Supporting Information (SI)

Modulation of the selectivity in anions recognition processes by complexing hydrogen- and halogen- bonding interactions

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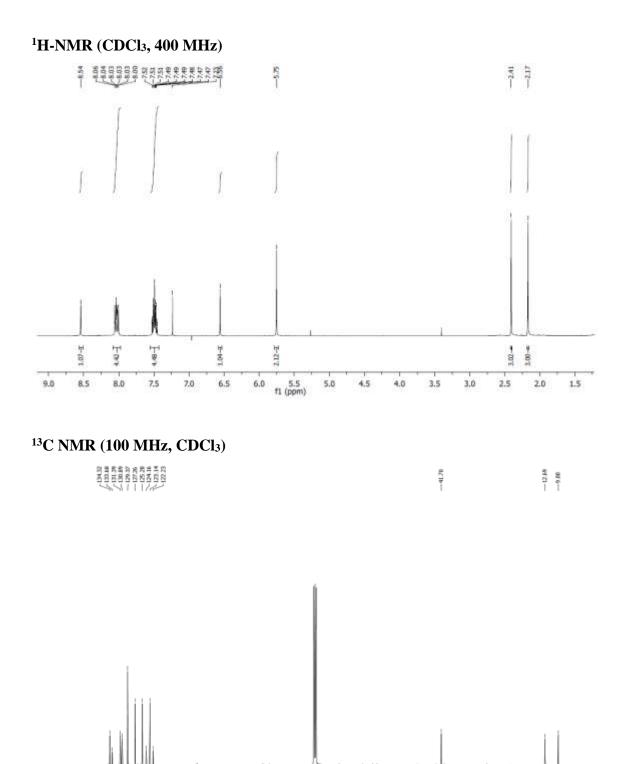
Table of contents

¹ H- and ¹³ C-NMR spectra of compounds 2-5	S 3
Figure S1. Changes in the fluorescence emission of 5 ²⁺ ·2PF ₆ ⁻ in CH ₃ CN	
upon addition of F ⁻ anions from 0 to 1.8 equiv	S 8
Figure S2. Changes in the fluorescence emission of 5 ²⁺ ·2PF ₆ ⁻ in CH ₃ CN	
upon addition of F ⁻ anions from 1.8 to 5 equiv	S 8
Figure S3. Changes in the fluorescence emission of $5^{2+} \cdot 2PF_6^-$ in CH ₃ CN upon addition of AcO ⁻ anion.	S9
Figure S4. Semilogarithmic plot for determining the detection limit of $5^{2+} \cdot 2PF_6^{-}$ towards H ₂ PO ₄ ⁻ .	S9

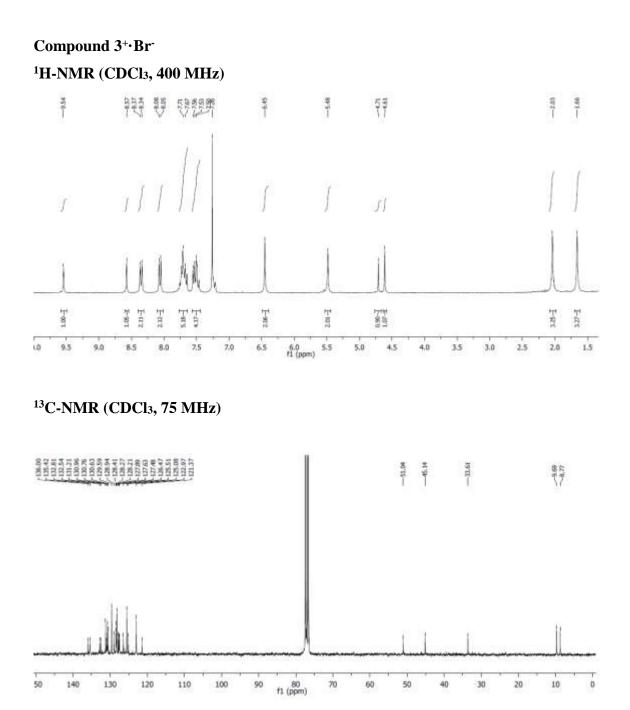
Figure S5. Semilogarithmic plot for determining the detection limit of

