

Article

Modification of Structure and Magnetic Properties in Coordination Assemblies Based on $[\text{Cu}(\text{cyclam})]^{2+}$ and $[\text{W}(\text{CN})_8]^{3-}$

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Abstract: The 1D $\{[\text{Cu}^{\text{II}}(\text{cyclam})]_3[\text{W}^{\text{V}}(\text{CN})_8]_2 \cdot 5\text{H}_2\text{O}\}_n$ ($1 \cdot 5\text{H}_2\text{O}$) (cyclam = 1,4,8,11-tetraazacyclotetradecane) coordination polymer of ladder topology can be obtained in water-alcohol solution from $[\text{Cu}(\text{cyclam})]^{2+}$ and $[\text{W}(\text{CN})_8]^{3-}$ building blocks. Upon dehydration, $1 \cdot 5\text{H}_2\text{O}$ undergoes a single-crystal-to-single-crystal structural transformation to the anhydrous $\{[\text{Cu}^{\text{II}}(\text{cyclam})]_3[\text{W}^{\text{V}}(\text{CN})_8]_2\}_n$ (**1**) form, which retains the same topology, but is characterized by shorter Cu-W distances and significantly more bent CN-bridges. The deformation of the coordination skeleton is reflected in magnetic properties: the predominant intra-chain interactions change from ferromagnetic in $1 \cdot 5\text{H}_2\text{O}$ to antiferromagnetic in **1**. The reaction between the same building blocks in water solution under slow diffusion conditions leads to the formation of a 0D $\{[\text{Cu}^{\text{II}}(\text{cyclam})(\text{H}_2\text{O})]_2[\text{Cu}^{\text{II}}(\text{cyclam})][\text{W}^{\text{V}}(\text{CN})_8]_2\} \cdot 3\text{H}_2\text{O}$ pentanuclear assembly ($2 \cdot 3\text{H}_2\text{O}$).

Keywords: molecular magnetism; octacyanotungstate(V); copper(II); cyclam; cyano bridge; magnetic properties

1. Introduction

One of the interesting features that can be achieved in molecular magnetism is the possibility to modify magnetic properties by sorption of small guest molecules. At the turn of the century magnetic sponges were defined as compounds that change structure and magnetic properties upon reversible removal of coordinated or non-coordinated water molecules [1–3]. Later, the term solvatomagnetism was introduced to describe changes in magnetic behavior upon sorption or desorption of different guest molecules.

We have previously shown that the solvatomagnetic effect can be repeatedly observed in CN-bridged networks containing the $[\text{Ni}(\text{cyclam})]^{2+/3+}$ cationic building block. By combining it with different polycyanometallates, we obtained two main families of coordination polymers of 1D and 2D connectivity. The 2D $\{[\text{Ni}^{\text{II}}(\text{cyclam})]_3[\text{M}(\text{CN})_n]_2\}_\infty$ ($\text{M} = \text{Cr}^{\text{III}}, \text{Fe}^{\text{III}}, n = 6$; $\text{M} = \text{W}^{\text{V}}, n = 8$) networks [4–7] show honeycomb-like topology with microporous channels running across the coordination layers. For these networks, we observed formation of different hydrates as well as the possibility to introduce small organic molecules. For each of the three compounds, four pseudopolymorphic forms differing in structure and magnetic properties have been characterized. The second family of solvatomagnetic polymers are non-porous 1D bimetallic alternating chains. The $\{[\text{Ni}^{\text{III}}(\text{cyclam})][\text{M}^{\text{III}}(\text{CN})_6]_2 \cdot 6\text{H}_2\text{O}\}_\infty$ ($\text{M}^{\text{III}} = \text{Cr}, \text{Fe}$) [8] chains can be reversibly dehydrated, which affects their magnetic properties. The $\{(\text{H}_3\text{O})[\text{Ni}^{\text{III}}(\text{cyclam})][\text{M}^{\text{II}}(\text{CN})_6]_2 \cdot 5\text{H}_2\text{O}\}_\infty$ ($\text{M}^{\text{II}} = \text{Fe}$, [9] Ru, Os [10]) chains, where the charge of the coordination skeleton is compensated by the presence of the H_3O^+ ions, are rare examples of sorption-driven charge-transfer. The removal of water causes partial electron transfer from M^{II} centers to Ni^{III} ions. The process is reversible for Fe^{II} and irreversible for Ru^{II} and Os^{II} —based compounds.

Cyclam complexes with different metal centers, including Ni^{II} [11,12], Cu^{II} [13–17] and Mn^{III} [18–20], were often used to build CN-bridged coordination polymers. However, apart from our studies, there are no reports on the desolvation effects in this group of compounds. Thus, we decided to examine the potential of cyclam complexes with other metal ions for the construction of solvatomagnetic materials. The manganese(III)-based alternating bimetallic chains $\{[\text{Ni}^{\text{III}}(\text{cyclam})][\text{M}^{\text{III}}(\text{CN})_6]_2 \cdot 6\text{H}_2\text{O}\}_\infty$ ($\text{M}^{\text{III}} = \text{Cr}, \text{Fe}$) exhibit loss of crystallization water accompanied by structural and magnetic changes [21]; however, in contrast to their Ni^{III}-based congeners [8], the dehydration process was irreversible. The Cu^{II} cationic building blocks often differ from those of other 3d metals, due to the strong Jahn-Teller effect. The $\{[\text{Cu}^{\text{II}}(\text{cyclam})]_3[\text{W}^{\text{V}}(\text{CN})_8]_2 \cdot 5\text{H}_2\text{O}\}_n$ compound [22] shows 1D ladder-like topology, unlike the series of Ni^{II}-based 2D honeycomb-like networks of the same stoichiometry described above. It was characterized in terms of structure and magnetic properties, but dehydration effect was not mentioned in the report [22]. Here, we present the post-synthetic modification of structure and magnetic properties of this ladder chain network and the synthesis of a new 0D assembly based on the same building blocks.

