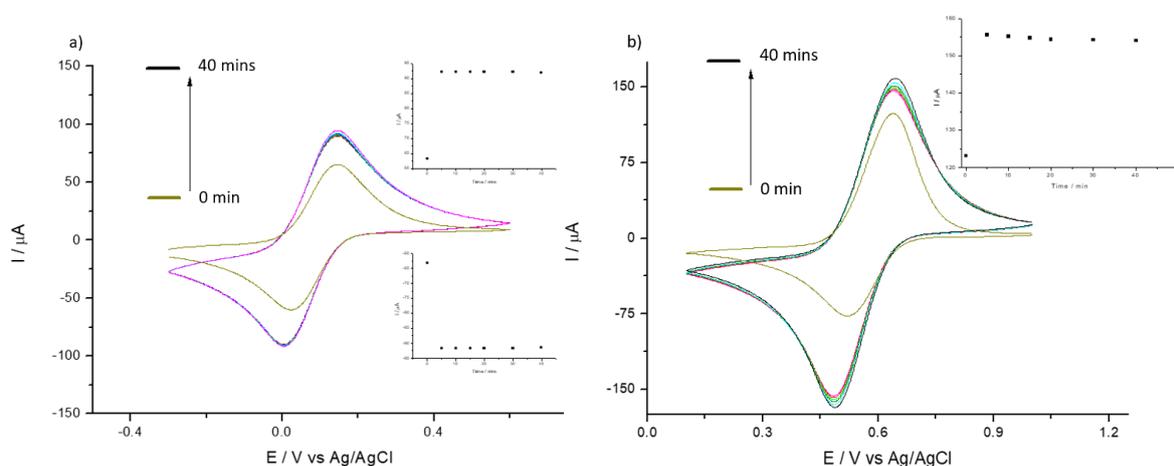


1 Electronic Supporting Information (ESI)

2 **Voltammetry at Hexamethyl-p-terphenyl**
 3 **poly(benzimidazolium) (HMT-PMBI)-Coated Glassy**
 4 **Carbon Electrodes: Charge Transport Properties and**
 5 **Investigation towards Detection of Uric and Ascorbic**
 6 **Acid**

7 Matthew Rees ¹, Andrew G. Wright ², Steven Holdcroft ² and Paolo Bertoncello ^{1,3,*}

8 **S1.**



9
 10 **Figure 1.** CVs of 1% HMT-PMBI coated film recorded during loading in an aqueous solution of 5 mM
 11 $K_4Fe(CN)_6$ (a) and 3 mM K_2IrCl_6 (b); supporting electrolyte 0.1 M NaCl. Scan rate of 100 mV s⁻¹.

12 **S2.**

13 The surface coverage values (Γ , mol cm⁻²) of HMT-PMBI coated films with incorporated the
 14 redox probes $K_4Fe(CN)_6$ or K_2IrCl_6 were calculated from the CVs recorded at low scan rates, after
 15 transferring the modified electrodes into the supporting electrolyte solution without the redox specie
 16 (Figure 1a,c), e.g., when the CVs display thin-layer characteristics using the relation:

17

$$\Gamma = \frac{Q}{nFA} \quad (1)$$

18 where Q (C) is the charge on the forward or reverse scan, n is the number of electrons transferred (n
 19 = 1 for both $K_4Fe(CN)_6$ and K_2IrCl_6), A (cm²) is the geometric area of the electrode, and F is the Faraday
 20 constant (96486 C mol⁻¹). These values can also be expressed in terms of concentration, C_0^* , by
 21 dividing the surface coverage with the thickness of HMT-PMBI coated film, Φ (cm), calculated in dry
 22 conditions using the profilometer, e.g.,

$$C_0^* = \frac{\Gamma}{\Phi} \quad (2)$$

23 The apparent diffusion coefficients were calculated by using two methods, e.g., (1) the Randles-
24 Sevcik and (2) the Anson's plot methods. For the method using the Randles-Sevcik equation[1], e.g.,

$$I_{p,a,c} = (2.69 \times 10^5) n^{3/2} A D_{app}^{1/2} C_0^* \nu^{1/2} \quad (3)$$