$5^{2+} \cdot 2PF_6^-$ towards AcO ⁻ .	S10
Figure S6. Semilogarithmic plot for determining the detection limit of	
$5^{2+} \cdot 2PF_6^-$ towards SO_4^{2-} .	S10
Figure S7. Job's plot for $5^{2+} \cdot 2PF_6^-$ and $H_2PO_4^{2-}$	S11
Figure S8. Job's plot for $5^{2+} \cdot 2PF_6^-$ and SO_4^{2-}	S11
Figure S9. Job's plot for $5^{2+} \cdot 2PF_6^-$ and AcO ⁻	S12
Figure S10. ¹ H NMR spectral changes of the receptor 5 ²⁺ ·2PF ₆ ⁻ during	
the addition of SO_4^{2-} anion.	S12
Figure S11. ¹ H NMR spectral changes of the receptor 5 ²⁺ ·2PF ₆ ⁻ during	
the addition of $HP_2O_7^{3-}$ anion.	S13
Figure S12. ¹ H NMR spectral changes of the receptor $5^{2+}\cdot 2PF_{6}$ during	
the addition of F^- anion.	S14
Figure S13. ¹ H NMR spectral changes observed in the methyl protons of	
the receptor $5^{2+}\cdot 2PF_6^-$ during the addition of $H_2PO_4^-$ ions.	S14
Figure S14. Titration profile of the receptor $5^{2+} \cdot 2PF_6^{-}$ and $H_2PO_4^{-}$ from	
the ¹ H-NMR experiments	S15
Figure S15. Titration profile of the receptor $5^{2+}\cdot 2PF_6^-$ and SO ₄ ²⁻ from	
the ¹ H-NMR experiments	S15
Figure S16. Mass spectrum of the imidazolone $6^+ \cdot PF_6^-$	S16

Compound 2

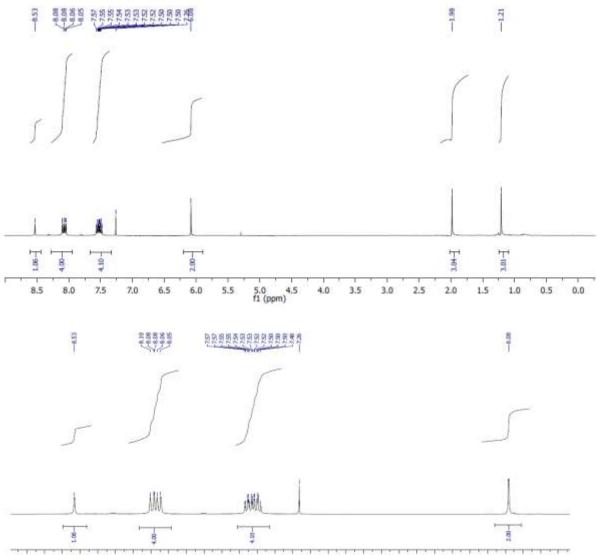


80 70 f1 (ppm) ò



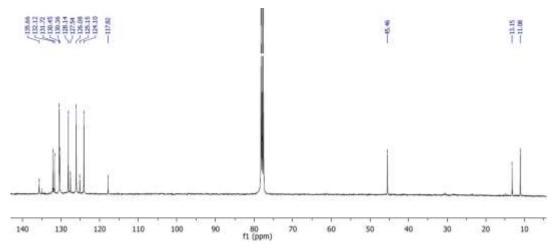
Compound 4

¹H-NMR (CDCl₃, 400 MHz)