2. Materials and Methods

The Na₃[W(CN)₈] \cdot 4H₂O precursor complex was prepared according to the published procedure [23]. Other reagents and solvents were commercially available and used as supplied. Measurements of powder X-ray diffraction were carried out on a PANanalytical X'Pert Pro powder diffractometer with the Cu K α radiation source. Elemental analysis for carbon, nitrogen and hydrogen were carried out on a Vario Micro Cube elemental analyser. IR spectra of single crystals were recorded using a Nicolet iN10 MX FTIR microscope operating in the transmission mode. The dehydration process was characterized by a dynamic vapor sorption method using an SMS DVS Resolution apparatus in the 40–0% RH range at 298 K. Magnetic measurements were performed on a Quantum Design MPMS3 magnetometer. Variable-temperature magnetic susceptibility was measured at an applied field of 1000 Oe in the temperature range of 300 K to 1.8 K and magnetization at 1.8 K up to 70 kOe. The samples were sealed in polyethylene bags to avoid dehydration at low pressure.

2.1. Synthesis of the [Cu(cyclam)](NO₃)₂ \cdot 6H₂O Precursor Complex

The compound was prepared according to the modified published procedure [24]. Cu(NO₃)₂ \cdot 3H₂O (1 mmol, 245.4 mg) was dissolved in 4 ml of absolute ethanol at 60 °C. To the above solution, solid cyclam (1 mmol, 200.0 mg) was added and immediate precipitation of violet product was observed. The suspension was left in the freezer for 1h and then filtered. Yield: 396.0 mg (79%).

2.2. Synthesis of $\{[\text{Cu}(\text{cyclam})]_3[\text{W}(\text{CN})_8]_2 \cdot 5\text{H}_2\text{O}\}_n$ (**1** \cdot 5H₂O)

The compound was prepared according to the modified published procedure [22]. A solution of [Cu(cyclam)](NO₃)₂ \cdot 6H₂O (0.13 mmol, 66.4 mg) in a water-ethanol mixture (1:2, 20 mL) was layered over a solution of Na₃[W(CN)₈] \cdot 4H₂O (0.07 mmol, 35.6 mg) in a water-ethanol mixture (2:1, 20 mL) in test tubes. Brownish-red needle-shaped crystals were formed after three days. (Found: C, 33.53; H, 4.976; N, 23.44; calc. C₄₆H₈₂Cu₃N₂₈O₅W₂: C, 33.17; H, 4.96; N, 23.55) IR ν CN: 2167 cm⁻¹, 2151 cm⁻¹, 2145 cm⁻¹, 2139 cm⁻¹.

2.3. Synthesis of $\{[\text{Cu}(\text{cyclam})]_3[\text{W}(\text{CN})_8]_2\}_n$ (**1**)

The anhydrous compound was obtained by drying **1** \cdot 5H₂O at 50 °C for 1 h. IR ν CN: 2166 cm⁻¹, 2161 cm⁻¹, 2151 cm⁻¹, 2144 cm⁻¹, 2139 cm⁻¹, 2134 cm⁻¹

2.4. Synthesis of $\{[\text{Cu}^{\text{II}}(\text{cyclam})(\text{H}_2\text{O})]_2[\text{Cu}^{\text{II}}(\text{cyclam})][\text{W}^{\text{V}}(\text{CN})_8]_2\} \cdot 3\text{H}_2\text{O}$ (**2** \cdot 3H₂O)

The solutions of [Cu(cyclam)](NO₃)₂ \cdot 6H₂O (0.05 mmol, 26.8 mg in 8 mL of H₂O) and Na₃[W(CN)₈] \cdot 4H₂O (0.03 mmol, 14.5 mg in 8 mL of H₂O) were placed in separate sections of an

H-tube and water was added to fill the horizontal connecting tube. Brownish-red plate-shaped crystals were formed after 2–3 weeks. (Found: C, 33.42; H, 4.76; N, 23.53; cal. $C_{46}H_{82}Cu_3N_{28}O_5W_2$: C, 33.17; H, 4.96; N, 23.55). IR ν_{CN} : 2166 cm^{-1} , 2151 cm^{-1} , 2144 cm^{-1} , 2139 cm^{-1} .

2.5. Structure Determination

Single crystal X-ray diffraction measurements for **1** and **2·3H₂O** were performed on a Bruker D8 QUEST diffractometer. The structures were solved by direct methods using SHELXT [25]. Refinement and further calculations were carried out using SHELXL [25]. The non-H atoms were refined anisotropically, apart from some O atoms of disordered crystallization water, which were refined isotropically. H—atoms were placed in idealized positions and refined using riding model. H—atoms were not considered for disordered water molecules in **2·3H₂O**; however, they were included in the calculation of the molecular weight. The unit cell of **1** was transformed in order to allow direct comparison to the structure of **1·5H₂O**, therefore it does not conform to the requirements of a reduced cell. Due to high absorption and relatively poor quality of the crystals of **1** after solid state transformation, some residual electron density peaks appear around W. The crystallographic data for **1** and **2·3H₂O** are presented in Table 1, with structure parameters for **1·5H₂O** [22] added for comparison. Graphics were created with Mercury 3.9.10. The analysis of coordination polyhedra was performed using SHAPE 2.1 [26]. CCDC 1885028 and 1885029 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via <http://www.ccdc.cam.ac.uk/conts/retrieving.html>

Table 1. Structure and refinement data for **1** and **2·3H₂O**; data for **1·5H₂O** included for comparison.