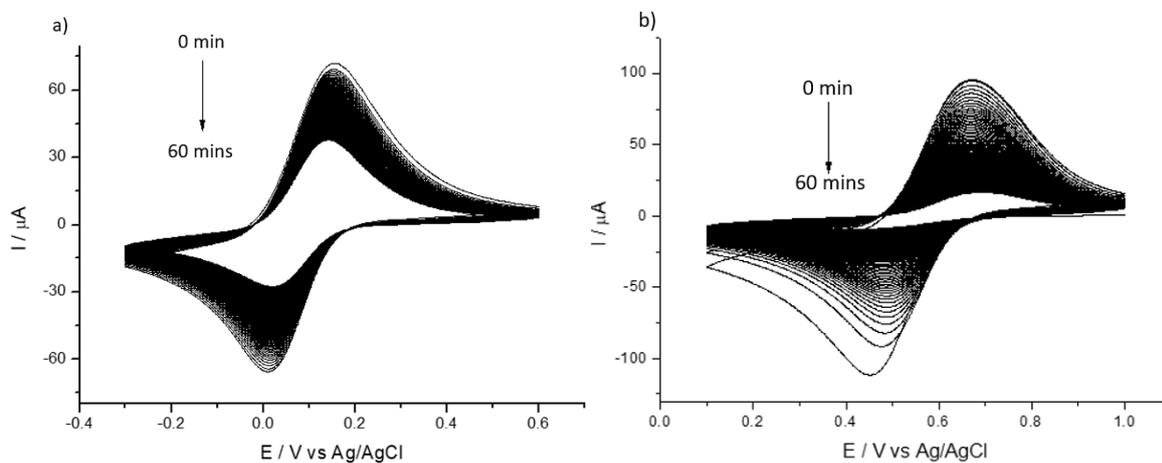
25 we plotted the anodic peak currents, $I_{p,a}$ in the case of $\text{Fe}(\text{CN})_6^{2-}$ or the cathodic peak current $I_{p,c}$ in the
26 case of IrCl_6^{2-} versus the square root of the scan rate, $\nu^{1/2}$, with the tacit assumption that the redox
27 process is reversible. To underline the fact that the ΔE_p for each recorded CV increased monotonically
28 with the scan rate and this will lead to an underestimation of D_{app} . The slope of these plots recorded
29 in the faster scan rate regime, combined with thickness of the polymer film evaluated using the
30 profilometer and the number of electroactive species obtained by coulometric integration of the
31 anodic (or cathodic) peak currents under thin-layer conditions allowed the evaluation of the apparent
32 diffusion coefficient values of HMT-PMBI coated electrodes. Instead, for the Anson's method,
33 potential-step chronocoulometry was used to determine the values of D_{app} from the slope of the plots
34 of the charge Q vs. the square root of time, $t^{1/2}$, using the relation[2]:

$$D_{app} = \left[\frac{S\Phi\pi^{1/2}}{2F\Gamma} \right]^2 = \left[\frac{S\pi^{1/2}}{2FC} \right]^2 \quad (4)$$

35 where S is the chronocoulometric slope ($\text{C cm}^{-2} \text{t}^{1/2}$), and Γ , Φ , and F with the conventional meaning
36 as previously mentioned.