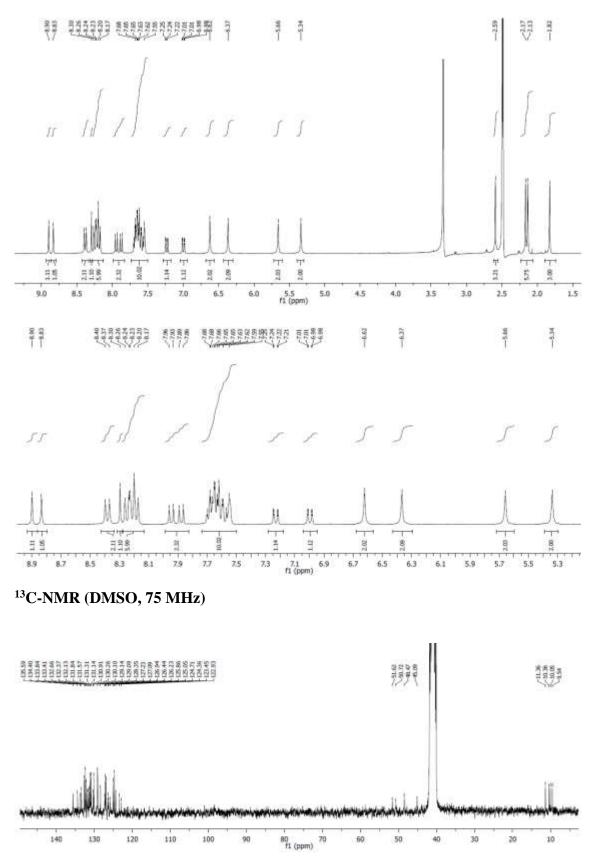


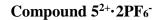
8.8 8.7 8.6 8.5 8.4 8.3 8.2 8.1 8.0 7.9 7.8 7.7 7.6 7.5 7.4 7.3 7.2 7.1 7.0 6.9 6.8 6.7 6.6 6.5 6.4 6.3 6.2 6.1 6.0 5.9 5.8 fl (ppm)

¹³C-NMR (CDCl₃, 100 MHz)

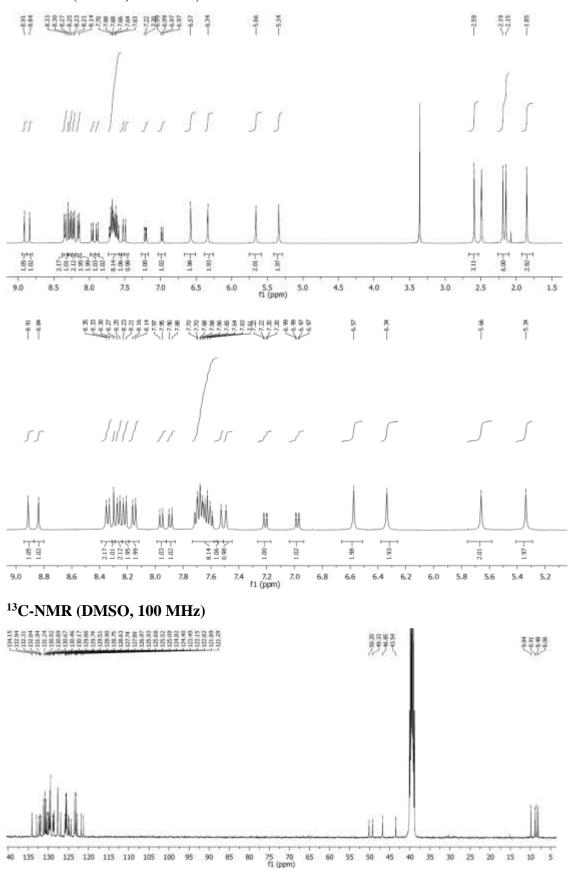


Compound 5²⁺·2Br⁻ ¹H-NMR (DMSO, 400 MHz)









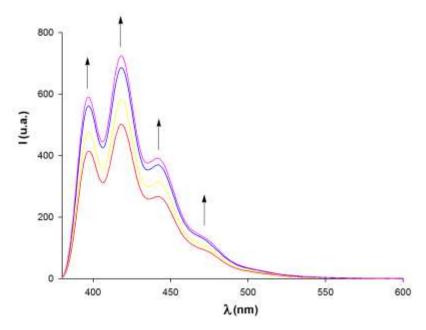


Figure S1. Changes in the fluorescence spectra of receptor $5^{2+}\cdot 2PF_6$ ($c = 1 \cdot 10^{-5}$ M in CH₃CN) upon addition of F⁻ anions from 0 to 1.8 equiv at 20 °C.