	1·5H₂O ¹	1	2·3H₂O
Empirical formula	$C_{46}H_{82}Cu_3N_{28}O_5W_2$	$C_{46}H_{72}Cu_3N_{28}W_2$	$C_{46}H_{82}Cu_3N_{28}O_5W_2$
Formula weight	1665.72	1575.57	1665.72
Crystal system	triclinic	triclinic	monoclinic
Space group	$P-1$	$P-1$	$P2_1/n$
a [Å]	8.9255(11)	9.0228(5)	14.0566(7)
b [Å]	10.0938(13)	10.8399(8)	12.8685(7)
c [Å]	19.549(3)	17.8734(13)	36.7198(19)
α [°]	78.342(5)	75.409(6)	90.0
β [°]	83.074(5)	87.810(4)	100.821(2)
γ [°]	69.472(5)	60.092(7)	90.0
V [Å ³]	1613.0(4)	1458.5(8)	6524.0(3)
Z	1	1	4
dc [g·cm ⁻³]	1.715	1.794	1.696
μ [mm ⁻¹]	4.590	5.064	4.539
F(000)	829	779	3276
θ range [°]	1.07–28.28	2.79–25.4	2.37–29.15
Reflns collected	21,064	5 261	17,526
No. of params	380	356	746
R_{int}	0.0693	0.0532	0.0362
Goof (F ²)	-	1.018	1.059
R_1 [$I > 2\sigma(I)$]	0.0693	0.0708	0.0446
ωR_2 (all data)	0.1662	0.1510	0.0973

¹ Data from reference [22].

3. Results and Discussion

3.1. Synthesis and Post-Synthetic Modification

The reaction between $[Cu(\text{cyclam})]^{2+}$ and $[W(\text{CN})_8]^{3-}$ building blocks in water-alcohol solution leads to the formation of $\{[Cu^{\text{II}}(\text{cyclam})]_3[W^{\text{V}}(\text{CN})_8]_2 \cdot 5\text{H}_2\text{O}\}_n$ (**1·5H₂O**) of the ladder-chain structure reported earlier [22]. The crystals of **1·5H₂O** are stable under ambient conditions, but at slightly

elevated temperatures (30–50 °C), they lose crystallization water to result in an anhydrous $\{[\text{Cu}^{\text{II}}(\text{cyclam})]_3[\text{W}^{\text{V}}(\text{CN})_8]_2\}_n$ (**1**) form. The process occurs with retention of crystallinity and thus the structure of **1** could be established from a single crystal XRD measurement. The single-crystal-to-single-crystal structural transformations upon desolvation are relatively rare [5,27,28]. In the family of the coordination polymers based on $[\text{Ni}(\text{cyclam})]^{2+/3+}$ ions and polycyanometallates, we observed this phenomenon only for the 2D $\{[\text{Ni}^{\text{II}}(\text{cyclam})]_3[\text{W}^{\text{V}}(\text{CN})_8]_2\}_\infty$ network [5]. The crystals of **1** exposed at 25 °C to humid air (40–60% RH) regain their original composition. However, upon rehydration, the crystals collapse and the PXRD results (Figure S1) show that the rehydration process leads to a mixture of phases, one of which is the original **1**·5H₂O form. We followed the dehydration process using the gravimetric dynamic vapor sorption method. The sample was subjected to decreasing relative humidity from 40% to 0% at a constant temperature of 298K. The sharp decrease in mass starts at about 15% RH (Figure S2), and takes place in two unresolved steps, which correspond to the subsequent loss of 3 and 2 water molecules.

When the reaction between the same building blocks is carried out in a water solution under slow diffusion conditions, a 0D pentanuclear assembly of the formula $\{[\text{Cu}^{\text{II}}(\text{cyclam})(\text{H}_2\text{O})]_2[\text{Cu}^{\text{II}}(\text{cyclam})][\text{W}^{\text{V}}(\text{CN})_8]_2 \cdot 3\text{H}_2\text{O}\}$ (**2**·3H₂O) is obtained. Interestingly, quick precipitation from the water solution does not lead to the same result. The analysis of the PXRD patterns (Figure S3) suggests that **1**·5H₂O is formed instead. Similar disparity between the course of slow diffusion and quick precipitation reactions was observed for the 3D $\{[\text{Ni}^{\text{II}}(\text{cyclam})]_2[\text{W}^{\text{IV}}(\text{CN})_8]\}_\infty$ network [29]. Due to a very low yield and long time taken to grow crystals, the dehydration of **2**·3H₂O was not studied, but it seems likely that both lattice and coordinated water could be removed, possibly with the formation of additional CN-bridges between neighboring molecules.