37

38 S3.



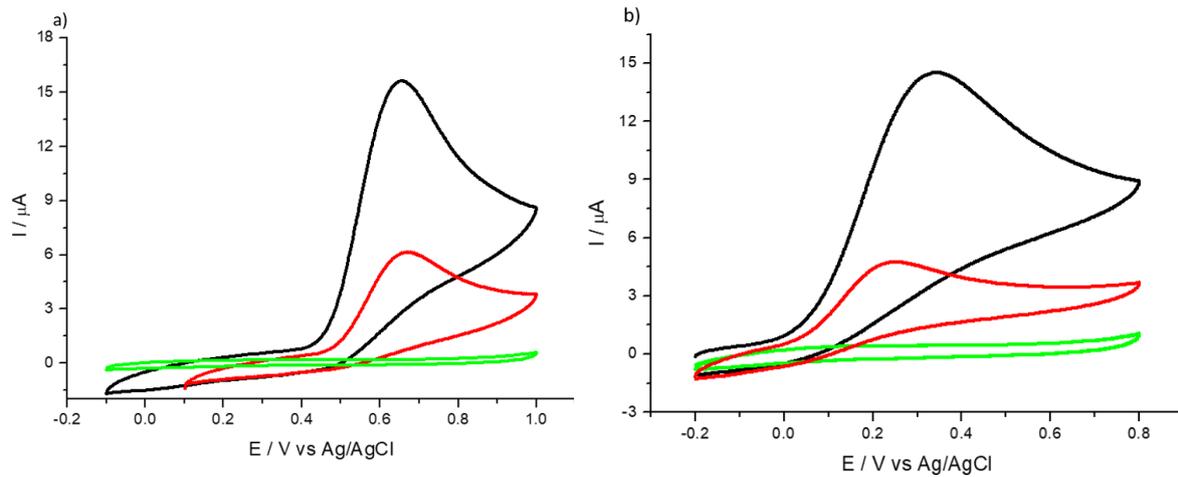
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40 **Figure 3.** CVs of 1% HMT-PMBI coated film loaded in 5 mM $\text{K}_4\text{Fe}(\text{CN})_6$ (a) and 3 mM K_2IrCl_6 (b) after
41 transferring to 0.1 M NaCl supporting electrolyte and continuous cycling for 1 hour. Scan rate: from
42 100 mV s^{-1} .

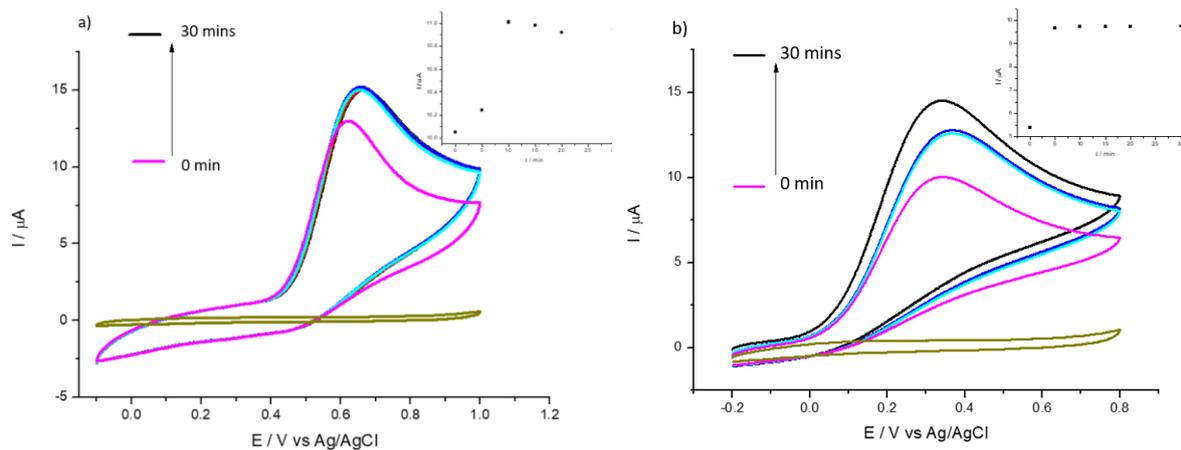
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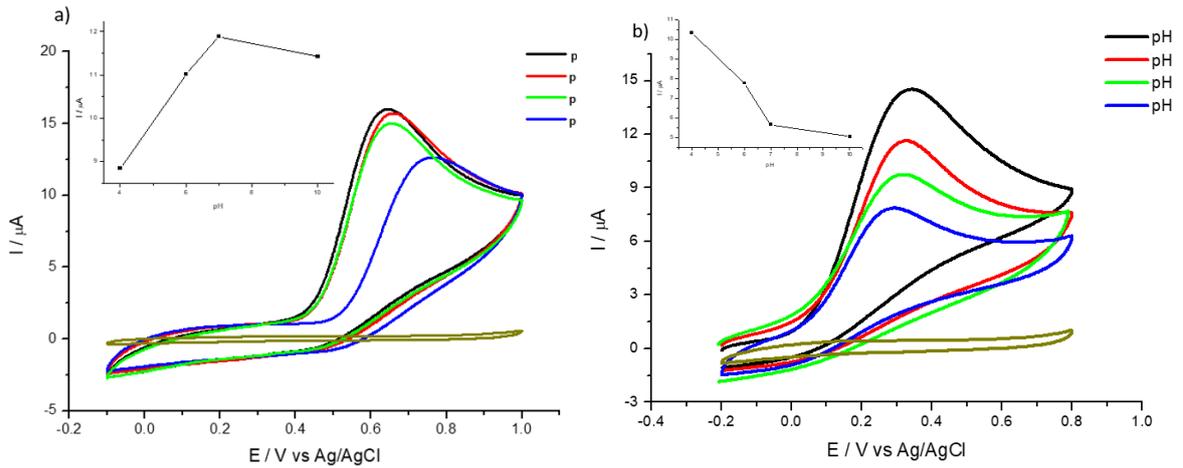
46 **S4.**

