure 2B), revealing that varying the polymer composition hardly changed the thermo-responsive behavior. This behavior was expected as the hydrophobicity of both monomers and the T_{CP} s of the homopolymers are comparable.

In the n PropOx–MestOx copolymer series, the T_{CP} s exponentially increased with increasing MestOx content (Figure 2C), also the 80% transmission values are plotted in this graph, showing comparable T_{CP} values for the copolymers and a clear difference for PMestOx. A similar trend was also reported for EtOx– n PropOx copolymers [14,28]. The hydrophobic n PropOx monomer seems to be more dominant in determining the overall polymer hydrophilicity and T_{CP} than the hydrophilic MestOx monomer.

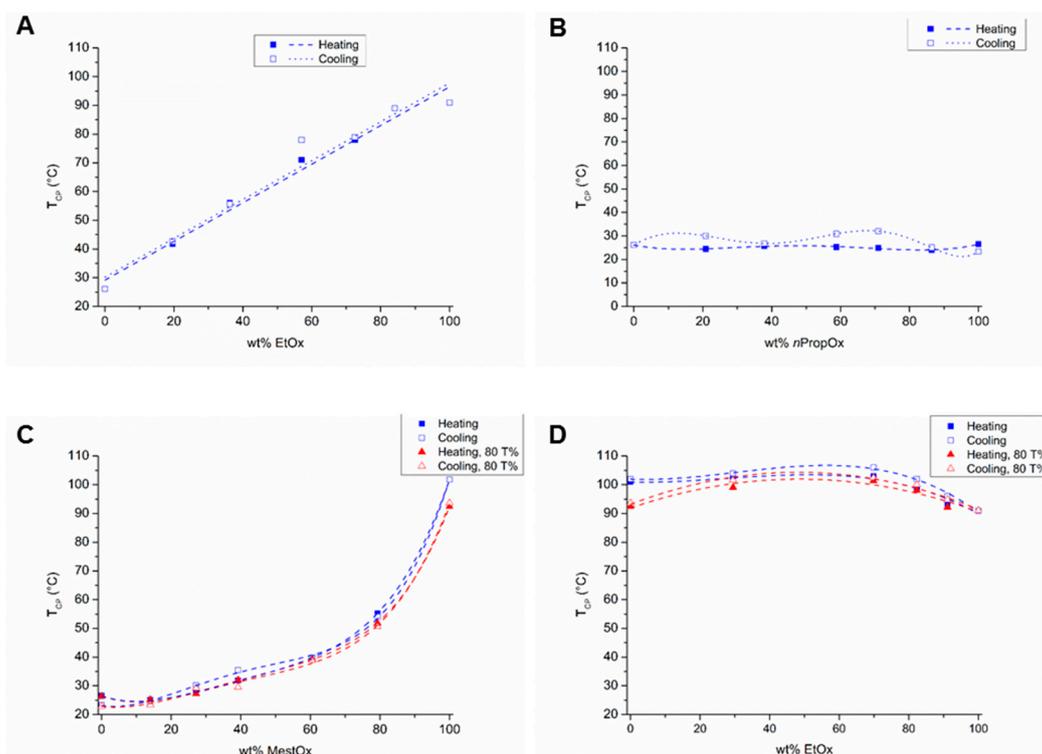


Figure 2. Cloud point temperatures (T_{CP} ; determined at 50% transmittance unless otherwise stated) as a function of polymer composition of (A) P(EtOx–C3MestOx), (B) P(n PropOx–C3MestOx), (C) P(n PropOx–MestOx, and (D) P(EtOx–MestOx); lines are added to guide the eye.

Copolymers of EtOx and MestOx displayed T_{CP} s in between 93 °C and 103 °C at EtOx–MestOx comonomer ratios of 90:10, 80:20, 70:30, and 30:70 (Figure 2D, Table 1), whereas surprisingly the copolymer with a 50:50 comonomer ratio did not show a transition up to 110 °C. Additionally, here the 80% transmission values are plotted in this graph, showing comparable T_{CP} values for the copolymers and a clear difference for PMestOx. This observation indicates more efficient polymer hydration and/or decreased polymer–polymer interactions, which both may be related to near random placement of the smaller ethyl and larger methoxycarbonyl ethyl side chains.

The thermal bulk properties of the homo- and copolymers were further investigated by differential scanning calorimetry (DSC) and thermogravimetric analysis (TGA). All polymers showed a clear glass transition (Figure 3, ESI Figure S24–S27, Table 1), indicating that they are amorphous. The T_g s of the homopolymers PMestOx ($T_g = 39$ °C) and PnPropOx ($T_g = 35$ °C) are comparable, while PC3MestOx has a T_g of -1 °C and PEtOx has a T_g of 54 °C. The rather low T_g of PC3MestOx indicates high chain mobility induced by the side chain.