3.2. Description of Structures

Upon dehydration of **1**·5H₂O, the triclinic system and space group P-1 are retained, but cell parameters change significantly (Table 1). In particular, period *c* shortens by 1.7 Å (8.7%) and the γ angle decreases by 9.38° (13.5%). The cell volume decreases by 155 Å³ (9.6%) per formula unit, which corresponds quite well to the volume expected for five H-bound water molecules. The connectivity of the coordination network is retained, with Cu and W CN-bridged centers forming a 1D ladder-like chain. Apart from the removed water molecules, the asymmetric unit of **1** (Figure 1) contains the same set of atoms as **1**·5H₂O, including Cu1 and W1 ions located in general positions and Cu2 ions in the inversion center. Each tungsten center (W1) is coordinated by eight cyanide ligands, three of which form bridges to copper centers (two Cu1 and one Cu2). The copper ions are connected with two tungsten centers through cyanide bridges coordinated in axial positions, with equatorial positions blocked by the cyclam ligand. The dehydration process strongly affects the geometry of CN-bridges and coordination polyhedra. The Continuous Shape Measure analysis [26] shows that the geometry of the $[\text{W}(\text{CN})_8]^{3-}$ ion in the hydrated form **1**·5H₂O is close to square antiprism, while for the anhydrous form (**1**), it changes to a triangular dodecahedron (Table S1). Both Cu ions in **1**·5H₂O show a highly distorted octahedral geometry (Table S2), due to the significant Jahn-Teller effect causing elongation of the axial bonds. The non-centrosymmetric Cu1 center is more distorted, due to the asymmetry of the axial bonds, one of which (Cu1-N1 of 2.698 Å) can be classified as a semi-coordination bond. Upon dehydration, the Cu2-N5 distances shorten by 0.053 Å, while the Cu2-N distances become more symmetrical (Table 2). As an effect, the octahedral geometry of both copper centers in the anhydrous form **1** is less distorted (Table S2). The angles of cyanide bridges (Cu1-N1-C1, Cu1-N4-C4, Cu2-N5-C5) decrease by 13.74–24.78° (Table 2). Therefore, dehydration leads to significant shortening of the distances between the CN-linked metal centers by 0.219–0.392 Å.

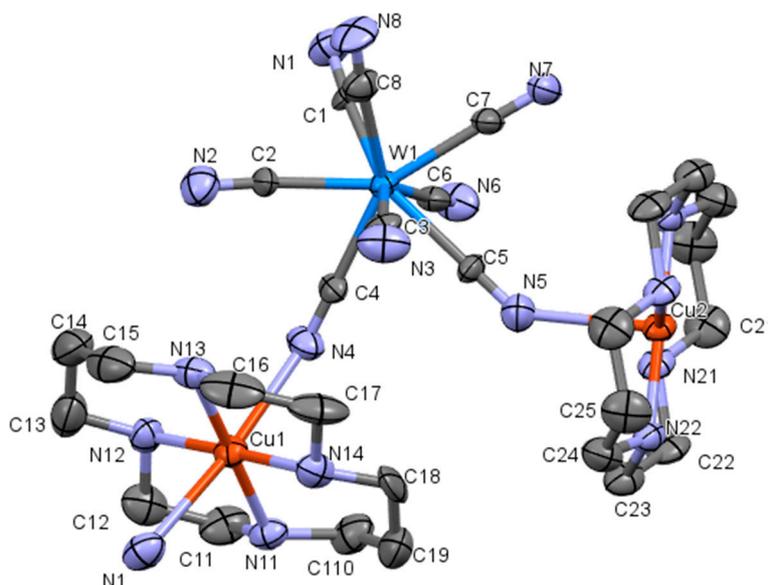


Figure 1. The asymmetric unit of **1** with atom numbering; ellipsoids at 50% probability.

Table 2. Selected bond lengths (Å) and angles (°) in **1·5H₂O** and **1**.

Bond	1·5H₂O ¹	1	Angle	1·5H₂O ¹	1
W1-C1	2.181	2.149	Cu1-N1-C1	143.0	129.3
W1-C2	2.157	2.129	Cu1-N4-C4	157.1	135.3
W1-C3	2.154	2.159	Cu2-N5-C5	157.1	142.1
W1-C4	2.169	2.155	N11-Cu1-N12	93.22	86.1
W1-C5	2.165	2.152	N12-Cu1-N13	86.44	95.5
W1-C6	2.152	2.159	N13-Cu1-N14	94.23	86.6
W1-C7	2.158	2.149	N14-Cu1-N11	85.93	91.9
W1-C8	2.169	2.214	N21-Cu2-N22	86.34	86.2
Cu1-N1	2.698	2.606	W1-C1-N1	176.75	176.3
Cu1-N4	2.510	2.585	W1-C4-N4	177.74	179.4
Cu2-N5	2.556	2.503	W1-C5-N5	174.43	176.5
Cu1-N11	2.004	2.037			
Cu1-N12	2.023	2.027			
Cu1-N13	2.024	2.018			
Cu1-N14	2.014	1.998			
Cu2-N21	2.015	2.011			
Cu2-N22	2.032	2.028			

¹ Data from reference [22].

Projections of structure along the crystallographic axes for **1·5H₂O** and **1** (Figure 2) show the ladder-like topology retained upon dehydration. In both structures, the ladders lie in the (11–1) plane and run along the [1–10] direction. However, the changes in the coordination geometry of the [W(CN)₈]^{3–} ion and different alignment of the cyclam ligand in both forms are noticeable. Likewise, the bending of the CN-bridges can be seen, particularly in the view along the *a* period. Upon dehydration, the distance between the neighboring chains in the [001] direction shortens, which is visible in the projections along the *b* and *c* axes. The distance between the Cu1 and W1 centers from the neighboring chains shortens by 0.725 Å. In the hydrated form **1·5H₂O**, the coordination ladders are bound into a 3D supramolecular network by H-bonds between cyanide ligands, NH groups of the cyclam molecules, and crystallization water [22]. Upon loss of crystallization water, the hydrogen bonds are rearranged (Figure S4). The ladders are bound into stacks along the [110] direction, resulting in the 2D supramolecular network. Conversely, there is no connection in the [001] direction, despite shorter distance between the ladders. Moreover, due to the bending of CN-bridges, intramolecular H-bonds appear. They link the terminal CN ligands with the NH groups of cyclam from the W1 and

Cu1 centers, and thus stabilize strong bending of the Cu1-N1-C1 bridge (Table 2). This may be the reason why the rehydration process leads to the collapse of crystals and is not fully reversible.