47 **Figure 4.** CVs of 1% HMT-PMBI coated electrode recorded in a solution containing 1mM UA at pH 7
 48 (a) and 1 mM AA at pH 4 (b) (black) and immediately after transferring in 0.1 M NaCl supporting
 49 electrolyte (red); Scan rate of 50 mV s^{-1} .

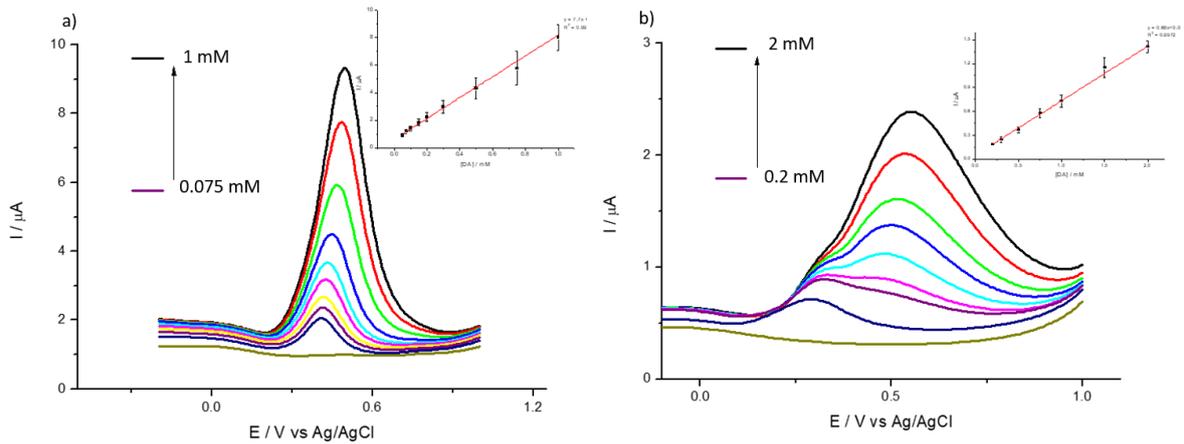
50 **S5.**

51 **Figure 5.** CVs of 1% HMT-PMBI coated electrode recorded in 1mM UA at pH 7 (a) and 1 mM AA at
 52 pH 4 (b) at different loading time; supporting electrolyte: 0.1 M NaCl; Scan rate of 50 mV s^{-1} .