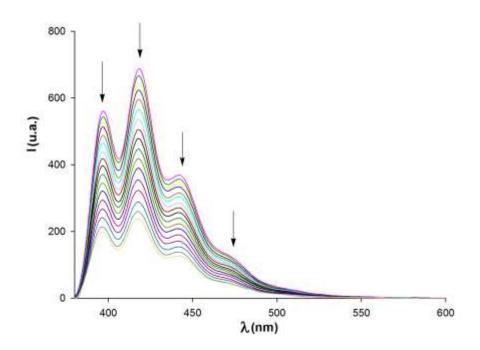


Figure S2. Changes in the fluorescence spectra of receptor $5^{2+}\cdot 2PF_6$ ($c = 1 \cdot 10^{-6}$ M in CH₃CN) upon addition of F⁻ anions from 1.8 to 5 equiv at 20 °C.

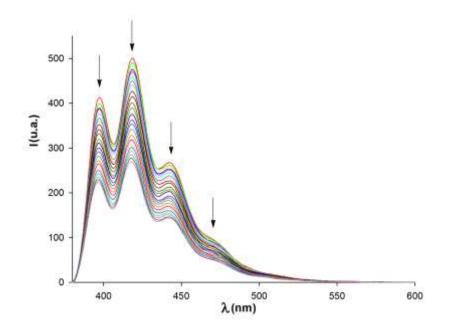


Figure S3. Changes in the fluorescence spectra of receptor $5^{2+}\cdot 2PF_6$ (c = $1\cdot 10^{-6}$ M in CH₃CN) upon addition of AcO⁻ anions at 20 °C.

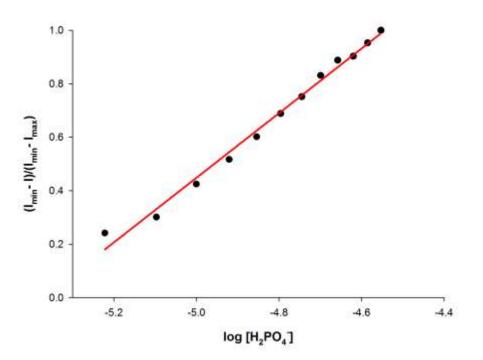


Figure S4. Fluorescence intensity of $5^{2+} \cdot 2PF_6^-$ (in CH₃CN) at each concentration of H₂PO₄⁻ added, normalized between the minimum fluorescence intensity, found at zero equiv of H₂PO₄⁻; and the maximum fluorescence intensity. The detection limit was found at [H₂PO₄⁻] = 3.90 x10⁻⁶M.

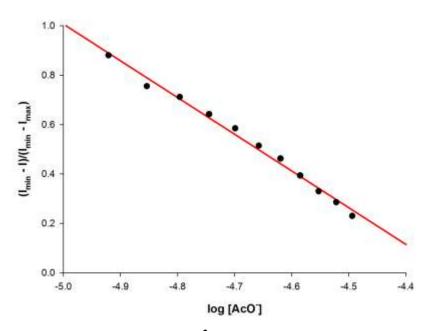


Figure S5. Fluorescence intensity of $5^{2+}\cdot 2PF_6^-$ (in CH₃CN) at each concentration of AcO⁻ added, normalized between the minimum fluorescence intensity, found at zero equiv of AcO⁻; and the maximum fluorescence intensity. The detection limit was found at [AcO⁻] = 4.59 x10⁻⁵M.

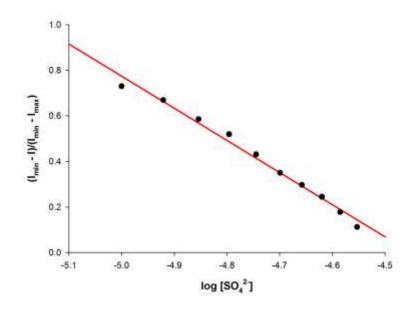


Figure S6. Fluorescence intensity of $5^{2+} \cdot 2PF_6^-$ (in CH₃CN) at each concentration of SO₄²⁻ added, normalized between the minimum fluorescence intensity, found at zero equiv of Pb²⁺; and the maximum fluorescence intensity. The detection limit was found at [SO₄²⁻]= 3.57x10⁻⁵ M.