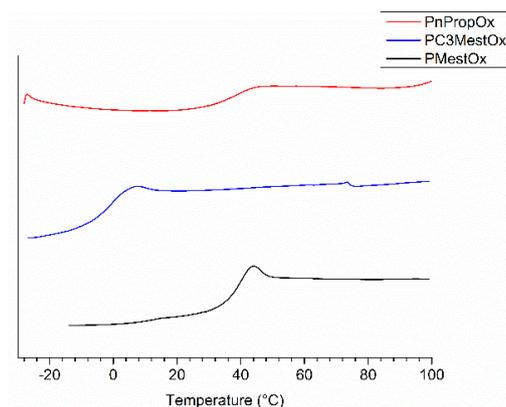


Figure 3. DSC traces of second heating run ($10 \text{ K} \cdot \text{min}^{-1}$) of PMestOx, PC3MestOx, and PnPropOx homopolymers.

The T_g s of the copolymers are plotted in Figure 4. In the P(EtOx–MestOx) copolymer series, the T_g s did not change much up to 40 wt % EtOx (Figure 4A) and from this point on it increased linearly towards the T_g of PEtOx. The T_g of the MestOx homopolymer is slightly higher than the T_g of P(EtOx–MestOx, 30:70), indicating that the polymer chain mobility is lower, which may be ascribed to more efficient (dipole-dipole) interactions between the polyamide backbones in the homopolymers. A very similar trend in T_g was observed for P(EtOx–C3MestOx) copolymers (Figure 4B), with the exception that after already 20 wt % incorporation of EtOx the glass transition temperature started to increase linearly with increasing amount of EtOx, possibly due to the larger difference in T_g of the homopolymers.

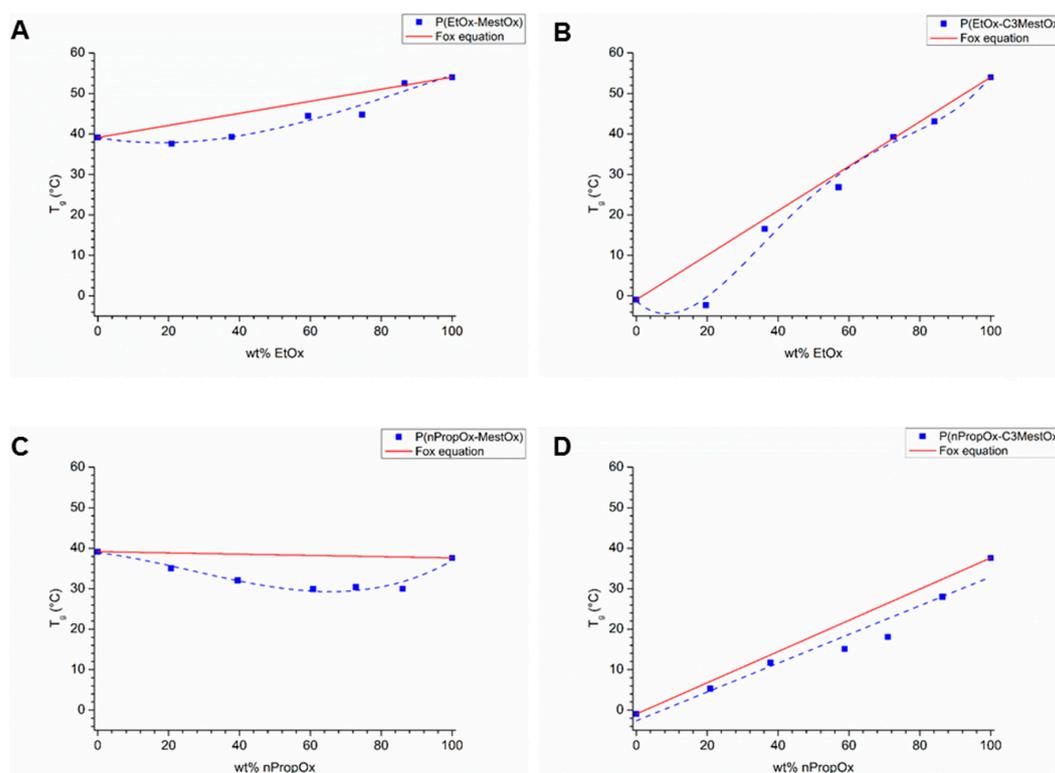


Figure 4. Glass transition temperatures as function of polymer composition of (A) P(EtOx–MestOx), (B) P(EtOx–C3MestOx), (C) P(nPropOx–MestOx), and (D) P(nPropOx–C3MestOx). Lines are added to guide the eye; the red line indicates the Fox equation.

For the P(*n*PropOx–MestOx) copolymers all measured T_g s were lower than those of the homopolymers (Figure 4C) and the T_g s only vary 5 °C within the composition range. A similar trend was recently reported for copolymers of 2-cyclopropyl-2-oxazoline (cPropOx) and EtOx [27], and this can be ascribed to suppression of interchain interactions in the copolymers. The T_g of *n*PropOx–C3MestOx copolymers increased linearly with increasing *n*PropOx content (Figure 4D), so with this combination of monomers the glass transition can be accurately tuned between the T_g s of the corresponding homopolymers by varying the monomer ratios.