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54 **S6.**

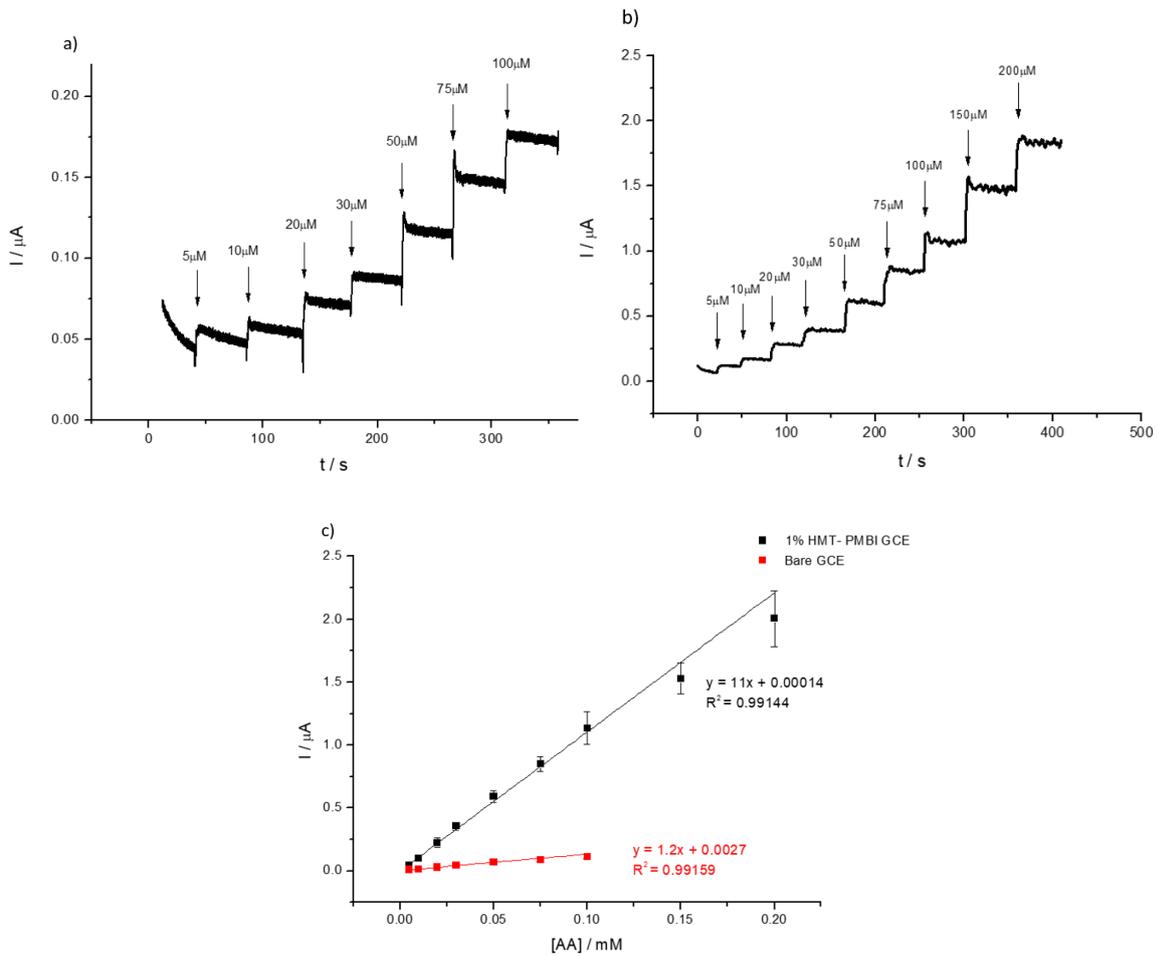
55 **Figure 6.** CVs of 1% HMT-PMBI coated electrodes recorded in 1 mM UA (a) and 1 mM AA (b) at
 56 different pH values; supporting electrolyte: 0.1 M NaCl; Scan rate of 50 mV s⁻¹.