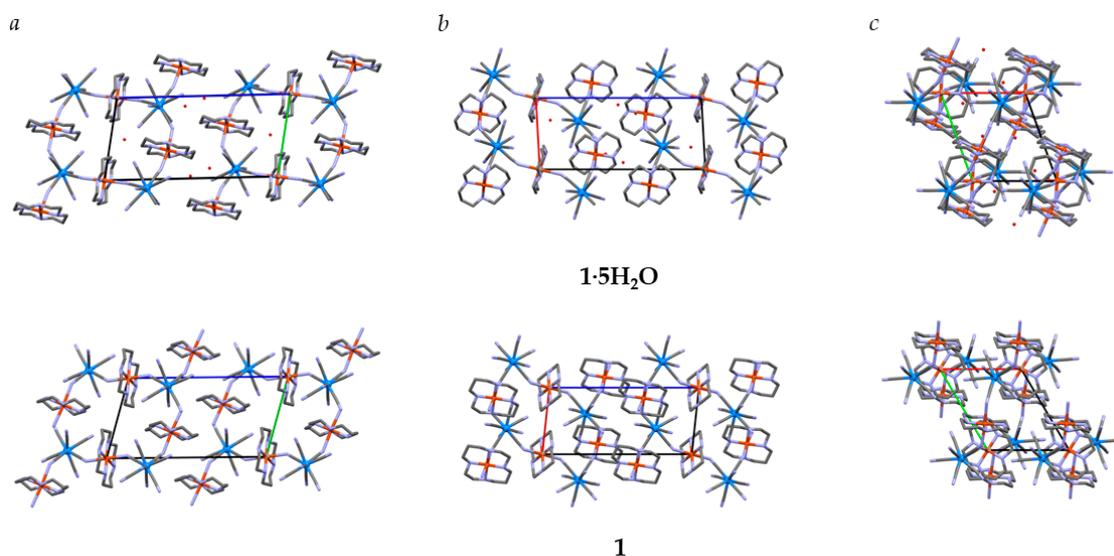


Figure 2. Projection of structures of $1 \cdot 5\text{H}_2\text{O}$ and **1** along (a–c) crystallographic directions.

The 0D $\{[\text{Cu}^{\text{II}}(\text{cyclam})(\text{H}_2\text{O})]_2[\text{Cu}^{\text{II}}(\text{cyclam})][\text{W}^{\text{V}}(\text{CN})_8]_2\} \cdot 3\text{H}_2\text{O}$ assembly (**2**· $3\text{H}_2\text{O}$) is obtained from the same building blocks as the $1 \cdot 5\text{H}_2\text{O}$ ladder chain, but under different synthetic conditions. The compound crystallizes in a monoclinic system, space group $P 2_1/n$ (Table 1). Its structure is composed of Z-shaped pentanuclear molecules containing two W and three Cu centers connected by CN-bridges with water molecules coordinated to the terminal Cu ions (Figure 3). There are two symmetrically independent molecules. Each of them is centrosymmetric with the central Cu ions located at the inversion centers. Therefore, the asymmetric unit consist of two tungsten and two copper centers (W1, W2, Cu1, Cu3) in general positions and two copper centers (Cu2 and Cu4) in special positions. Each W atom is coordinated by eight cyanide ligands, two of which form bridges to two Cu centers (W1 with Cu1 and Cu2; W2 with Cu3 and Cu4). The coordination geometry of both W centers is close to ideal square antiprism (Table S1). All Cu ions have cyclam ligand coordinated in equatorial positions. The centrosymmetric Cu2 and Cu4 atoms are linked by two cyanide bridges, whereas the Cu1 and Cu3 centers located at the ends of the Z-shaped molecule are coordinated by one cyanide bridge and one water molecule in axial positions. The octahedral coordination geometry is distorted due to the Jahn-Teller effect (Table S2). There is a large disproportion of the extent of this distortion between the centrosymmetric Cu centers: Cu2 with the very long and bent NC-bridges (Table 3) shows very large distortion, while Cu4 shows the lowest distortion among the Cu centers in **1** and **2**.

The average length of the Cu-N bond in the CN-bridge is 2.543 Å, which is shorter than for $1 \cdot 5\text{H}_2\text{O}$ (2.588 Å) or **1** (2.565 Å). The average value of the Cu-N-C angles of the CN-bridges (147.87°) is closer to those observed in $1 \cdot 5\text{H}_2\text{O}$ (153.58°) than those in the anhydrous form **1** (132.31°). The distances between the CN-bridged metal centers are shorter for the central Cu (W1-Cu2, W2-Cu4) than for the terminal ones (W1-C1, W2-Cu3) by about 0.25 Å, due to the stronger bending of the central CN bridges.