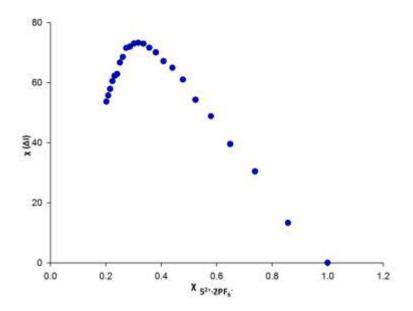


Figure S7. Job plot experiment, using the emission titration data, with a maximum at 0.33 indicating 1:2 stoichiometry for receptor $5^{2+} \cdot 2PF_6^-$ and $H_2PO_4^-$ anions.

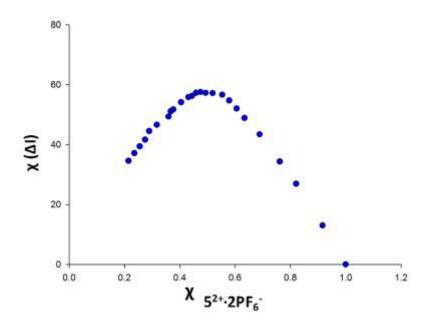


Figure S8. Job plot experiment, using the emission titration data, with a maximum at 0.5 indicating 1:1 stoichiometry for receptor $5^{2+} \cdot 2PF_6^-$ and SO_4^{2-} anions.

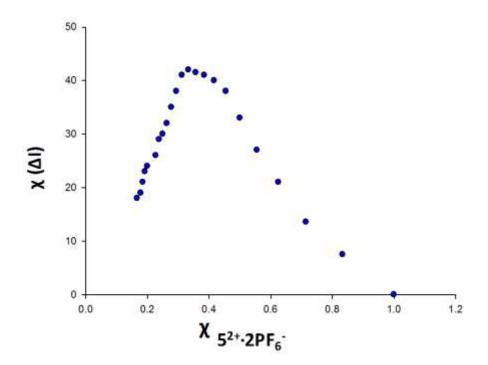


Figure S9. Job plot experiment, using the emission titration data, with a maximum at 0.33 indicating 1:2 stoichiometry for receptor $5^{2+} \cdot 2PF_6^-$ and AcO⁻ anions.

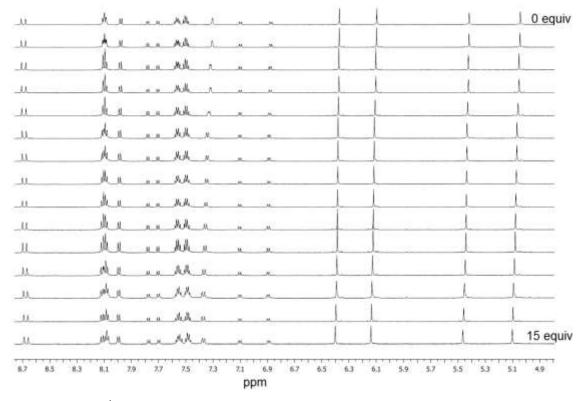


Figure S10. ¹H NMR spectral changes observed in the receptor $5^{2+}\cdot 2PF_6^-$ in CD₃CN/CD₃OD (8:2, v/v) during the addition of up to 15 equivalents of SO₄²⁻ ions.

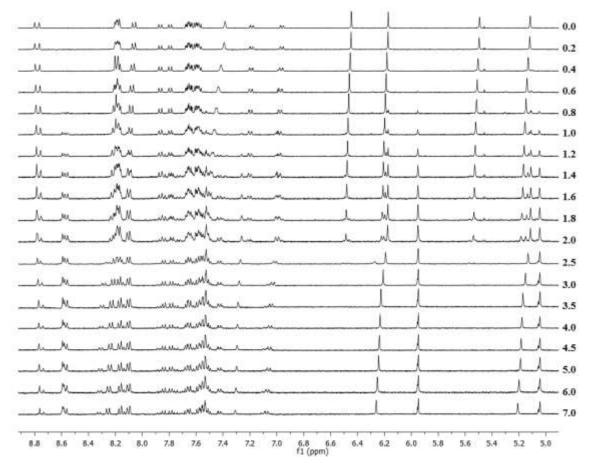


Figure S11. ¹H NMR spectral changes observed in the receptor $5^{2+}\cdot 2PF_6^-$ in CD₃CN/CD₃OD (8:2, v/v) during the addition of up to 15 equivalents of HP₂O₇³⁻ ions.