57 **S7.**

58 **Figure 7.** DPVs of bare GCE (a) and of 1% HMT-PMBI (b) coated electrode recorded in 0.1M NaCl in
 59 the presence of constant 0.05 mM AA and various concentrations of DA, from 75 μM to 1 mM for (a)
 60 and 0.2 mM to 2 mM for (b). Scan rate of 10 mV s⁻¹. Inset: plot of peak currents vs. concentration of
 61 DA.

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63 S8.



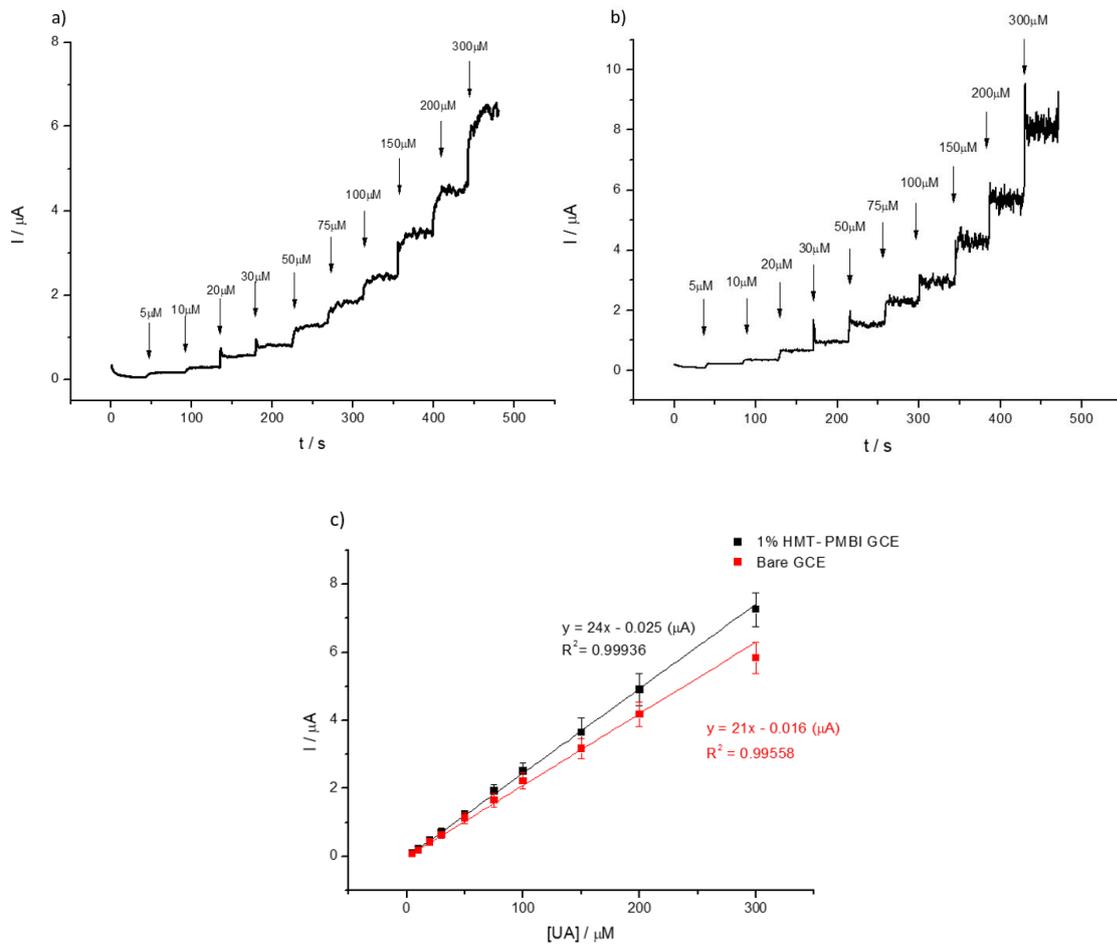
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66 **Figure 8.** Chronoamperometric (i - t) response of bare (a) and of 1% HMT-PMBI (b) coated electrode
 67 obtained with successive concentration of AA from $5 \mu\text{M}$ to 0.2 mM recorded in 0.1 M NaCl
 68 supporting electrolyte (pH 4), applied potential 0.3 V . (c) Calibration plot as a function of AA
 69 concentration as in (a, b). Error bars calculated from 3 repeat measurements.