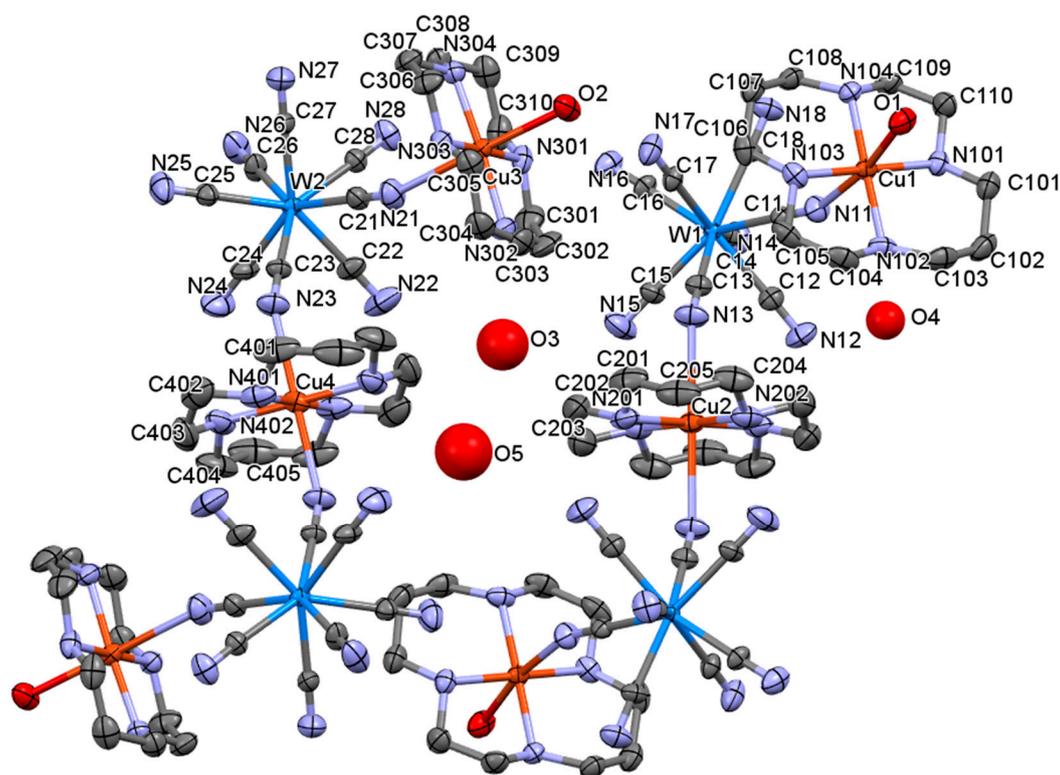


Figure 3. Structure of 2·3H₂O with atom numbering; ellipsoids at 50% probability.

Table 3. Selected bond lengths (Å) and angles (°) in 2·3H₂O.

Cu1-N11	2.523	Cu1-N11-C11	157.9
Cu1-O1	2.518	Cu2-N13-C13	137.2
Cu2-N13	2.596	Cu3-N21-C21	153.9
Cu3-N21	2.594	Cu4-N23-C23	142.8
Cu3-O2	2.422	N101-Cu1-N102	93.3
Cu4-N23	2.457	N102-Cu1-N103	86.0
Cu1-N101	2.013	N103-Cu1-N104	95.0
Cu1-N102	2.037	N104-Cu1-N101	85.7
Cu1-N103	2.011	N201-Cu2-N202	86.0
Cu1-N104	2.023	N301-Cu3-N302	85.8
Cu2-N201	2.008	N302-Cu3-N303	92.8
Cu2-N202	2.019	N303-Cu3-N304	86.4
Cu3-N301	2.021	N304-Cu3-N301	95.0
Cu3-N302	2.029	N401-Cu4-N402	85.6
Cu3-N303	2.008	W1-C11-N11	178.5
Cu3-N304	2.013	W1-C13-N13	177.2
Cu4-N401	2.024	W2-C21-N21	179.6
Cu4-N402	2.016	W2-C23-N23	176.5

Projection along the *a* and *b* crystallographic directions (Figure 4) shows that the symmetrically independent molecules (marked red and blue) are arranged in independent layers in the (10–1) and (20–2) planes. They are bound into a 3D supramolecular network by the net of hydrogen bonds between cyanide ligands, NH groups of cyclam, aqua ligands, and crystallization water (Figure S5). The molecules of the same symmetry are bound within the layers by H-bonds formed by coordinated water, which donates two protons to two CN ligands of two neighboring molecules. The layers are linked by direct bonds between CN ligands and NH groups, as well as by H-bonds between CN-ligands mediated by crystallization water.