TGA analysis revealed that the PMestOx homopolymer was thermally stable up to 250 °C (Figure 5A), while all other homo- and copolymers were stable at least up to 300 °C (Figure 5, Table 1, and Figures S28–S30 in Supplementary Information). The lower stability of PMestOx is rather surprising, especially compared to PC3MestOx and may indicate that the close proximity of the secondary amide to the ester group decreases the stability of the latter. In Table 1 the temperature for 5% and 50% weight loss are listed, whereby the 5 wt % loss was determined after the loss of solvents. However, not all polymers show a plateau after loss of solvent and for these the start of the flatter area was used for calculation of the 5% weight loss temperature. The 50% loss is reported as the 50% loss of the total weight. In general, when more (C3)MestOx was incorporated into the copolymers, the thermal stability slightly decreased, because the P(C3)MestOx homopolymers are less stable than PEtOx or P*n*PropOx homopolymers, most likely due to decomposition of the methyl ester units.

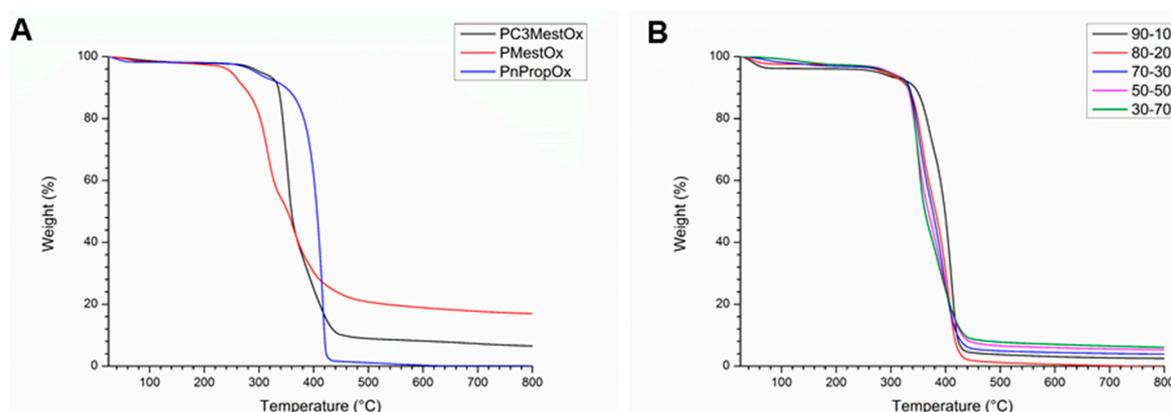


Figure 5. Thermogravimetric analysis of (A) PMestOx, PC3MestOx, and P*n*PropOx homopolymers and (B) P(EtOx–C3MestOx) copolymers.

4. Conclusions

Homo- and copolymers of MestOx and C3MestOx with EtOx and *n*PropOx were successfully synthesized. All polymers, except P(EtOx₄₉–MestOx₅₁), showed T_{CP} s that varied in between 24 °C and 108 °C depending on the composition, at a concentration of 5 mg/mL in water. Interestingly, PC3MestOx has a low T_{CP} of 26 °C making it interesting for biomedical applications. DSC measurements revealed that all polymers are amorphous with T_g s between –1 °C and 54 °C. For P(EtOx–MestOx), P(EtOx–C3MestOx), and P(*n*PropOx–C3MestOx) copolymers the glass transition temperature can be tuned in between the T_g s of the homopolymers. However, for P(*n*PropOx–MestOx) copolymers, the T_g s of all copolymers are lower than the ones of both homopolymers, indicating suppression of interchain interactions. All homo- and copolymers are thermally stable up to at least 250 °C, whereby copolymers with more (C3)MestOx were found to be less stable. Nonetheless, the stabilities will allow thermal processing, e.g. hot-melt extrusion, at temperatures well above room temperature.

Supplementary Information: The supplementary information is available online at: www.mdpi.com/2073-4360/7/10/1494/s1.

Acknowledgments: The authors acknowledge the Polymer Innovation Program (PIPM10006, General adhesive Tissue Tape) and the Special Research Fund of Ghent University; in particular PB for a joint PhD fellowship.

Author Contributions: The work was designed by Petra J. M. Bouten, Jan C. M. van Hest and Richard Hoogenboom. Petra J. M. Bouten performed the kinetic experiments and turbidimetry studies. Kathleen Lava performed the thermal analysis. The article was written by Petra J. M. Bouten; Kathleen Lava, Jan C. M. van Hest and Richard Hoogenboom corrected this paper.

Conflicts of Interest: The authors declare no conflict of interest.

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