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71 S9.

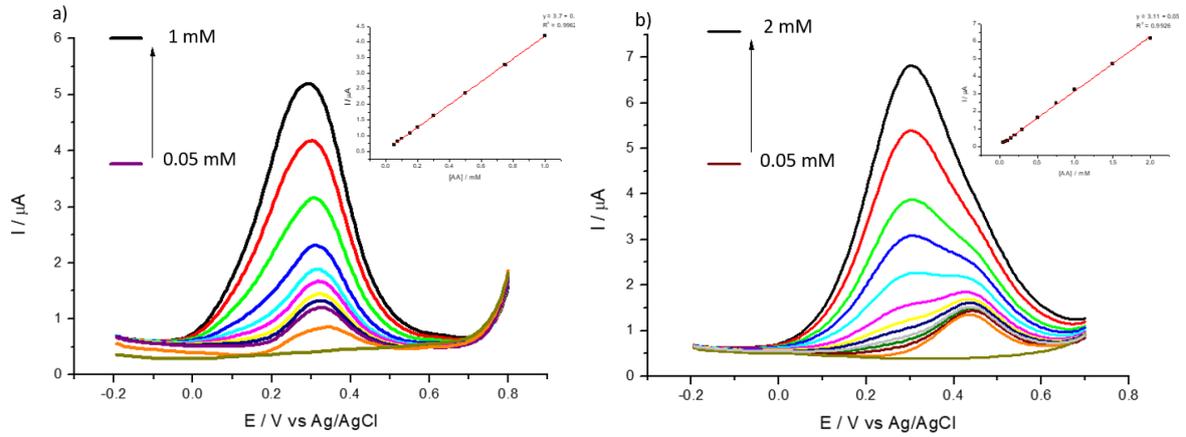


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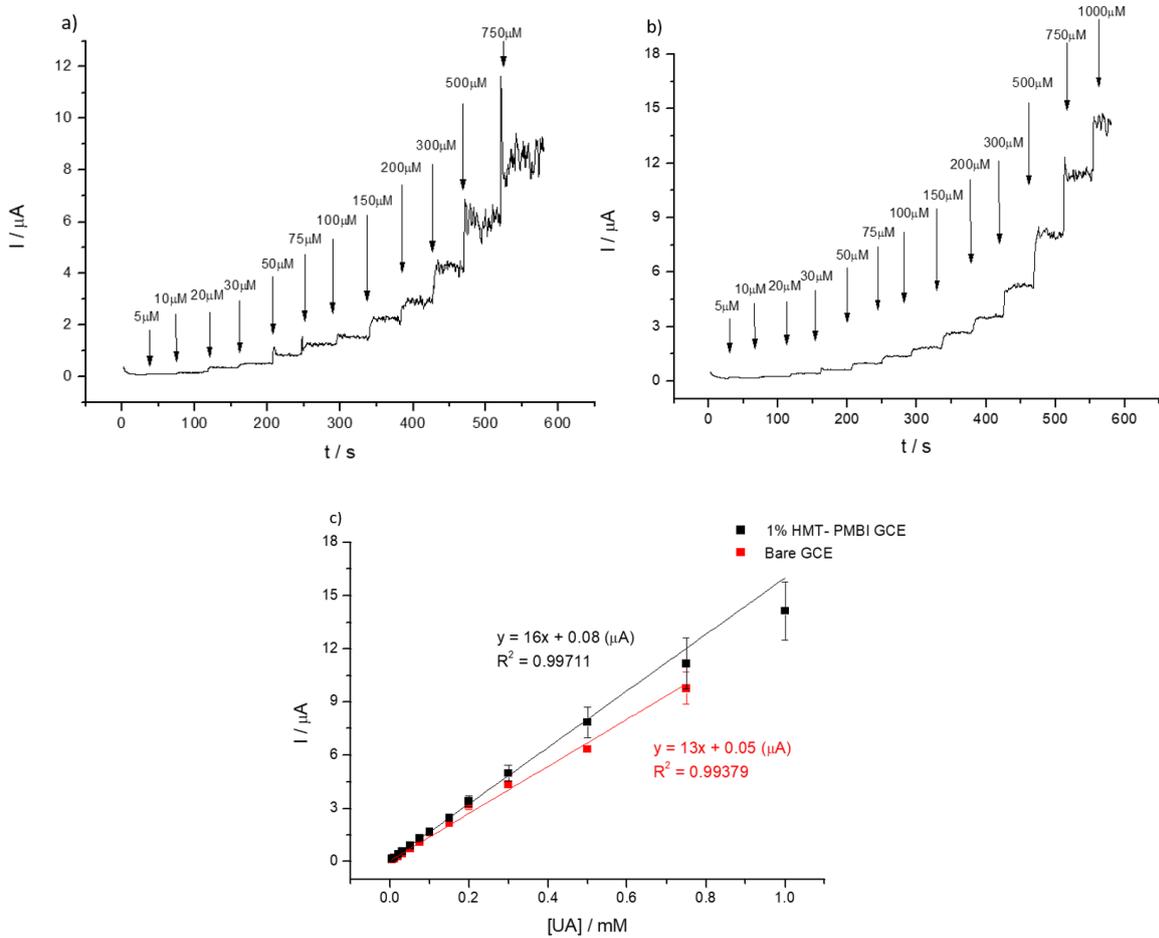
74 **Figure 9.** Chronoamperometric (i - t) response of bare (a) and of 1% HMT-PMBI (b) coated electrode
 75 obtained with successive concentration of UA from 5 μM to 0.3 mM recorded in 0.1 M NaCl
 76 supporting electrolyte (pH 7), applied potential 0.6 V. (c) Calibration plot as a function of UA
 77 concentration as in (a, b). Error bars calculated from 3 repeat measurements.

78

79 **S10.**

80 **Figure 10.** DPVs of bare GCE (a) and of 1% HMT-PMBI (b) coated electrode recorded in Surine® in
 81 the presence of constant 0.05 mM UA and various concentrations of AA, from 0.05 mM to 1 mM for
 82 (a) and 0.05 mM to 2 mM for (b). Scan rate of 10 mV s⁻¹. Inset: plot of peak currents vs. concentration
 83 of UA.

84

85 **S11.**

86

87

88 **Figure 11.** Chronoamperometric (i - t) response of bare GCE (a) and 1% HMT-PMBI (b) coated GCE
 89 applying 0.6 V respectively obtained with successive concentration of UA from 5 μM –0.75 mM of (a)
 90 and 5 μM –1 mM for (b) in Surine® (pH 6.8).

91 References

- 92 1. Bard, A.J. and L.R. Faulkner, *Electrochemical Methods: Fundamentals and Applications*. 2001, John Wiley & Sons
 93 Inc.: New York. p. 226-260.
 94 2. Buttry, D.A. and F.C. Anson, *EFFECTS OF ELECTRON EXCHANGE AND SINGLE-FILE DIFFUSION ON*
 95 *CHARGE PROPAGATION IN NAFION FILMS CONTAINING REDOX COUPLES*. Journal of the American
 96 Chemical Society, 1983. **105**(4): p. 685-689.
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