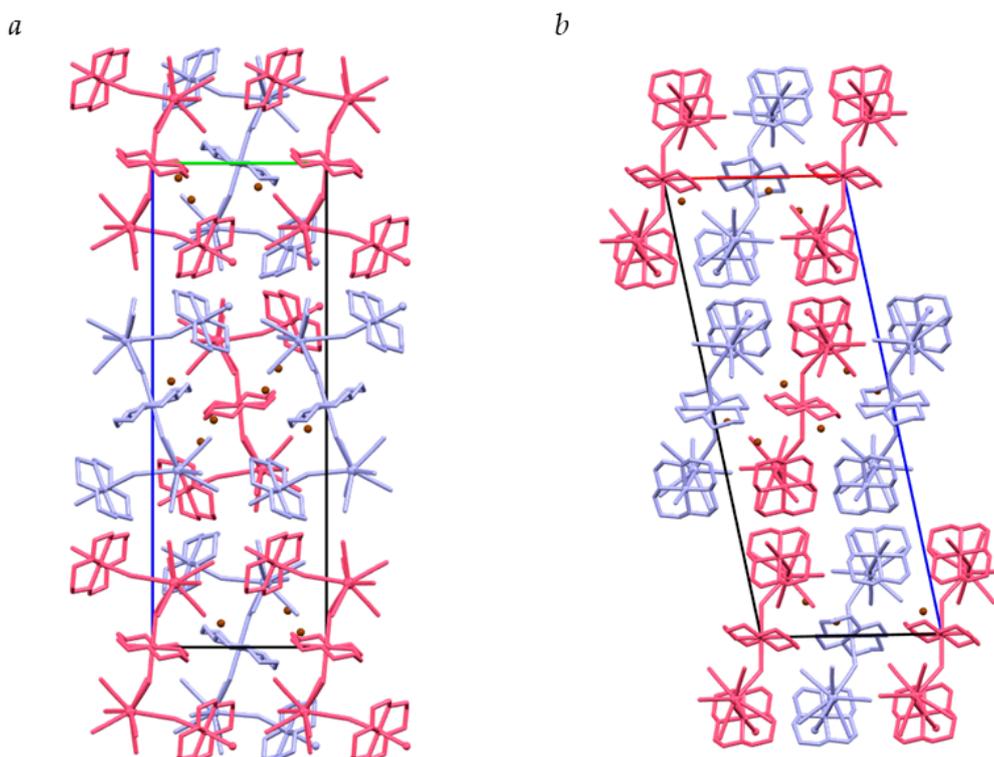


Figure 4. Projection of the structure of $2 \cdot 3\text{H}_2\text{O}$ along (a,b) crystallographic directions; symmetrically independent molecules marked pink and violet.

3.3. Magnetic Properties

All compounds were characterized by DC susceptibility measurements at the applied field of 1000 Oe in the temperature range 1.8–300 K (Figure 5) and magnetization at 1.8 K in the field up to 70 kOe (Figure 6). At the high temperature limit, the value of χT for all compounds is close to the Curie constant of $1.88 \text{ cm}^3\text{K/mol}$, expected for three Cu^{II} and two W^{V} ions with $s = 1/2$ and $g = 2$, and it remains constant down to about 60 K. Below that temperature, differences in magnetic behavior between the compounds become apparent. For $1 \cdot 5\text{H}_2\text{O}$, where intra-chain interactions are predominantly ferromagnetic [22], the χT curve rises slightly to reach the maximum of $2.03 \text{ cm}^3\text{K/mol}$ at 7.5 K. For the dehydrated sample **1**, the χT value decreases below 60 K down to $0.87 \text{ cm}^3\text{K/mol}$ at 1.8 K, indicating predominantly antiferromagnetic interaction within the chains. This change in magnetic behavior can be attributed to the bending of the CN-bridges. As was shown before [30–32] for d^9 configuration, the bridge geometry strongly affects the character and strength of magnetic superexchange through the π^* orbitals of CN^- ions, which changes from ferromagnetic for a linear bridge to antiferromagnetic for strongly bent bridges. The antiferromagnetic character of intra-chain interactions in **1** is also visible in the magnetization at 1.8 K. The increase of the magnetic moment in an increasing field is much slower for the dehydrated sample **1** than for $1 \cdot 5\text{H}_2\text{O}$ and it reaches only $2.98 \text{ N}\beta$ at 70 kOe, which is far from the expected saturation value of $5 \text{ N}\beta$. The magnetic characteristic that we obtained for $1 \cdot 5\text{H}_2\text{O}$ is very similar, but not identical with the published data [22]. Slightly lower values of magnetization and high-temperature susceptibility are most probably caused by the relatively low accuracy of small sample weight assessment. The difference in the χT course in the low temperature range, where we observed a lower maximum at a slightly higher temperature, may be due to partial loss of water and onset of the structural transformation in one of the samples.

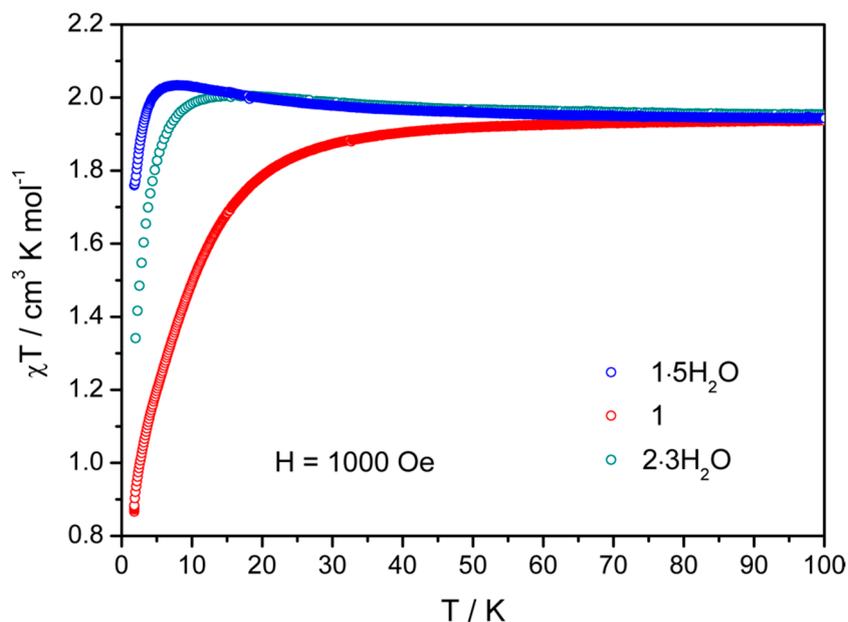


Figure 5. Temperature dependence of the DC magnetic susceptibility at 1000 Oe for $1.5\text{H}_2\text{O}$, **1** and $2.3\text{H}_2\text{O}$.