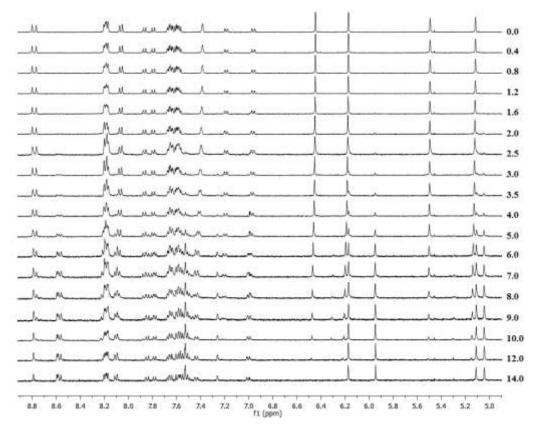


Figure S12. ¹H NMR spectral changes observed in the receptor $5^{2+}\cdot 2PF_6^-$ in CD₃CN/CD₃OD (8:2, v/v) during the addition of up to 15 equivalents of F⁻ ions.

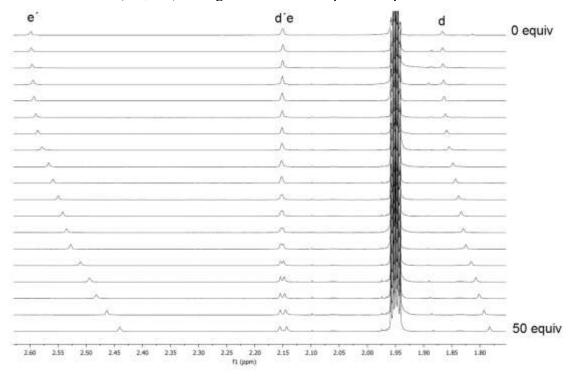


Figure S13. ¹H NMR spectral changes observed in the methyl protons of the receptor $5^{2+}\cdot 2PF_6^-$ in CD₃CN/CD₃OD (8:2, v/v) during the addition of up to 50 equivalents of H₂PO₄⁻ ions.

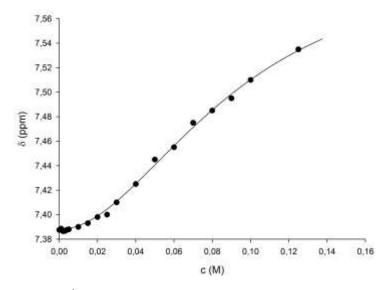


Figure S14. Changes in ¹H NMR spectra in the H_a proton of the receptor **5**²⁺•**2PF**6⁻upon addition of increasing amounts of H₂PO₄⁻ anions. Points represent experimental data, continuous lines represent calculate curve.

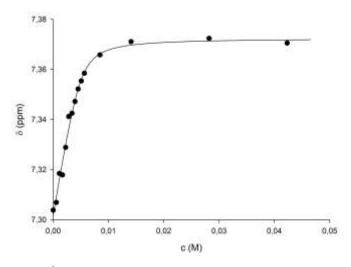


Figure S15. Changes in ¹H NMR spectra in the H_a proton of the receptor **5**²⁺•**2PF**₆⁻upon addition of increasing amounts of SO₄²⁻ anions. Points represent experimental data, continuous lines represent calculate curve.

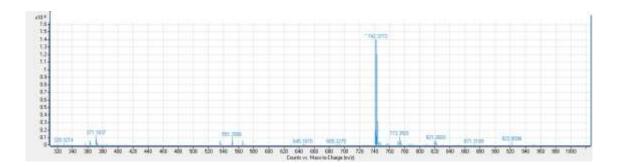


Figure S16. Mass spectrum of the solution of $5^{2+} \cdot 2PF_6^-$ in CH₃CN after the addition of 3 equivalent of HP₂O₇³⁻ ions, indicating the formation of the imidazolone $6^+ \cdot PF_6^-$ C₅₂H₄₅N₄O⁺ (M⁺ = 741.36)