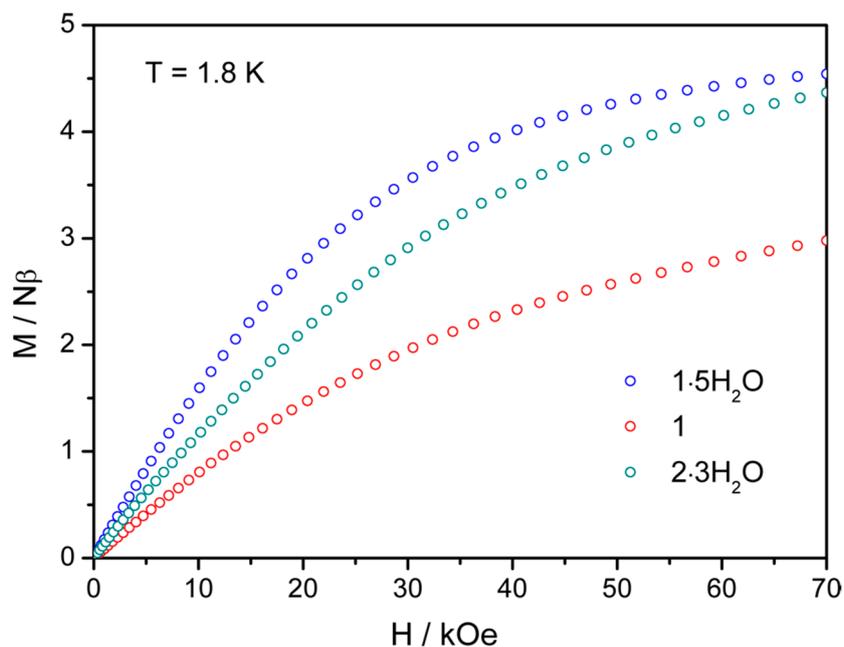


Figure 6. Field dependence of magnetization at 1.8 K for $1.5\text{H}_2\text{O}$, **1** and $2.3\text{H}_2\text{O}$.

The magnetic behavior of $2.3\text{H}_2\text{O}$ is similar to that of $1.5\text{H}_2\text{O}$. In the χT dependence on T (Figure 5), a slight increase below 60 K is observed with a broad maximum of $2.01 \text{ cm}^3\text{K/mol}$ at about 14 K and an abrupt drop below 7 K. The magnetization curve lies slightly below the one of $1.5\text{H}_2\text{O}$ and at 70 kOe reaches $4.37 \text{ N}\beta$ (Figure 6). These results suggest the presence of weak intramolecular ferromagnetic interactions. This is consistent with the observed geometry of CN-bridges, which in $2.3\text{H}_2\text{O}$ is similar to that in $1.5\text{H}_2\text{O}$.

4. Conclusions

We have shown that similarly to $[\text{Ni}(\text{cyclam})]^{2+/3+}$ complexes, the $[\text{Cu}(\text{cyclam})]^{2+}$ ion can be used as a cationic building block in combination with polycyanometallates to construct

solvatomagnetic coordination networks. The 1D $\{[\text{Cu}^{\text{II}}(\text{cyclam})]_3[\text{W}^{\text{V}}(\text{CN})_8]_2 \cdot 5\text{H}_2\text{O}\}_n$ ($1 \cdot 5\text{H}_2\text{O}$) polymer exhibits a rare single-crystal-to-single-crystal structural transformation to the anhydrous $\{[\text{Cu}^{\text{II}}(\text{cyclam})]_3[\text{W}^{\text{V}}(\text{CN})_8]_2\}_n$ (**1**) form. Although the topology of the compound is retained, the modification of bonds geometry and intermolecular interactions results in a noticeable change of magnetic properties, switching the predominant intra-chain interactions from ferromagnetic in $1 \cdot 5\text{H}_2\text{O}$ to antiferromagnetic in **1**. The dehydration process is not fully reversible, probably due to the formation of intra-chain H-bonds, which stabilize strong bending of the CN-bridges. The rehydration causes collapse of the crystals and partial degradation of the sample. Apart from the post-synthetic modification of $1 \cdot 5\text{H}_2\text{O}$, we have also shown that the same building block can be arranged in a different structure by slight modification of synthetic conditions. The $\{[\text{Cu}^{\text{II}}(\text{cyclam})(\text{H}_2\text{O})]_2[\text{Cu}^{\text{II}}(\text{cyclam})][\text{W}^{\text{V}}(\text{CN})_8]_2\} \cdot 3\text{H}_2\text{O}$ pentanuclear Z-shaped molecule ($2 \cdot 3\text{H}_2\text{O}$), obtained from water under slow diffusion conditions, exhibits weak ferromagnetic interactions through CN-bridges.

Supplementary Materials: The following are available online at <http://www.mdpi.com/2073-4352/9/1/45/s1>, Figure S1: PXRD pattern for the rehydrated sample of **1** in comparison to $1 \cdot 5\text{H}_2\text{O}$ and **1**; Figure S2: Mass loss upon dehydration of $1 \cdot 5\text{H}_2\text{O}$ monitored by dynamic vapor sorption method; Figure S3: PXRD pattern for sample obtained from $[\text{Cu}(\text{cyclam})](\text{NO}_3)_2$ and $\text{Na}_3[\text{W}(\text{CN})_8]$ by quick precipitation from water solution in comparison to $1 \cdot 5\text{H}_2\text{O}$ and $2 \cdot 3\text{H}_2\text{O}$; Table S1: Continuous shape measure parameters for octa-coordinated W centers in the structures of $1 \cdot 5\text{H}_2\text{O}$, **1** and $2 \cdot 3\text{H}_2\text{O}$; Table S2: Continuous shape measure parameters for hexa-coordinated Cu centers of $1 \cdot 5\text{H}_2\text{O}$, **1** and $2 \cdot 3\text{H}_2\text{O}$; Figure S4: Inter- and intra-chain H-bonds in **1**; Figure S5: Inter- and intra-molecular H-bonds in $2 \cdot 3\text{H}_2\text{O}$.

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